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Biosurfactant-mediated one-step synthesis of hydrophobic functional imogolite nanotubes

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Abstract

Imogolites are aluminosilicate nanotubes that attract a particular interest due to their insulating properties, surface reactivity and shape-selective properties in catalysis. Although being available in nature, synthetic protocols have been developed since the 70's to obtain large amounts of this particular mineral without purification processes. In particular, surface functionalized imogolites are currently studied for their application as dispersant in hydrophobic matrixes like polymers. With respect to known synthetic processes, natural imogolite is produced by using either milli- or deci-molar amounts of aluminium salts in solution while surface functionalized imogolites are systematically prepared in two steps: imogolite synthesis followed by surface functionalization mainly using phosphonic acids. Here, we propose a new one-step synthesis involving a cheap, widely-used, non-toxic surfactant, glycerol α-monolaurate (MG)

and centimolar aluminium concentrations. In the presence of Al^{3+} cations, MG forms an emulsion, and subsequent addition of tetraethoxysilane leads to the formation of imogolite nanotubes within 3 days at neutral pH under hydrothermal conditions. An interesting, new, lamellar long-range order of the nanotubes can then be achieved after 10 days. Under similar conditions, blank tests performed on MG-free systems result in the formation of proto-imogolite. Based on the combination of highly selective ^{29}Si, ^{27}Al and ^{1}H solid-state NMR experiments and additional synthesis using lauric acid (LA) and glycerol, a mechanism of formation is proposed based on MG dissociation and LA binding on the outside surface of the nanotubes. As a result, our process provides imogolite with unprecedented hydrophobic surface properties that offer new perspectives for these materials as nanoscale fillers in, e.g., polymeric media.

Introduction

Synthetic metal oxide Single-Walled NanoTubes (SWNTs) are expected to possess a number of interesting and unique properties (surface reactivity, mechanical strength, tailored length, insulating properties, etc.) and have generated a great deal of interest and potential use in a variety of applications such as molecular separation and storage, chemical sensing, drug delivery applications, photonics, biotechnology, catalysis, composite materials and anion/cation retention from water.[1] Although much less popular than the conventional carbon nanotubes, they represent attractive materials due to their potentially wide range of tunable compositions and properties accessible by low-temperature liquid-phase chemistry. An important goal in this area is to achieve a precise size control at the nanoscale over the three dimensions to get access to unique properties such as tunable band gaps, ballistic transport of charge/heat/mass, and quantum confinement phenomena. [2]

Although the variety of such SWNT is not yet extensive, an attractive aspect of these materials is that they can be synthesized by liquid-phase synthesis under moderate conditions, often hydrothermal or solvothermal. An interesting member of this family is imogolite,[3] a natural aluminosilicate-based SWNT, whose synthesis has attracted substantial interest in recent years.[4,5] The cylindrical wall of the imogolite nanotubes can be visualized as a rolled-up sheet of gibbsite (aluminum hydroxide), with isolated silanol (Si-OH) groups linked to the inner surface of the nanotube wall. The nanotube wall is structurally ordered and is composed of an aluminum octahedron (AlO_6) with three oxygen atoms shared by silicon (SiO_4) tetrahedra, while the other three oxygen atoms are shared with other Al octahedra in the nanotube wall. The nanotubes have an inner diameter of about 1.0 nm, while the outer diameter of the nanotubes varies with relative composition of Si atoms in the structure. The length of the tubes is in the 400 to 1000 nm range. Imogolite empirical formula was found to be $(OH)_3Al_2O_3$-SiOH, with a [Al]/[Si] = 2. These para-crystalline hydrous silicate polymers are most commonly associated with weathering of non-crystalline volcanic materials, which have also been identified in pumice deposits, coating of primary particles, stream sediments, and soil derived from igneous and sedimentary rock sandstone. Taken together,[3,5]

previous reports dedicated to the synthesis, structure and properties of imogolite lead to the following statements: i) synthetic imogolite always appears as SWNT with monodispersed diameter size; ii) imogolite formation is generally preceded and/or accompanied by that of proto-imogolite, whose XRD and NMR signatures are undistinguishable from imogolite but whose morphology is not well-defined (particles, platelets); TEM and FT-IR are commonly used to make the distinction between imogolite and proto-imogolite; c) boehmite and amorphous aluminosilicates are common side-products of the synthetic procedures; d) even if the theoretical Al/Si molar or atomic ratio is equal to 2, the experimental ratio may vary within a 15% range, depending on the synthesis method; e) nanotube packing occurs in a monoclinic space group rather than in a hexagonal one, as one would expect; f) imogolite outer surface is hydrophilic and its isoelectric point is close to pH= 9.

Preparation of the synthetic imogolite generally requires rigorous experimental conditions.[4,5] The sol-gel route employing an aluminium salt (mostly aluminium perchlorate or aluminium chloride) in presence of tetraethoxysilane (TEOS) in millimolar concentrations in water at T< 100°C is the most explored pathway. TEOS is preferred to sodium silicate for its slower condensation kinetics while perchlorates are preferred to chlorides due to their lower complexing ability, both reducing the uncontrolled homo-condensation towards amorphous aluminium hydroxides and silica. The crystallization procedure is optimized at temperatures slightly below 100 ºC. Below 90°C, the formation rate strongly decreases while above 100 ºC boehmite rods or gibbsite platelets are formed. Synthesis of pure imogolite from more concentrated solution is also possible but with usually longer reactions times, going from 10 to 60 days to obtain long-range ordered nanotube packing.[5c,5d] The functionalization of the outer surface of imogolites constitutes a new field of research due to the interest of dispersing imogolite nanotubes in hydrophobic media. Attempts to use organosilanes is not as efficient as expected and phosphonic acids were used instead. Using this procedure, poly(methylmethacrylate) and octadecyl groups were recently grafted on the surface of freshly synthesized imogolite nanotubes providing hydrophobic properties[5h,i]. So far, no one-pot synthesis of surface modified imogolite synthesis has been reported, to the best of our knowledge.

In the present work, we have investigated, for the first time to our knowledge, an alternative hybrid organic/inorganic strategy to control the formation and organization of imogolite SWNTs. Indeed, the use of organic-templating strategies was already explored to obtain metal oxide nanotubes but all reported structures are multiwalled and are several tens of nanometers in outer diameter.[6] Our organic additive, glycerol α-monolaurate (MG) was selected for two main reasons: 1) monoacylglycerol derivatives (fatty acid chain covalently bonded to a glycerol head group through an ester linkage) are cheap, non toxic, commercial emulsifiers derived from both natural (animal, vegetable) and synthetic routes and routinely used in food products (bakery products, beverages, ice cream, chewing gum, shortening, whipped toppings, margarine, and confections)[7-8]; 2) both glycerol and carboxylic acids can complex Al^{3+} cations in solution and in the solid state,[9] potentially influencing morphological changes and/or process kinetics. In line with our concern to establish a low-environmental impact synthesis, we kept conditions as simple as possible: i) use of chlorides instead of more toxic perchlorates; ii) synthesis at quasi-neutral pH and limited pH changes; iii) centimolar concentration range ($[Al^{3+}] > 0.05$ mol·L^{-1}). Use of TEOS as silica precursor and of hydrothermal (T= 98°C) conditions are the only factors making our approach similar to previous works.

Through this approach, it is possible to obtain functionalized imogolite SWNTs in less than two weeks under neutral pH conditions. At the moment, boehmite constitutes the main impurity but preliminary results show that an optimization of the synthesis should drastically reduce its overall content. In addition, we show that this process yields to surface-modified imogolite nanotubes with hydrophobic properties that make them promising candidates as nanoscale anisotropic fillers in polymer-based composite materials. Finally, surface modification induced an interesting two-dimensional packing of the nanotubes never reported before, as shown by small angle X-ray diffraction and TEM.

