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**Effects of anions size on the redox behavior of poly(o-toluidine) in acid solutions.
An in situ vis-NIR cyclic spectroelectrogravimetry study.**

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ABSTRACT

The combination of different in situ techniques allows a more complete analysis combining several physical and chemical processes of polymer-modified electrodes to be obtained. The hybrid technique -cyclic electrogravimetry coupled with visible-near infrared spectroscopy- was used to study in depth the electrochemistry of poly(o-toluidine) (POT) thin films. In this work, a POT modified electrode was polarized between the pernigraniline and leucoemeraldine forms in 0.5 M HNO₃ and 0.5 M HClO₄ solutions. The electrochromic properties of POT-ClO₄⁻ and POT-NO₃⁻ were localized in the experimental potential range associated to the transfer of charged species during the redox reactions. The analysis of all the experimental results has allowed the main contribution of proton, anion and water transfers to be separated in three redox transitions together with the electrochromic properties of POT. Semi-quantitatively, the better electrochemical and electrochromic properties of both films take place predominantly at potentials where the higher accumulation of the radical cations (polarons and bipolarons) in the polymer structure involves great lattice reorganization. The analysis of this work is a complete objective manner to discriminate the better conditions where a film can be used for a specific technical application.

Keywords: Intrinsically conducting polymer, Reflectance, Spectroelectrochemistry, Electrochemical quartz crystal microbalance.

1. INTRODUCTION

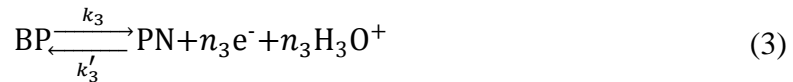
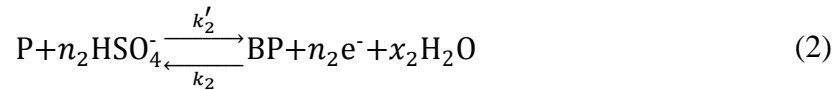
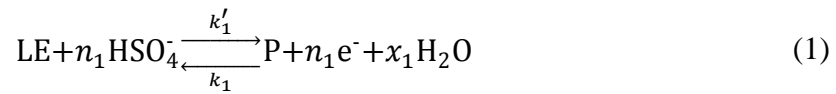
Over the past several decades, spectroelectrochemistry has revealed as a powerful tool for the elucidation of electron transfer processes [1,2]. One direction in this field of study is to modify the kind of interaction between matter and radiated energy. Interesting information about different physical properties can be obtained by analyzing altogether the information provided by the coupling of different in situ techniques. The high reflectance gold electrodes of electrochemical quartz crystal microbalances (EQCM) allow the spectroscopic radiation to be recorded together with the mass changes of a modified electrode during a cyclic voltammogram. In this manner, the cyclic spectroelectrogravimetry could be a good candidate to obtain a complete analysis during the continuous polarization of polymer-modified electrodes. It avoids any question of polymer history effects when sequential experiments and subsequent correlation of the data neglect temporal structural changes, such as relaxation within polymers [3,4]. Afterward, a subsequent intensive mathematical analysis is necessary to extract the interesting information from the large amount of experimental data. The coupling of different in situ techniques has been shown to be a powerful approach for studying electrochemically active materials as it allows different physical and chemical properties to be correlated [5–7].

Intrinsically conducting polymers (ICPs) have been the subject of many research efforts because of their environmental stability, good processability and their low relative cost [8–14]. The electrical conduction through the film needs to pass some potential barriers: 1) the electron transfer at the electrode|film interface, b) the electron transport through the film which will be neglected in this layer by supposing that deposited polymers are thin enough and finally c) the exchange of charged species (cations and anions) at the film|solution interface necessary for the electrical charge balance. During these processes, free solvent transfer between the solution and the polymer often occurs.

The understanding of the electrochemical behavior of these permselective polymers is essential and provides an important motivation for the present study since the potentiometric response, electrical conductivity, insertion kinetics of ions and/or free water molecules depends significantly on the size of incorporated ions [15,16]. Spectroscopy will allow an unequivocally identification of the electroactive species causing these transfers during the complex electrochemical redox reactions of conducting polymers [5]. Elucidation of these mechanisms will help to improve their properties of

interest in several technologies like electrochromic devices [17,18], photogalvanic cells [19], artificial muscles [20,21], light-emitting electrochemical cells [22,23], biosensors [24–28], biofuel cells-based devices [29,30].

Among ICPs, poly(o-toluidine) (POT) is one of the most studied polymers [31–39]. Previously in a diprotic acid [5], three different redox reaction steps were isolated between the non-conductive forms of POT: the oxidized form (pernigraniline salt, PN) and the reduced form (leucoemeraldine salt, LE). The intermediated conductive forms of POT (polarons, P, and bipolarons, BP) were considered as single entities with specific and independently measurable electrochemical properties. The reaction scheme proposed was:



where n_i represents the number of electrons or available sites inside the polymer for the insertion/expulsion of ionic species and x_i is the number of water molecules exchanged during the electrochemical reaction. The redox reactions involving the anions (HSO_4^-) and water molecules are faster than the redox reaction involving hydrated protons, in other words, $k'_1 > k_3$ and $k'_2 > k_3$ in the oxidation direction and $k_1 > k'_3$ and $k_2 > k'_3$ in the reduction direction [40].

In this work, POT films were generated in 0.5 M H_2SO_4 aqueous solutions on a gold electrode. These modified electrodes were studied by in situ spectroelectrogravimetry in two different aqueous solutions of 0.5 M monoprotic acids: HClO_4 and HNO_3 . The influence of these anions on the electrochemical response of the film and on the transfers of the other species during the electrochemical transitions was compared.

2. THEORETICAL APPROACH

The simplest model used to interpret voltammetric curves of electroactive films is assume that the reduced and the oxidized forms of the film are adsorbed and the Nernst equation is satisfied [41]. Simultaneously, the redox process involves insertion/expulsion of charged species for electrical stability. As an example for anions, it comes:



where A^- is the anion, $\langle P \rangle$ is the site in the host polymer and $\langle P^+, A^- \rangle$ is occupied site in the polymer matrix doped with anions. A similar approach can be made for the exchange of cations and solvent [42,43].

At the polymer|solution interface, the flux of anions is $J_{A^-} = -d_f(dC_{A^-}/dt)$ where d_f is the polymer film thickness and C_{A^-} the concentration of anions. For very thin films, the insertion/expulsion of species at the polymer|solution interface is considered the rate-limiting step and the change of C_{A^-} is:

$$\frac{dC_{A^-}}{dt} = -k_{A^-}(C_{A^-} - C_{A^-}^{min}) + k'_{A^-}(C_{A^-}^{max} - C_{A^-})C_{A^-}^{sol} \quad (5)$$

where $C_{A^-}^{max}$ and $C_{A^-}^{min}$ are the maximum and minimum concentrations of the electrochromic sites where anions can be inserted and $C_{A^-}^{sol}$ is the outer concentration of anions. This equation can be simplified as:

$$\frac{dC_{A^-}}{dt} = -k_{A^-}(C_{A^-}) + k'_{A^-}(C_{A^-}^{max} - C_{A^-})C_{A^-}^{sol} \quad (6)$$

when $C_{A^-}^{min}=0$.

For a potential sweep where eq (4) occurs, the current dq/dt considering q in C cm^{-2} is equal to:

$$\frac{dq}{dt} = -v d_f F \frac{dC_{A^-}}{dE} \quad (7)$$

where $E = E_0 - vt$, v is the scan rate, F is the Faraday constant and

$$\frac{dC_{A^-}}{dE} = k'_{A^-} k_{A^-} C_{A^-}^{sol} \frac{(b'_{A^-} - b_{A^-}) C_{A^-}^{max}}{(k'_{A^-} C_{A^-}^{sol} + k_{A^-})^2} \quad (8)$$

Here, the kinetic constants are supposed to be activated by the potential (the classical Tafel law), such as:

$$k'_{A^-} C_{A^-}^{sol} = k'_{A^-0} C_{A^-}^{sol} \exp(b'_{A^-} E) = k'_{A^-00} \exp(b'_{A^-} (E - E_p)) \quad (9)$$

$$k_{A^-} = k_{A^-0} \exp(b_{A^-} E) = k_{A^-00} \exp(b_{A^-} (E - E_p)) \quad (10)$$

where

$$b_{A^-} - b'_{A^-} = \frac{nF}{RT} \quad (11)$$

and R is the gas constant, T is the absolute temperature, E_p is the peak potential where the maximum current is reached. nF/RT is about 40 V^{-1} at 298 K for a single electron transfer ($n=1$) [44,45].

Finally, we can obtain

$$\frac{dq}{dt} = vd_f \frac{F^2 C_{A^-}^{max} \exp\left((b_{A^-} - b'_{A^-})(E - E_p)\right)}{RT \left[1 + \exp\left((b_{A^-} - b'_{A^-})(E - E_p)\right)\right]^2} \quad (12)$$

which may be mathematically rewritten by considering $(e^{-x} + e^{+x})/2 = \cosh(x)$ as:

$$\frac{dq}{dt} = vd_f \frac{F(b_{A^-} - b'_{A^-})C_{A^-}^{max}}{\cosh^2\left[\frac{(b_{A^-} - b'_{A^-})}{2}(E - E_p)\right]} \quad (13)$$

$$\text{where } vd_f F(b_{A^-} - b'_{A^-})C_{A^-}^{max} = \left(\frac{dq}{dt}\right)_{peak}$$

This equation can be easily convert in changes of mass considering the molar mass of anion, M_{A^-} and the direction of anion transfer δ_{A^-} during the electrochemical reaction as.

$$\frac{dm}{dt} = vd_f \frac{\delta_{A^-} M_{A^-} (b_{A^-} - b'_{A^-}) C_{A^-}^{max}}{\cosh^2\left[\frac{(b_{A^-} - b'_{A^-})}{2}(E - E_p)\right]} \quad (14)$$

$$\text{where } vd_f \delta_{A^-} M_{A^-} (b_{A^-} - b'_{A^-}) C_{A^-}^{max} = \left(\frac{dm}{dt}\right)_{peak}$$

In this case $\delta_{A^-}=+1$ since the anions are inserted during the oxidation. On the contrary, δ is negative for species inserted during the reduction (i.e. cations).

