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# Experimental evidence of NO SCR mechanism in the presence of the BEA zeolite with framework and extra-framework cobalt species

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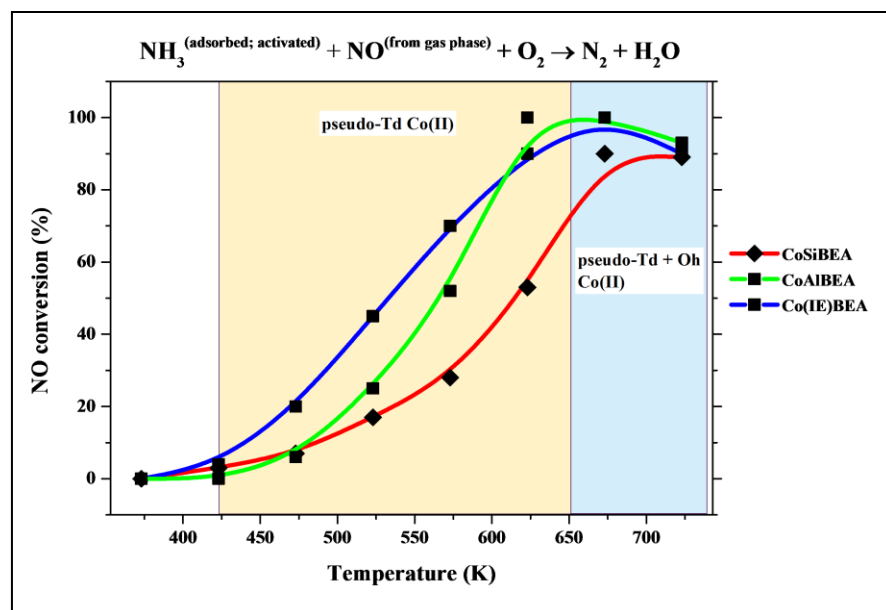
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## Graphical abstract



## Highlights

- Co introduced in Beta supports effectively activates  $\text{NH}_3$  for reaction with NO.
- The highest reactivity of ammonia in SCR of NO found for Co(IE)BEA catalyst.
- During SCR of NO ammonia and NO compete for the same adsorption sites.
- SCR of NO process involves the reaction of chemisorbed  $\text{NH}_3$  with NO from gas phase.

## Abstract

Temperature-programmed desorption (TPD), temperature-programmed surface reaction (TPSR) and stop flow-TPD techniques as well as FT-IR spectroscopy of pre-adsorbed  $\text{NH}_3$  and NO are applied to characterize active sites present in Co-containing BEA zeolites. Cobalt is introduced into BEA zeolite by different methods: (i) two-step postsynthesis – PS, (ii) conventional wet impregnation - Imp and (iii) ion exchange - IE. The obtained catalysts are

labelled as Co(PS)SiBEA (postsynthesis), Co(IMP)AlBEA (wet impregnation) and Co(IE)AlBEA (ion exchange). All the obtained catalysts are found to be active in the process of selective catalytic reduction (SCR) of NO with ammonia. The catalytic activity depends on the method used for cobalt introduction and thus on cobalt centres speciation. The reactivity of NH<sub>3</sub> and NO adsorbed on Co(IMP)AlBEA and Co(IE)AlBEA is higher than that measured for the Co(PS)SiBEA catalysts. The temperature-programmed studies and FT-IR with NO and/or NH<sub>3</sub> adsorption for the Co-containing samples show that during SCR reaction NO and NH<sub>3</sub> molecules compete for the same adsorption sites. The main reaction pathways of the DeNO<sub>x</sub> process performed in the presence of Co(IMP)AlBEA and Co(IE)AlBEA involves the reaction between chemisorbed ammonia and NO from gas phase, according to the Eley-Riedel mechanism.

**Keywords:** cobalt zeolite, NO, SCR, TPSR, FT-IR

## 1. Introduction

The emission of nitrogen oxides (NO<sub>x</sub>) from stationary and mobile sources causes serious environmental problems. Ammonia based selective catalytic reduction (SCR) of NO<sub>x</sub> is a widely practiced technology for diesel engine vehicles to reduce NO<sub>x</sub> (NO, NO<sub>2</sub>) to N<sub>2</sub> [1]. In this technique, NO<sub>x</sub> is continuously reduced by NH<sub>3</sub> on the V<sub>2</sub>O<sub>5</sub>/WO<sub>3</sub>/TiO<sub>2</sub> catalysts [2,3]. However, V<sub>2</sub>O<sub>5</sub>-based catalysts have inadequate stability at the higher temperatures encountered in high load vehicle operation and are not suitable for systems including a particulate filter that actively regenerates at higher temperatures. Hence, various catalysts, including noble metals based systems [4], ion-exchanged zeolites [5-9], and metal oxides [10,11] have been extensively studied. However, reaction mechanism, which determines the overall activity of the catalysts, is still a subject of a debate. The standard SCR reaction between NO<sub>x</sub> and NH<sub>3</sub> occurs in the presence of oxygen. Several papers analysing the mechanism and the kinetics of the SCR of NO reaction over various zeolite catalysts have been already published [12-20]. It was shown that Co(II) ion-exchanged zeolites, such as Co-ZSM-5 [21-23], Co-ferrierite [24-26], Co-BEA [27,28] and Co-mordenite [29,30], exhibit high activity in the SCR of NO<sub>x</sub>. The need for more active and selective catalysts initiated a number of studies focused on understanding the reaction mechanism [31-33]. In order to describe the possible reaction pathways, the reaction surface intermediates

and the nature of the active sites must be identified. The controversial results published by different authors indicate the complexity of the problem [31,32,34]. It is well known that the DeNO<sub>x</sub> efficiency is greatly related to loading, dispersion and state of cobalt species that are usually influenced by the preparation method, surface acidity related to the Si/Al ratio and type of the zeolite porous structure [35-38]. Tetrahedrally and octahedrally coordinated Co(II) ions are considered as main active sites of the SCR process, while Co<sub>3</sub>O<sub>4</sub> is reported to be active in hydrocarbons combustion and in NO to NO<sub>2</sub> oxidation [38].

In general, it is postulated that the reduction of NO takes place via the interaction between surface NO<sub>x</sub> species with the reducing agent (e.g. NH<sub>3</sub> or CH<sub>4</sub>) to form surface intermediates, which then decompose to the reaction products [31,32,34,39,40]. However, the structure of these species is ambiguous. Nitrite-nitrate complexes were postulated as surface NO<sub>x</sub> species by several authors [41,42]. In contrast, Resini et al. [32] proposed that the reaction proceeds through a kind of surface nitrate via a redox mechanism over Co(II)/Co(III) active sites. Lobree et al. [34] substantiated surface nitrile (CN) species as important reaction intermediates. According to Chen [31], the prevailing reaction path involves the reaction of either ammonia or amine intermediate with both NO and NO<sub>2</sub>. In catalysts such as Co-ZSM-5, the transition metal species were suggested to play the crucial role in the initiating the reduction of NO [31]. Sun et al. [43] presented a mechanistic model for the metal-exchange zeolite catalysts, in which the preferred path for NO<sub>x</sub> reduction with ammonia occurs via ammonium nitrite, which decomposes to N<sub>2</sub> and H<sub>2</sub>O. Other studies have suggested that the reduction of NO and NO<sub>2</sub> by NH<sub>3</sub> involves the formation of HNO<sub>2</sub> and HNO<sub>3</sub> [44-47]. It is reported that both NO<sub>x</sub> and NH<sub>3</sub> could adsorb on the surface of ion-exchanged zeolites [46,47].

In order to obtain more clear indication concerning the nature of the active species and reaction intermediates involved in the reaction, as well as on the nature of the actual catalyst active sites, spectroscopic and surface reactivity techniques can be applied. It was shown that NO does not adsorb strongly on the catalyst surface, whereas ammonia is strongly adsorbed [48,49]. The SCR NO reaction involves a strongly adsorbed ammonia species and gaseous or weakly bonded NO molecules. Moreover, it was postulated that the rate of the SCR NO reaction depends on the ammonia surface concentration for low NH<sub>3</sub> coverage, whereas a much weaker dependence, if any, exists at high ammonia coverage [48,49]. In spite of many data resulting from

experimental studies focused on mechanistic issues and reaction pathways, the SCR mechanism on the Co-BEA zeolite catalysts is still the subject of the debate.

In this study, the Co-containing BEA zeolite catalysts were prepared by different methods: two-step postsynthesis, wet impregnation and conventional ion exchange. The obtained samples exhibited high activity and selectivity in SCR NO with NH<sub>3</sub>. Thus, the studies of the reaction mechanism for this series of the catalysts were undertaken to determine the role of cobalt species type on catalytic performance of the catalysts.

## 2. Experimental

### 2.1. Catalysts preparation

Three series of Co-containing zeolites were prepared by two-step postsynthesis (PS), conventional wet impregnation (IMP) and ion exchange (IE) procedures.

Co(PS)SiBEA zeolite (with Co loading of 1.0 wt %) was prepared by two-step postsynthesis procedure reported earlier [6-9]. In the first step, 2 g of HAIBEA zeolite, obtained by calcination of tetraethyl-ammonium BEA (TEABEA) zeolite at 823 K for 15 h provided by RIPP (China) was treated with 13 mol L<sup>-1</sup> HNO<sub>3</sub> solution under stirring (353 K, 4 h) to remove aluminium from the zeolite structure. In the second step, 2 g of resulting SiBEA obtained after filtration were dispersed in aqueous solutions (pH = 2.8) containing 1.7 · 10<sup>-3</sup> mol L<sup>-1</sup> of Co(NO<sub>3</sub>)<sub>2</sub> · 6H<sub>2</sub>O and stirred at room temperature for 24 h. Then, the obtained suspensions were stirred in evaporator under vacuum of a water pump in air at 353 K for 2 h until water was evaporated. The solid with the cobalt content of 1.0 wt % was labelled as Co(PS)SiBEA.

Co(IMP)AIBEA zeolite (with Co loading of 1 wt %) was prepared by conventional wet impregnation method. In this case, 2 g of HAIBEA was dispersed in aqueous solution (pH = 3.3) containing 1.7 · 10<sup>-3</sup> mol L<sup>-1</sup> of Co(NO<sub>3</sub>)<sub>2</sub> · 6H<sub>2</sub>O and stirred at room temperature for 24 h. Then, the suspension was stirred in evaporator under vacuum of a water pump at 353 K for 2 h in air until water was evaporated. The solid with cobalt content of 1.0 wt % was labelled as Co(IMP)AIBEA.

Co(IE)AIBEA sample (with Co loading of 1.4 wt %) was prepared by ion exchange method. Firstly, tetraethylammonium form of BEA zeolite was calcined (823 K, 15 h) to obtain HAIBEA. Organic-free BEA was treated two times with a 0.1 mol L<sup>-1</sup> NH<sub>4</sub>NO<sub>3</sub> solution at 343 K for 3 h in order to exchange K<sup>+</sup> and Na<sup>+</sup> ions present in industrial BEA zeolite for NH<sub>4</sub><sup>+</sup>. Then,

the solid was washed with distilled water, dried overnight at 363 K and labelled as  $\text{NH}_4\text{AlBEA}$ . The latter was dispersed in a  $1.0 \cdot 10^{-2} \text{ mol L}^{-1}$  aqueous solution of  $\text{Co}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$  at 343 K (pH = 5.2). Then, after filtration, washing with distilled water the sample was dried. The obtained sample was labelled as  $\text{Co}(\text{IE})\text{AlBEA}$ .