Experimental Section

Preparation of the materials

Glycerol α-monolaurate (MG) was synthesized following a previously-reported procedure[10] on hydrophobic bed in continuous reactor at 80°C giving a 86 % hight purity. The resulting powder was equilibrated in 100 mL distilled water under stirring at room temperature overnight. The desired amount of $AlCl_3.6H_2O$ was then added under stirring at pH ca. 2.5 at 80 °C. The solution turned into an emulsion and stirring was kept for 6 additional hours to homogenize the system. Finally, tetraethylorthosilicate (TEOS) was added under stirring and pH of the system was adjusted to ca. 7 using 1M NaOH solution about one hour after TEOS addition. The mixture was allowed to stir at the same temperature for 3 h after which the sample undergoes a hydrothermal treatment at 98 °C for a period of 72 h or 240 h. A white solid product was recovered by filtration, washed with distilled water and dried at room temperature under atmospheric conditions. Similar experiments were performed with lauric acid (LA) and glycerol, both purchased from Aldrich. Molar ratios are: additive:H_2O:$AlCl_3$:TEOS = 1: 5550: x: 3.5 where additive is MG, LA or glycerol and $1<x<7.5$. Blank experiments refer to those conducted without any additive.

Characterization of the materials

Chemical analysis of the alumina-silica samples was obtained by Energy Dispersive X-ray Spectroscopy (EDAX) using an Oxford X-Max (area: 20 mm^2) detector installed on a Hitachi S3400N Scanning Electron Microscope. Calibration of the instrument was performed on the Ti Kα at 4.509 keV. Powder X-ray diffraction (XRD) study was performed on Bruker D8 Advance diffractometer using Cu-Kα radiation (λ = 1.5418 Å) at 45 kV and 40 mA, 0.05° step size and 60 s step time over a range of 1° to 80°. Small angle X-ray scattering (SAXS) measurements were performed on a pinhole type Rigaku Nanoviewer instrument using Cu-Kα radiation equipped with a CCD detector over a range of 1 to 5 degrees (2θ) respectively. The applied voltage and filament current were 40 kV and 50 mA respectively. The finely powdered samples were sealed in metal cells with polyethylene terephthalate films and kept in a temperature controlled sample chamber for measurements. Bragg spacing is calculated by $d = 2\pi/q$, where q is the magnitude of the scattering vector. Transmission electron microscopy (TEM) images were

acquired with a FEI Tecnai 120 Twin microscope operating at 120 kV and equipped with a high resolution Gatan Orius CCD 4k x 4k numeric camera. Sample holder is equipped with an alpha-tilt (+70°/-70°) goniometer to observe the sample at different tilt angles and avoid 2D projections artefacts. Adsorption/desorption isotherms were recorded on a Belsorp-Max Instrument (BEL-Japan Inc.) using nitrogen as probe gas.

Solid-State Nuclear Magnetic Resonance (NMR) experiments were performed on a Bruker AVANCE III 300 (7.05 T) spectrometer using a wide-bore superconducting magnet. ^{1}H and ^{27}Al Magic Angle Spinning (MAS) NMR spectra were recorded with a 4 mm MAS probe at a spinning frequency of 14 kHz. For the ^{1}H and ^{27}Al spectra, 32 and 4200 transients were collected with $\pi/2$ and $\pi/12$ pulses of 6.5 and 0.95 μs and at repetition times of 10 s and 500 ms respectively. The chemical shifts were referenced to TMS and $Al(NO_3)_3$ at 0 ppm respectively for ^{1}H and ^{27}Al. For the $^{13}C\{^{1}H\}$ Cross-Polarization (CP), one pulse ^{13}C MAS, ^{29}Si CP$\{^{1}H\}$ and one pulse ^{29}Si MAS NMR studies, 7 mm zirconia rotors spinning at 5 kHz were employed. A total of 4096 transients were collected with a proton $\pi/2$ pulse of 6.5 μs and recycle delays of 1 s for the CP spectra of both nuclei. A cross-polarization contact time of 3 ms was applied for all the experiments. For ^{13}C and ^{29}Si MAS NMR experiments 4096 and 1024 transients were collected with a $\pi/4$ pulse of 2.65 μs and at a relaxation delay of 10 and 60 s respectively. The chemical shift values were calibrated with reference to tetramethylsilane (TMS) at 0 ppm. The ^{1}H-^{29}Si-^{1}H double CP were performed on a Bruker AVANCE III 300 (7.05 T) spectrometer. Details for double CP experiment are given in ref 11. The ^{27}Al-^{1}H CP experiments were performed on a Bruker AVANCE III 700 (16.45 T) spectrometer using a 3.2 mm rotor spinning at 23 kHz.

Results and Discussion

Nanotube preparation

Imogolite was obtained by modification of earlier procedures,[4,5] involving the mixing of acidic solutions of an aluminium salt with tetraethoxysilane, followed by neutralization and hydrothermal treatment. Main variations compared to the literature were the selection of aluminium choride, the use of centimolar precursor concentration and the addition of glycerol α-monolaurate (MG).

During washing procedures, we observe a different behavior for samples prepared with or without biosurfactant: whereas the purely inorganic preparation was easily dispersed in water, as expected for hydrophilic aluminosilicate nanotubes, the sample prepared with MG remained located at the air-water interface, suggesting that it exhibits an hydrophobic character. To confirm this hypothesis, 1-nonanol was added to the water solution. As shown in Figure 1, after phase separation, the MG-free product was still in the bottom aqueous phase whereas the surfactant-templated sample was located in the top organic layer. This unexpected and unprecedented behavior of imogolite urged us to investigate in details the resulting materials first from an inorganic and organic point of view and then at the hybrid interface. Additional experiments led us to propose a reaction mechanism that is compatible with the one-step formation of hydrophobic imogolite nanotubes

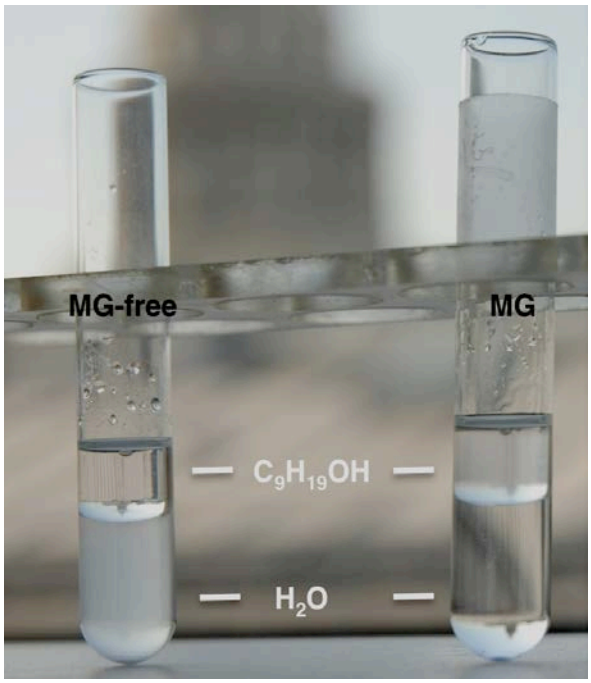
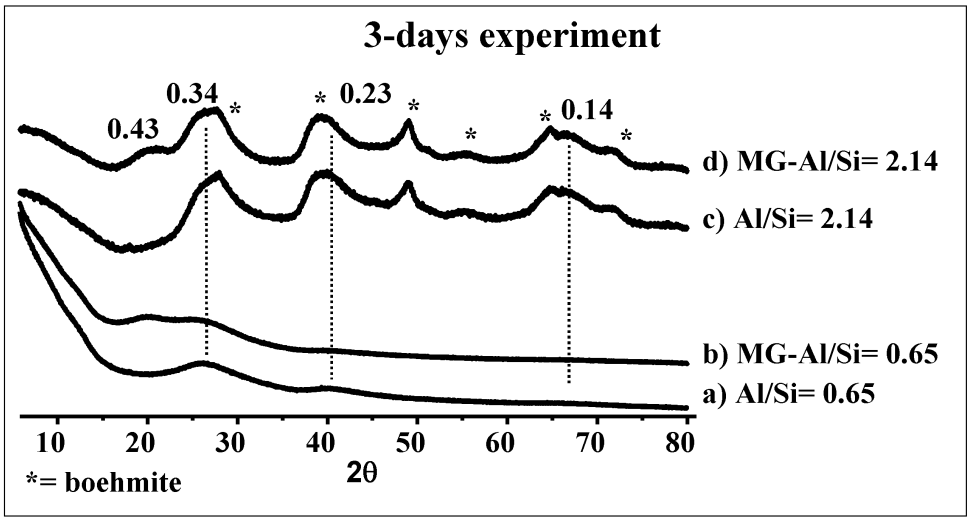
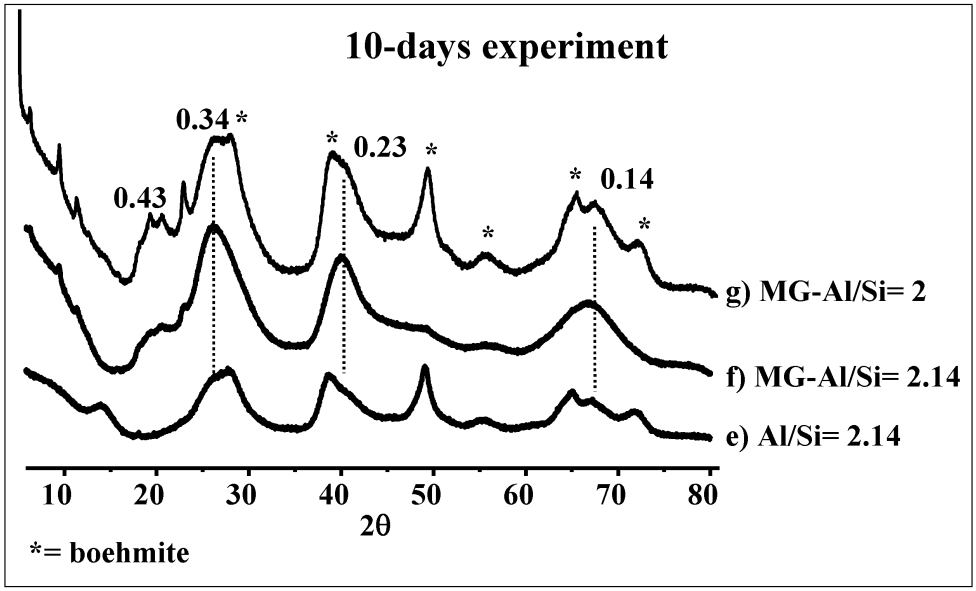