Likewise, the absorbance changes at each wavelength λ , dA^λ/dt of an electrochromic film can be related to the charge when the Lambert-Beer's law is satisfied:

$$\frac{dA^\lambda}{dt} = \frac{1}{F} \varepsilon_{A^-}^\lambda l \frac{dq}{dt} \quad (15)$$

where $\varepsilon_{A^-}^\lambda$ is the absorbance coefficient of the electrochromic sites and l the length of the light pass. So,

$$\frac{dA^\lambda}{dt} = vd_f \frac{\varepsilon_{A^-}^\lambda l (b_{A^-} - b'_{A^-}) C_{A^-}^{max}}{\cosh^2\left[\frac{(b_{A^-} - b'_{A^-})}{2}(E - E_p)\right]} \quad (16)$$

$$\text{where } vd_f \varepsilon_{A^-}^\lambda l (b_{A^-} - b'_{A^-}) C_{A^-}^{max} = \left(\frac{dA^\lambda}{dt}\right)_{peak}$$

By some algebra, three instantaneous crossed functions may be calculated. From eq (14) and eq (13), it is easy to obtain a good estimation of the molar mass of the anion participating [46]:

$$F \left(\frac{dm/dt}{dq/dt}\right) = F \left(\frac{dm}{dq}\right) = \delta_{A^-} M_{A^-} \quad (17)$$

From eq (16) and eq (13), the absorbance coefficient is obtained (also called electrochromic efficiency) [47]:

$$F \left(\frac{dA^\lambda/dt}{dq/dt} \right) = F \left(\frac{dA^\lambda}{dq} \right) = l\varepsilon_A^\lambda \quad (18)$$

In this function, $l\varepsilon_A^\lambda$ is positive when the absorbance increases during the oxidation reaction ($dq/dt > 0$) or diminishes during the reduction reaction ($dq/dt < 0$). On the contrary, $l\varepsilon_A^\lambda$ is negative when the absorbance increases during the reduction reaction or diminishes during the oxidation reaction.

It is also possible to obtain a different electrochromic efficiency at a given wavelength related to the changes of film mass owing to the anion transfer between the polymer and the solution from eq (16) and eq (14) [5]:

$$\frac{dA^\lambda/dt}{dm/dt} = \frac{dA^\lambda}{dm} = \frac{l\varepsilon_A^\lambda}{\delta_A \cdot M_A} \quad (19)$$

Here, positive mass electrochromic efficiency can be found when absorbance increases during the insertion of anions ($dm/dt > 0$) or when absorbance diminishes during the expulsion of anions ($dm/dt < 0$). Likewise, this electrochromic efficiency is negative if the absorbance increases during the expulsion of anions or if the absorbance diminishes during the insertion of anions. In this manner, we remove the contributions due to catalytic reactions if these reactions do not affect both the color and the mass.

3. EXPERIMENTAL

3.1. Film preparation

The electrochemical polymer deposition and characterization was controlled by cyclic voltammetry (CV) through an AUTOLAB potentiostat-galvanostat set up (PGSTAT302) and an Electrochemical Quartz Crystal Microbalance (RQCM, Maxtek, Inc.). The three-electrode cell involves the Ag|AgCl|KCl_(sat) reference electrode (RE-1C, Bas Inc., Japan), a platinum counter electrode and the gold working electrode with an area of 0.3 cm². It was patterned on a 9 MHz quartz crystal resonator (TEMEX, France). All solutions were prepared with deionized and double-distilled water. The POT polymerization solution was 0.5 M H₂SO₄ (Fisher scientific, for trace analysis) and 0.2 M o-toluidine (Fluka) used as received. POT was formed through 100 voltammetric cycles between -0.1 V and +0.9 V with a 100 mV s⁻¹ scan rate.

3.2. Film characterization

The resulting modified electrode was cycled between -0.1 V and 0.9 V vs SCE in 0.5 M HNO₃ (Normapur AnalR) or in 0.5 M HClO₄ (Normapur RP) at 100 mV s⁻¹ until a stable cyclic voltammogram was obtained (after 10 cycles) assuming the complete substitution of bisulfate anions inserted during the polymerization process by the corresponding anion in POT (pH≈0.45 in both solutions).

After that, in order to monitor the chemical changes of POT during a cyclic voltammogram, a vis-NIR absorbance spectrum between 380 nm and 1100 nm was simultaneously collected in a PC through the EPP2000 Portable Fiber Optic Spectrometer (StellarNet Inc.) each 9 mV (about 11 spectra/second) using a 2x2 cm high-transmittance glass cell from Hellma[®] (OG quality) as the electrochemical cell. The color response received was recorded as number of counts at each wavelength, n_{counts}^{λ} , which is directly converted into light intensity at a given wavelength, I^{λ} , reflected on the gold electrode of the EQCM. In this work, it is preferred to convert into an apparent time derivative of absorbance at a specific wavelength (A^{λ}) which was shown to be better for comparing with mass and electrical charge by means of the Lambert-Beer's law:

$$A^{\lambda} = -\log(I^{\lambda}/I_0^{\lambda}) = -\log(I^{\lambda}) + \log(I_0^{\lambda}) \quad (20)$$

$$\frac{dA^{\lambda}}{dt} = \frac{d(-\log(I^{\lambda}) + \log(I_0^{\lambda}))}{dt} = -\frac{d(\ln(I^{\lambda}))}{\ln(10) \cdot dt} = \frac{1}{2.303I^{\lambda}} \frac{dI^{\lambda}}{dt} \quad (21)$$

Since, the intensity of transmitted light, I^{λ} , is proportional to the number of counts, n_{counts}^{λ} , received by the diode at this wavelength, it can be written:

$$\frac{dA^{\lambda}}{dt} = \frac{1}{2.303I^{\lambda}} \frac{dI^{\lambda}}{dt} = \frac{1}{2.303n_{counts}^{\lambda}} \frac{dn_{counts}^{\lambda}}{dt} \quad (22)$$

The analysis of the raw data was carried out using the software Mathcad 14 (Mathsoft) where all data collected in the time domain were treated by using the Savitzky–Golay filter.

4. RESULTS AND DISCUSSION

4.1. Optical characterization

In a redox system like POT films in H₂SO₄ solutions, the total current is the sum of the partial currents relative to the contribution of 3 overlapped electrochemical

processes in the oxidation direction and the reduction direction of a voltammogram scan (eq. (1)-(3)):

$$\left. \frac{dq_T}{dt} \right|_{ox} = \frac{dq_{LE \rightarrow P}}{dt} + \frac{dq_{P \rightarrow BP}}{dt} + \frac{dq_{BP \rightarrow PN}}{dt}$$

and

$$\left. \frac{dq_T}{dt} \right|_{red} = \frac{dq_{PN \rightarrow BP}}{dt} + \frac{dq_{BP \rightarrow P}}{dt} + \frac{dq_{P \rightarrow LE}}{dt}$$
(23)

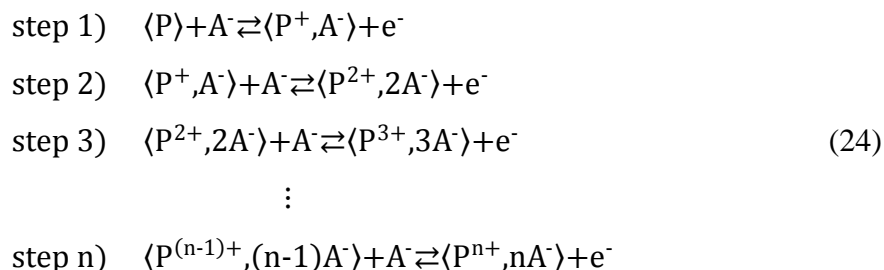
The voltammetric response of POT films involves two successive quasireversible redox reactions corresponding to eq (1) and (3) [48]. Only the coupling of complementary techniques allows the redox reaction in eq (2) to be discriminated. By Vis-NIR spectroscopy, the formation or elimination of all different redox species of POT can be followed at specific wavelengths [5]. However, a relationship one wavelength-one specie is not always possible since several species could contribute to the absorbance at a given wavelength in ICPs, and therefore it is necessary to select the correct wavelength for proper identification of electrochromic species. With the intention to localize the redox transitions in the potential scale, POT films were studied by using the derivative voltabsorptometric method. This method allows the potential dependence of the absorption bands to be effectively monitored by discriminating the different contributions of the overlapped signals given at close redox potentials [49].

<Figure 1>

Figure 1 shows the three-dimensional (3D) surface plot of the time derivative of absorbance (dA^λ/dt) between 400 nm and 900 nm together with the corresponding stable voltammogram for two forms of POT with different anionic composition (POT-NO₃⁻ and POT-ClO₄⁻). As it was expected, two main voltammetric peak couples are observed centered at 0.1-0.2 V and at 0.5-0.6 V depending to the anion in solution. Some mountains ($dA^\lambda/dt > 0$) and valleys ($dA^\lambda/dt < 0$) observed in the 3D plot coincide with the voltammetric peaks: the changes of absorbance derivative about 435 nm, and between 700 and 900 nm can be associated with the more cathodic peak couple and the changes of absorbance derivative between 450 and 600 nm with the more anodic peak couple. In addition, the absorption band at wavelengths around 700 nm with maximum and minimum takes place at potentials between voltammetric peaks.

The wide range of wavelengths obtained to explain the same electrochemical transition is due to the complexity of the electrochemical system studied [50,51]. The electrochemical oxidation or reduction of film goes through n consecutive stoichiometric

steps related to n consecutive ionization potentials where the structural composition of the polymeric backbone (single or double links, radicalization, protonation or deprotonation) and the surrounding environment (solvation and counterions) inside of film change. An example taking into account the eq (4) is given here:



At every step, the structure and anion composition of polymer change simultaneously involving little hypsochromic and bathochromic shifts in optical spectra [47,52]. The step 1 can be followed by the changes of absorbance at λ_1 , the step 2 by λ_2 and step 3 by λ_3 . Therefore, a single electrochemical process can be followed with a set of characteristics wavelengths. In POT films, three electrochemical processes are overlapped and this complicates the good election of interesting wavelengths. With the intention of explaining the most of the electrochemical processes in both POT forms and for practical discussions, only four wavelengths were accurately selected taking into account the bibliography and previous results [5,39,53–60]. Figure 2 shows the curves of dA^{435nm}/dt , dA^{810nm}/dt , dA^{700nm}/dt and dA^{540nm}/dt during the oxidation scan following the evolution of isolated polarons (435 nm), the coil structures of polarons when the anions interact strongly with the $C=NH^+=C$ group (810 nm), bipolaron forms (700 nm) and the pernigraniline forms (540 nm) of POT, respectively.

In the oxidation direction for both POT forms, the peak around 0.25 V corresponds to the oxidation of LE form since dA^{435nm}/dt and dA^{810nm}/dt are positives, which indicate the formation of isolated polarons and the formation of coil structures, respectively. That corresponds to the oxidation semi-reaction of eq (1). Afterwards, the formation of coils structures is the principal form in POT since the isolated polarons disappears (dA^{435nm}/dt is negative) at 0.45 V. At the same time, bipolaron structures are formed as the positive changes of absorbance at 700 nm show (oxidation semi-reaction of eq (2)). Finally, the coil structures disappears (dA^{810nm}/dt is negative) and the PN form appears (dA^{540nm}/dt are positive) corresponding to the oxidation semi-reaction of eq (3). The formation of different structures during redox transitions of POT

is not affected by the anion in solution since the behavior is similar than that of POT in H₂SO₄ solution in similar experimental conditions.

As it can be seen in Figure 2, only the experimental dA^{435nm}/dt curves show a typical profile due to the formation and elimination of only one specie, in this case the isolated polarons. Other experimental curves show more complexes profiles due to the overlapping of at least three electrochromic transitions. Consequently, the simplistic association between one wavelength-one electrochromic species is not always valid.