Before physicochemical characterization and catalytic tests,  $\text{Co}(\text{PS})\text{SiBEA}$ ,  $\text{Co}(\text{IMP})\text{AlBEA}$  and  $\text{Co}(\text{IE})\text{AlBEA}$  were calcined at 773 K for 2 h.

## 2.2. Catalyst testing

$\text{Co}(\text{PS})\text{SiBEA}$ ,  $\text{Co}(\text{IMP})\text{AlBEA}$  and  $\text{Co}(\text{IE})\text{AlBEA}$  were studied as catalysts for SCR of NO with ammonia. Catalytic experiments were performed in a fixed-bed flow microreactor system. The reactant concentrations were continuously measured using a quadrupole mass spectrometer RGA 200 Prevac, with Faraday cup (FC) detector, connected directly to the reactor outlet. Sensitivity of the detector is equal  $2 \times 10^{-4} \text{ Torr}^{-1}$  (measured with  $\text{N}_2$  28 amu with 1 amu full peak width 10 % height, 70 eV electron energy, 12 eV ion energy and 1 mA electron emission current). The minimum detectable partial pressure is about  $5 \times 10^{-11} \text{ Torr}$ . Prior to the reaction, each sample (100 mg) of the catalyst was outgassed in a flow of pure helium at 823 K for 30 min. The following composition of the gas mixture was used:  $[\text{NO}] = [\text{NH}_3] = 0.25 \text{ vol. \%}$ ,  $[\text{O}_2] = 2.5 \text{ vol. \%}$  and  $[\text{He}] = 97 \text{ vol. \%}$ . The reaction was studied in the temperature range between 373 K and 723 K. Total flow rate of the reaction mixture was  $40 \text{ mL min}^{-1}$ , with a weight hourly space velocity (WHSV) of about  $24.000 \text{ mL h}^{-1} \text{ g}^{-1}$ .

## 2.3. Catalysts characterization

The obtained zeolite catalysts were characterized with respect to: (i) structure (XRD, FT-IR); (ii) texture (low-temperature  $\text{N}_2$  adsorption) and (iii) coordination and aggregation of deposited cobalt species (DR UV-vis, XPS, TPR). Detailed description of the experimental systems used as well as experimental procedures and analysis of the obtained results can be found in our previous paper [9]. While, in this paper a summary of the catalyst characteristics and the selected examples of the characterization studies are presented.

## 2.4. Studies of the reaction mechanism

The mechanism of the SCR of NO with ammonia over BEA zeolite modified with cobalt species was studied by various temperature-programmed techniques and FT-IR analysis of the samples pre-adsorbed with the reaction substrates (NO, NH<sub>3</sub>).

#### ***2.4.1 Temperature-programmed desorption***

The temperature-programmed desorption of NH<sub>3</sub> (NH<sub>3</sub>-TPD) or NO (NO-TPD) were carried out in the temperature range of 343 - 773 K in a fixed bed continuous flow microreactor system. The reaction temperature was measured by a K-type thermocouple located in a quartz capillary immersed in the catalyst bed. The molecules desorbing from the samples were monitored on-line by a quadrupole mass spectrometer (RGA 200 Prevac) connected to the reactor outlet via a heated line. Before TPD experiments, the zeolite sample (50 mg) was outgassed at 673 K for 1 h in a flow of pure helium (20 mL min<sup>-1</sup>). Subsequently, the sample was cooled down to 343 K and saturated for about 30 min in a flow of 1 vol. % NH<sub>3</sub> diluted in He or 1 vol. % NO diluted in He (20 mL min<sup>-1</sup>). Then, the catalyst was purged in a pure helium flow until a constant baseline level was attained. Desorption was carried out with the linear heating rate (10 K min<sup>-1</sup>) in a flow of pure helium (20 mL min<sup>-1</sup>). Calibration of the quadrupole mass spectrometer with commercial mixtures allowed to recalculate the detector signal into a desorption rate of NH<sub>3</sub> and NO. Apart from standard TPD measurements, desorption experiments for the samples pre-adsorbed with both NH<sub>3</sub> and NO were also done. The outgassed catalyst (673 K, 1 h, helium flow - 20 mL min<sup>-1</sup>) was saturated with 1 vol. % NO diluted in He (20 mL min<sup>-1</sup>) for about 30 minutes at 343 K. Then, the sample was purged in a helium flow until a constant baseline level was attained. Subsequently, the catalyst was saturated with 1 vol. % NH<sub>3</sub> diluted in He (20 mL min<sup>-1</sup>) for about 30 minutes at the same temperature. After purging the sample in a flow of pure helium, the TPD measurements were performed. Desorption was carried out with the linear heating rate of 10 K min<sup>-1</sup> in a flow of pure helium (20 mL min<sup>-1</sup>). These experiments are referred as NO-NH<sub>3</sub>-TPD. In the case of the reverse order of adsorbed gases (first-NH<sub>3</sub>, second-NO), the experiments are referred as NH<sub>3</sub>-NO-TPD.

#### ***2.4.2. Temperature-programmed surface reaction***

In the first step of temperature-programmed surface reaction experiment (NH<sub>3</sub>-TPSR), outgassed zeolite sample (673 K, 1 h, helium flow - 20 mL min<sup>-1</sup>) was exposed to ammonia



according to the procedure applied in the  $\text{NH}_3$ -TPD measurements. In the next step the reactor was purged in a flow of pure helium ( $20 \text{ mL min}^{-1}$ ), and then the catalyst was heated up to about 773 K with the linear increase of temperature ( $10 \text{ K min}^{-1}$ ) in a flow of 1 vol. % NO diluted in He ( $20 \text{ mL min}^{-1}$ ). The emitted gases were analysed by a quadrupole mass spectrometer connected on-line with the reactor outlet.

The reactivity of NO molecules adsorbed on the zeolite samples was studied by temperature-programmed surface reaction measurements (NO-TPSR). Prior to the NO-TPSR experiments, the zeolite sample was outgassed (673 K, 1 h, helium flow -  $20 \text{ mL min}^{-1}$ ) and then cooled down to 343 K. In the next step the catalyst was saturated at 343 K for 30 min with a gas mixture containing 1 vol. % of NO diluted in helium ( $20 \text{ mL min}^{-1}$ ), and subsequently purged with pure helium until a constant base level was attained. Finally, the sample was heated up to about 773 K with a linear increase of temperature ( $10 \text{ K min}^{-1}$ ) in a flow of 1 vol. %  $\text{NH}_3$  diluted in He ( $20 \text{ mL min}^{-1}$ ). A quadrupole mass spectrometer was used for analysis of gases.

#### **2.4.3 Stop-flow TPD**

Prior to the stop-flow TPD experiments, the catalysts (100 mg) were activated in a flow of pure helium at 823 K for 30 min. Then the reactor temperature decreased to 623 K and flow of helium was changed for the reaction mixture consisting of 0.25 vol. % NO, 0.25 vol. %  $\text{NH}_3$  and 2.5 vol. %  $\text{O}_2$  diluted in helium, which was used as a balancing gas at a total flow rate of  $40 \text{ mL min}^{-1}$ . The catalytic test was carried out at 623 K for 1 h. Subsequently, a flow of reacting mixture was exchanged for pure helium. The catalyst was purged in a helium flow until a constant baseline level was attained. Afterwards, in order to identify species adsorbed on the catalyst surface under reaction conditions, the TPD run was carried out with a linear heating rate of  $10 \text{ K min}^{-1}$  in a flow of pure helium ( $20 \text{ mL min}^{-1}$ ) to about 823 K. A quadrupole mass spectrometer was used for analysis of gases.

#### **2.4.4 FT-IR studies of the samples pre-adsorbed with reaction substrates**

The interaction of  $\text{NH}_3$  or NO with the surface of the catalysts as well as their reactivity was studied by FT-IR spectroscopy. The zeolite samples in the form of self-supported wafers were placed in the home-made quartz cell closed with  $\text{CaF}_2$  windows, transparent for the infra-

red light. Cell was connected to the vacuum line and at the same time allowed for the adsorption of the gases in static condition. Wafers could be heated both during activation and adsorption. For the present measurement wafers were activated under vacuum at 723 K and FTIR spectra were recorded at 373 K with Bruker Tensor 27 spectrometer, equipped with an MCT detector. The course of typical experiment was as follow:  $\text{NH}_3$  or NO (purified before adsorption by freeze-and-thaw) was adsorbed on the activated sample at 373 K and IR spectra were taken after consecutive desorption for 20 minutes at 373, 473 and 573 K (all spectra recorded at 373 K). Additional FTIR experiments in which NO was adsorbed on the samples with pre-adsorbed  $\text{NH}_3$  and  $\text{NH}_3$  was adsorbed on the samples with pre-adsorbed NO were also done. The samples with both  $\text{NH}_3$  and NO adsorbed were also evacuated at 373, 473 and 573 K and kept at each temperature for 20 minutes under vacuum before recording the spectrum at 373 K.

### 3. Results and discussion

The most important physicochemical properties of the samples are presented in Table 1. As it was reported in a previous papers [6-9], all materials are characterized by high BET surface area of 642 - 790  $\text{m}^2 \text{g}^{-1}$  and micropore volume of 0.21 - 0.24  $\text{cm}^3 \text{g}^{-1}$ , which are typical of the BEA structure and indicate that textural properties of BEA zeolite are preserved upon dealumination and introduction of cobalt into the zeolite structure by different methods. XRD patterns of all the samples are typical of BEA zeolite. For the Co-containing catalysts, dealumination of zeolite BEA by nitric acid does not affect its crystallinity (Fig. 1). Cobalt introduction does not lead to the appearance of XRD peaks characteristic of cobalt oxides (Fig. 1). The significant increase of  $d_{302}$  spacing observed for the Co(PS)SiBEA sample upon incorporation of cobalt into SiBEA, from 3.887 Å ( $2\theta = 22.86^\circ$ ) for SiBEA to 3.898 Å ( $2\theta = 22.79^\circ$ ) for Co(PS)SiBEA, indicates expansion of the zeolite matrix and introduction of cobalt into the framework positions of BEA zeolite (Fig. 1) [9]. Introduction of cobalt into HAIBEA zeolite by conventional wet impregnation and ion exchange methods leads to smaller increase in the  $d_{302}$  spacing than that observed for introduction of cobalt into SiBEA support (Fig. 1), from 3.940 Å ( $2\theta = 22.58^\circ$ ) (HAIBEA) to 3.941 Å ( $2\theta = 22.55^\circ$ ) (Co(IMP)AlBEA and Co(IE)AlBEA) suggesting that only limited amount of cobalt may be successfully incorporated into the framework of HAIBEA zeolite by these methods [9] and as a result, the extraframework cobalt species should dominate over the framework ones.