Figure 1. Behaviour of aluminosilicate powders in a water/nonanol biphasic system: MG-free sample is easily dispersed in the bottom aqueous media (left) whereas the MG-templated sample localizes in the top organic layer.

Characterization of the inorganic structure

3-days aging. XRD patterns of MG and MG-free materials at increasing Al/Si molar ratios after 3 and 10 days aging are given in Figure 2. At low Al/Si ratio (0.65) and after 3 days of aging at 98°C, the main broad diffraction peaks for MG (b) and MG-free (a) samples are at d = 0.43 nm, 0.34 nm and 0.23 nm (Figure 2a,b). These values can be related to the silicate/gibbsite layers belonging to the proto-imogolite structure, considered to be the precursor of imogolite.[3-5] At similar aging time but higher Al^{+3} concentration values (Al/Si= 2.14, Figure 2c,d), additional peaks are observed: the d= 0.14 nm distance corresponds to the (004) diffraction plane of imogolite while reflections at d= 0.32, 0.24, 0.18, 0.16 and 0.13 nm indicate the additional presence of boehmite, our major impurity.[12]





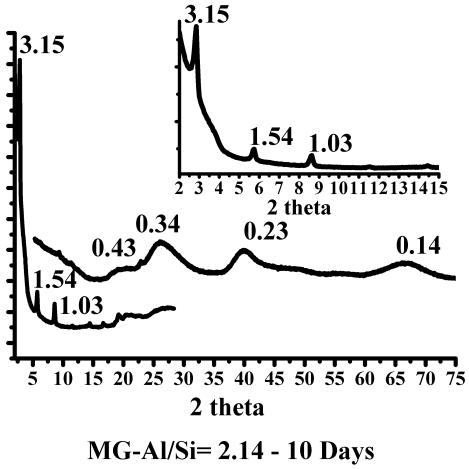


Figure 2. Powder X-Ray patterns of MG (b,d,f,g) and MG-free (a,c,e) alumina-silica samples obtained after 3 days (a-d) and 10 days (e-g) aging time: (a,b) Al/Si = 0.65; (c,f) Al/Si = 2.14; (g) Al/Si = 2. Small and wide angle XRD of MG-Al/Si= 2.14 (10

days) sample is given in the inset. The given *d*-spacing values, in nm, and dotted lines refer to imogolite while asterisks indicate the presence of boehmite.

The evolution from amorphous alumino-silicates to imogolite as a function of Al/Si molar ratio (3 days aging time) can also be followed by ^{29}Si CP and ^{27}Al solid state MAS NMR experiments in Figure 3. At low Al/Si values (2a,b), a broad peak centered at -93 ppm characterizes an amorphous alumino-silicate network made of Al sites, ranging from $Q^3(4Al)$ to $Q^4(4Al)$.[13] At increasing Al^{3+} values (3c,d), a characteristic peak at -78 ppm appears and, at Al/Si= 2.14 (Figure 2d), it constitutes more than 30% of the overall silicon spectrum. This peak, in good agreement with previous studies,[13a] is extremely rare in aluminosilicates[13b] and is commonly assigned to $(AlO)_3SiOH$ species in the imogolite environment, that is silicon tetrahedron connected to a gibbsite layer.

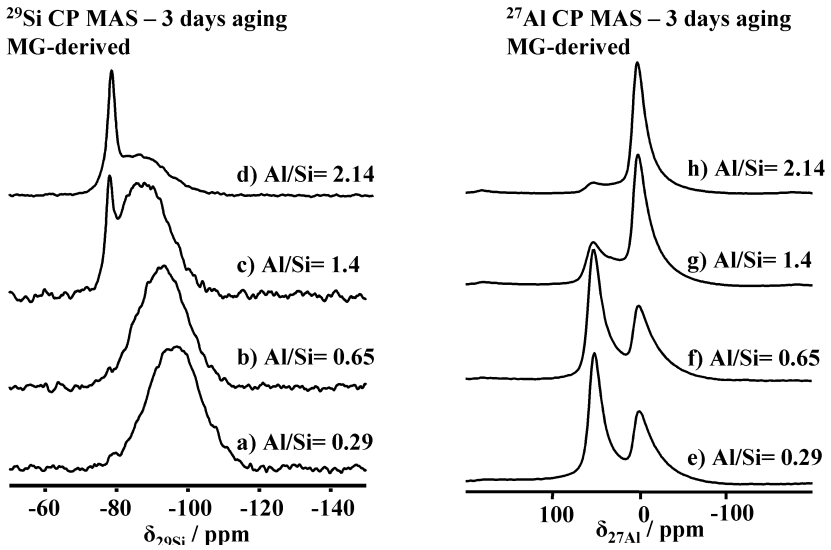


Figure 3. (a-d) ^{29}Si CP and (e-h) ^{27}Al MAS solid state NMR experiments performed on MG-derived imogolite samples after 3 days aging and at different Al/Si ratios.

A similar behavior occurs for the ^{27}Al spectra evolution (Figure 3e-h), where, at low aluminum amount (Figure 3e-g), a mixture of octahedral (δ= 2.5 ± 0.5 ppm) and tetrahedral (δ= 55 ± 0.5 ppm) sites characterizes the material. At Al/Si values close to 2 (3h), integration of the ^{27}Al spectrum shows that

more than 90% of the material is composed of octahedral aluminum sites, as expected for imogolite, allophanes and boehmite compounds.[14] This is characteristic of the octahedral environment of aluminium polymerized in a gibbsite-like environment. A similar trend can be observed in MG-free samples (data not shown), showing that evolution from amorphous alumino-silicate to proto-imogolite does not actually depend on MG.

