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Dendritic Nanostructured Copper Oxide Electrocatalyst For Oxygen Evolution Reaction

Tran Ngoc Huan^[a], Gwenaëlle Rouse^[b], Sandrine Zanna^[c], Ivan T. Lucas^[d], Xiangzhen Xu^[e], Nicolas Menguy^[f], Victor Mougel^{[a]*}, Marc Fontecave^{[a]*}

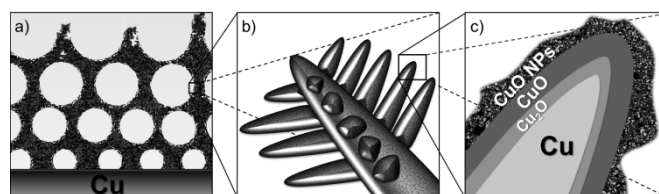
Abstract: In order to use water as the source of electrons for proton or CO₂ reduction within electrocatalytic devices, catalysts are required for facilitating the proton-coupled multi-electron Oxygen Evolution Reaction (OER, 2H₂O → O₂ + 4H⁺ + 4e⁻). These catalysts, ideally based on cheap and earth abundant metals, have to display high activity at low overpotential, good stability and selectivity. While numerous examples of Co, Mn, Ni catalysts were recently reported for water oxidation, only few examples were reported using copper, despite promising efficiencies. We report here a rationally designed nanostructured copper/copper oxide electrocatalyst for OER. This material derives from conductive copper foam passivated by a copper oxide layer and further nanostructured by electrodeposition of CuO nanoparticles. The generated electrodes are highly efficient for catalyzing selective water oxidation to dioxygen with an overpotential of 290 mV at 10 mA/cm² in 1M NaOH solution.

Water splitting into oxygen and hydrogen is one of the most appealing strategies for storing renewable energies in the form of chemical energy.^[1-3] However the water oxidation half reaction, also called oxygen evolution reaction (OER), providing the

electrons required for proton reduction, implies the loss of 4 electrons and four protons from two water molecules with concomitant formation of an O-O bond. This complex process is thus fraught with slow kinetics and requires efficient and stable catalysts. Metal oxides, preferably based on non-noble transition metals (Co, Ni, Mn...), are the most popular candidates and thus have been extensively investigated with renewed interest during the last decade.^[4-10] In that context, few recent examples have indicated that CuO-based materials electrodeposited from molecular Cu complexes or simple Cu salts could display interesting performances for water oxidation.^[11-17] While these are still outperformed by state of the art catalysts^[6-10] (Tables S1 and S2), the low cost and large abundance of copper justifies pursuing studies on rationally designed copper oxide electrodes. With the aim of lowering the overpotential observed for water oxidation with such systems, we targeted the preparation of a copper-based material constituted of a highly conductive support – to limit resistivity of the electrode – possessing a high surface area and a large number of active sites – to facilitate charge transfer and increase overall reaction rates. Meeting these last two requirements typically involves the use of nanoparticles, but the large number of grain boundaries in such systems significantly increases the overall resistivity and overpotential, and their dense packing when immobilized at an electrode reduces the efficiency of mass transfer. Both of these concerns could be addressed by the immobilisation of a thin layer of active nanoparticles on a high surface area conductive support.

In this paper, we report the three-step preparation of a copper-supported high surface area porous dendritic Cu/Cu₂O/CuO material further nanostructured with a thin layer of electrodeposited CuO nanoparticles (Scheme 1). This noble metal free system presents unprecedented water oxidation electrocatalytic activity and stability at low overpotential in a wide range of conditions.^[18]

Scheme 1. Strategy for nanostructured copper oxide electrodes. (a) Macropores resulting from the H₂ bubbles template. (b) Typical dendritic structure of the material. (c) Sectional view of a dendrite.



As a conductive support, we used a robust high surface area three-dimensional porous dendritic Cu material, **1**, generated via a hydrogen evolution assisted electrodeposition approach.^[19-21] For that purpose we followed a previously reported procedure involving the electrodeposition of CuSO₄ on a Cu plate in acidic

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conditions using a large current density ($0.5 \text{ A}\cdot\text{cm}^{-2}$, see ESI for details).^[21] The porous structure of this material, resulting from the combination of copper dendritic growth and the concomitant generation of a foam due to intense hydrogen bubbling, (scheme 1 a-b) is particularly suited for electrocatalytic applications, as the macropores allow for a fast diffusion of species to the micropores, hence avoiding mass transfer limitations.^[19-21]