In all as obtained catalysts cobalt is present mainly as pseudo-tetrahedral Co(II), however it should be noted that in Co(IMP)AlBEA and Co(IE)AlBEA samples obtained by wet impregnation and ion exchange methods, also some amounts of extra-framework Co(II) in octahedral coordination are present (Fig. 2), as evidenced by the combined use of DR UV-vis and XPS methods, reported earlier [9]. The Co(IMP)AlBEA and Co(IE)AlBEA samples have a significantly higher concentration of the acidic sites than Co(PS)SiBEA, as evidenced by FT-IR of adsorbed CO and TPD of ammonia. The differences in acidity of the catalysts are related to the presence of Al and Co species in the zeolite structure, which can act as Brønsted or Lewis acidic sites and influence catalytic properties of the obtained samples [9].

The as obtained Co-containing zeolites were studied as catalysts for the selective reduction of NO with ammonia (DeNO<sub>x</sub>, SCR). The results of these studies are presented in Fig 3. The fully dealuminated BEA (SiBEA) shows a very low activity and NO conversion, which does not exceed 7 % in the studied temperature range. The NO conversion substantially increases after incorporation of cobalt in the framework of SiBEA as shown for Co(PS)SiBEA (Fig. 3). For this sample, the reaction starts already at 423 K (Fig. 3) and the NO conversion reaches 90 % with the selectivity to N<sub>2</sub> above 90 % at temperature 673 K. It suggests that well dispersed framework pseudo-tetrahedral Co(II) are responsible for high activity of the catalyst in the SCR of NO process. For HAlBEA, the reaction starts at about 673 K and the NO conversion increase to 35 - 40 % at 723 K (Fig. 3). An introduction of cobalt into the HAlBEA support significantly increases its catalytic activity. The NO conversion starts at 473 K and increases to about 100 % at 623 K with selectivity to nitrogen above 95 % in the presence of the Co(IMP)AlBEA catalyst (Fig. 3). It seems that high conversion of NO measured over the Co(IMP)AlBEA catalyst is related to the co-presence of well dispersed pseudo-tetrahedral and octahedral Co(II) species. Figure 3 shows also the results of activity tests in the DeNO<sub>x</sub> process for the catalyst prepared by ion exchange method. The NO conversion measured for Co(IE)AlBEA increases in the range of 423 - 673 K reaching about 100 %. At higher temperatures the side process of direct ammonia oxidation significantly decreased the effectiveness of the DeNO<sub>x</sub> process (Fig. 3). The selectivity to nitrogen, measured for Co(IE)AlBEA, is above 90 % in the broad temperature range (473 - 723 K). The catalytic activities of all the Co-containing zeolite catalysts are high, however, it should be noted that the catalysts with both pseudo-tetrahedral and octahedral Co(II) are more active in the low-middle temperature range (423 - 653 K), while the sample containing only

framework pseudo-tetrahedral cobalt species (Co(PS)SiBEA), shows higher activity at elevated temperatures ( $T > 653$  K) (Fig. 3). Thus, it seems that because of different contribution and state of cobalt species, well dispersed framework pseudo-tetrahedral Co(II) for Co(PS)SiBEA and well dispersed octahedral Co(II) in cationic extra-framework positions for Co(IMP)AIBEa and Co(IE)AIBEa, the obtained catalysts are characterised by different catalytic activities in the temperature range of 423 - 723 K. It should be noted that commercial catalysts for this process, based on the  $V_2O_5$ - $TiO_2$  oxide system, operate in a significantly narrower temperature range of 523-673 K. Therefore, these preliminary results obtained in our studies, are very promising. Earlier, Iwasaki et al. [50] studied catalytic performance of Fe/ZSM-5 prepared by various methods (IMP, IE, CVD) and shown that the iron-containing ZSM-5 catalysts prepared by CVD method, demonstrate high activity in SCR-NO ( $\sim 100\%$  NO conversion at about 573 K). The similar results were reported by Delahay et al. [5]. Moreover, the results of the catalytic tests in SCR NO provided by Metkar et al., [16,51] and other authors [52-54] are characterized by a very similar performance of the dual-layer and dual-brick Fe/Cu-zeolite catalysts. However, it should be noted that the reaction conditions are different in both cases. Moreover, nitrogen oxide conversions measured for Co(IMP)AIBEa ( $\sim 100\%$  NO conversion at 623 K) and Co(IE)AIBEa ( $\sim 90\%$  NO conversion at 623 K), proves their high catalytic activity in SCR-NO with ammonia. Therefore, not only CVD but also IE or IMP methods result in well-defined catalysts, containing well dispersed Co(II) in pseudo-tetrahedral and octahedral coordination. It suggests that such Co(II) sites play major role in the SCR-NO reaction. The Co(IMP)AIBEa and Co(IE)AIBEa samples achieve high NO conversions over a broader temperature range and seems to be a strong alternative to the high-performance catalysts reported earlier [52-54]. Analysis of the catalytic results of Co(IMP)AIBEa and Co(IE)AIBEa in SCR-NO with  $NH_3$  (Fig. 3), have shown that ion exchange as well as wet impregnation procedure allow to obtain the high-performance catalysts. The aim of the presented below studies is to obtain more insight into the mechanism of the DeNO<sub>x</sub> process over cobalt modified BEA zeolite.

### 3.1. $NH_3$ -TPD

The  $NH_3$ -TPD patterns of the zeolite samples are shown in Fig. 4. Apart from ammonia no other gas was emitted during the TPD runs. The ammonia desorption profile obtained for SiBEA zeolite is represented by a symmetric maximum at about 402 K. The SiBEA sample is

characterized by low acidity, represented by a relatively low concentration of adsorbed ammonia ( $0.099 \mu\text{mol m}^{-2}$ ). An introduction of cobalt into SiBEA by two-step postsynthesis method results in an increase of acidic sites concentration from  $0.099 \mu\text{mol m}^{-2}$  (for SiBEA) to  $0.135 \mu\text{mol m}^{-2}$  (for Co(PS)SiBEA). Therefore, it seems that cobalt ions incorporated into the zeolite support are responsible for the formation of additional centres for ammonia adsorption [55-62]. Moreover, incorporation of Co ions into the SiBEA support shifts TPD peaks to higher temperatures (Fig. 4), that suggests a slight increase in the strength of acidic sites in Co(PS)SiBEA [55]. For the samples prepared by wet impregnation and ion exchange methods, the TPD patterns can be divided into two ranges: 430 - 500 K and 500 - 625 K (Fig. 4). It should be noted that introduction of cobalt into the HAIBEA support results in decrease of acidic sites concentration by more than a half, from  $1.882 \mu\text{mol m}^{-2}$  (HAIBEA) to  $0.834 \mu\text{mol m}^{-2}$  (Co(IMP)AIBEA). An introduction of Co by ion exchange method into the HAIBEA support is also accompanied by a decrease in adsorbed ammonia concentration from  $1.882 \mu\text{mol m}^{-2}$  (HAIBEA) to  $1.519 \mu\text{mol m}^{-2}$  (Co(IE)AIBEA).

### 3.2. NO-TPD

The NO desorption profiles obtained for the HAIBEA, Co(PS)SiBEA, Co(IMP)AIBEA and Co(IE)AIBEA samples are shown in Fig. 5. First of all, it should be noted that the estimated content of nitrogen oxide adsorbed on HAIBEA, Co(PS)SiBEA, Co(IMP)AIBEA and Co(IE)AIBEA is significantly lower in comparison to chemisorbed ammonia, and is equal to  $0.011$ ,  $0.002$ ,  $0.005$  and  $0.008 \mu\text{mol m}^{-2}$ , respectively. The NO-TPD profile of nitrogen oxide evolution from HAIBEA zeolite contains two desorption peaks at 396 and 558 K. The NO desorption from the Co(PS)SiBEA sample occurs in the temperature range of 353 - 573 K and consists of one maximum at 463 K (Fig. 5), thus NO molecules are relatively weakly bound to the catalyst surface. The NO desorption from the Co(IMP)AIBEA proceeds in two stages (Fig. 5). The weakly chemisorbed NO molecules are desorbed in the range of 353 - 423 K, while nitrogen oxide strongly bounded to the catalyst surface is released in the region of 423 - 593 K (Fig. 5). The NO-TPD desorption profile of the Co(IE)AIBEA sample consists of two maxima at 390 and 558 K. It should be noted that desorption of NO from Co(IMP)AIBEA and Co(IE)AIBEA (containing both cobalt and aluminium) proceeds in two steps, indicating that NO forms species bounded to the surface with different strength and suggests that NO adsorbs at two different types

of active sites. Since the peak at 558 K is also observed for HAlBEA, one can conclude that NO is probably bonded to aluminium sites with higher strength [63,64]. On the other hand for Co(PS)SiBEA (containing only cobalt species), only one NO desorption peak is found. Thus, it could be suggested that cobalt is responsible for the formation of weaker sites, which can be identified as pseudo-tetrahedral Co(II) framework sites [63,64].

### 3.3. FT-IR of $\text{NH}_3$ or NO adsorbed on Co(PS)SiBEA, Co(IMP)AlBEA and Co(IE)AlBEA

In order to identify the nature of the surface species, the FT-IR spectra were recorded for the Co(PS)SiBEA, Co(IMP)AlBEA and Co(IE)AlBEA catalysts on which  $\text{NH}_3$  or NO were adsorbed.

Fig. 6 shows the spectra recorded for the samples after NO adsorption and subsequent desorption at different temperatures (373, 473 and 573 K). The FT-IR spectra recorded for Co(PS)SiBEA exhibit characteristic bands at 1875 and 1795  $\text{cm}^{-1}$  (Fig. 6). These bands are due to nitrosyl complexes on cobalt cations [64-68]. In agreement with literature data [64-68], the bands at 1875 and 1795  $\text{cm}^{-1}$  can be assigned to Co(II)-(NO)<sub>2</sub> type dinitrosyls. The weak band at 1620  $\text{cm}^{-1}$  present in the FT-IR spectrum for Co(PS)SiBEA (Fig. 6), indicates the presence of small amount of bridged NO (NO<sub>2</sub>) species [69]. FT-IR spectra recorded for the Co(IMP)AlBEA and Co(IE)AlBEA samples, exhibit bands at 1895, 1815 and 1620  $\text{cm}^{-1}$ , which can be assigned to Co(II)-(NO)<sub>2</sub> type dinitrosyls (1895, 1815  $\text{cm}^{-1}$ ) and bridged NO species (1620  $\text{cm}^{-1}$ ) [66-69], respectively. For Co(IMP)AlBEA and Co(IE)AlBEA, additional IR band at 1930  $\text{cm}^{-1}$  is observed (Fig. 6), which can be attributed to linear Co<sup>n+</sup>-NO mononitrosyls [67]. In this case the possible oxidation state of cobalt ions is 3+ since, as a rule, Co(II)-NO compounds absorb at lower frequencies than Co(III)-NO (Fig. 6) [67]. The Co(III) species are formed as a result of NO adsorption, by *in situ* oxidation of Co(II) species, since their presence was not confirmed by UV-Vis and XPS studies for the samples not contacted with NO. It should also be mentioned that no band was observed at the region of NO<sup>+</sup> vibration. Obviously, since no NO<sup>+</sup> was observed, NO adsorption did not disturb the vibration of the Si-OH-Al groups.