As mentioned earlier, the only way to distinguish between imogolite and proto-imogolite is either TEM or FT-IR spectroscopy. TEM images (Figure 4a,b) show that MG-derived powder is composed of a bundle of fibers whose characteristic dimensions, length L > 50 nm and diameter d = 2.5 nm, confirm them as imogolite nanotubes. The SAED pattern extracted from Figure 4a is given in Figure 4c. The d-spacings calculated from the analysis of the diffraction rings numbered from (1) to (4) provide the following values: $d_{(1)}$= 0.32 nm, $d_{(2)}$= 0.21 nm, $d_{(3)}$= 0.18 nm, $d_{(4)}$= 0.14 nm. $d_{(1)}$, $d_{(2)}$ and $d_{(4)}$ can be attributed to the (071), (004) and (006) (hkl) diffraction planes observed for imgolite.[5a] On the contrary, the $d_{(3)}$= 0.18 nm value corresponds to the (200) plane of the boehmite impurity, as discussed before.

MG-free samples (Figure 4d) only show agglomerates of particles whose morphology is not fibrous. Therefore, at a Al/Si = 2.14 ratio, the presence of MG favors the formation of imogolite nanotubes within a 3 days aging period. However, the absence of significant peaks in the low-angle domain of the XRD pattern suggests that the nanotubes are completely disordered with no bundle organization at this stage.

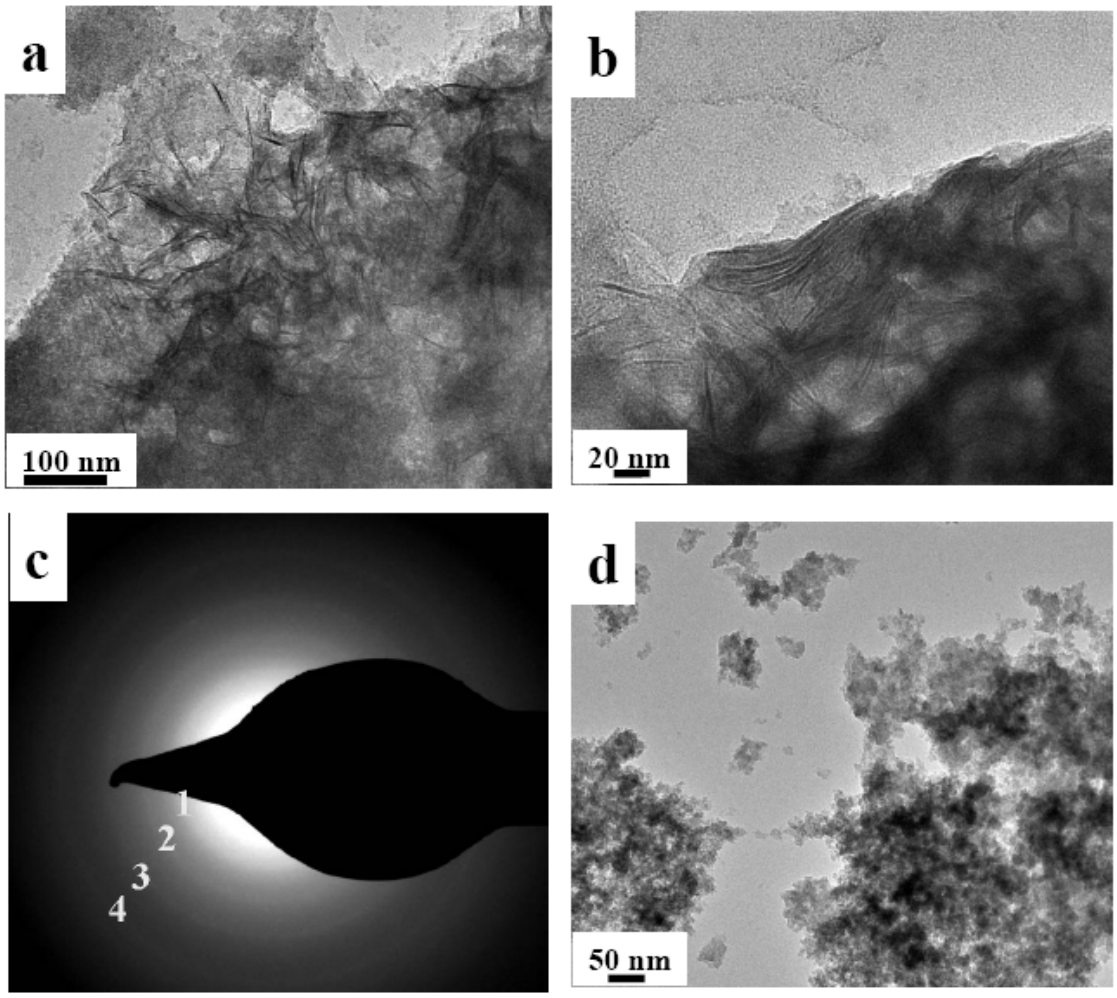


Figure 4. (a, b) Transmission electron microscopy images obtained for MG-derived imogolite (tubular morphology); c) SAED pattern obtained on image displayed in (a); (d) MG-free proto-imogolite (particle morphology). The Al/Si ratio is 2.14 and aging time is 3 days for all samples.

10-days aging. In a previous report, it was demonstrated that longer aging times favors three-dimensional nanotube packing in Al-rich samples.[5c] In our system, XRD patterns of the 10-days aged MG and MG-free samples for Al/Si = 2.14, given in Figure 2(e,f), show the striking effect of MG on the imogolite nanotube stacking. For MG-free samples, no significant change occurs in the XRD pattern with respect to the 3-days aged sample material, both at low (diffraction pattern not shown) and wide (Figure 2e) angles. On the contrary, in the presence of MG, more defined diffraction peaks appear in the low-angle region. In particular, as shown in Figure 2, a series of sharp low-angle diffraction peaks characterizes the MG-Al/Si= 2.14 – 10 days sample. The calculated d-values indicate the presence of a

clear lamellar stacking. This result is highly unexpected for imogolite for two reasons: 1) imogolite bundles are known to pack in a monoclinic crystalline system[5]; 2) imogolite peaks at low angle are much broader than what is currently found in MG-derived system. The absence of the characteristic imogolite low-angle pattern and the presence of a lamellar pattern seem to be in contrast with the presence of imogolite nanotubes but rather support the presence of a layered aluminosilicate or a lamellar mesostructure material.

Before proceeding with a deeper analysis of the 10 days samples, we observe that the amount of residual boehmite is particularly low in this sample (Figure 2f), showing that our approach may lead to rather pure imogolite provided that synthetic conditions are optimized through a more systematic study. We did not explore this aspect any further in this work.

Very interestingly, nitrogen adsorption/desorption isotherms are of particular use to exclude the presence of a mesostructure lamellar solid. As shown in Figure 1 in supplementary information, the MG-Al/Si= 2.14 – 10 days sample displays combined micro and mesoporosity, as indicated, respectively, by the nitrogen uptake below P/P_0= 0.2 and above P/P_0= 0.5, where a hysteresis loop occurs. In particular, after calcination at 350°C, the amount of microporosity increases while the hysteresis loop remains intact. The presence of a hysteresis loop for the MG-containing and calcined sample excludes the presence of a lamellar structure. In fact, mesostructured, surfactant-containing, materials do not generally display any porosity and, most importantly, the lamellar framework collapses after calcination, which is not the case here. Any additional hypothesis of a layered aluminosilicate is also excluded by the sharp ^{29}Si NMR peak at δ= -78 ppm, which cannot be attributed to any flat aluminosilicate surface.[13b]

Transmission electron microscopy images obtained for the MG and MG-free samples after 10 days are shown in Figure 5 (a-d). Evolution with time of the MG-free sample goes from morphologically undefined proto-imogolite (3 days aging, Figure 4d) to disordered imogolite nanotubes (10 days aging, Figure 5d). MG-containing samples, on the contrary, show imogolite nanotubes already at 3 days aging

(Figure 4a,b) and an interesting planar lamellar packing after 10-days aging, as shown on several sections of the samples in Figure 5a-c). Throughout the sample, a large number of apparent sheets are actually composed of individual nanotubes arranged together in a parallel fashion. No three-dimensional order seems to exist between the nanotube sheets but superposition occurs rather in a non-organized way. Further confirmation of the existence of the nanotubes can be found in Figure 5e-f, where we show an image of the same section, containing both individual nanotubes and sheets, before and after a 20° tilting of the sample holder. This manipulation allows to discriminate between a true, uni-dimensional, nanotube and a sideviewed, two-dimensional, platelet, which could actually be confused due to the projection of a 3D object on a two-dimensional surface. The black arrows point towards individual nanotubes before and after tilting and, as one can see, the indicated one does not change in lateral size upon tilting. This is of course expected for a nanotube but not for a platelet. TEM observations perfectly corroborate low-angle peaks in XRD shown in Figure 2 and attributed now to a long-range order of the imogolite nanotubes within a single sheet.