Metallic copper is typically not stable under the conditions required for water oxidation,^[22] but recent work have proven that surface CuO films on Cu electrodes are very effective at preventing anodic corrosion under catalytic water oxidation conditions.^[15] In order to generate such a passivation film on the dendritic material, **1** was annealed in air at 300°C during 30 min, affording material **2**. Powder X-Ray Diffraction (XRD) patterns of the electrode material after annealing showed the presence of crystalline Cu, Cu_2O and small amounts of CuO, while only Cu and trace amounts of Cu_2O ^[23] could be observed before the high temperature treatment (Figure 1 d). X-ray Photoelectron Spectroscopy (XPS) of **2** presents Cu $2p_{3/2}$ and $2p_{1/2}$ peaks at 933.6 and 953.8 eV and a Cu-Auger peak at 918.0 eV (Figure S1), indicating that CuO is the main surface species (within a 6 nm layer),^[24] in agreement with the analysis of the O 1s region of the XPS spectra, revealing the presence of a main O^{2-} signal at 529.6 eV. Raman spectroscopy of the material (Figure S2) revealed the presence of Cu_2O and CuO signals.^[25] Scanning Transmission Electron Microscopy - X-ray Energy-Dispersive spectroscopy (STEM-XEDS) analyses (Figure S3) confirmed the increase in the oxygen content after thermal treatment, while Scanning electron microscopy (SEM) and Transmission Electron Microscopy (TEM) analyses, Figures 1 a-c and S4, showed that the initial porous and dendritic structure of the solid foam was retained after annealing while being covered by a thin amorphous layer. This, combined with the observation of a significant amount of metallic copper after annealing, demonstrates that the material is likely only oxidized in surface, keeping a metallic copper core and hence a good conductivity.

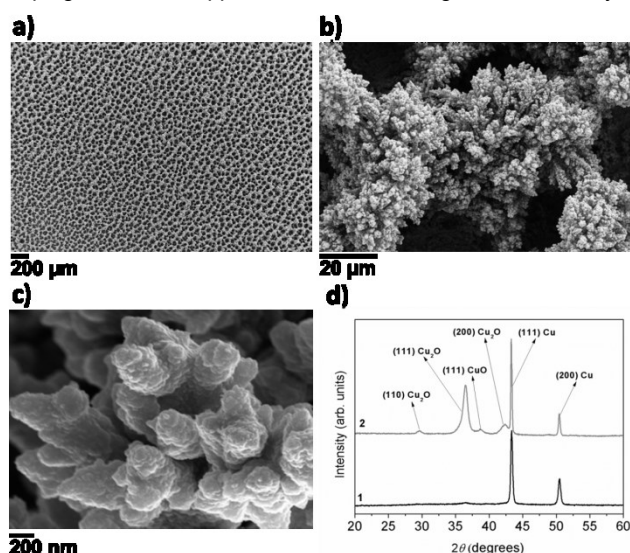


Figure 1. (a)-(c) SEM images of electrode **2**. (d) Powder XRD patterns ($\lambda_{\text{Cu}} = 1.54 \text{ \AA}$) of materials **1** & **2**.

Finally, nanostructuring of the surface of material **2** was achieved following a methodology recently reported by our group and others^[12, 26-29] and based on the electrodeposition of

nanoparticles (NPs) using Cu complexes bearing N-containing ligands as precursors. Here NP electrodeposition was achieved applying 2 Cyclic Voltammetry (CV) cycles (between -0.5 and 1 V vs Ag/AgCl) to a $\text{Cu}(\text{imidazole})_4\text{Cl}_2$ (0.2 mM) solution in acetonitrile (3% water v/v) (See ESI for details).^[30]

SEM characterization of the resulting material, named **3**, clearly showed that this treatment did not alter the macroporosity and the dendritic structure of the material but resulted in a significant modification of the surface (Figure 2 a). Further analysis by TEM revealed the formation of a thin layer of small nanoparticles (< 10 nm) on top of the dendrites (Figure 2 b-c). XPS of **3** is very similar to **2**, exhibiting two peaks at 933.7 and 953.8 eV and a Cu-Auger peak at 917.7 eV, characteristic of CuO (Figure S5). Powder XRD analysis of **3** revealed a largely increased ratio of CuO with respect to **2** (Figure 2 d), and fitting of the broad CuO peaks using Scherrer's equation indicated an average particle size of 6 nm, in line with the dimensions of the nanoparticles observed by TEM. Raman spectra before and after deposition of the nanoparticles proved very similar, the only noticeable difference being an increase of the CuO signal at 645 cm^{-1} with respect to the Cu_2O signals at 220, 425 and 510 cm^{-1} ^[25, 31] (Figure S2), in agreement with the XPS and XRD analysis. These data suggest that the deposited nanoparticles are mainly composed of CuO. In addition, XPS and XEDS measurements confirmed the absence of nitrogen and chlorine in this layer (Figures S5-S6), thus ruling out the presence of imidazole and Chlorine ligands within this CuO film.^[32]