Like in current (eq (23)), the absorbance derivative at a given wavelength is the sum of the variation of the concentration considering all processes involved. So, the total absorbance derivative at 435 nm can be expressed as following [5]:

$$\left. \frac{dA_T^{435nm}}{dt} \right|_{ox} = \frac{dA_{LE \rightarrow P}^{435nm}}{dt} + \frac{dA_{P \rightarrow BP}^{435nm}}{dt}$$

and

$$\left. \frac{dA_T^{435nm}}{dt} \right|_{red} = \frac{dA_{BP \rightarrow P}^{435nm}}{dt} + \frac{dA_{P \rightarrow LE}^{435nm}}{dt}$$
(25)

where the formation of isolated structures allows the LE→P transition in the oxidation directions or the BP→P transition in the reduction direction to be followed. On the contrary, the disappearances of these structures are related to the P→BP or the P→LE transitions.

At other selected wavelength ($\lambda = 540$ nm, 700 nm and 810 nm), the contribution of BP↔PN transitions to the total absorbance derivative must be also considered due to the mixture of species formed during the n redox steps (eq. (24)) .

$$\left. \frac{dA_T^\lambda}{dt} \right|_{ox} = \frac{dA_{LE \rightarrow P}^\lambda}{dt} + \frac{dA_{P \rightarrow BP}^\lambda}{dt} + \frac{dA_{BP \rightarrow PN}^\lambda}{dt}$$

and

$$\left. \frac{dA_T^\lambda}{dt} \right|_{red} = \frac{dA_{PN \rightarrow BP}^\lambda}{dt} + \frac{dA_{BP \rightarrow P}^\lambda}{dt} + \frac{dA_{P \rightarrow LE}^\lambda}{dt}$$
(26)

<Figure 2>

Accordingly, with the expression in eq (16) and the above expressions for the total absorbance derivative at these four wavelengths, we can try to fit the shape of experimental dA^λ/dt curves. Figure 2 shows experimental dA^λ/dt in the oxidation scan direction, which can be deconvoluted in two or three partial curves in both solutions with good results.

<Table 1>

As it can be seen, spectroelectrochemistry is a good technique for the full assignment of the different transitions to the color changes and in particular in this work, a better understanding of the influence of the anion size. However, a large amount of information can be obtained from this methodology. In this work, we propose to take into consideration this data complexity, not as a hindrance, but as an advantage, which will allow the influence of the nature of the anion in the solution and the effect on other transfers to be studied in depth during the electrochemical reactions of POT under potentiodynamic conditions. Table 1 collects numerical values obtained from the fitting of these experimental curves to the model proposed by eq (16). In this case, the extinction coefficient of species at a given wavelength is unknown and for this reason, Table 1 only shows $(dA^\lambda/dt)_{peak}$.

Owing to the amount of information obtained, the deconvolution procedure allows the partial contribution of interest at a given wavelength to be hold. As it was commented above, each wavelength has been mainly associated to an electrochromic transition: A^{435nm} is associated to the formation of isolated polarons, A^{810nm} to coil structures, A^{700nm} to bipolaron forms and A^{540nm} to the pernigraniline forms of POT. Accordingly, we can discriminate the information of interest and we can study only the most interesting evolution of $dA_{LE\leftrightarrow P}^{435nm}/dt$, $dA_{LE\leftrightarrow P}^{810nm}/dt$, $dA_{P\leftrightarrow BP}^{700nm}/dt$ and $dA_{BP\leftrightarrow PN}^{540nm}/dt$.

As it is observed from data in Table 1, both forms of POT have remarkable differences during the electrochromic redox transitions. The separation in potential scale between the oxidation peak and the corresponding reduction peak for each electrochromic partial contribution of dA^λ/dt is lower for POT- NO_3^- (in average 30 mV) than POT- ClO_4^- (in average 100 mV). In voltammetry simulations, the peak separation is reduced when the uncompensated resistance of electrochemical system decreases [61]. The higher mobility of NO_3^- should reduce the uncompensated resistance in the electrochemical cell reducing the peak separation contrarily to larger ions as ClO_4^- . The lower peak separation is observed by the $P\leftrightarrow BP$ transitions in both POT forms (dA^{710nm}/dt). For POT- NO_3^- , it is practically 0 V and for POT- ClO_4^- , it is about 60 mV. This separation could be expected if the $P\leftrightarrow BP$ transition is an “easy” structure phase transition like a reorganization of two polaronic sites to form a bipolaronic site inside of polymeric structure by means of a thermodynamic equilibrium ($2P\leftrightarrow BP$).

For each transition, an estimation of its reversibility can be obtained from the $(dA^\lambda/dt)_{peak}^{ox}/(dA^\lambda/dt)_{peak}^{red}$ ratio. In both forms of POT, the electrochromic LE \leftrightarrow P and P \leftrightarrow BP transitions shows a good reversibility close to 1 at 435 nm, 810 nm and 700 nm. The main difference is found in the BP \leftrightarrow PN transitions. In HClO₄ solution, POT-ClO₄⁻ shows a peaks ratio close to 1. However, POT-NO₃⁻ shows the reduction peak for the PN \rightarrow BP transition $((dA_{PN\rightarrow BP}^{540nm}/dt)_{peak}^{red})$ about two times higher than the oxidation peak for the BP \rightarrow PN transition $((dA_{BP\rightarrow PN}^{540nm}/dt)_{peak}^{ox})$.

Previous results are also in accordance with the symmetry showed for the voltammetric peak of POT in these solutions. In this last transition, $(b - b')$ are 70 V⁻¹ and 53 V⁻¹ for the POT-NO₃⁻ and POT-ClO₄⁻, respectively. These values are higher with respect to the other values, which are close to 20 V⁻¹. From $(b - b')$, the width of the half height $(\Delta E_{p,i(1/2)} = (E_{p,i(1/2)} - E_{p,i(-1/2)}))$ of simulated curves of all separated transitions can be calculated [62]. Based on the Frumkin-type isotherm model [41,63–65], $(b - b')$ should be 40 V⁻¹ for a single electron transfer at 298 K which correspond to a width at half height of 90.6 mV per electron transfer (90.6/n mV) considering non lateral interactions between oxidized and reduced forms of structure film on working electrode during a monoelectronic transfer. 20 V⁻¹ are considered as a classic result of this kind of films which correspond to a width at half height of 180/n mV (a wider peak than this one expected without interactions) [5,66]. This fact points to repulsive lateral interactions between random distributed electrochromic sites in permselective redox films adsorbed on an electrode. In the opposite side, the PN \rightarrow BP transitions in both solutions present a narrower peak than this one expected without interactions $((b - b') > 40 \text{ V}^{-1})$. In this case, we can assume attractive interactions between electrochromic sites much more favored by the size of inserted NO₃⁻ in the POT-NO₃⁻ form than by inserted ClO₄⁻ in the POT-ClO₄⁻. Moreover, the insertion of proton during the PN \rightarrow BP transitions also could be favored by the high concentration of protons in solution (pH \approx 0.45).

In spite of this, we still have some difficulties to give an exclusive explanation since alternative perspectives could be considered: the existence of a previous chemically controlled step of the single-electron transfers [67] or an apparent transfer of 0.5 electrons may also be considered with the same numerical result.

4.2. Voltammetric considerations

As it was seen, the POT electrochemistry is defined by the anions exchange because they are able to modify the macromolecular structure of POT as shown by the spectroscopic results. Figure 3 shows the cyclic voltammograms of POT in both acid solutions. From spectroscopic deconvolutions, the complete electrochemical process of both POT-ClO₄⁻ and POT-NO₃⁻ forms can also be described by a sequence of three discrete but overlapped redox transitions. The redox model proposed in eq (1)-(3) can be fitted using eq (13) and (23) and taking into account the results of spectroscopic deconvolutions.

<Figure 3>

Table 2 collects the values of the CV deconvolutions. Good correlations are obtained between CV and dA^λ/dt deconvolutions. It must be taken into account that the charge flux is due to the global electrochemical reactions in the film (all n steps in eq (24)) and the absorbance changes at a single wavelength is due only to one specific structure of film (one step in eq (24)). Taking this into account, the couple of peaks around 0.2 V corresponds to the LE \leftrightarrow P transitions (eq (1)) and the BP \leftrightarrow PN transitions (eq (3)) are around 0.55 V. In addition, the P \leftrightarrow BP transitions (eq (2)) can be taken into account in the voltammogram from spectroscopic results (around 0.45 V).

<Table 2>

As it was expected, the size of ClO₄⁻ increases the peak potential separation of LE \leftrightarrow P and BP \leftrightarrow PN transitions (about 140 mV) if it is compared with the smaller size of NO₃⁻ (about 70 mV). Likewise, P \leftrightarrow BP transitions are separated about 0 V in HNO₃ solution and about 100 mV in HClO₄ solutions. The thermodynamic equilibrium assumed during this transitions (2P \leftrightarrow BP) can be facilitated by the presence of smaller anions inside de polymer, as NO₃⁻ does.

Table 2 also shows the surface concentration of electroactive sites (Γ_{max}) calculated from peak intensity (eq (13)) since the thickness of POT films is unknown. Theoretically, the same number of electroactive sites is expected to be reduced or oxidized, however, it is not observed in both POT forms. On the one hand, the amount of oxidized sites during the BP \rightarrow PN transition is about 3.8 $\mu\text{mol cm}^{-2}$ and on the other hand, the amount of reduced sites during the PN \rightarrow BP transition is about 3 $\mu\text{mol cm}^{-2}$. In these electrochemical transitions, protons transfers are involved, therefore, a chemical re-protonation of recently oxidized sites (deprotonated) during the BP \rightarrow PN transition could

increase the charge flux necessary for a complete oxidation of film owing to the low solution pH [68]. Likewise, the number of reduced sites is larger (about $4.1 \mu\text{mol cm}^{-2}$) than the number of oxidized sites (about $2.5 \mu\text{mol cm}^{-2}$) in both POT forms for the electrochemical transitions where anions are involved ($\text{LE} \leftrightarrow \text{P}$ and $\text{P} \leftrightarrow \text{BP}$ transitions). Therefore, the recently reduced sites during the $\text{BP} \rightarrow \text{P} \rightarrow \text{PN}$ transition can be oxidized by oxygen in the solution, in this way, the charge flux necessary for the complete reduction of film increases. Similar experiments with higher pH solutions or solutions without oxygen should be necessary to confirm these assumptions. Anyway, the number of electroactive sites is similar independently of anion in solution, therefore, the structure of POT generated in H_2SO_4 solution is not a limit factor to the transfer and transport of both anions. Therefore, the properties of POT are due to the anion-structure interactions during the insertion or expulsion of anions.

4.3. Mass evolution

From these results, we can speculate about the transfer of species. The total mass derivative can be expressed as the sum of the mass flux due to the three transitions:

$$\left. \frac{dm_T}{dt} \right|_{ox} = \frac{dm_{\text{LE} \rightarrow \text{P}}}{dt} + \frac{dm_{\text{P} \rightarrow \text{BP}}}{dt} + \frac{dm_{\text{BP} \rightarrow \text{PN}}}{dt}$$

and

$$\left. \frac{dm_T}{dt} \right|_{red} = \frac{dm_{\text{PN} \rightarrow \text{BP}}}{dt} + \frac{dm_{\text{BP} \rightarrow \text{P}}}{dt} + \frac{dm_{\text{P} \rightarrow \text{LE}}}{dt}$$
(27)

where an anion-free solvent coupling transfer in opposite directions takes place during the $\text{LE} \leftrightarrow \text{P}$ and $\text{P} \leftrightarrow \text{BP}$ transitions (exclusion effect), and protons transfer takes place during the $\text{BP} \leftrightarrow \text{PN}$ transitions. We tried to fit the experimental dm/dt curves in both solutions taking into account eq (27) in a similar way than Figure 2 and Figure 3.