This specific arrangement was never reported for imogolite and, at the moment, we can qualitatively attribute it to the presence of the MG on the surface of the tubes, as shown hereafter. Non-modified imogolite bundles are known to pack in a monoclinic crystal structure[5] but, unfortunately, we could not find any report describing the packing of surface-modified imogolite. Park and coll.[5i] have reported the two-dimensional alignment of octadecylphosphonic acid modified imogolite but that was only possible by using post-synthesis modification of fresh imogolite combined with a Langmuir-Blodgett technique. Nevertheless, despite the lack of experimental details, they briefly comment on the different packing of pure imogolite in bulk with modified imogolite nanotubes aligned on the surface as observed by XRD.

The faster formation of functionalized nanotubes in 2D lamellar packing in presence of MG is highly reproducible and also occurs at the stoichiometric Al/Si= 2 ratio. In this situation, the final material after 10 days aging shows a good quality XRD pattern (Figure 2g), typical for imogolite at wide angles, and of a lamellar phase, as described above, at low angles.

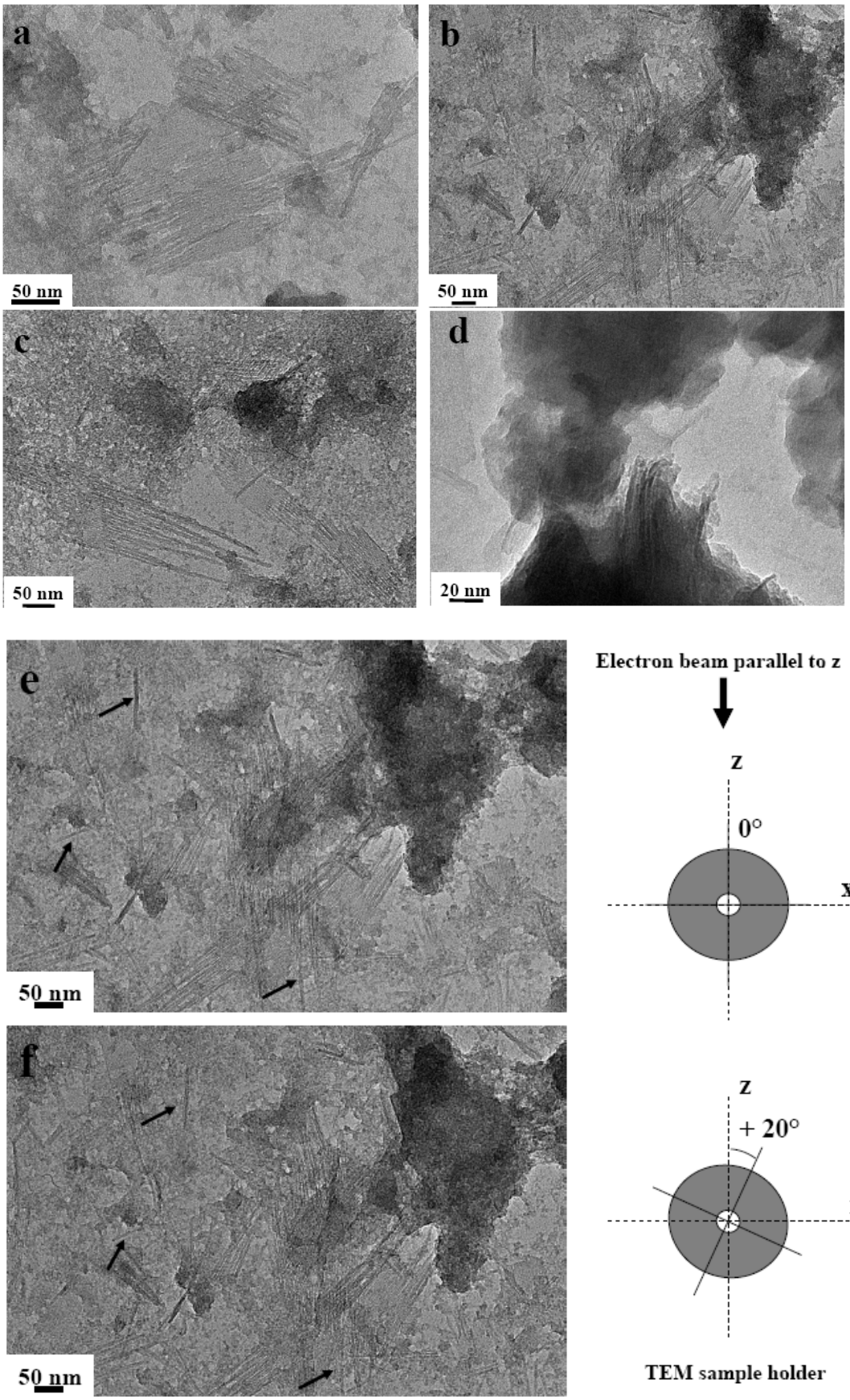

a
50 nm
b
50 nm
c
50 nm
d
20 nm
e
50 nm
f
50 nm
Electron beam parallel to z
z
0°
x
z
+ 20°
x
TEM sample holder

Figure 5. Transmission electron microscopy images obtained for (a-c) MG and (d) MG-free imogolite samples upon 10 days aging at Al/Si= 2.14. (e-f) Highlight on images observed at two different tilt angles of the sample holder with respect to the electron beam z-direction: 0° for (e) and +20° for (f). Arrows in both figure put in evidence the individual nanotubes before and after tilting showing no change in their thickness.

As a complementary source of information, the Al/Si ratio of the samples was measured by SEM/EDAX and averaged on different sections of the material. The experimental values were very similar to the initial stoichiometric content, i.e. 2.1 ± 0.1 when MG was present. In contrast, for an initial Al excess (Al/Si = 2.14), MG-containing samples showed lower value (1.7 ± 0.1) whereas MG-free powders contain a more important Al excess (2.5 ± 0.1). These discrepancies (about 20%) could probably be justified by the presence of impurities in the final sample.

Characterization of the organic component

XRD and TEM have so far shown the important role of MG in favouring the formation of imogolite over proto-imogolite phase. In these conditions, TGA analyses (not shown) also revealed that the organic fraction constitutes between 20% and 30 wt% of the final materials. To understand better this process, we used a combination of multinuclear solid-state NMR experiments. Figure 6 shows the ^{13}C CP solid-state NMR spectra of MG-derived imogolite samples after 3 (Figure 6b) and 10 (Figure 6c) days aging. In a classical imogolite synthesis, one would not expect the presence of any major carbon species; here, both samples show the classical fingerprint of the carbon backbone of MG, as compared to the ^{13}C solution spectrum of this compound (Figure 6a).