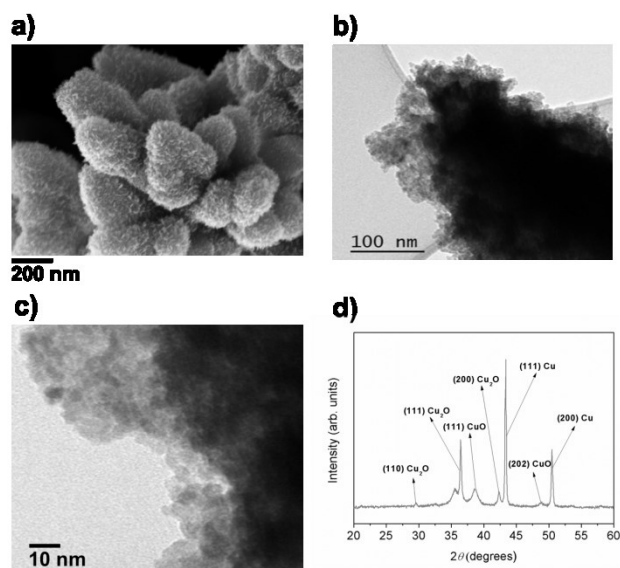


Figure 2. (a) SEM and (b)-(c) TEM images of electrode **3**. Powder XRD pattern ($\lambda_{\text{Cu}} = 1.54 \text{ \AA}$) of electrode **3**

In order to gain deeper insight onto the multilayer structure and chemical composition of this material, Focused Ion Beam (FIB) cross sections (Figure S6) of **3** were analyzed using High Resolution TEM (HRTEM) combined with SAED analysis and elemental mapping using STEM-XEDS (Figure 3). The XEDS maps of the cross section of the dendrite show a core of metallic copper (Figure 3 d) covered by two successive layers containing increasing amounts of oxygen (Figure 3 b-c) and identified by SAED patterns as Cu_2O and CuO (Figure 3 e-h). The presence of these surface CuO/ Cu_2O layers was confirmed by XPS depth analysis of the sample through Ar^+ etching of the surface (Figure S7).

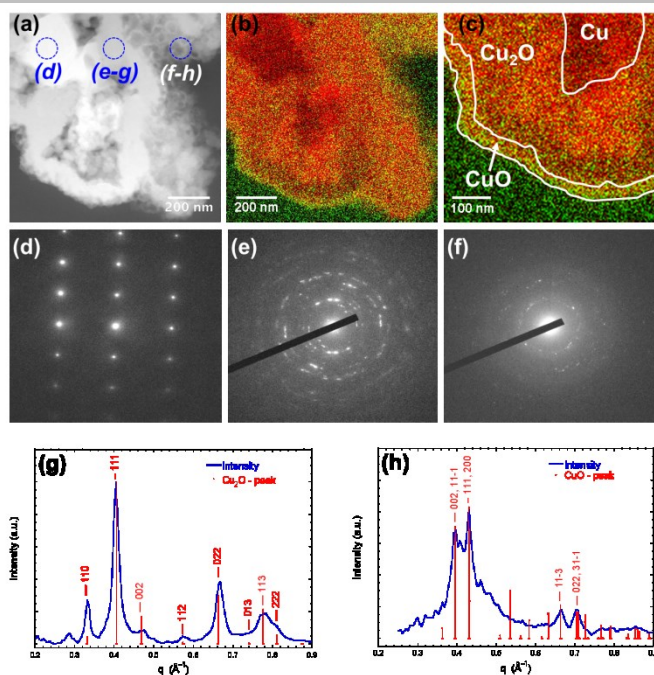


Figure 3. TEM analysis of an FIB cross section of material **3**: (a) STEM-HAADF (High-Angle Annular Dark-Field) image indicating areas where SAED patterns were recorded in blue circles (b)-(c) STEM-XEDS analyses (Cu in red, O in green) (d)(e)(f) SAED patterns of the zones indicated in (a) showing presence of metallic copper ([112] zone axis), Cu₂O (Intensity profile analysis of the SAED pattern presented in (e)) or CuO (Intensity profile analysis of the SAED pattern presented in (f)).