As shown in Fig. 6, after  $\text{NH}_3$  adsorption FT-IR spectra for the Co(IMP)AlBEA and Co(IE)AlBEA samples, contain bands at 1455, 1620 and 1690  $\text{cm}^{-1}$ . The bands observed at 1455 and 1690  $\text{cm}^{-1}$  can be attributed to  $\text{NH}_4^+$  species formed by the reaction of  $\text{NH}_3$  with Brønsted acidic sites (symmetric and antisymmetric N-H bending modes), while band at 1620  $\text{cm}^{-1}$ , proves

the presence of Lewis acidity due to the formation of coordinatively bonded  $\text{NH}_3$  species [17,18,70-72] (Fig. 6). FT-IR spectrum of  $\text{Co(PS)SiBEA}$  after  $\text{NH}_3$  adsorption contains only one band at  $1620\text{ cm}^{-1}$ . It suggests that, in the  $\text{Co(PS)SiBEA}$  sample only Lewis acid sites originating from the presence of the framework cobalt species are present [17,18,70-72]. It should be noted that ammonia is much stronger bonded to the zeolite surface than nitric oxide because the bands related to ammonia bonded to Brønsted and Lewis acidic sites are still present in the FT-IR spectra of the catalysts evacuated at 573 K (Fig. 6).

### 3.4. *NO-NH<sub>3</sub>-TPD, NH<sub>3</sub>-NO-TPD and FT-IR with NO + NH<sub>3</sub> or NH<sub>3</sub> + NO*

Fig. 7 shows the results of  $\text{NO-NH}_3$ -TPD measurements (first adsorption of  $\text{NO}$  and then  $\text{NH}_3$ ) obtained for the  $\text{Co(PS)SiBEA}$ ,  $\text{Co(IMP)AlBEA}$  and  $\text{Co(IE)AlBEA}$  catalysts. For all the studied samples, ammonia desorption patterns are similar to that measured in  $\text{NH}_3$ -TPD (Fig. 4). It should be noted that apart from  $\text{NH}_3$  and  $\text{NO}$  desorption, also evolution of  $\text{N}_2$  and  $\text{H}_2\text{O}$ , which are the products of the SCR reaction, is detected in the range of 373 - 593 K. The evolution of water vapour, which is spread to higher temperatures, suggests that  $\text{H}_2\text{O}$  molecules do not desorb directly to gas phase but are bounded to the catalyst surface [70]. It should be noted that not all ammonia and nitrogen oxide molecules adsorbed on the studied catalysts are properly activated to be converted to  $\text{DeNO}_x$  products and desorb in unchanged forms. For all the studied samples, the main peak of ammonia desorption, located at about 395 - 442 K, is related to  $\text{NH}_3$  adsorbed on Lewis acidic sites, whereas the shoulder peak at 625 K can be assigned to  $\text{NH}_3$  adsorbed on Brønsted acidic centres, this is confirmed by the FT-IR studies (Fig. 9) because the maximum at  $1620\text{ cm}^{-1}$  of coordinatively bonded  $\text{NH}_3$  disappears at lower desorption temperatures than the maximum at  $1450\text{ cm}^{-1}$ , characteristic of  $\text{NH}_4^+$  cations [58-60]. Nitrogen oxide molecules desorb mainly from framework cobalt sites at 350 - 600 K [63,64]. For all the catalysts studied, the maxima of ammonia desorption fall at lower temperatures for zeolites with pre-adsorbed  $\text{NO}$  in comparison to  $\text{NH}_3$ -TPD profiles for pure supports (Fig. 4). It may suggest that ammonia and nitrogen oxide adsorb on the same adsorption sites and  $\text{NO}$ , which is adsorbed as first, occupied part of stronger acidic sites and is not removed from these sites during subsequent ammonia adsorption. This is also confirmed by the fact that the intensity of the  $\text{NH}_4^+$  band (Fig. 9) is not influenced by the presence of the pre-adsorbed  $\text{NO}$  while the intensity of  $1620\text{ cm}^{-1}$  band is reduced (the case of  $\text{Co(IE)AlBEA}$ ). Interestingly, in the case of  $\text{NH}_3$  adsorption right after  $\text{NO}$

adsorption, the band at  $1930\text{ cm}^{-1}$  characteristic of Co(III)-NO adducts is absent in this zeolite and the intensities of all dinitrosyls band is increased. This means that in the presence of  $\text{NH}_3$  the oxidation of cobalt sites by NO is reduced and that in fact the presence of ammonia influences the properties of the redox sites in the catalyst prepared by this method. The strength of NO and  $\text{NH}_3$  bonding in this zeolite is similar.

For Co(IMP)AIBEA adsorption of  $\text{NH}_3$  causes instantaneous NO desorption suggesting that ammonia is adsorbed at the same adsorption sites as NO and the former is bonded much stronger. Co(PS)SiBEA shows the reverse behaviour. In this case the presence of NO predominantly blocks the adsorption sites for ammonia. As evidenced by very low intensity of the band at  $1620\text{ cm}^{-1}$  assigned to coordinatively bonded ammonia for this sample in comparison with Co(IMP)AIBEA and Co(IE)AIBEA samples. The band at  $1620\text{ cm}^{-1}$  may originate both from NO and  $\text{NH}_3$  adsorption, but since the N-H stretching maxima are present in the region  $3400$  to  $3200\text{ cm}^{-1}$ , this suggest that the band  $1620\text{ cm}^{-1}$  may be assigned only to coordinatively bonded ammonia. Thus, for Co(PS)SiBEA ammonia is bonded on different adsorption sites than NO and the presence of adsorbed  $\text{NH}_3$  is only slightly influencing the redox properties of framework cobalt sites.

Fig. 8 shows the results of  $\text{NH}_3$ -NO-TPD (first adsorption of  $\text{NH}_3$  and then NO) measurements for Co(PS)SiBEA, Co(IMP)AIBEA and Co(IE)AIBEA. For Co(PS)SiBEA, the nitrogen and water vapour are formed in the temperature range from  $373$  to  $553\text{ K}$  (Fig. 8), while the evolution of these products in case of Co(IMP)AIBEA and Co(IE)AIBEA is spread in the range of  $393$  -  $653\text{ K}$  (Fig. 8). The profiles of nitrogen and water vapour evolution obtained for Co(IMP)AIBEA and Co(IE)AIBEA, contain two maxima, which proves that part of water is emitted directly to the gas phase (maxima at  $433$  and  $441\text{ K}$ ), while another part of  $\text{H}_2\text{O}$  molecules is bounded to the catalyst surface (maxima at  $502$  and  $512\text{ K}$ ) (Fig. 8) (e.g. in the form of -OH groups), and is removed from the catalysts at higher temperatures [14]. It should be underlined that the quantity of chemisorbed NO estimated in  $\text{NH}_3$ -NO-TPD is lower than that assessed in NO- $\text{NH}_3$ -TPD, which suggests that  $\text{NH}_3$  molecules (adsorbed prior to NO adsorption) partially block active sites for NO adsorption and proves the hypothesis that ammonia and nitrogen oxide molecules compete for the same adsorption sites [17,48,49].

The FT-IR spectra recorded for the samples pre-adsorbed with  $\text{NH}_3$  before NO adsorption are shown in Fig. 10. In case of Co(IMP)AIBEA and Co(IE)AIBEA, after  $\text{NH}_3$  adsorption the



presence of the bands at 1455, 1690 and 1620  $\text{cm}^{-1}$  proves the presence of Brønsted and Lewis acidic sites, respectively [18,57,70-72]. The band at 1805  $\text{cm}^{-1}$ , observed in FT-IR spectra for Co(IMP)AlBEA and Co(IE)AlBEA after NO adsorption, can be assigned to Co(II)-(NO)<sub>2</sub> type dinitrosyls [66-69], the symmetric band with much lower absorption coefficient is almost invisible. The Co(PS)SiBEA sample contains Lewis acidic sites (Co(II) species) characterized by the NH<sub>3</sub> adsorption band at 1620  $\text{cm}^{-1}$  [17,18,70,71] and with NO adsorption bands at 1790 and 1875  $\text{cm}^{-1}$  [66-69] (Fig. 10).

Analysis of the obtained spectra recorded for the samples, shows that NO adsorption on the samples with pre-adsorbed NH<sub>3</sub> is very limited, and replacement of ammonia by NO on BEA zeolite is very difficult. Thus, the obtained results show that NO and NH<sub>3</sub> molecules compete for the same adsorption sites, and ammonia is able to remove NO from the active centres [13,15,17,48,49]. The only exception are framework Co(II) sites in Co(PS)SiBEA, which adsorbs NO much stronger than ammonia. In fact, there is negligible amount of ammonia adsorbed in this zeolite, due to the absence of aluminium-based Brønsted and Lewis acid sites. Moreover, combined temperature-programmed studies and FT-IR with NO and/or NH<sub>3</sub> adsorption of the reaction mechanism show that the DeNO<sub>x</sub> process performed in the presence of Co(IMP)AlBEA and Co(IE)AlBEA involves the reaction between chemisorbed ammonia and NO from gas phase (Eley-Riedel mechanism) [73,74]. In case of Co(PS)SiBEA a significant role of strongly adsorbed NO on framework cobalt species in the DeNO<sub>x</sub> mechanism cannot be excluded.

General conclusions from the FT-IR analysis may be summarized as follows. There are two type of cobalt centres present in the investigated zeolites: (i) Framework, pseudo-tetrahedral Co(II) species, the most abundant in Co(PS)SiBEA zeolite, in which cobalt occupies the position taken before by the framework Al atoms. Such centres are able to bond NO much stronger than NH<sub>3</sub> and are not easily oxidized by NO (ii) Extra-framework, probably octahedral Co(II) species, the most abundant in Co(IE)AlBEA. They bond NH<sub>3</sub> much strongly than NO (NO is instantaneously flushed by NH<sub>3</sub>). Presence of NH<sub>3</sub> is influencing their redox properties, since they are easily oxidized to Co(III) in the absence of ammonia and oxidation is stopped in the NH<sub>3</sub> presence.

Both types of centres are present in the Co(IMP)AlBEA and thus this zeolite behaviour is intermediate – some of the cobalt centres preferentially adsorb NO and some – ammonia. The presence of NH<sub>3</sub> is influencing the properties of extraframework cobalt(II) only.