In a first approximation, if the potential peaks of deconvoluted current intensity and/or absorbance derivative are chosen (Table 1 and Table 2), four simulated curves were necessary to obtain good fittings of dm/dt . A new strategy to simulate the experimental mass derivative was searched. From current deconvolutions, it was assumed that all passed charge in the simulated curves of dq/dt leads to the creation of redox centers. Therefore, the ionic transfer can be calculated using eq (14) [5,48]. Once obtained dm/dt partial curves for the ionic transfers, the mass flux of solvent can be calculated as:

$$\frac{dm_{\text{H}_2\text{O}}}{dt} = \frac{dm_T}{dt} - \frac{dm_{\text{A}^-}}{dt} - \frac{dm_{\text{H}_3\text{O}^+}}{dt} \quad (28)$$

In the oxidation direction:

$$\left. \frac{dm_{\text{A}^-}}{dt} \right|_{\text{ox}} = \frac{1}{F} \delta_{\text{A}^-} M_{\text{A}^-} \left(\frac{dq_{\text{LE} \rightarrow \text{P}}}{dt} + \frac{dq_{\text{P} \rightarrow \text{BP}}}{dt} \right)$$

and

$$\left. \frac{dm_{\text{H}_3\text{O}^+}}{dt} \right|_{\text{ox}} = \frac{1}{F} \delta_{\text{H}_3\text{O}^+} M_{\text{H}_3\text{O}^+} \left(\frac{dq_{\text{BP} \rightarrow \text{PN}}}{dt} \right) \quad (29)$$

and in the reduction direction:

$$\left. \frac{dm_{\text{A}^-}}{dt} \right|_{\text{red}} = \frac{1}{F} \delta_{\text{A}^-} M_{\text{A}^-} \left(\frac{dq_{\text{BP} \rightarrow \text{P}}}{dt} + \frac{dq_{\text{P} \rightarrow \text{LE}}}{dt} \right)$$

and

$$\left. \frac{dm_{\text{H}_3\text{O}^+}}{dt} \right|_{\text{red}} = \frac{1}{F} \delta_{\text{H}_3\text{O}^+} M_{\text{H}_3\text{O}^+} \left(\frac{dq_{\text{PN} \rightarrow \text{BP}}}{dt} \right) \quad (30)$$

where $\delta_{\text{A}^-}=+1$ and A^- refers to both anions of this work where the molar masses are 62 g mol^{-1} and 99.5 g mol^{-1} for NO_3^- and ClO_4^- , respectively. Our previous results indicate the transfer of protons stoichiometrically accompanied by one water molecule during the $\text{PN} \leftrightarrow \text{BP}$ transitions [5,40]. Therefore, we consider $M_{\text{H}_3\text{O}^+}=19 \text{ g mol}^{-1}$ and $\delta_{\text{H}_3\text{O}^+}=-1$ due to the exit of H_3O^+ during the oxidation and the insertion of H_3O^+ during the reduction.

<Figure 4>

Figure 4 shows the thus-calculated mass fluxes during the redox cycling of POT- ClO_4^- and POT- NO_3^- . dm/dt partial curves for H_2O also show a similar profile like dm/dt partial curves for the corresponding anion as it was expected if we consider the exclusion effect [69–72]. However, a slight shift can be observed (10-20 mV) between the potential peaks of dm/dt partial curves of free water and anions. Structural reconfiguration owing to both the insertion/expulsion of anions and the electrochemical processes of POT are accompanied by changes in free volume of film where the free water molecules can be inserted [73,74]. In our model [40], the oxidation state of POT film controls the transfer of ions by charge balance but the kinetic of free water transfer is regulated by the exclusion effect of anions and the structural rearrangement of the macromolecular backbone of POT film.

Once $dm_{\text{H}_2\text{O}}/dt$ is obtained in both solutions, the partial curves may be similarly simulated using eq (14) considering two independent curves, which correspond to the electrochemical transitions where the water transfers are associated:

$$\left. \frac{dm_{\text{H}_2\text{O}}}{dt} \right|_{ox} = \left. \frac{dm_{\text{H}_2\text{O}}}{dt} \right|_{\text{LE} \rightarrow \text{P}} + \left. \frac{dm_{\text{H}_2\text{O}}}{dt} \right|_{\text{P} \rightarrow \text{BP}}$$

and

$$\left. \frac{dm_{\text{H}_2\text{O}}}{dt} \right|_{red} = \left. \frac{dm_{\text{H}_2\text{O}}}{dt} \right|_{\text{BP} \rightarrow \text{P}} + \left. \frac{dm_{\text{H}_2\text{O}}}{dt} \right|_{\text{P} \rightarrow \text{LE}}$$
(31)

in the oxidation direction for the LE→P→BP transitions and in the reduction direction for the BP→P→LE transitions. The quantities obtained are in Table 3.

<Table 3>

As it was expected, the total amount of water excluded by ClO_4^- is superior to that excluded by NO_3^- (approximately twice). Using the commercial software ChemBio3D Ultra v. 12.0 ChemBioOffice 2010, the volume of species can be estimated from the Connolly Solvent Excluded Volume. Thus, the volume of NO_3^- is about 32 \AA^3 , about 47 \AA^3 for ClO_4^- and about 10 \AA^3 for H_2O . Taking into account this volume, about 3 water molecules can be theoretically excluded by each NO_3^- and about 5 by each ClO_4^- in accordance with the ratio of amount of excluded water by the anions above commented. From surface concentration calculated of anions and water molecules, a global water/anion ratio for each transition can be obtained ($|\Gamma_{max}|_{\text{H}_2\text{O}}/|\Gamma_{max}|_{\text{A}^-}$). In general, two water molecules are transferred in opposite direction in presence of NO_3^- and about four in presence of ClO_4^- . The mechanical expansion/contraction of POT structure during the redox processes causes lower values than the theoretical ones. The reconfiguration of POT films during the anion transfers prevents a pure exclusion effect because the volume of anions is not in accordance with the volume of excluded water molecules. In spite of this, we obtain a more dehydrated film of oxidized POT inserting ClO_4^- that with the insertion of NO_3^- independently of the changes of free volume inside film. Finally, $(b - b')$ for partial $dm_{\text{H}_2\text{O}}/dt$ is higher in the POT- NO_3^- than in the POT- ClO_4^- in both transitions (Table 3). Therefore, the interactions owing to the transfer of anions affect the free water transfer, especially during the oxidation direction when the inserted anions are expelled from film.

<Figure 5>

4.4. Crossed functions analysis

The crossed functions using the simulated partial curves of current, absorbance derivative and mass derivative will allow the electrochemical behavior of POT films to

be understood in depth. Figure 6 shows the evolution of the apparent molar mass of transferred species ($F(dm/dq)$, eq. (17)), the electrochromic efficiency ($F(dA^\lambda/dq)$, eq. (18)) and the mass electrochromic efficiency ((dA^λ/dm) , eq. (19)).

$F(dm/dq)$, was calculated considering the partial contribution of both current and mass of anions and free water by using:

$$\frac{F\left(\frac{dm_{A^-}}{dt} + \frac{dm_{H_2O}}{dt}\right)}{\frac{dq}{dt}} \quad (32)$$

In this manner, an apparent molar mass is calculated but we can evaluate the instantaneous evolution of anion/water ratio during the cyclic voltammograms.

Moreover, we can discriminate A^{435nm} results. The changes of absorbance at 435 nm can also be used to analysis LE \leftrightarrow P transitions but the formation of isolated polarons take place at initial stages of the oxidation transition and at final stages of the reduction transition. This fact is graphically observed in Figure 5 where the partial simulated curves of LE \rightarrow P transition of POT-ClO₄⁻ are showed. On the contrary, the formation of coil structures when the anions interact strongly with the C=NH⁺=C group could take place during all these electrochemical transitions. At lower pH, Hillman et al showed a H₃O⁺ release preceding the anion uptake during oxidation scan by using probe beam deflection technique and EQCM [48]. The variation of A^{810nm} in all the range of potentials confirms the only participation of anions during the LE \leftrightarrow P transitions in both POT forms and dm_{A^-}/dt calculated as in eq (29) and (30). Thus, only the simulated partial $dA_{LE\leftrightarrow P}^{810nm}/dt$ curves was used to study the LE \leftrightarrow P transitions, $dA_{P\leftrightarrow BP}^{700nm}/dt$ to study the P \leftrightarrow BP transitions and $dA_{BP\leftrightarrow PN}^{540nm}/dt$ to study the BP \leftrightarrow PN transitions with the electrochromic transfer functions.

<Figure 6>

Figure 6a shows the crossed functions corresponding to the LE \leftrightarrow P transitions (with a formal potential, $E^{o'}$, about 0.210 V in HNO₃ solution and about 0.170 V in HClO₄ solution). As it is expected for the exclusion effect between anions and free water molecules, $F(dm/dq)$ shows lower values than $M_{NO_3^-}$ and $M_{ClO_4^-}$, except for initial and final states of the formation of coil structures during these transitions because $F(dm/dq)$ tends to theoretical values of anions in both POT-NO₃⁻ and POT-ClO₄⁻, respectively. Moreover, it is observed an asymmetry between the LE \rightarrow PN transition (insertion of anions together with the coil structure formation) and the PN \rightarrow LE transition (expulsion

of anions and relaxation of the structure). In the oxidation direction, the insertion of anions in both POT forms exclude about 3 water molecules close to the peak potential of LE→P transition, on the contrary, the exit of NO₃⁻ allow the insertion of 2 water molecules and apparently 4 water molecules for each ClO₄⁻ considering a monoelectronic transfer. $F(dA^{810nm}/dq)$ curves show a better electrochromic efficiency of POT-ClO₄⁻ than POT-NO₃⁻. This fact could be due to the different anion-backbone interactions. ClO₄⁻ involves a great change of color per charge unit than for NO₃⁻. However, the mass electrochromic efficiency (dA^{810nm}/dm) is higher in the POT-NO₃⁻ than POT-ClO₄⁻ during LE→P transition owing to the smaller size of NO₃⁻. In the reduction direction, POT-ClO₄⁻ involves higher changes of color at 810 nm per mass unit than in POT-NO₃⁻ because the mass of ClO₄⁻ when this is expelled from POT is highly compensated by the insertion of water molecules.

The P↔BP transitions have a $E^{0'}$ about 0.450 V for both POT forms (Figure 6b) owing to the structural reorganization assumed above. In this occasion, ClO₄⁻ exclude one water molecule more than NO₃⁻ in both transitions directions. However, one extra water molecule is excluded during the reduction transition with respect to the oxidation transition in both POT forms. As before, $F(dA^{700nm}/dq)$ of POT-ClO₄⁻ reaches values higher than the electrochromic efficiency obtained of POT-NO₃⁻. However, the dA^{700nm}/dm is higher for the POT-ClO₄⁻ during the P→BP transition and dA^{700nm}/dm is higher for the POT-NO₃⁻ during the BP→P transition due to the perfect mass balance between the expelled NO₃⁻ and inserted water molecules ($F(dm/dq)$ close to 0 g mol⁻¹).