These new materials were evaluated for electrocatalytic water oxidation. Representative linear sweep voltammograms (LSV) in 1.0 M aqueous NaOH solution are shown in Figure 4. Under these conditions **3** proved to be the most efficient catalytic material, with a remarkable overpotential of only 290 mV at 10 mA.cm⁻², this value being to our knowledge the lowest for copper-based materials (Table S2). In order to assess the influence of the high Electrochemical Surface Area (ECSA) of studied materials (values given in Table S3) on their catalytic properties, we normalized the LSV plots by the ECSA of the electrodes (Figures S8) and compared them to a state of the art porous copper oxide OER catalyst^[14] (Figure S9). We thus observed that catalyst **3**, with an ECSA-normalized overpotential of 340 mV/cm²_{ECSA} at 1 mA, was more active than all the previously reported copper based catalysts even after its high surface area is taken into account (Table S2). As for many water oxidation catalysts, the Tafel plot for **3** is complex and present a strong potential dependency of the Tafel slope.^[33, 34] A Tafel slope of 64 mV/decade could be measured between 250 and 400 mV overpotential, a comparable value with many other OER catalysts (Figure 4, inset).^[16, 31] One should however be cautious using the Tafel slope value as a benchmarking parameter since, as in the case of OER catalysts, the latter greatly depends on which elementary step of the reaction is the actual limiting one.^[33] As shown in Figure 4, the porous dendritic materials **1** and **2** are significantly less active (overpotential of 460 mV and 370 mV at 10 mA.cm⁻² for **1** and **2** respectively), highlighting the strong beneficial influence of the CuO nanoparticle film.

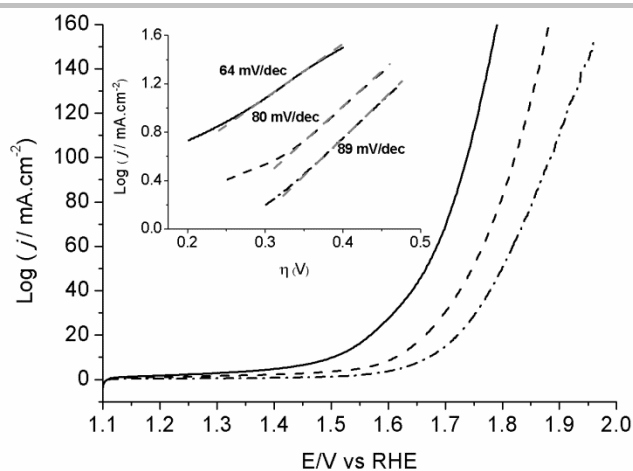


Figure 4: LSV in 1.0M NaOH of electrodes **1** (dashed-dot line), **2** (dashed line) and **3** (solid line), using a scan rate of 10 mV.s⁻¹ (currents are uncorrected for resistive losses incurred within the electrolyte, all current densities are based on projected geometric area). Corresponding Tafel plots and resulting slopes are given in the inset.

In order to get insights into the relative contributions of the nanoparticle film and the dendritic copper oxide material on activity, we generated a support of identical morphology and surface area by sputter deposition of a 10 nm film of gold on material **2**. Structural integrity of this new material, named **2-Au**, was confirmed by SEM images (Figure S10) and by the fact that identical ECSA was measured for **2** and **2-Au**. **2-Au** was found to be less active than its parent material **2**, as witnessed by a larger overpotential of 410 mV at 10 mA.cm⁻² in 1.0 M NaOH (vs. 370 mV for **2**). A CuO NP layer was then deposited on the gold film via the procedure used for **3** preparation, resulting in a new material called **2-Au-CuO** (See ESI for details). LSV of these new materials in 1.0 M NaOH solution are given in Figure S10 d. As in the case of **3**, the CuO NP layer provided significantly improved activity with an overpotential of 340 mV at 10 mA.cm⁻² for **2-Au-CuO** in 1.0 M NaOH. Yet, **2-Au-CuO** is less active than **3**, indicating that the copper oxide dendritic support in **3** plays a synergistic role with the CuO NP.