### 3.5. NO-TPSR

The reactivity of pre-adsorbed nitrogen oxide was studied by temperature-programmed surface reaction (NO-TPSR) performed in a continuous flow of helium-diluted ammonia. Results obtained for all the studied samples are shown in Fig. 11. It should be noted that significant amounts of ammonia are adsorbed on the catalyst surface just after a flow of NH<sub>3</sub>/He gas mixture was switched on [60,70]. In spite of pre-adsorbed NO, the samples adsorb ammonia from the gas phase up to about 353 - 393 K without NO desorption (Fig. 8), what suggests that NH<sub>3</sub> molecules are firstly chemisorbed on sites non-occupied by NO molecules. At higher temperature (403 - 523 K), ammonia replaces weakly adsorbed NO from surface sites (Fig. 11) [48,49,60]. It should be noted that NO desorption patterns for all the catalysts are similar to that obtained in NO-TPD (Fig. 5) or NO-NH<sub>3</sub>-TPD (Fig. 7). For all the studied samples, the reaction between NH<sub>3</sub> and NO adsorbed species occurred, forming N<sub>2</sub> and H<sub>2</sub>O. In case of Co(PS)SiBEA, the evolution of N<sub>2</sub> and H<sub>2</sub>O is detected in the temperature range of 353 - 523 K (Fig. 11), while in the presence of the Co(IMP)AlBEA and Co(IE)AlBEA samples, nitrogen and water vapour are realised in the temperature range of 393 - 653 K (Fig. 11), which indicates that surface nitrates are converted by ammonia into the DeNO<sub>x</sub> products [70]. It is worth noting that nitrous oxide production is not detected for the studied catalysts, which is consistent with the earlier studies on Fe-ZSM-5 [37].

### 3.6. NH<sub>3</sub>-TPSR

The reactivity of NH<sub>3</sub> pre-adsorbed on the catalyst surface was characterized by temperature-programmed surface reaction (NH<sub>3</sub>-TPSR) performed in a continuous flow of helium-diluted NO. The NH<sub>3</sub>-TPSR profiles are shown in Fig. 12. The ammonia desorption patterns obtained for all the samples are spread in the range of 353 - 533 K. It is shown in Fig. 12, that the reaction between NO and pre-adsorbed NH<sub>3</sub> species occurred, producing N<sub>2</sub> and H<sub>2</sub>O [60,70]. The NH<sub>3</sub> desorption pattern obtained for Co(PS)SiBEA is spread in the range of 353 - 453 K with maximum at 405 K (Fig. 12). The evolution of nitrogen in case of the Co(PS)SiBEA sample begins at temperature as low as 353 K (Fig. 12), while the emission of H<sub>2</sub>O molecules is detected in the range from 353 to 753 K with two maxima at 412 and 648 K (Fig. 12), respectively [60,70]. Similar TPSR patterns are recorded for the Co(IMP)AlBEA and Co(IE)AlBEA samples, however, maxima of nitrogen and water vapour evolution are found in

different positions in relation to that found for the Co(PS)SiBEA sample (Fig. 12). In all NH<sub>3</sub>-TPSR experiments, evolution of H<sub>2</sub>O molecules occurs at higher temperature than emission of nitrogen, suggesting that N<sub>2</sub> immediately desorbs from the catalyst surface, while water is bonded much strongly to the catalyst surface, and remains on the surface to higher temperature [60,70]. It is worth to notice that the rate of the nitrogen and water formation proceeds with an increase in reaction temperature until ammonia is completely removed from the catalyst surface [60,70]. Moreover, for all the studied catalysts, only the fraction of chemisorbed ammonia is converted into the DeNO<sub>x</sub> products, while majority of NH<sub>3</sub> desorbs from the zeolite surface. This fact suggests that only part of chemisorbed ammonia is activated properly to be converted into the DeNO<sub>x</sub> products [60].

### **3.7. Stop flow-TPD**

In order to check what type of species is present on the catalyst surface under reaction conditions, stop flow-TPD experiments were performed for the studied samples. After isothermal catalytic test performed at 623 K for 1 hour, the catalyst sample was purged in a flow of helium and then TPD measurement in a flow of pure helium was done in order to identify species present on the catalyst surface. The results of this study are shown in Fig. 13. It should be underline that chemisorbed ammonia molecules are the main species present under reaction conditions in case of the studied catalysts. Ammonia desorption from Co(PS)SiBEA (sample without Al but containing Co species) occurs in the range of 643 - 673 K, while for Co(IMP)AlBEA and Co(IE)AlBEA (containing both Al and Co species), only above 683 K (Fig. 13). These differences can be related to different acidity of cobalt and aluminium species. Evolution of nitrogen oxide for all the studied catalysts is detected in very similar ranges as desorption of ammonia from these samples, however amount of desorbing NO is significantly lower comparing to ammonia desorption [13,15,17,48,49]. In case of Co(PS)SiBEA chemisorbed ammonia and nitrogen oxide present low reactivity and therefore, only slight amount of nitrogen is formed during the experiment. Larger amounts of nitrogen is formed as results the reaction between NH<sub>3</sub> and NO chemisorbed on the surface of Co(IMP)AlBEA and Co(IE)AlBEA. The results of the presented studies lead to the conclusion that the main reaction pathway proceeds by adsorption and activation of NH<sub>3</sub> molecules on the zeolite catalysts surface [13,15,17,48,49]. It must be stressed that only part of adsorbed ammonia is properly activated for the DeNO<sub>x</sub> process, while

part of them are only spectators [13,15,17]. In case all studied samples only small amount of NO molecules, in comparison to ammonia, is adsorbed, possibly in the form of nitrates, on the catalyst surface under reaction conditions. Thus, probably the main reaction proceeds between chemisorbed ammonia and NO from gas phases (or loosely bounded to the catalyst surface) [13,15,17].

#### 4. Conclusions

The modification of the BEA supports with cobalt results in their catalytic activation in the DeNO<sub>x</sub> process. It was shown that cobalt introduced in the zeolitic supports effectively activates chemisorbed NH<sub>3</sub> molecules for reaction with NO.

The highest reactivity of ammonia in the SCR-NO reaction, with the high selectivity towards nitrogen, was found for the Co(IE)AlBEA catalyst, prepared by ion exchange method.

Total conversion of chemisorbed NH<sub>3</sub> is not observed for any zeolite catalyst. Probably, not all ammonia and nitrogen oxide molecules adsorbed on the studied catalysts are properly activated to be converted to DeNO<sub>x</sub> products and desorb in the unchanged forms. The surface concentration of chemisorbed NO is lower in comparison to chemisorbed NH<sub>3</sub> in case of all the samples studied.

The results of temperature-programmed studies and FT-IR with NO and/or NH<sub>3</sub> adsorption for the Co-containing samples, show that during the SCR reaction, ammonia and nitrogen oxide adsorb on the same adsorption sites. In general, ammonia is stronger bound to the catalyst surface in comparison with NO and is able to remove nitrogen oxide molecules from the adsorption sites. An exception is NO adsorbed on stronger Lewis acid sites, probably framework Co(II) species, which was not removed from these sites during subsequent ammonia adsorption on the Co(PS)SiBEA sample. Analogously, NH<sub>3</sub> molecules adsorbed prior to NO adsorption probably block active sites for NO adsorption on Co(IMP)AlBEA. In case of Co(IE)AlBEA, it is partially blocking of active sites by ammonia, while for Co(PS)SiBEA zeolite, this effect does not appear. It proves the hypothesis that ammonia and nitrogen oxide molecules compete for the same adsorption sites.

The results of FT-IR studies for the catalysts pre-adsorbed with both NO and NH<sub>3</sub>, clearly show that nitrogen oxide is desorbed from Co(IMP)AlBEA immediately upon contact with

ammonia, while in case of Co(PS)SiBEA and Co(IE)AlBEA, ammonia does not displace completely nitrogen oxide.

Comparison of the results of catalytic tests and results of the reaction mechanism studies leads to the conclusion that extra-framework cobalt species are more catalytically active in comparison to Co(II) cations incorporated to the BEA zeolite framework. In case of the extra-framework cobalt species the reaction proceeded according to Eley-Riedel mechanism including reaction between chemisorbed ammonia and NO from gas phase. On the other hand, in case of the Co(PS)SiBEA catalyst an important role of NO adsorbed on Co(II) cations incorporated into the BEA zeolite framework in the mechanism of the DeNO<sub>x</sub> process cannot be excluded.

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Figure captions

**Figure 1.** XRD patterns recorded at room temperature of HAIBEA, Co(PS)SiBEA, Co(IMP)AlBEA and Co(IE)AlBEA.

**Figure 2.** DR UV-vis spectra recorded at ambient atmosphere of Co(PS)SiBEA, Co(IMP)AlBEA and Co(IE)AlBEA.

**Figure 3.** NO conversion and N<sub>2</sub> selectivity in SCR of NO with NH<sub>3</sub> of SiBEA, HAIBEA, Co(PS)SiBEA, Co(IMP)AlBEA and Co(IE)AlBEA.

**Figure 4.** NH<sub>3</sub>-TPD profiles of SiBEA, HAIBEA, Co(PS)SiBEA, Co(IMP)AlBEA and Co(IE)AlBEA.

**Figure 5.** NO-TPD of HAIBEA, Co(PS)SiBEA, Co(IMP)AlBEA and Co(IE)AlBEA.

**Figure 6.** FTIR spectra of NO (first line) and NH<sub>3</sub> (second line) adsorbed on Co(IE)AlBEA, Co(IMP)AlBEA and Co(PS)SiBEA zeolites activated under vacuum at 723 K. All spectra at 373 K, normalized to 10 mg of catalyst.

**Figure 7.** The results of NO-NH<sub>3</sub>-TPD experiments obtained for Co(PS)SiBEA, Co(IMP)AlBEA and Co(IE)AlBEA.

**Figure 8.** The results of NH<sub>3</sub>-NO-TPD experiments obtained for Co(PS)SiBEA, Co(IMP)AlBEA and Co(IE)AlBEA.

**Figure 9.** FTIR spectra of NH<sub>3</sub> adsorption on Co(IE)AlBEA, Co(IMP)AlBEA and Co(PS)SiBEA zeolites with pre-adsorbed NO. All spectra at 373 K, normalized to 10 mg of catalyst.

**Figure 10.** FTIR spectra of NO adsorption on Co(IE)AlBEA, Co(IMP)AlBEA and Co(PS)SiBEA zeolites with pre-adsorbed NH<sub>3</sub>. All spectra at 373 K, normalized to 10 mg of catalyst.