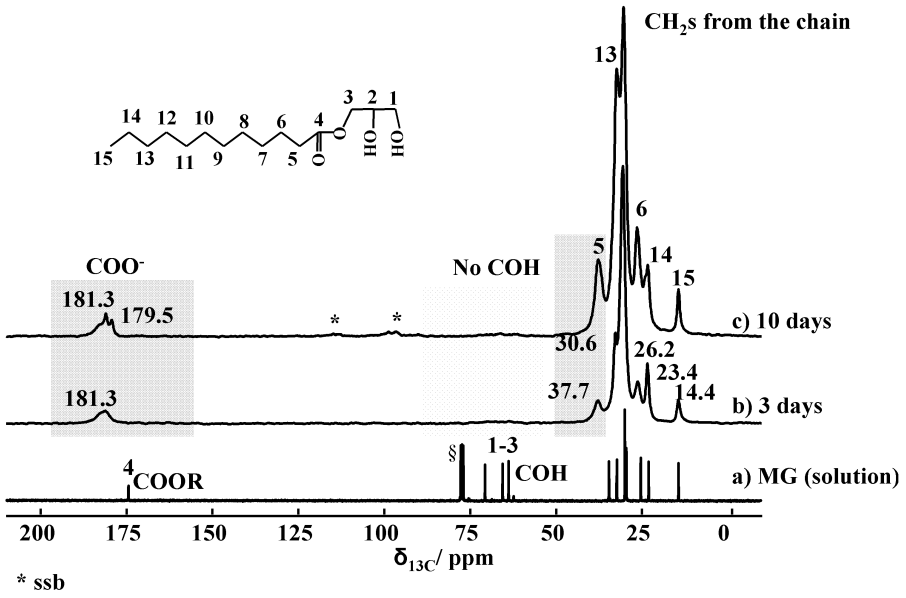


Figure 6. ^{13}C solution NMR spectrum of MG (a) and ^{13}C CP MAS NMR spectra of MG-derived imogolite upon 3 (b) and 10 days (c) aging. Highlighted regions are discussed in the text. § refers to $CDCl_3$. *: Spinning Side Bands

Close comparison between solution and solid-state spectra depicts some important differences:

(i) The glycerol moiety (between 60 and 75 ppm) of MG is lost during the synthetic process, as indicated by the lack of peaks n°1-3 in the solid-state spectra (Figure 6b,c). This is probably due to the hydrolysis of the ester bond between lauric acid and glycerol due to the acidic conditions used at the stage of $AlCl_3$ addition.

(ii) The ester group in MG has a typical chemical shift value of δ = 175 ppm (Figure 6a). The chemical shift of the same group in the imogolite samples is now around δ= 180 ppm, suggesting that the ester has been hydrolyzed into carboxylate groups COO^-. This might result from the neutral pH conditions used for TEOS condensation and/or from complexation with Al^{3+} ions.

(iii) The chemical shift of peak n°5, corresponding to the α-CH_2 with respect to the ester function, shifts from 34 ppm (Figure 6a) to 37 ppm (Figure 6b,c), which may be due to the change in the chemical surrounding occurring on the carbonyl group. For instance, values at 37 ppm were reported for α-CH_2 group of aliphatic esters.[15] Additionally, the relative intensities of peaks n° 5, 13 and 6 are affected at various aging times in imogolite. It was previously reported that adsorption of SDS on alumina provokes

relative intensities variations with increasing SDS, but the authors could not provide a clear explanation.[16] This point is still to be clarified.

Investigation of the hybrid interface

^{13}C CP MAS NMR experiments show that MG undergoes hydrolysis during the material forming process. Glycerol, being water-soluble, is washed away and the remaining fatty acid alone (lauric acid) in its basic form is retained in the final material. Possible conformational changes and the unshielded value δ= 180 ppm for the COO^- group suggest that MG lays on the nanotube surface. This idea is confirmed by ^{29}Si- or ^{27}Al-filtered ^1H NMR solid-state experiments using both $^1H\{^{29}Si\{^1H\}\}$ double cross polarization and $^1H\{^{27}Al\}$ cross polarization. The double cross polarization sequence uses the proton bath to polarize all atoms through an initial CP step at an optimized contact time, t_{CP1}; then, ^{29}Si magnetization is transferred to its nearest proton neighbors by employing a second, variable, contact time, t_{CP2}. The shortest the t_{CP2} value, the closest ^{29}Si-to-proton distance is actually explored.

Figure 7a shows the single-pulse and $^1H\{^{29}Si\{^1H\}\}$double-CP experiments on the MG-derived sample after 3 days aging time. In the single pulse experiment, the methylene $(CH_2)_x$ chain and the CH_3 goups are detected between 0.5 and 2 ppm while H-bonded H_2O is observed at 5 ppm. For double-CP experiments, the $^{29}Si\{^1H\}$ contact time (t_{CP1}) is set to 3 ms while the second contact time (t_{CP2}) varies between 500 μs and 10 ms. These values are optimized to explore both close and large internuclear distances. At all t_{CP2}, only the H-bonded water (5 ppm) is detected while no MG-related signal is observed. The fatty acid is therefore not in close proximity with silicon, which excludes the direct functionalization of silicate layers of imogolite. However if an additional 10 ms delay time is set after cross-polarization, favoring uncontrolled long-range spin-spin interactions, the CH_2 backbone signal is observed. The residual fatty acid is therefore intimately mixed with the imogolite sample and it is not phase separated.