Interestingly, **3** was remarkably active also at lower pH values, (Figure S11 and table S2), presenting an overpotential of 570 mV at 5 mA.cm⁻² in a pH 9.2 potassium borate (KBi) buffer solution, 470 mV at 10 mA.cm⁻² in a pH 11 sodium carbonate solution and 365 mV at 10 mA.cm⁻² in 0.1 M NaOH solution. In order to investigate the long term stability of the catalyst **3**, controlled current electrolysis was carried out at 10 mA.cm⁻² in a pH 11 sodium carbonate buffer solution as well as in 0.1 or 1.0 M NaOH solutions (Figure 5). In the case of both pH 11 and 0.1 M NaOH solutions the potential remained remarkably stable during the 48 hours of the reaction at values of 1.70 V or 1.62 V vs RHE, respectively, corresponding to overpotentials of only 470 mV and 390 mV. In 1.0 M NaOH, a stable potential of 1.53 V (corresponding to an overpotential of 300 mV) was observed for 10 h, but the electrode then quickly degraded, as witnessed by the strong increase of overpotential. The loss of activity was accompanied by a degradation of the surface structuration, the dendritic framework being largely damaged. Such a fast decline in catalytic activity had been similarly reported for a CuO OER electrocatalyst.^[13] In contrast, at lower pH conditions (pH 11 and 0.1 M NaOH) both dendritic framework and surface structuration were maintained even after 48 h operation (SEM pictures and

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XPS analyses of the electrodes after long-term electrolysis are given in Figures S12-S14).^[35] In addition, TEM, STEM-HAADF and STEM-XEDS analyses of the catalyst after long term electrolysis in 0.1 M NaOH showed that the catalyst structure was essentially unchanged, the thin CuO nanoparticles layer remaining at the surface of a dense crystalline dendritic structure (Figure S15). As the only noticeable difference with the sample before electrolysis, the nanoparticle layer at the surface appeared more homogeneous in thickness.

Interestingly, the fast decline of catalytic activity of material **3** observed during long-term electrolysis at 10 mA/cm² in 1.0 M NaOH was not observed when catalytic tests were carried out at lower current densities, as previously observed with other CuO catalysts.^[13] For example, catalytic activity remained stable over more than 100 h in 1.0 M NaOH when electrolysis was carried out at 1 mA/cm² (Figure 5). This observation, together with the unexpected stability of copper oxides occasionally observed in concentrated hydroxide solutions,^[36, 37] highlights the complex behavior of copper oxide materials in highly basic media, whose origin remains to be identified. In all cases, quantification of O₂ formed during long term electrolyses by Gas Chromatography revealed faradic yields over 90% (see ESI for details).

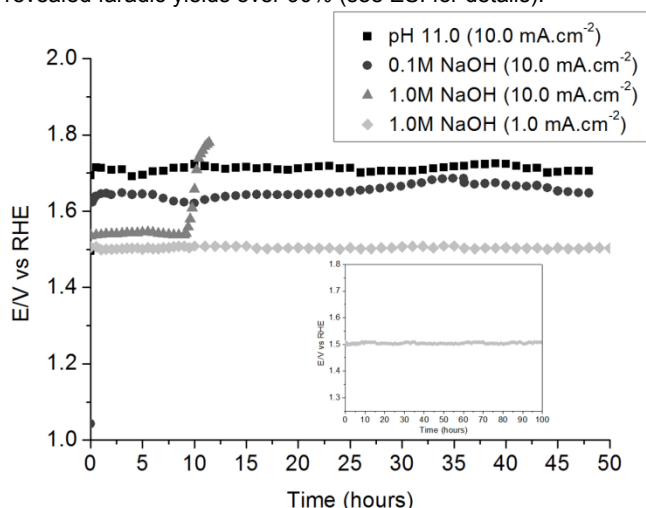


Figure 5: Plot of catalytic potential at fixed current density during long-term electrolysis for material **3** at pH 11.0 (10 mA.cm⁻²) and in 0.1 M (10 mA.cm⁻²) and 1 M NaOH (1 and 10 mA.cm⁻²) aqueous solution (See ESI for details). Inset: 100 h electrolysis in 1 M NaOH (1 mA.cm⁻²).

The results reported herein demonstrate that highly active water oxidation catalysts can be generated using nanostructured copper oxide dendritic electrodes. The dense copper oxide layer at the surface of the dendritic supports prevents the corrosion of the basal copper substrate, allowing for good and stable conductivity of the material. This new type of electrodes can be generated using simple inorganic salts and thermal treatments, allowing an easy and cost effective preparation using earth abundant copper as the sole metal, including the electrode support. These electrodes stand among the best materials for water oxidation reported in the literature, and the very versatile methods used for their generation could likely be extended to other metals.

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Keywords: copper oxide • water oxidation • porous electrode • electrocatalysis • electrodeposition

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