**Figure 11.** The results of NO-TPSR measurements for Co(PS)SiBEA, Co(IMP)AlBEA and Co(IE)AlBEA.

**Figure 12.** The results of NH<sub>3</sub>-TPSR measurements for Co(PS)SiBEA, Co(IMP)AlBEA and Co(IE)AlBEA.

**Figure 13.** The results of stop-flow TPD experiments for Co(PS)SiBEA, Co(IMP)AlBEA and Co(IE)AlBEA.

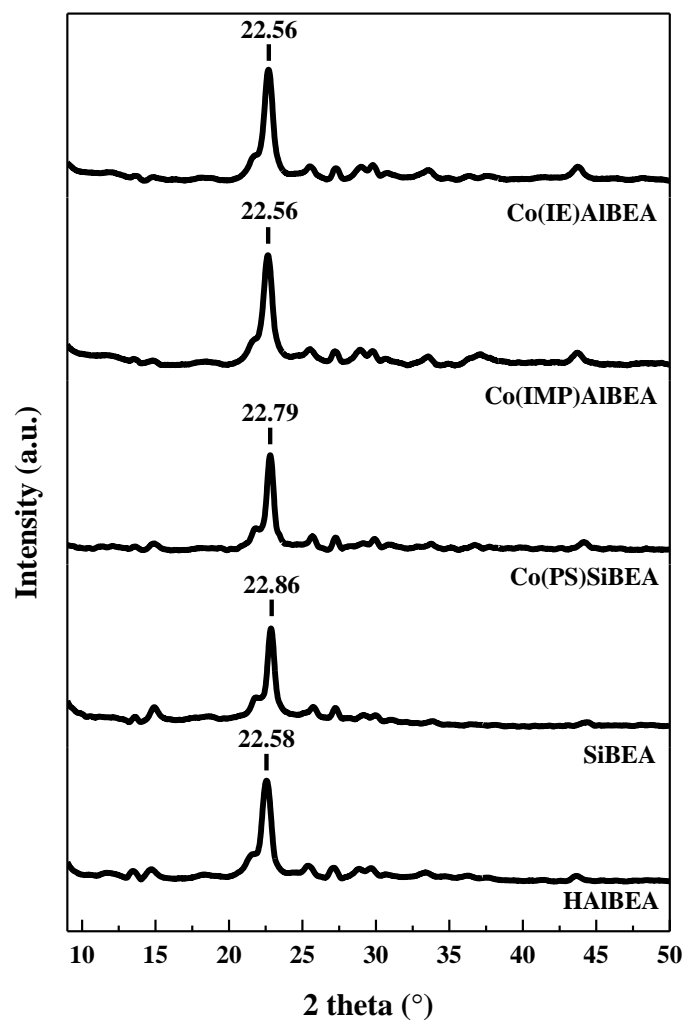
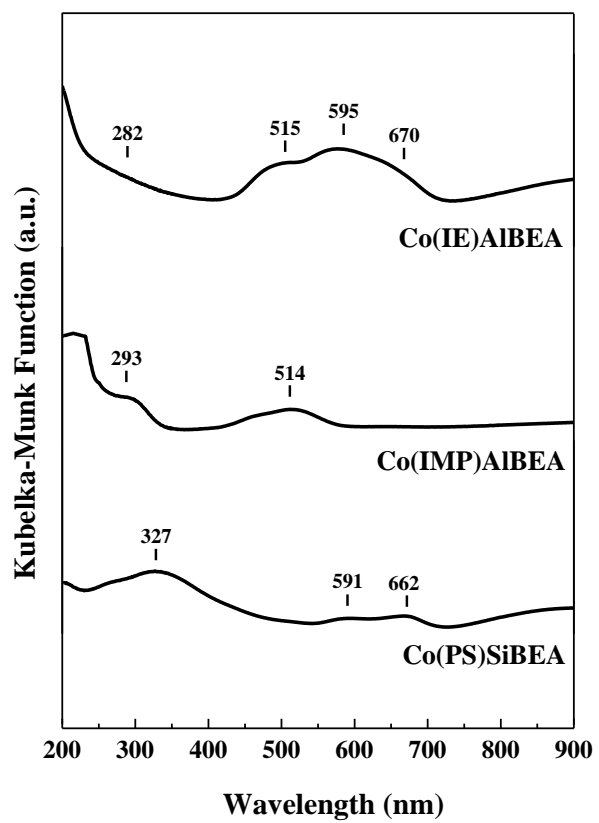


Fig. 1

**Fig. 2**

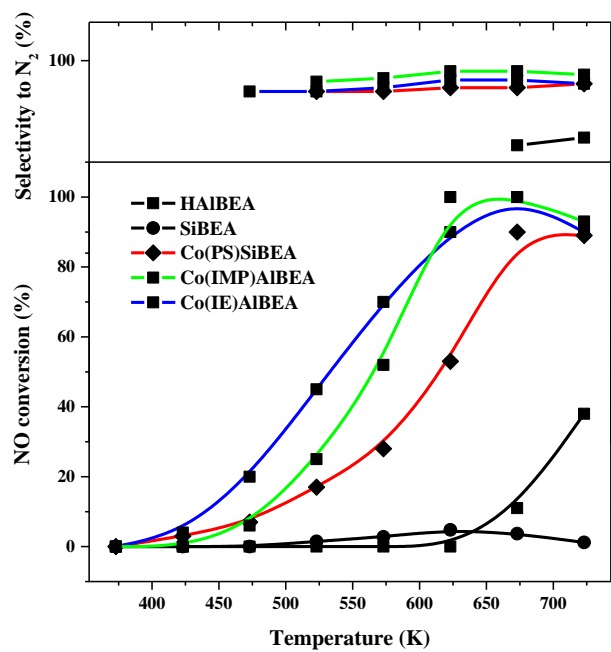


Fig. 3

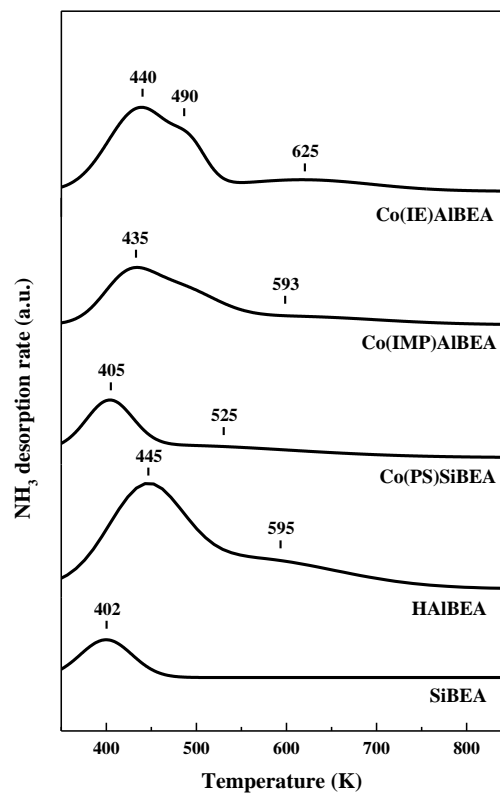
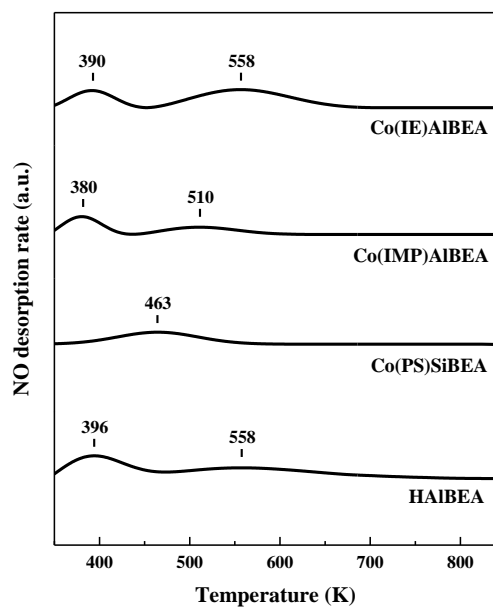