The best electrochromic characteristics observed in POT films takes place at potentials of these transitions due to a higher accumulation of polaronic forms (P or BP) in the polymer lattice. The amount of P electroactive sites in the polymer can be calculated by integration of $(dq_{LE↔P}/dt)$ partial curves from the most cathodic potentials to 0.45 V in the oxidation direction and from the most anodic potentials up to 0.45 V in the reduction directions. Similarly, BP electroactive sites can be calculated from the charge of $(dq_{P↔BP}/dt)$ partial curves. In the oxidation direction, around 2-3 polarons sites per bipolaron site seems to be the conducting form of both POT forms. On the contrary, 38 bipolaron sites per polaron site are the conducting form of POT during the reduction direction for POT-NO₃⁻ (16 BP per P for POT-ClO₄⁻). In this manner, we can offer an accurate evaluation of the percolation threshold potential where electric connection of the conducting species generates a well-connected network.[75]. This knowledge is of special

interest for their possible technological application in sensor devices enhancing the sensitivity when the films operate in the vicinity of the percolation threshold [76]. Moreover, the mechanical reorganization of film allows a fast transfer of anions and free solvent between the film and the solution [40]. That is in accordance with molecular models showing that the fully polaronic form is a serious structural perturbation with respect to the other forms of the polymer [77,78].

Finally for the BP \leftrightarrow PN transitions, the crossed functions of BP \leftrightarrow PN transitions are showed in Figure 6c with a $E^{0'}$ about 0.50 V in both solutions. $F(dm/dq)$ is constantly -19 g mol^{-1} during the BP \leftrightarrow PN transitions owing to the assumed model where the hydrated protons are the only cation necessary to the protonation/deprotonation of POT during these transitions. The electrochromic efficiencies show similar values in both POT forms but lower than the efficiencies in other redox transitions. In general, both anions do not affect the proton transfer and, in consequence, the electrochemistry of BP \leftrightarrow PN transitions. However, the peak potential of PN \rightarrow BP transition of POT-ClO $_4^-$ is close to 0.4 V. The tetrahedral molecular geometry of ClO $_4^-$ inserted in POT during the oxidation reactions could affect the insertion of protons during this redox transition.

By cyclic spectroelectrogravimetry, both forms of POT are intensively compared and specific differences can be observed during the oxidation and reduction reactions of films. The cyclic spectroelectrogravimetry provides a more complete analysis of POT electrochemistry than spectroelectrochemistry techniques or EQCM. Color changer or, in other words, structural changes of polymers can be correlated with mass changes owing to the species transfers during the electrochemical reactions. In this manner, we can neglect temporal structural changes caused by sequential experiments (history effects) and the findings shown here would be much more difficult to achieve.

5. CONCLUSION

The combination of several techniques in situ allows electrochemically complex materials to be described in depth. The complementarity of voltammetry, electrogravimetry and Vis-NIR spectra (cyclic spectroelectrogravimetry) allows the three electrochemical transitions of POT films to be separated. This analysis has been done in two monoprotic acid solutions. The size of anions plays an important role causing new electrochemical properties of the hosting film because they interact with the molecular structure of film. During the P \leftrightarrow BP transitions, the internal reorganization of species

could improve the transfer of species maybe by repulsion forces between electrically charged species (anions, cations and polaronic/bipolaronic forms of POT) enhancing simultaneously the electronic conductance of POT. This fact could explain the results of $(b - b')$ around 20 V^{-1} taking into account the Frumkin-type isotherm. Better electrochromic behaviors are observed during these transitions of both POT forms. However, the electrochromic properties of POT depend on the direction of anions transfer. Better electrochromic efficiencies ($F(dA^\lambda/dq)$) are observed for POT- ClO_4^- during the oxidation $\text{LE} \rightarrow \text{P} \rightarrow \text{BP}$ transitions (anion insertion) and better mass electrochromic efficiencies (dA^λ/dm) are observed for POT- NO_3^- during the reduction $\text{BP} \rightarrow \text{P} \rightarrow \text{LE}$ transitions (anion expulsion). Moreover, the size of inserted ClO_4^- may be the main problem for the protonation of POT during the $\text{PN} \rightarrow \text{BP}$ transition. On the other hand, ClO_4^- is able to exclude more water molecules than NO_3^- , which allows the formation of less hydrated films if they are oxidized in spite of the structure reorganization during the $\text{LE} \leftrightarrow \text{P} \leftrightarrow \text{BP}$ transitions.

The application of the crossed functions analysis allows these changes and processes to be analyzed separately in situ during a cyclic voltammetry experiment minimizing possible errors due to the different history of the films if a sequential analysis is done. This information opens new perspectives mainly for the development of new methodologies allowing a better evaluation and understanding of ion insertion and structural changes in the conducting POT and more widely in other electroactive materials, which can be used for specific tailoring of signal transducer devices. In the future, it can contribute to the development of some devices in the better electrochemical conditions such as smart ion-high selective electrodes.

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TABLES

Table 1. Parameters obtained from the fitting of experimental absorbance derivative curves to eq (16) and eq (26) in Figure 2.

Transitions	HNO ₃ solution											
	dA^{435nm}/dt			dA^{540nm}/dt			dA^{700nm}/dt			dA^{810nm}/dt		
	$(dA/dt)_{peak}$	$b - b'$	E_p	$(dA/dt)_{peak}$	$b - b'$	E_p	$(dA/dt)_{peak}$	$b - b'$	E_p	$(dA/dt)_{peak}$	$b - b'$	E_p
	a.u. s ⁻¹	V ⁻¹	V	a.u. s ⁻¹	V ⁻¹	V	a.u. s ⁻¹	V ⁻¹	V	a.u. s ⁻¹	V ⁻¹	V
LE→P	2.4x10 ⁻⁴	23	0.226	1.6x10 ⁻⁴	20	0.270	2.8x10 ⁻⁴	27	0.238	4.3x10 ⁻⁴	21	0.226
P→BP	-1.4x10 ⁻⁴	19	0.438	3.1x10 ⁻⁴	23	0.468	5.9x10 ⁻⁴	15	0.438	2.3x10 ⁻⁴	15	0.438
BP→PN	-	-	-	3.5x10 ⁻⁴	35	0.565	-3.4x10 ⁻⁴	15	0.600	-2.4x10 ⁻⁴	20	0.580
P→LE	-2.4x10 ⁻⁴	23	0.140	-1.2x10 ⁻⁴	15	0.200	-2.6x10 ⁻⁴	15	0.2057	-3.2x10 ⁻⁴	15	0.178
BP→P	1.6x10 ⁻⁴	17	0.440	-3.2x10 ⁻⁴	28	0.450	-5.0x10 ⁻⁴	20	0.440	-2.5x10 ⁻⁴	18	0.460
PN→BP	-	-	-	-7.4x10 ⁻⁴	70	0.532	10.0x10 ⁻⁴	61	0.550	7.5x10 ⁻⁴	40	0.550
	HClO ₄ solution											
LE→P	2.9x10 ⁻⁴	23	0.200	1.5x10 ⁻⁴	17	0.28	3.0x10 ⁻⁴	14	0.260	4.1x10 ⁻⁴	14	0.237
P→BP	-2.4x10 ⁻⁴	23	0.44	2.6x10 ⁻⁴	24	0.475	9.0x10 ⁻⁴	19	0.500	4.1x10 ⁻⁴	24	0.470
BP→PN	-	-	-	3.8x10 ⁻⁴	34	0.582	-7.0x10 ⁻⁴	18	0.610	-4.5x10 ⁻⁴	25	0.630
P→LE	-2.6x10 ⁻⁴	16	0.012	-1.5x10 ⁻⁴	12	0.192	-3.2x10 ⁻⁴	12	0.150	-4.1x10 ⁻⁴	12	0.128
BP→P	1.4x10 ⁻⁴	16	0.300	-2.4x10 ⁻⁴	29	0.371	-6.4x10 ⁻⁴	16	0.433	-4.0x10 ⁻⁴	21	0.460
PN→BP	-	-	-	-5.0x10 ⁻⁴	52	0.446	10.3x10 ⁻⁴	29	0.504	8.3x10 ⁻⁴	27	0.500

Table 2. Parameters obtained from the fitting of experimental voltammetric curves to eq. (13) in Figure 3.

Transitions	HNO ₃ solution			HClO ₄ solution		
	$ \Gamma_{max} $ $\mu\text{mol cm}^{-2}$	$b - b'$ V^{-1}	E_p V	$ \Gamma_{max} $ $\mu\text{mol cm}^{-2}$	$b - b'$ V^{-1}	E_p V
LE→P	2.3	25	0.259	2.9	18	0.245
P→BP	2.3	15	0.460	2.8	16	0.480
BP→PN	4.3	17	0.570	3.3	25	0.605
P→LE	3.7	15	0.172	4.9	12	0.097
BP→P	4.4	18	0.457	3.7	13	0.378
PN→BP	3.2	52	0.521	2.7	33	0.471

Table 3. Parameters used in the simulated curves of dm/dt for H_2O using eq (31) from the data of Figure 4.

Transitions	HNO ₃ solution			HClO ₄ solution		
	$ \Gamma_{max} $ $\mu\text{mol cm}^{-2}$	$b - b'$ V^{-1}	E_p V	$ \Gamma_{max} $ $\mu\text{mol cm}^{-2}$	$b - b'$ V^{-1}	E_p V
LE→P	4.8	31	0.245	5.1	28	0.224
P→BP	1.6	33	0.485	10.8	17	0.495
P→LE	5.6	20	0.155	18.6	13	0.090
BP→P	13.2	21	0.457	13.9	14	0.370

FIGURES LEGENDS

Figure 1. 3D surface of the dA^λ/dt -potential-wavelength for the gold electrode modified by POT in .5 M HNO₃ (a) and 0.5 M HClO₄ (b) and corresponding cyclic voltammogram performed with a 100 mV s⁻¹ potential scan rate. Eleven Vis-NIR spectra were collected each second. The mountains (solid lines) and valleys (dotted lines) of both profiles represent increasing ($dA^\lambda/dt > 0$) and decreasing ($dA^\lambda/dt < 0$) regions of absorbance, respectively.

Figure 2. Experimental and simulated curves of dA^λ/dt at 435 nm, at 540 nm, at 700 nm and at 810 nm using eq (16) with the parameters given in Table 1 for the change of absorbance of the gold electrode modified by POT in 0.5 M HNO₃ (a) and 0.5 M HClO₄ (b) with a 100 mV s⁻¹ potential scan rate.

Figure 3. Cyclic voltammogram of POT performed at 100 mV s⁻¹ scan rate of POT gold modified electrode in 0.5 M HNO₃ (a) and 0.5 M HClO₄ (b). The parameters of simulated curves are in Table 2 using eq (13)

Figure 4. Experimental and simulated curves of dm/dt using eq (14) with the parameters given in Table 3 (inset in (a)), for the cyclic voltammetry of gold electrode modified by POT in 0.5 M HNO₃ (a) and 0.5 M HClO₄ (b) with a 100 mV s⁻¹ potential scan rate.