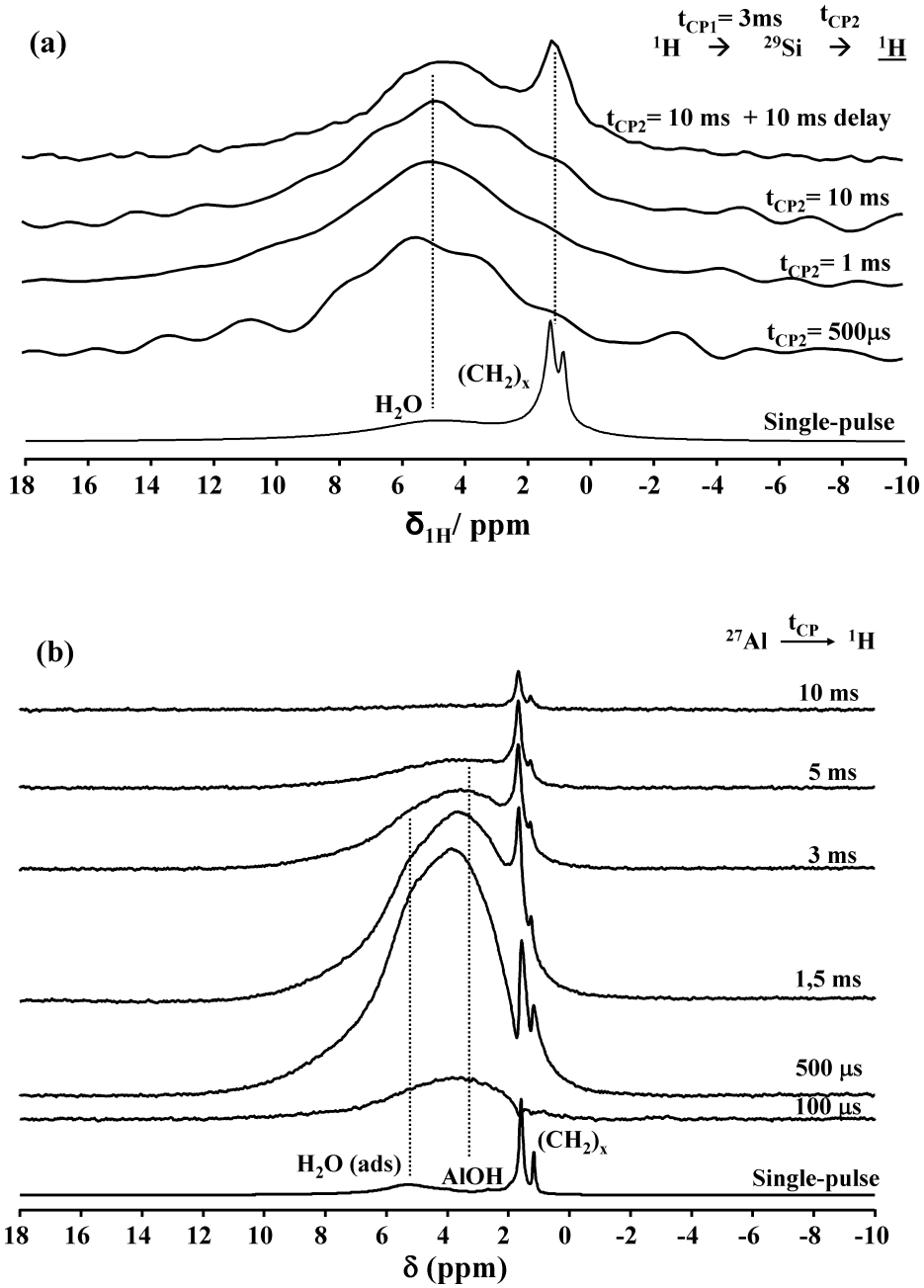


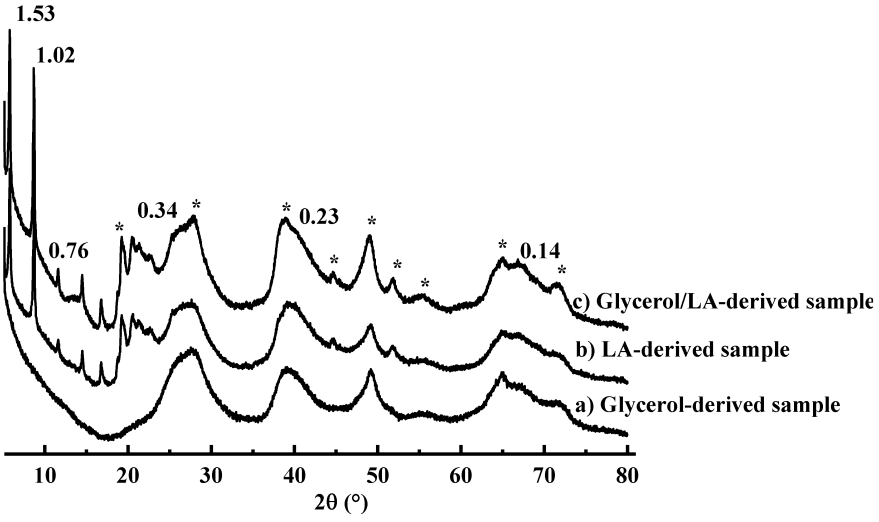
Figure 7. (a) $^1H\{^{29}Si\{^1H\}\}$ double CP MAS NMR (B_0= 7.05 T; ν_{MAS}= 14 kHz) spectra of MG-derived imogolite upon 3 days aging time. t_{CP1}= 3 ms while t_{CP2} varies as indicated next to each spectrum. (b) $^1H\{^{27}Al\}$ CP MAS NMR (B_0= 16.45 T; ν_{MAS}= 23 kHz) spectra of MG-derived imogolite upon 3 days aging time. Contact times are provided next to each spectrum. The corresponding single pulse ^1H MAS NMR spectra, recorded under comparable conditions (B_0 and ν_{MAS}) for experiments shown in (a) and (b), are shown at the bottom of each set of experiments.

As a complementary experiment, the proton signal was also detected after a $^1H\{^{27}Al\}$ cross polarization step. $^1H\{^{27}Al\}$ CP experiments provide a clear-cut answer on the localization of the residual fatty acid. The proton signal related to the fatty acid backbone (Figure 7b) appears at t_{CP2}= 500 μs and its

intensity increases continuously up to 10 ms. Additionally, this experiment partially filters the SiOH signal with respect to AlOH, δ= 3.6 ppm. The interpretation of these experiments is quite straightforward: the fatty acid is very close to the Al^{3+} centre, undoubtedly chelated by the carboxylate group (that cannot be detected directly via ^{1}H NMR experiments). These experiments are consistent with the specific nanotube morphology and structure parameters of imogolite, whose silicate layer is localized inside the tube while the gibbsite sheet is outside. Chemical and steric considerations help understanding that the C_{12} fatty acid preferentially complexes aluminium atoms on the nanotube external surface rather than the inner silicon atoms. Indeed, the inner tube diameter is only 1 nm, thus smaller than the theoretical length (1.6 nm) of the lauric acid molecule in its *all-trans* configuration.[17] All together, these data suggest that laurate residues functionalize the external surface of the nanotubes, providing an hydrophobic character to the whole sample, in agreement with our initial observations (Figure 1).

Insights in the reaction mechanisms.

Previous experiments show that MG hydrolyzes during reaction and its laurate moiety finally ends as complexing agent for Al^{3+} sites. In order to verify the role of glycerol versus laurate in the formation mechanism, we have performed three test experiments with pure lauric acid, pure glycerol and a mixture of lauric acid and glycerol at 10 days aging time.



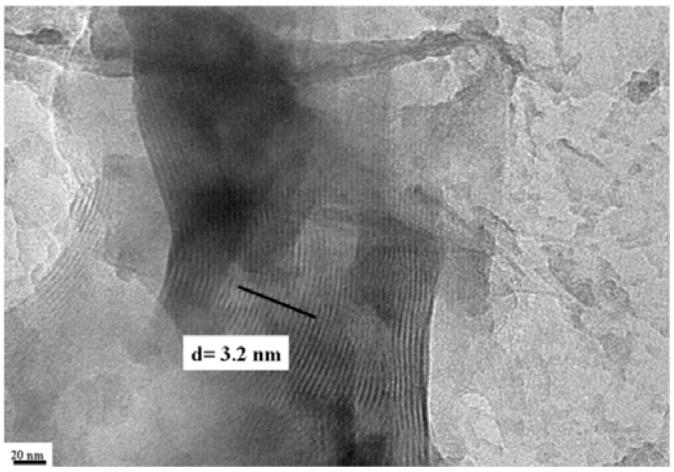


Figure 8. Wide angle X-ray diffraction patterns of (a) glycerol, (b) Lauric acid (LA) and (c) glycerol/LA derived samples obtained after 10 days aging time (asterisks indicate the peaks due to boehmite structure). Values are given in nm and correspond to associated *d*-spacings. (C) TEM image of lauric acid-derived imogolite at 10 days aging time.

Figure 8 shows the wide angle X-ray diffraction patterns obtained for these samples: both samples containing lauric (b,c) acid have the typical pattern of imogolite nanotubes, whose packing is, once again, constituted of a 2D arrangement of nanotubes as shown by the diffractions peaks below 2θ= 15° and where the inter-tube distance, *d*, corresponds to 3.03 nm. On the contrary, no evidence of stacked imogolite nanotubes is obtained when glycerol is used alone (Figure 8a). These data are confirmed by the TEM image of the lauric acid-derived sample in Figure 8, showing a long-range order nanotube packing with an inter-tube distance of d = 3.2 nm, consistent with XRD data. These experiments support previous data on the MG system and indicate that lauric acid plays a key role in the formation of imogolite.

The synthesis of imogolite nanotubes has been reported so far under several conditions.[4,5] In general, good quality imogolite can be obtained at millimolar Al^{3+} concentrations in presence of TEOS as silica source, two factors that contribute to favor hetero- (formation of Al-O-Si bonds) with respect to homo-condensation (formation of Al-O-Al and Si-O-Si bonds). Functionalization of the imogolite outer surface could only be achieved by post-synthesis functionalization with phosphonic acids. In our system, we combine relatively high Al^{3+} concentration probably due to the stabilizing action of the surfactant. In a second stage, when proto-imogolite forms, laurate strongly interacts with gibbsite sheets, thus favoring

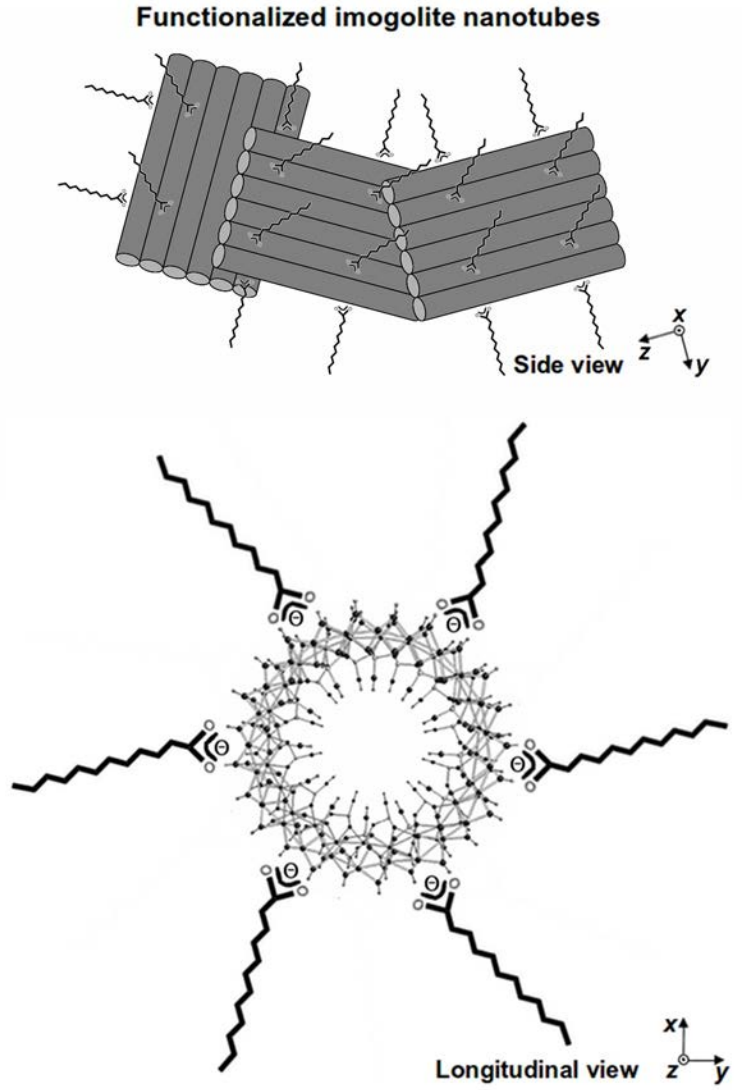
the formation of curved surfaces, already imposed by the silicate tetrahedra.[5a,18] Laurate-functionalized imogolite nanotubes are then formed. Finally, side-by-side interactions between nanotubes are assisted by van der Waals interactions between laurate alkyl chains, that favor nanotube packing. At the moment, even if we do not have a clear explanation of the reason why packing occurs only within a plane and not in the 3rd dimension, providing so-called "raft-like" objects, we can nevertheless formulate some hypothesis.