Fig. 4

**Fig. 5**

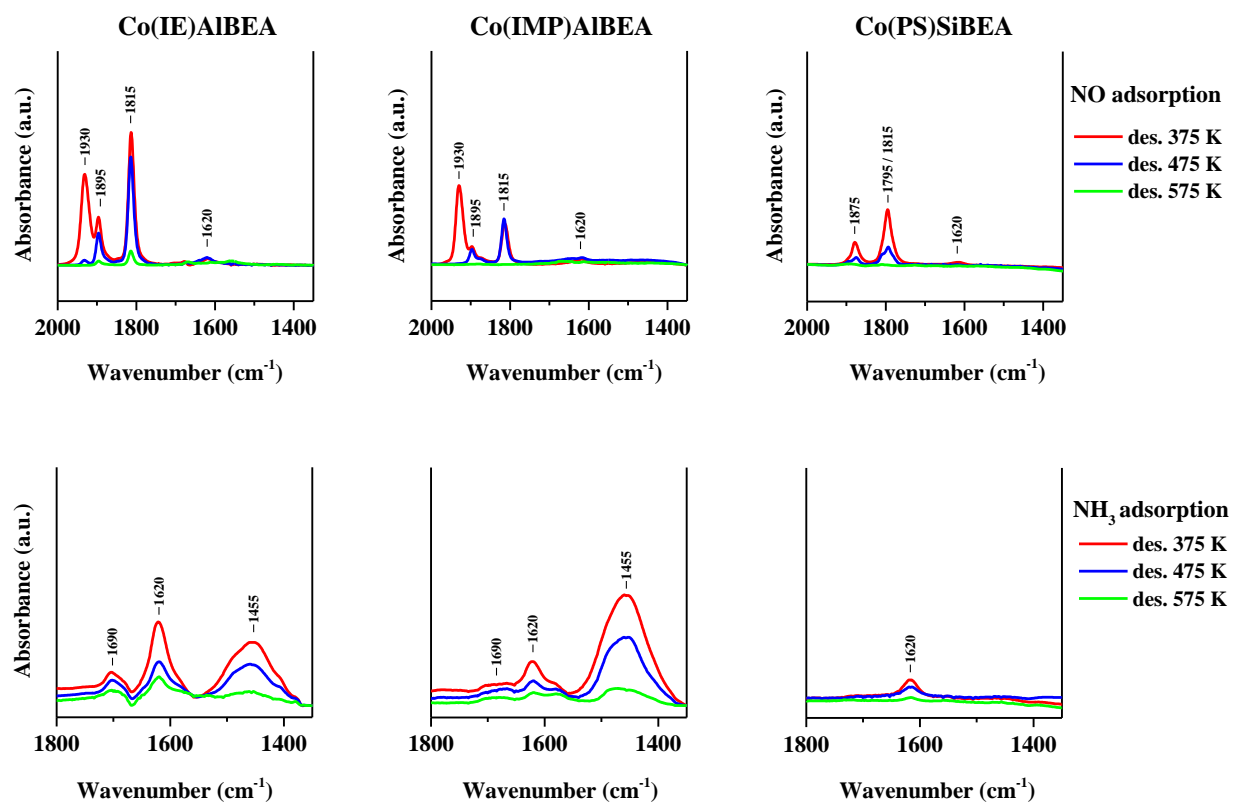


Fig. 6



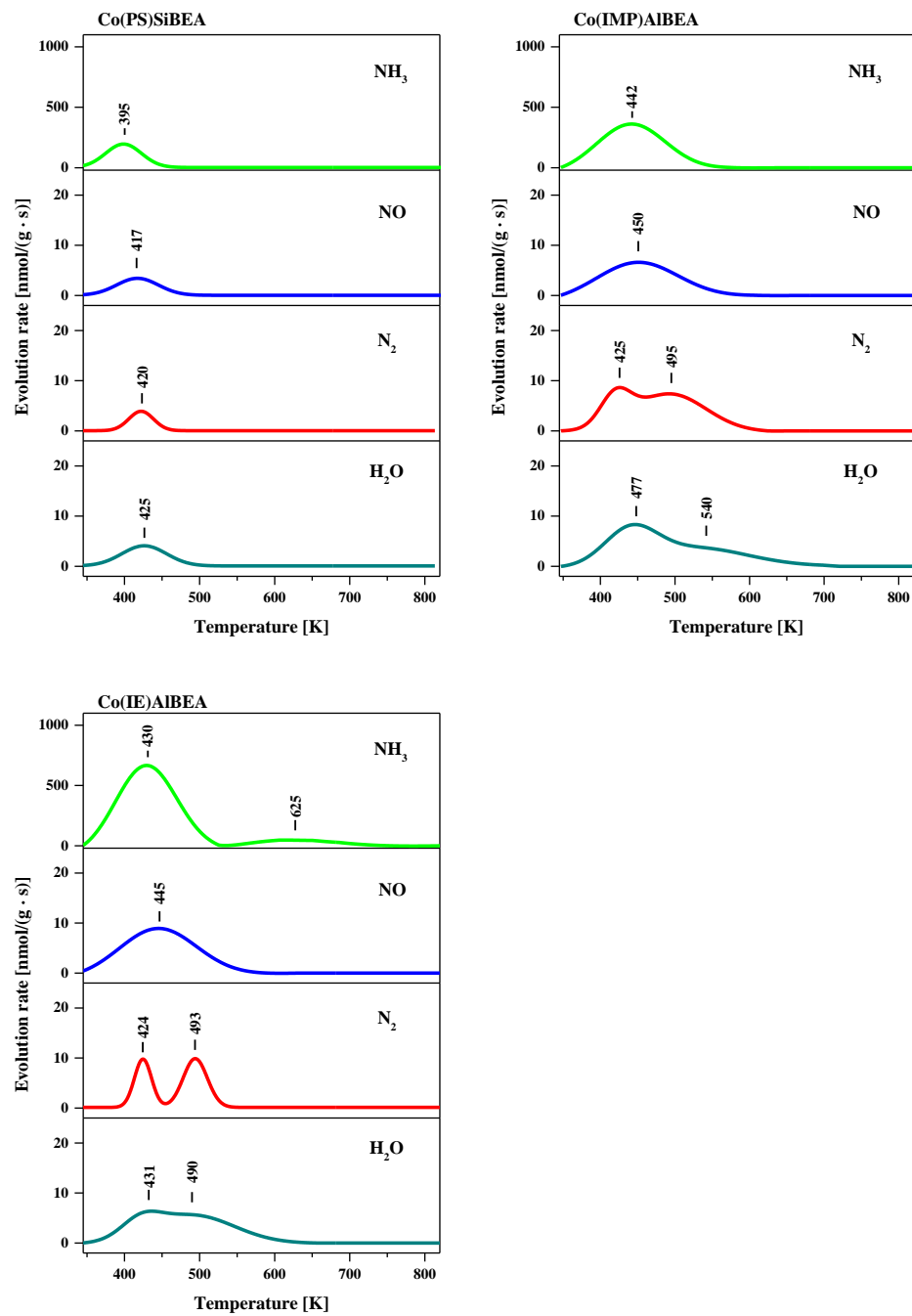


Fig. 7

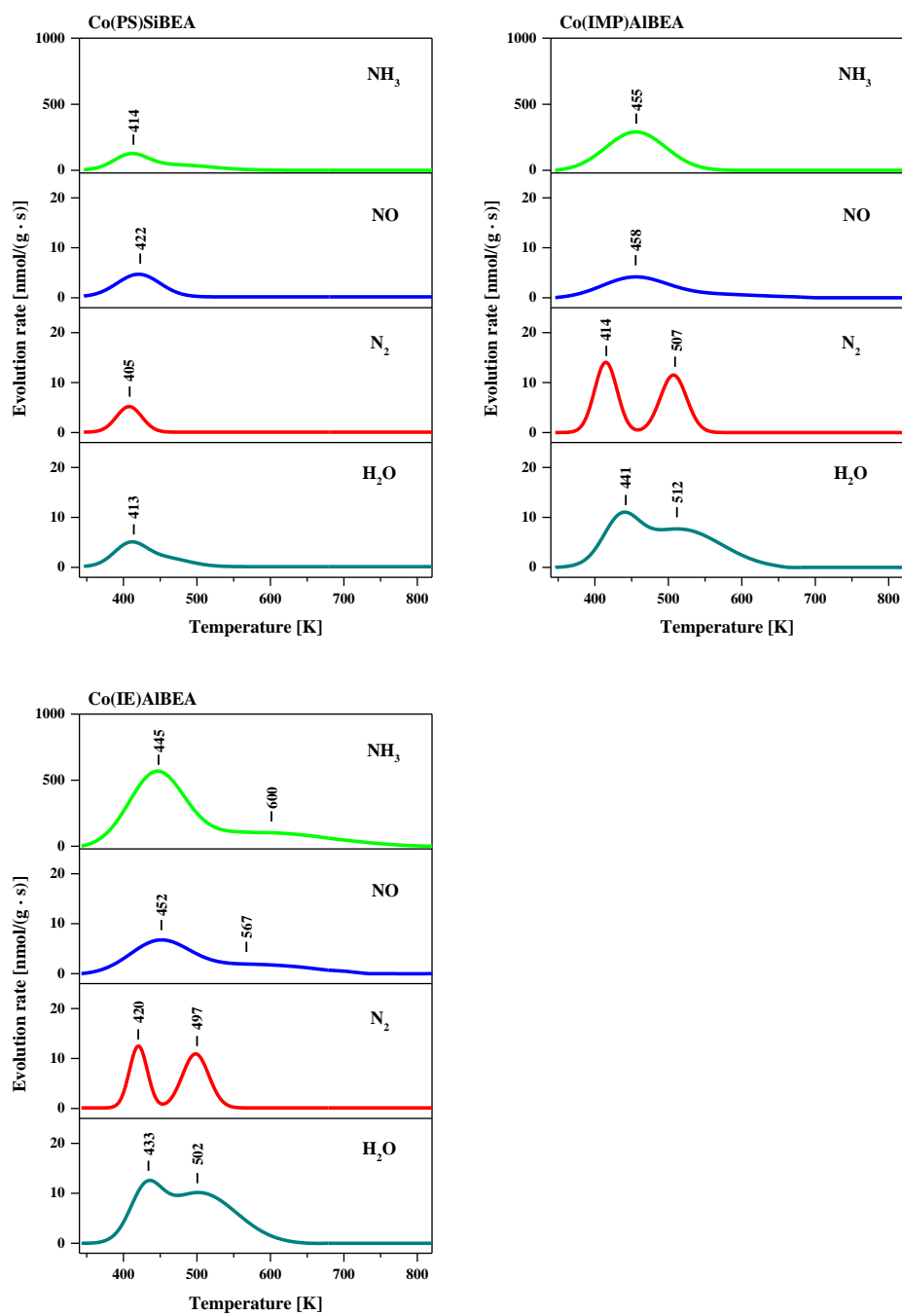
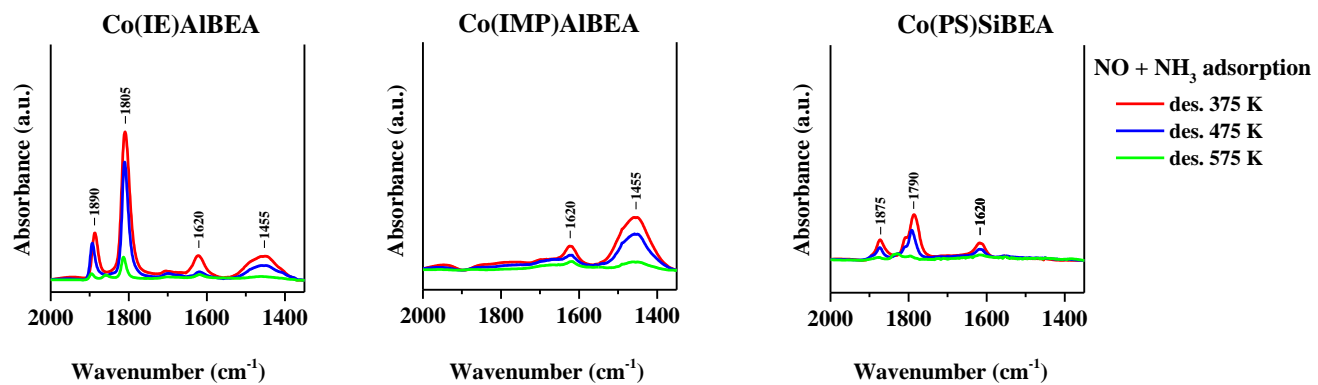