Figure 5. dq/dt , dA^{435nm}/dt and dA^{800nm}/dt partial curves corresponding to the LE→P transition of POT in 0.5 M HClO₄ solution using the parameters in Table 1 and Table 2 for simulation by eq (13) and (16).

Figure 6. Crossed functions on potential scale for POT in 0.5 M HNO₃ (circles) and 0.5 M HClO₄ (squares) in the oxidation (open symbols) and reduction (closed symbols) direction of potential scan for the LE↔P transitions at 810 nm (a), the P↔BP transitions at 700 nm (b) and BP↔PN transitions at 540 nm (c).

FIGURES

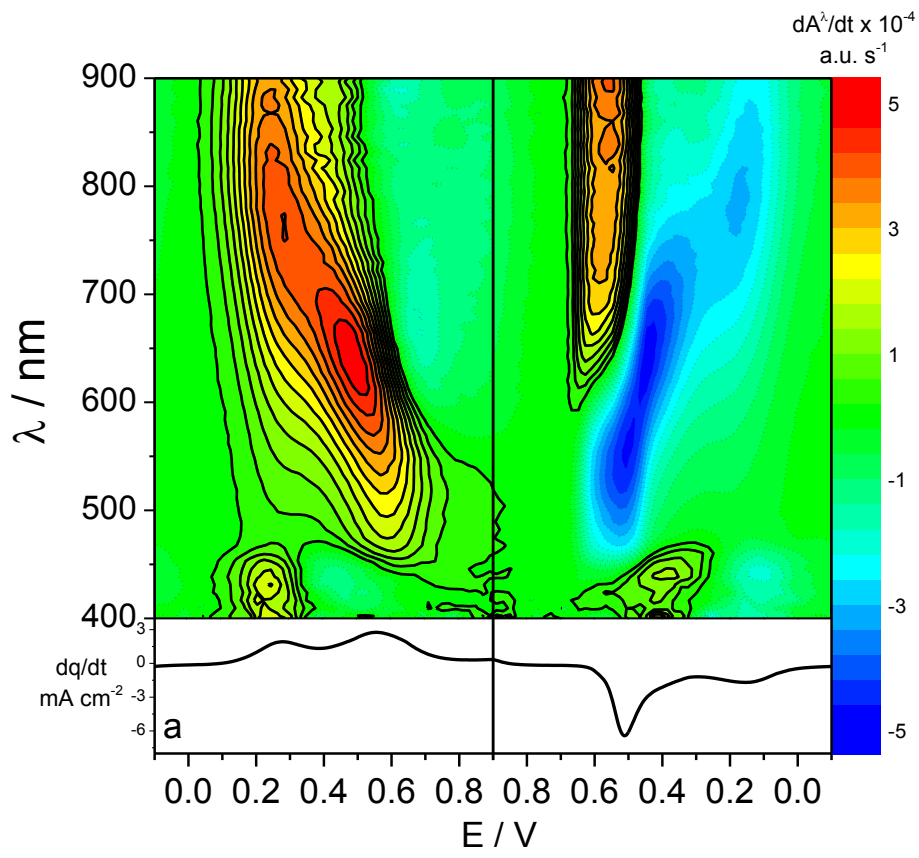


Figure 1a

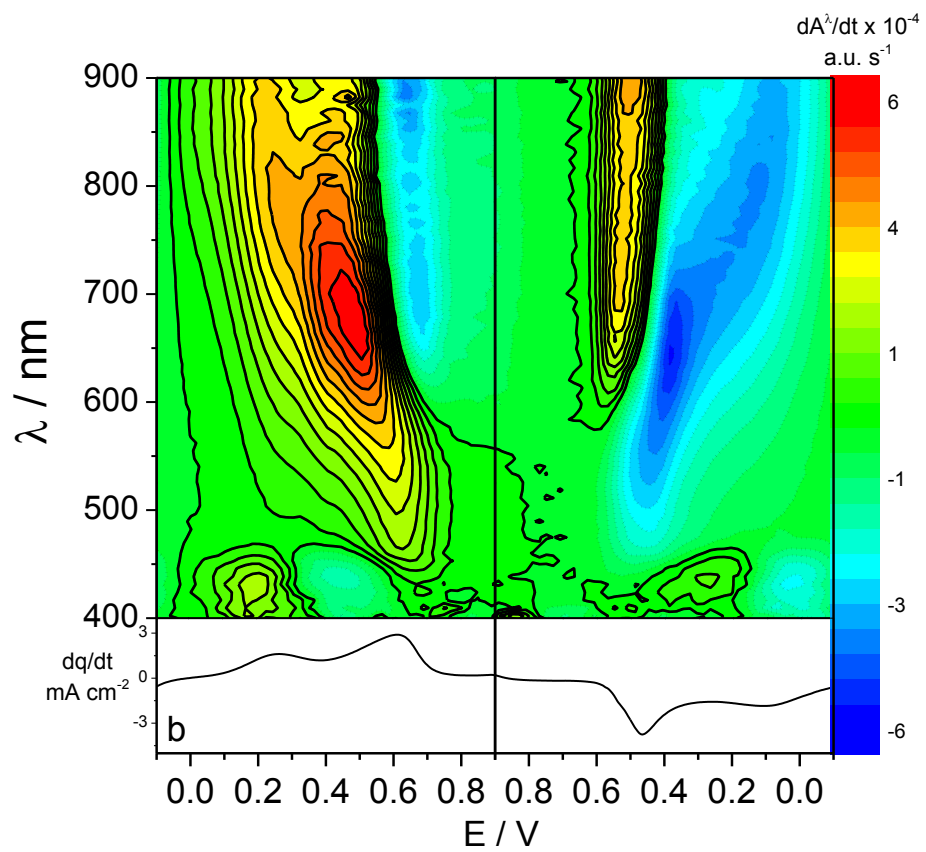


Figure 1b

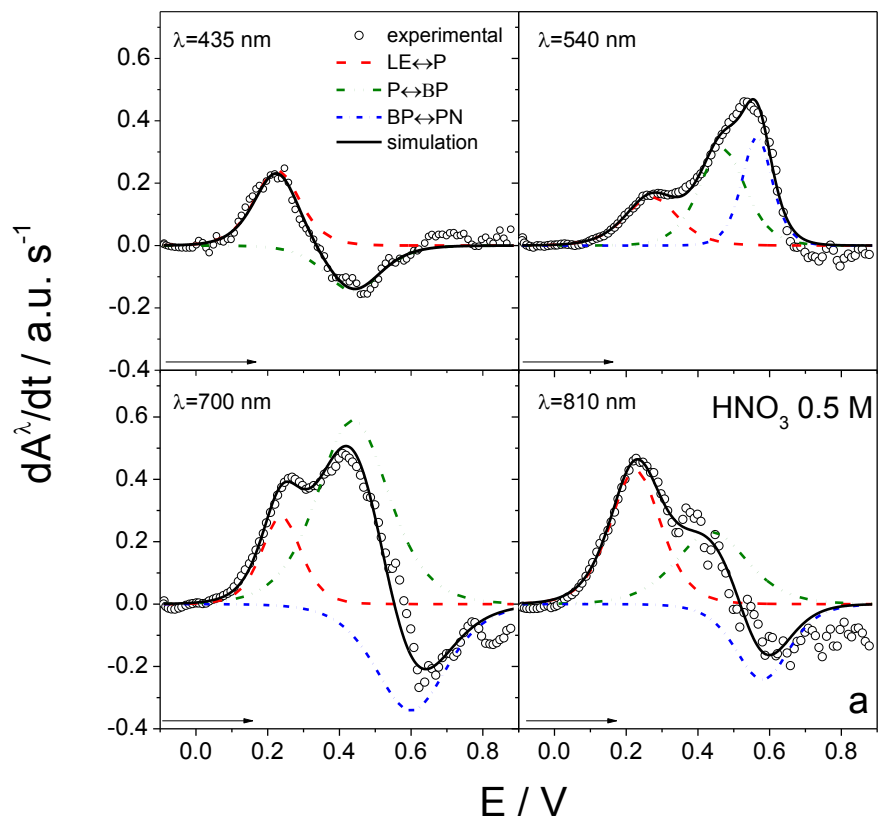


Figure 2a

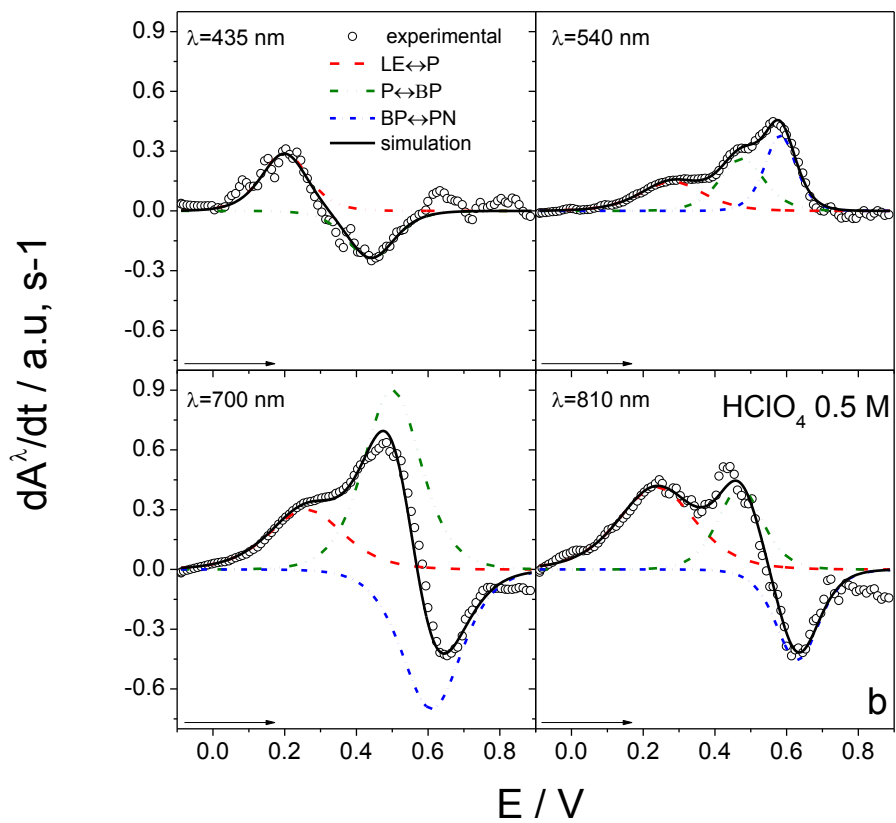


Figure 2b

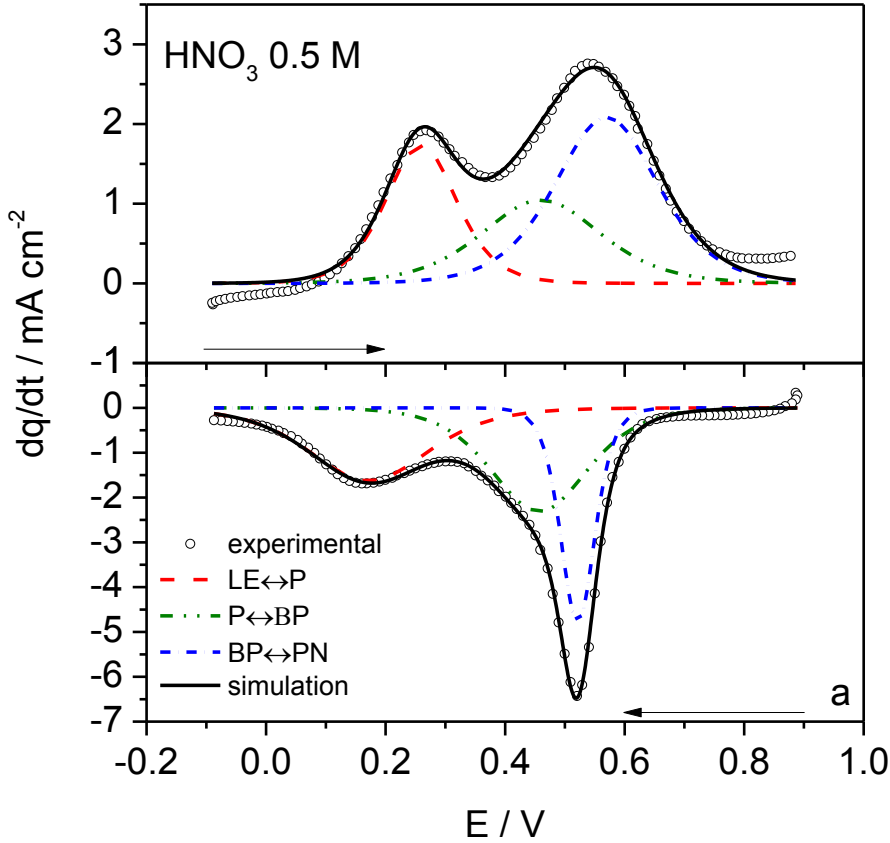