In the domain of carbon nanotubes, it was shown that packing (hexagonal vs. collier-like) can be influenced by their chemical surrounding.[19,20] In particular, Jeong at al.[20] have shown that single-walled carbon nanotubes can pack in two-dimensional plate-like sheets using a smectic liquid crystal template. In our system, the amount of MG is not enough to justify the formation of a liquid crystal. Nevertheless, in the field of surfactant-templated mesostructured silica and aluminosilica materials, the formation of lyotropic hybrid organic/inorganic phases occurs both under a liquid crystal like approach (high concentration of surfactant)[21] and under dilute conditions, where self-organisation and precipitation is driven in solution by the cooperative formation of inorganic-organic interfaces, as proposed by Monnier et al.[22] Chmelka and co. also reported on the formation of mesostructured aluminosilicate by the same approach and, interestingly, their advanced solid state NMR study showed that incorporation of aluminum in the inorganic framework did not induce a ^{29}Si chemical shift at -78 ppm,[23] which supports the fact that in our case the formation of a classical mesoporous aluminosilicate is most-likely excluded.

Plate, "raft" and nanosheets (nomenclature varies according to the authors) aluminosilicate materials have been described earlier. Barret[24] and Zhao,[25] for instance, reported, respectively, about the formation of pure and Ga-doped boehmite "rafts". Nevertheless, in their case, the "raft" population was described as being uniform in size and composed of either "units" or layered materials; in particular, XRD did not show any small-angle reflection at all, excluding any long-range order in the mesoscale range. In our work, the 2D plates are not homogeneous in size and are composed of nanotubes of different length, as clearly shown by TEM experiments at different tilt angles (Figure 5 e,f). Interestingly, "raft"-like objects have been reported for imogolite before,[26] and in particular several authors have pointed out the liquid

crystal behavior of imogolite nanotubes at high concentration, as predicted by the Onsager and Flory theory.[26c]

In our case, we think that the plate-like sheets composed of imogolite nanotubes is driven by a synergic interaction between the laurate and the imogolite nanotubes and assembly may occurs during the drying process at the liquid/air interface, as predicted by several authors for this type of objects.[27] Scheme 1 proposes a picture summarizing the main achievement of this work. In order to describe properly the sharp small-angle reflections in the XRD pattern, one has to assume that each plate contains a large number of nanotube units. Finally, uncontrolled side reactions, probably due to non-optimized MG/Al/Si stoichiometry, lead to the formation of boehmite.



Scheme 1. Schematic side and longitudinal tentative representation of the hydrophobic imogolite nanotubes. The outer surface of the tube is composed of gibbsite sheets whose Al^{3+} sites are complexed by lauric acid. The longitudinal view is adapted

from Ref. 5e. The size of the plates in this drawing is not to scale. In particular one has to consider the presence of a large number of nanotubes in order to account for the sharp XRD reflections observed in the small-angle range.

Conclusions

Among inorganic nanotubes, imogolite represents a very unique case due to its natural occurrence. Therefore, in contrast to other metal oxide phases whose morphology and dimensions can be controlled by templating approaches,[28] the addition of organics is not expected to change the structure of individual nanotubes but may control their kinetics of formation and their packing, as demonstrated here. Of particular interest is the possibility to achieve functionalization of the imogolite surface in one-step and achieve a new type of nanotube arrangement, never observed before for imogolite; additionally, this is done in a short time and starting from high inorganic precursor concentration. This was possible by the selection of a surfactant molecule that exhibits good affinity for the aluminium species, both in the initial precursor solution and within the aluminosilicate framework. An additional, unexpected outcome of this specific affinity is that a direct surface functionalization of the nanotube is occurring, turning the hydrophilic aluminosilicate particles into a solvent-dispersable powder. Together with the fact that our synthetic pathway complies with several key requirements of green chemistry principles applied to inorganic materials (solvent-free, near neutral pH, biosurfactant, use of chloride instead of perchlorate aluminium salt,..),[29] it can be expected that these results open new perspectives for the use of imogolite as a functional filler in the field of (nano)-composite materials.[30]

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Supplementary data

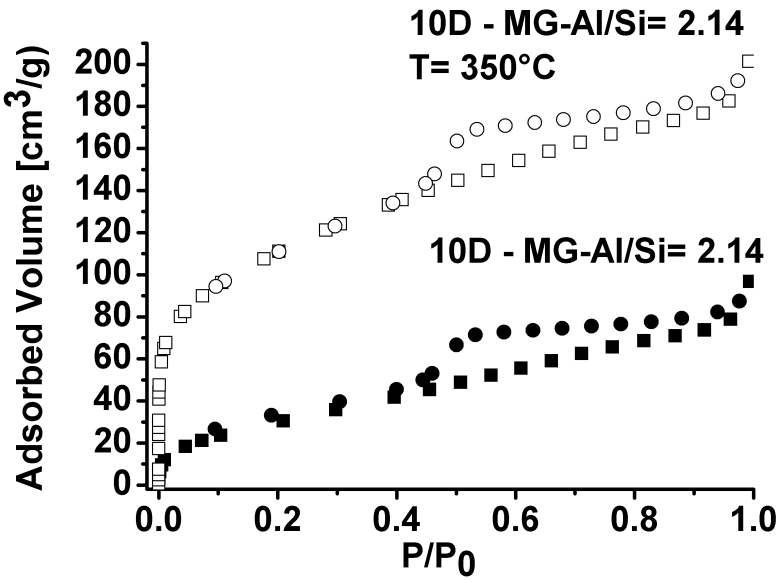
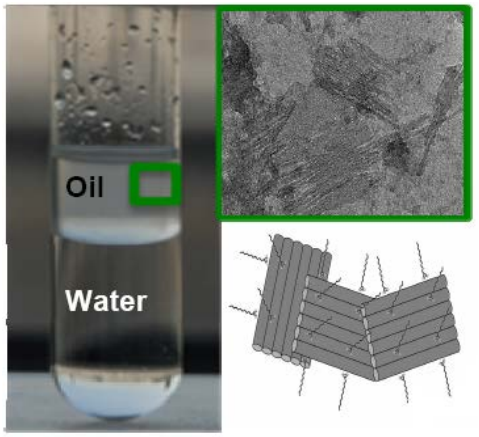


Figure 1 (SI) – Nitrogen adsorption/desorption isotherms recorded on the 10-days MG-containing sample (10D-MGAl/Si= 2.14) both on the as synthesized (filled symbols) and calcined (T= 350°C, empty symbols) materials.

Entry for the Table of Contents



One-step functionalized imogolite single wall nanotubes, obtained from centimolar precursor solutions within 10 days at neutral pH using a biosurfactant-mediated route, demonstrate an unusual 2D lamellar order packing and highly-demanded hydrophobic surfaces.