Fig. 8

**Fig. 9**

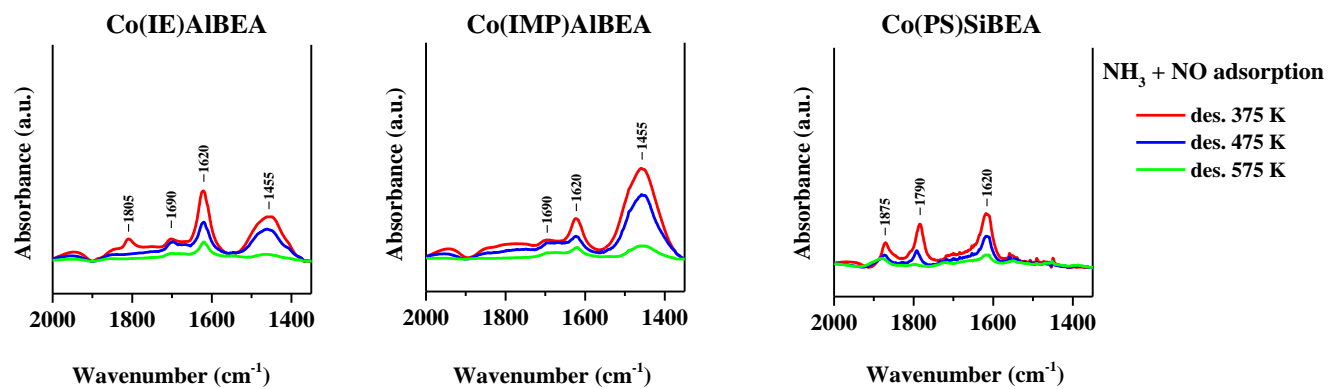


Fig. 10

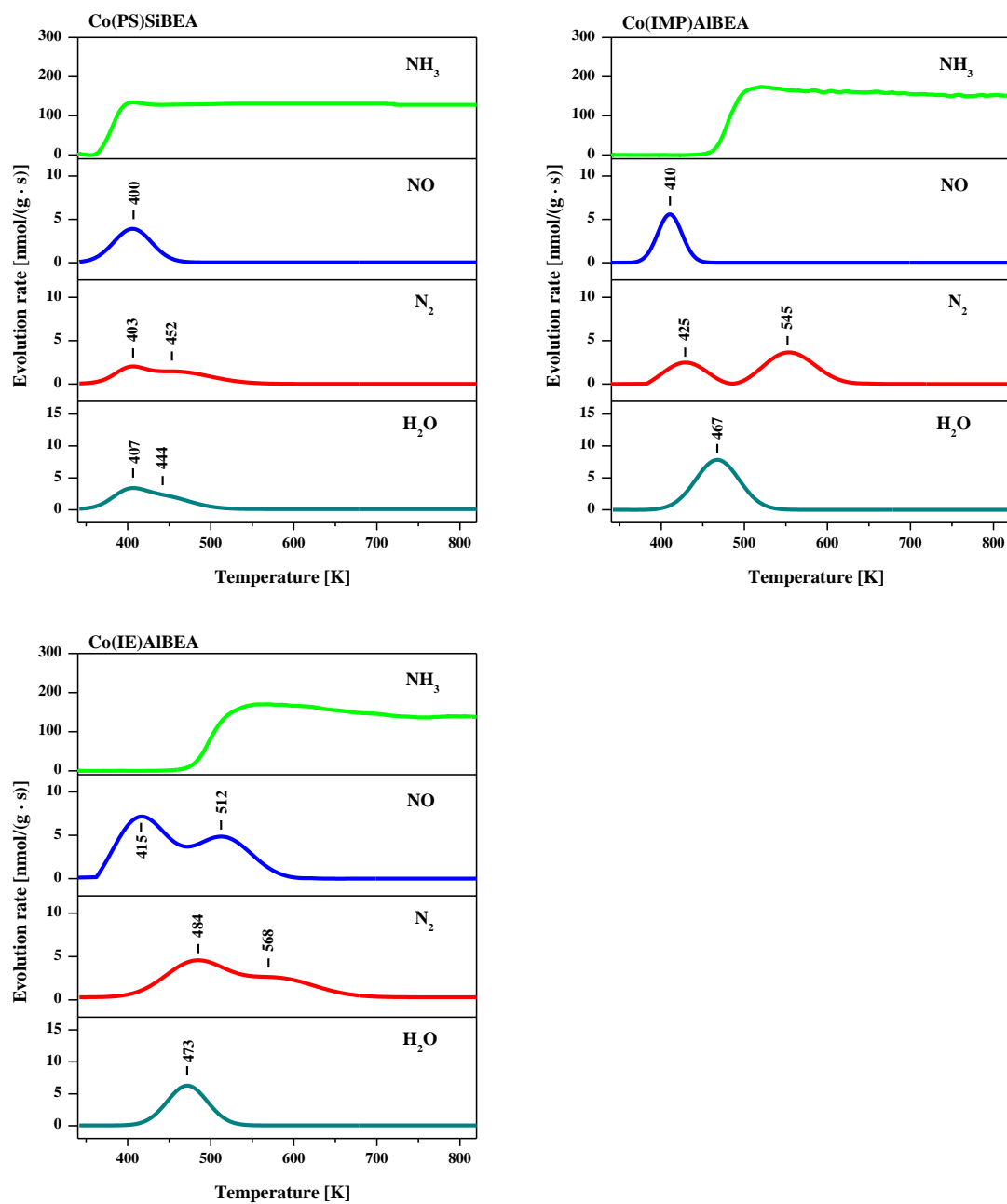


Fig. 11

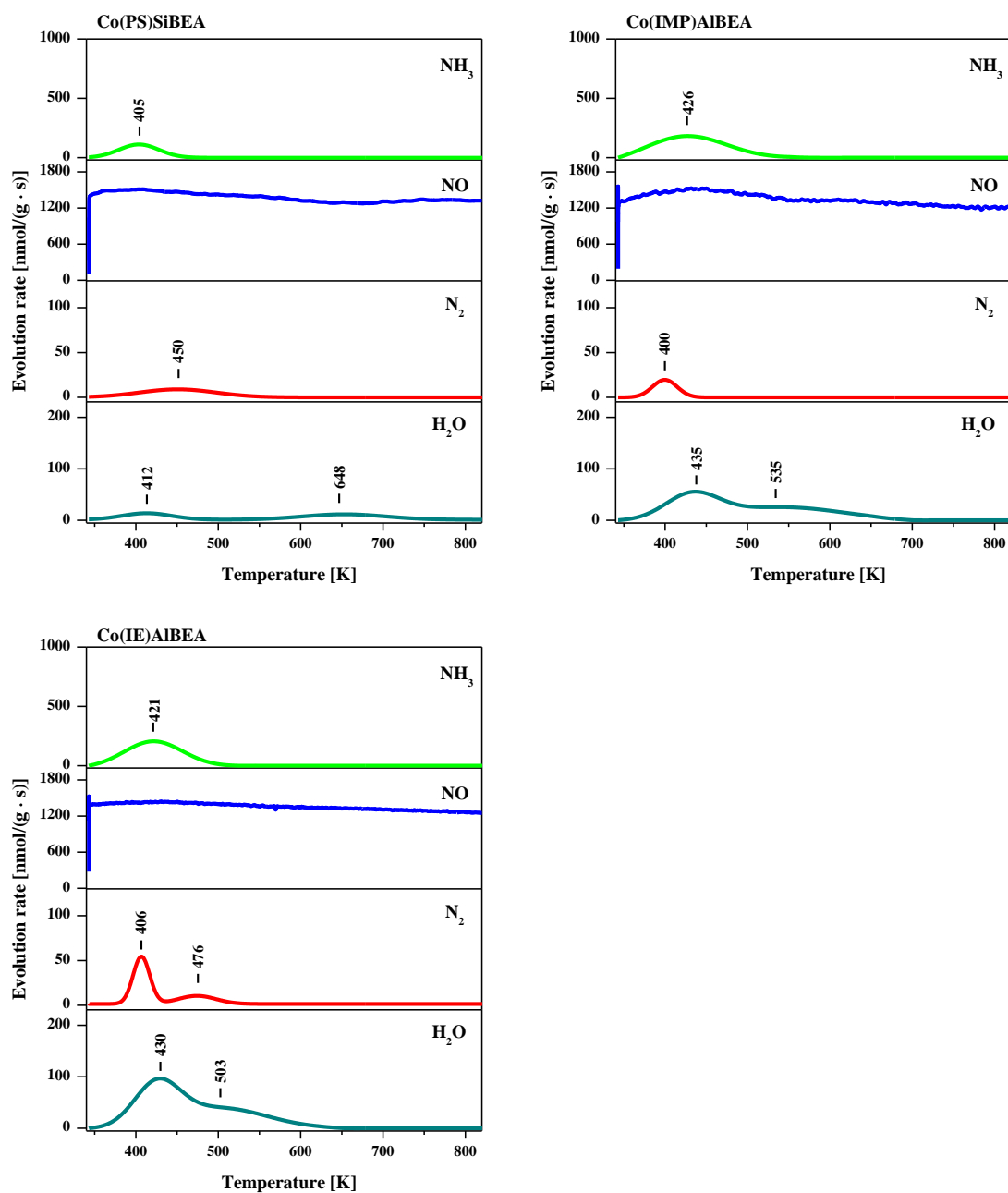


Fig. 12

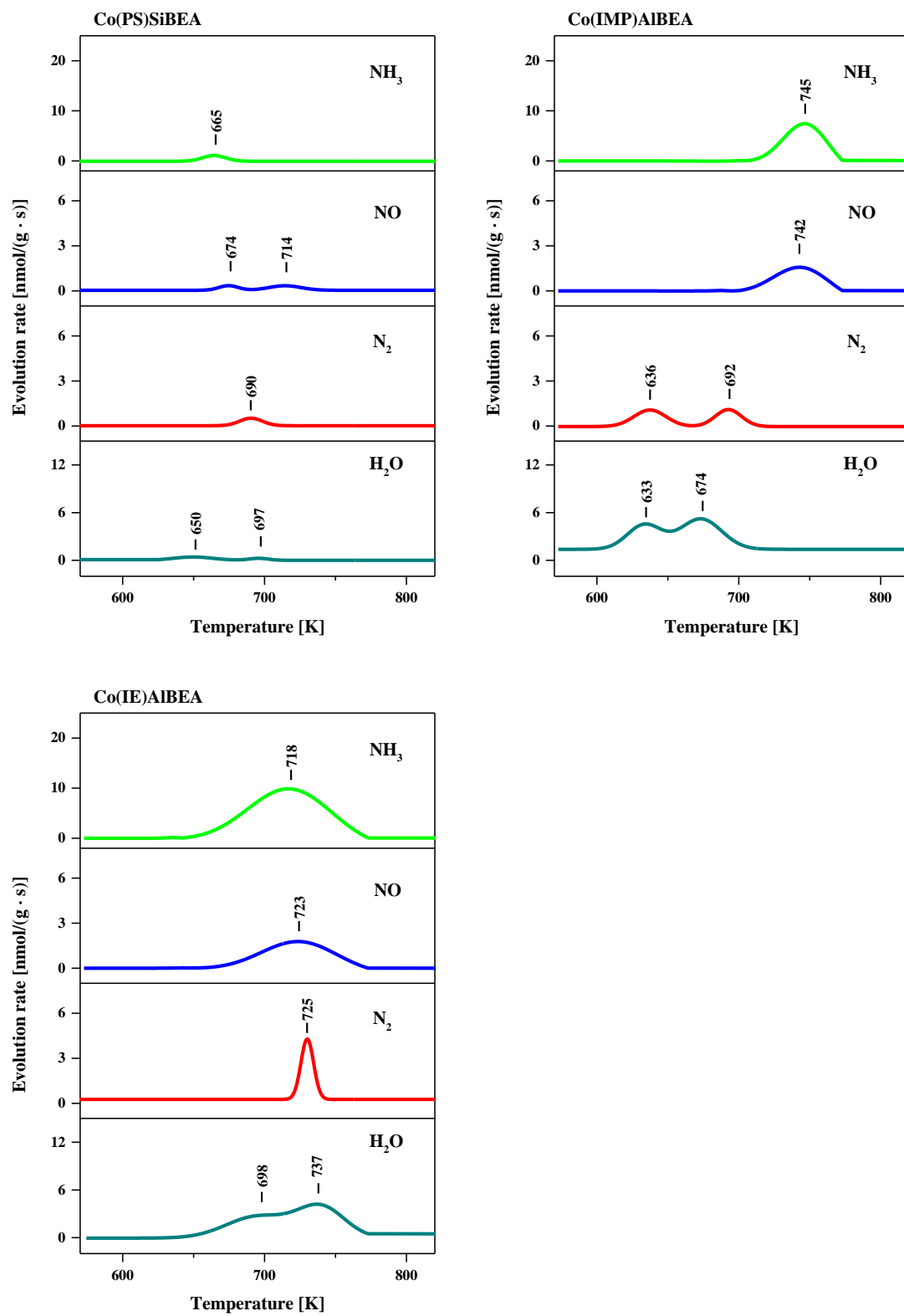


Fig. 13

**Table 1.** The physicochemical properties of the HAIBEA and SiBEA supports and Co(PS)SiBEA, Co(IMP)AIBEA and Co(IE)AIBEA catalysts.

SAMPLE	Modification method	Position of $d_{302}$ XRD peak ( $^{\circ}$ )	BET surface area ( $m^2/g$ )	Micropore volume ( $cm^3/g$ )	Speciation of Co species
HAIBEA	-	22.58	728	0.22	-
SiBEA	-	22.86	789	0.24	-
Co(PS)SiBEA	Two-step postsynthesis	22.79	645	0.21	pseudo-Td Co(II)
Co(IMP)AIBEA	Wet impregnation	22.56	641	0.21	pseudo-Td Co(II), Oh Co(II)
Co(IE)AIBEA	Ion exchange	22.56	649	0.21	pseudo-Td Co(II), Oh Co(II)