Figure 3a

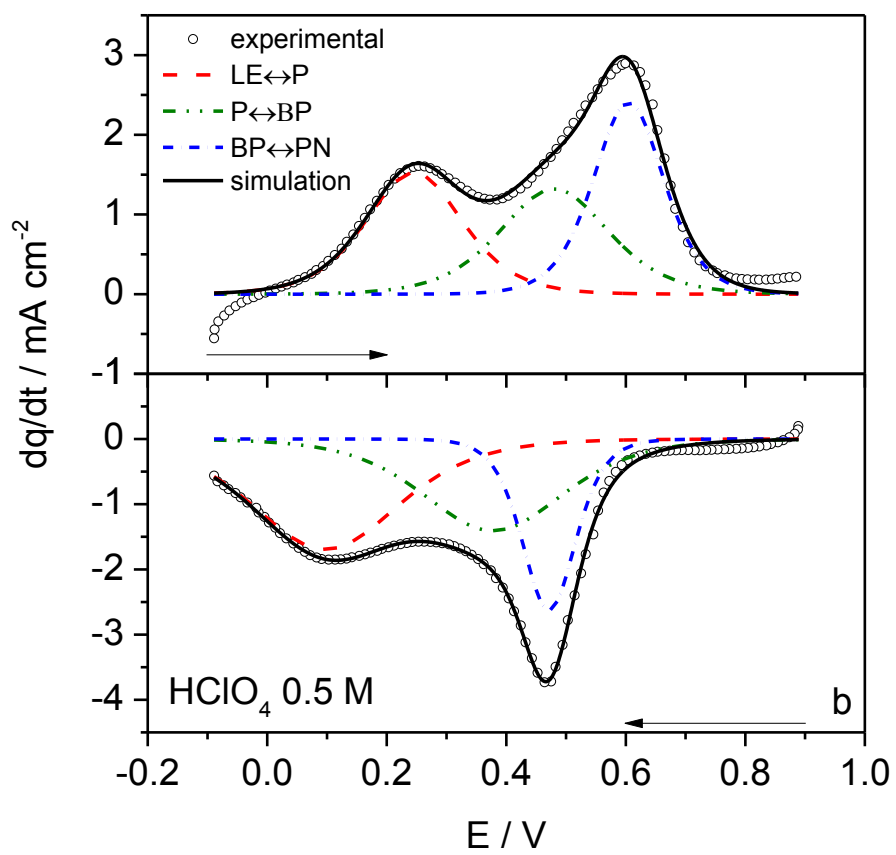


Figure 3b

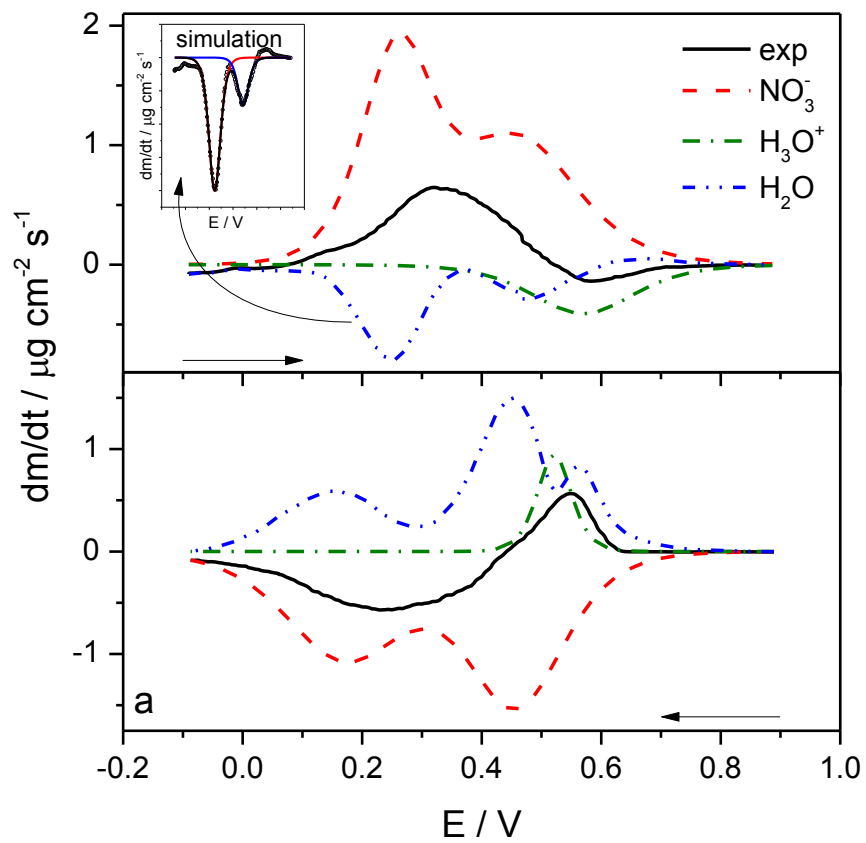


Figure 4a

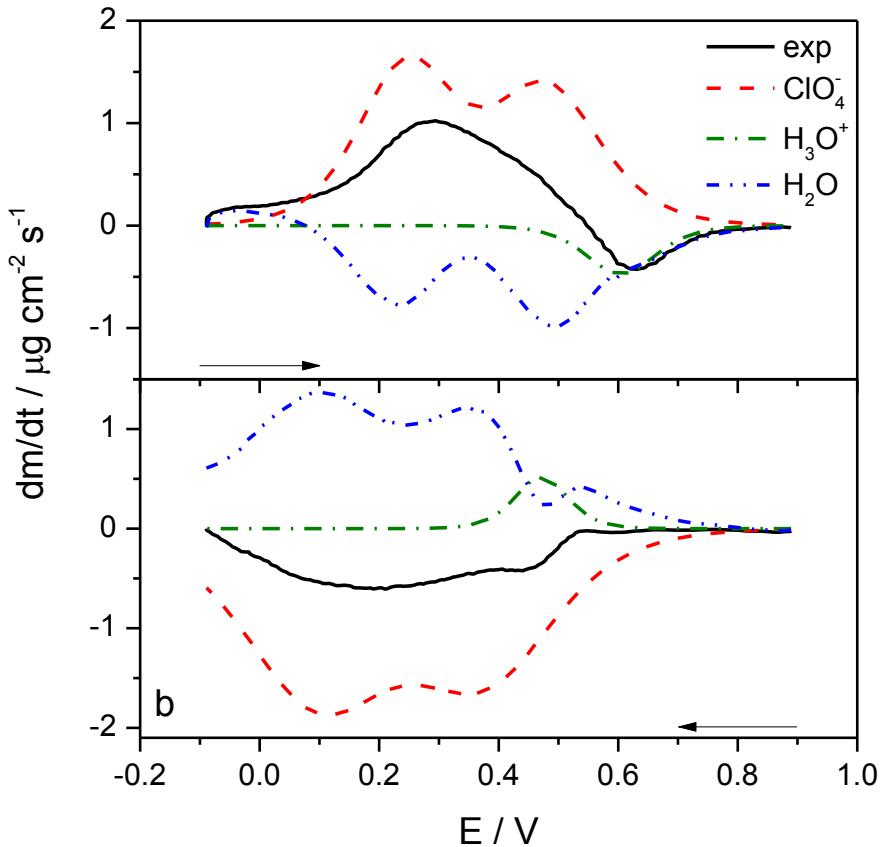


Figure 4b

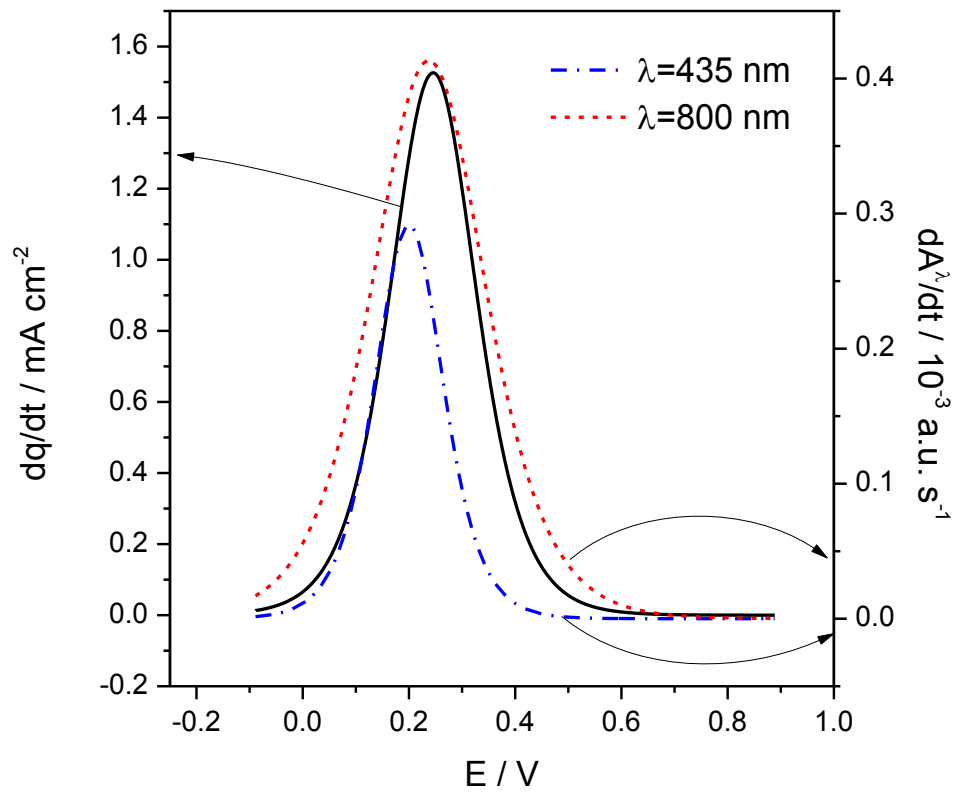


Figure 5

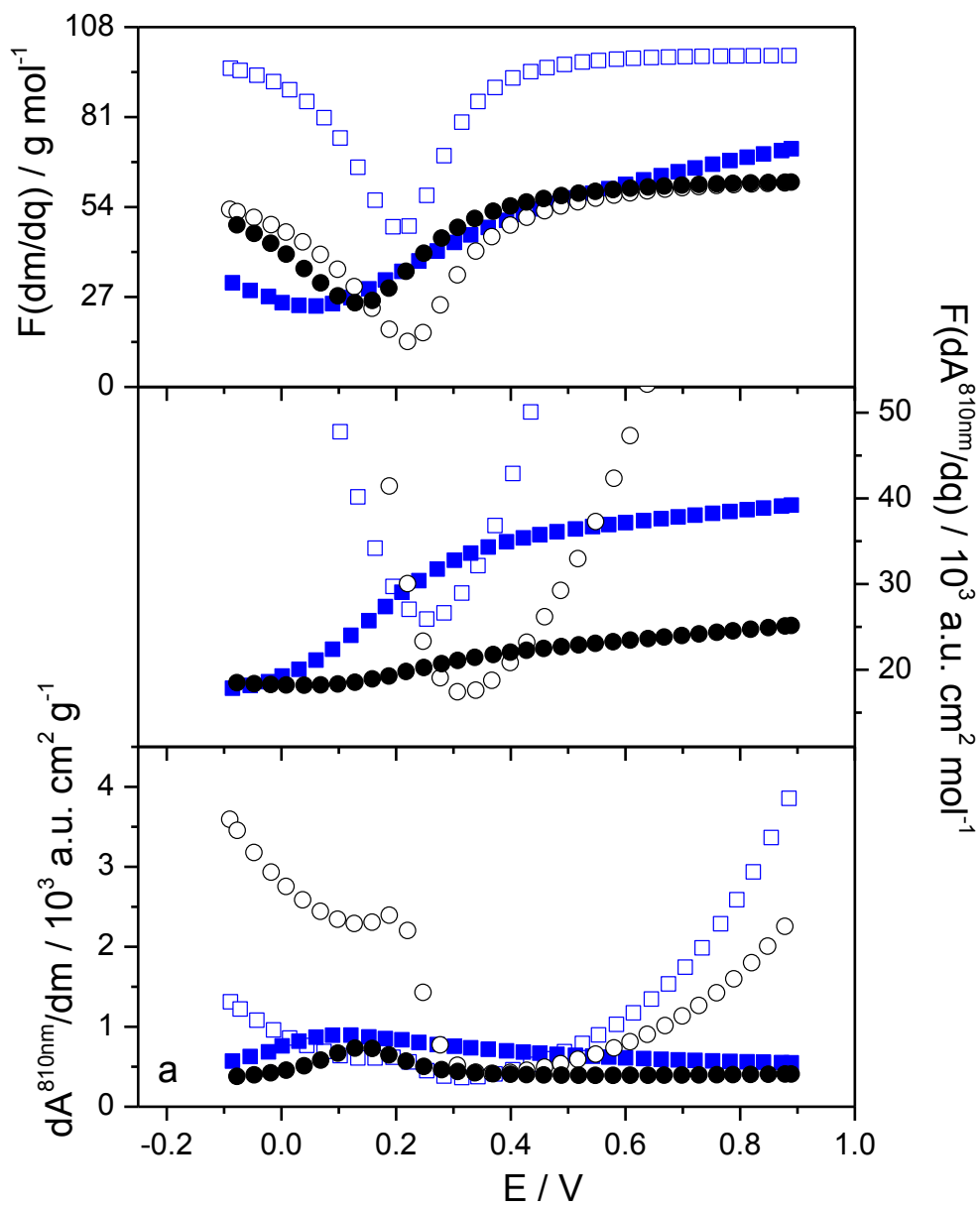


Figure 6a

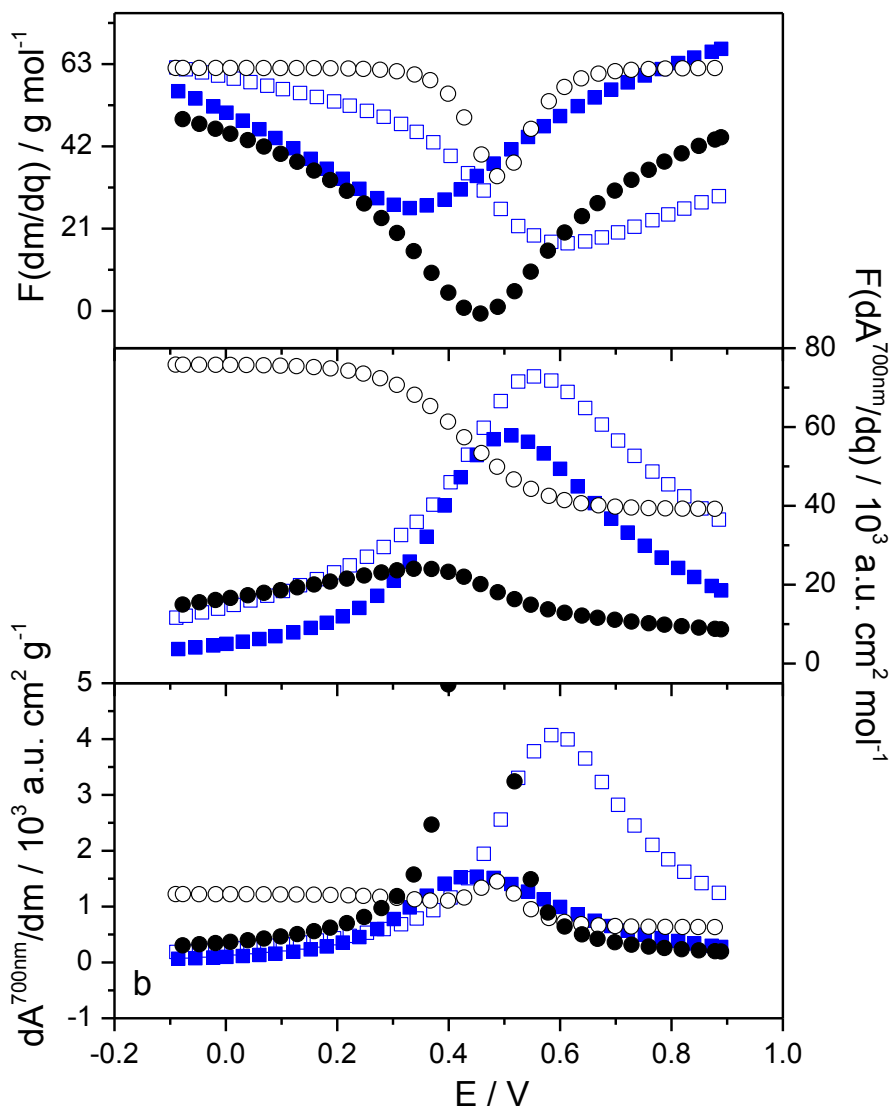


Figure 6b

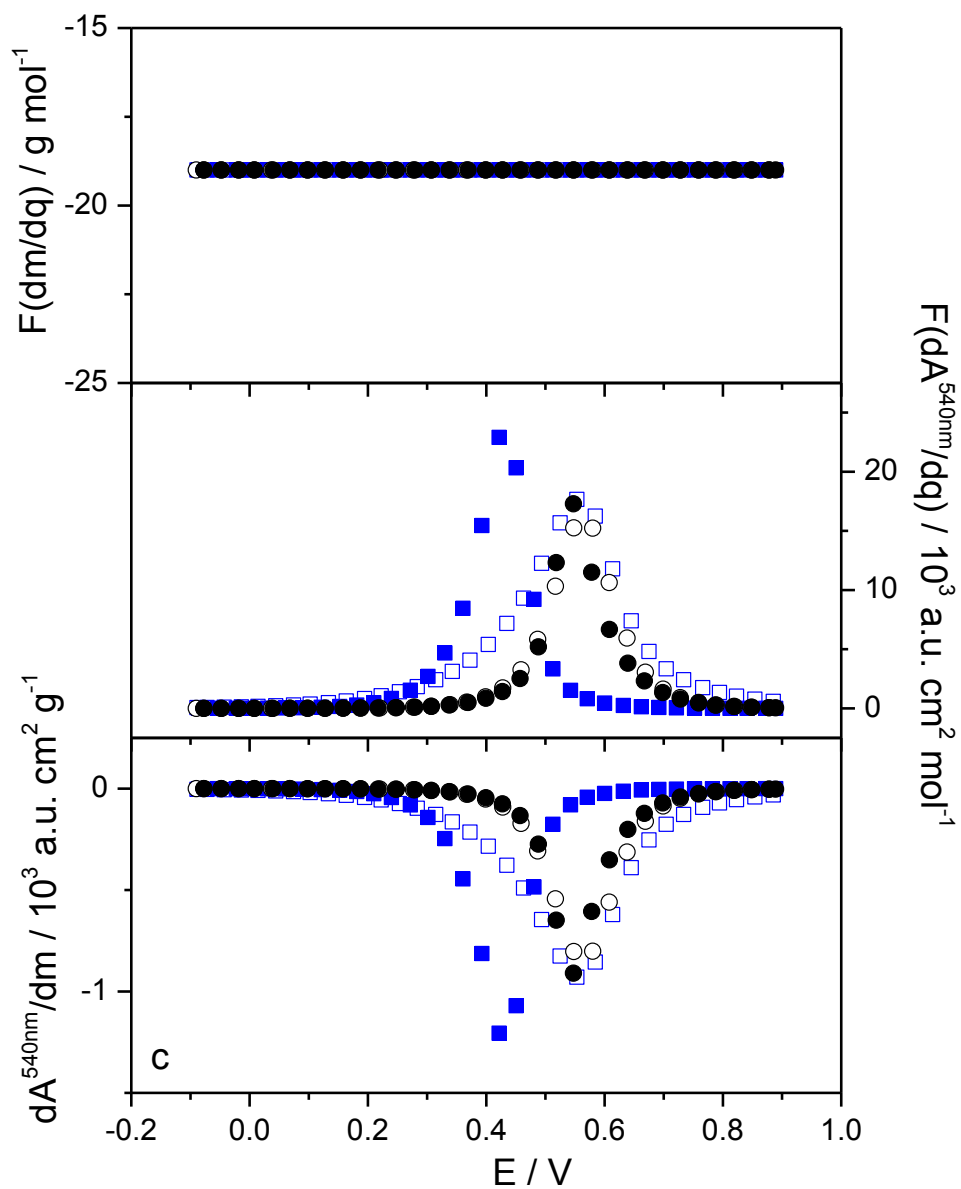


Figure 6c