



HAL
open science

Coordination-Driven Folding in Multi-Zn II -Porphyrin Arrays Constructed on a Pillar[5]arene Scaffold

Thi Minh Nguyet Trinh, Iwona Nierengarten, Haifa Ben aziza, Eric Meichsner, Michel Holler, Matthieu Chessé, Rym Abidi, Christian Bijani, Yannick Coppel, Emmanuel Maisonhaute, et al.

► **To cite this version:**

Thi Minh Nguyet Trinh, Iwona Nierengarten, Haifa Ben aziza, Eric Meichsner, Michel Holler, et al.. Coordination-Driven Folding in Multi-Zn II -Porphyrin Arrays Constructed on a Pillar[5]arene Scaffold. *Chemistry - A European Journal*, 2017, 23 (46), pp.11011-11021. 10.1002/chem.201701622 . hal-01609037

HAL Id: hal-01609037

<https://hal.sorbonne-universite.fr/hal-01609037>

Submitted on 11 Oct 2017

HAL is a multi-disciplinary open access archive for the deposit and dissemination of scientific research documents, whether they are published or not. The documents may come from teaching and research institutions in France or abroad, or from public or private research centers.

L'archive ouverte pluridisciplinaire **HAL**, est destinée au dépôt et à la diffusion de documents scientifiques de niveau recherche, publiés ou non, émanant des établissements d'enseignement et de recherche français ou étrangers, des laboratoires publics ou privés.

Coordination-driven folding in multi-Zn(II)porphyrin arrays constructed on a pillar[5]arene scaffold

Thi Minh Nguyet Trinh,^[a] Iwona Nierengarten,^[a] Haifa Ben Aziza,^[a,b] Eric Meichsner,^[a] Michel Holler,^[a] Matthieu Chessé,^[a] Rym Abidi,^[b] Christian Bijani,^[d] Yannick Coppel,^[d] Emmanuel Maisonhaute,^{*,[c]} Béatrice Delavaux-Nicot,^{*,[d]} and Jean-François Nierengarten^{*,[a]}

Dedicated to Prof. Vincenzo Balzani on the occasion of his 80th birthday

Abstract: Pillar[5]arene derivatives bearing peripheral porphyrin subunits have been efficiently prepared from a decaazide pillar[5]arene building block (**17**) and Zn(II)-porphyrin derivatives bearing a terminal alkyne function (**9** and **16**). For the resulting deca-Zn(II)-porphyrin arrays (**18** and **20**), variable temperature NMR studies revealed an intramolecular complexation of the peripheral Zn(II)-porphyrin moieties by 1,2,3-triazole subunits. As a result, the molecules adopt a folded conformation. This was further confirmed by UV/vis spectroscopy and cyclic voltammetry. In addition, we have also demonstrated that the coordination-driven unfolding of **18** and **20** can be controlled by an external chemical stimulus. Specifically, addition of an imidazole derivative (**22**) to solution of **18** or **20** breaks the intramolecular coordination at the origin of the folding. The resulting molecular motions triggered by the addition of the imidazole ligand mimics the blooming of a flower.

Introduction

Pillar[5]arene are macrocyclic compounds at the forefront of modern supramolecular chemistry.^[1] Whereas their paracyclophane skeleton was already reported in the 80s,^[2] the synthesis of such compounds was difficult thus limiting both their availability and applicability.^[2] This problem was solved in 2008 when Ogoshi and co-workers reported an efficient one-step synthesis. In their seminal paper, they have shown that cyclopentameric macrocyclic compounds can be prepared in

because of their tubular shape. Following this report, the field of pillar[*n*]arene has generated significant research efforts focused on both the improvements of their synthetic methods^[4] and their incorporation in supramolecular structures.^[1] Owing to the presence of ten peripheral substituents, pillar[5]arenes are also attractive compact scaffolds for the preparation of multivalent compounds for biological applications.^[5-6] Examples include glycopillar[5]arenes for the specific binding to bacterial lectins^[5] and polycationic compounds for gene delivery.^[6] On the other hand, nanomaterials with a controlled distribution of functional groups on the macrocyclic framework have been also reported.^[7] As part of this research, we now report the preparation of pillar[5]arene derivatives bearing ten peripheral Zn(II)-porphyrin subunits. These compounds have been efficiently prepared from a clickable pillar[5]arene building block and appropriate Zn(II)-porphyrin derivatives under copper-catalyzed alkyne-azide cycloaddition (CuAAC) conditions. An unprecedented folding due to intramolecular coordination of the 1,2,3-triazole linkers to the peripheral Zn(II)-porphyrin moieties has been evidenced for the resulting decatriazoles. Interestingly, this coordination-driven folding can be controlled by an external chemical stimulus, namely addition of an imidazole ligand. As a result, unfolding of the molecules occurs and the resulting molecular motions mimic the blooming of a flower.

Results and Discussion

Synthesis. The preparation of the first clickable porphyrin building block is depicted in Scheme 1. Treatment of alcohol **1** with *p*-toluenesulfonyl chloride (*p*-TsCl) gave tosylate **2** in 77% yield. Reaction of *p*-hydroxybenzaldehyde (**3**) with tosylate **2** under classical Williamson etherification conditions (K₂CO₃, DMF, 80°C) afforded **4** in 85% yield. Porphyrin **7** was obtained by using the reaction conditions developed by Lindsey for the synthesis of sterically hindered porphyrins.^[8] A key feature of these conditions involves BF₃-ethanol co-catalysis. The condensation of **4** (1 equiv.), mesaldehyde (**5**, 3 equiv.) and pyrrole (**6**, 4 equiv.) was performed in CHCl₃ (commercial CHCl₃ containing 0.75% ethanol as stabilizer) at room temperature in the presence of BF₃.Et₂O. After 3 h, *p*-chloranil (tetrachlorobenzoquinone) was added to irreversibly convert the porphyrinogen to the porphyrin. Free base porphyrin **7** was subsequently isolated in 15% yield by tedious chromatographic separations. Metalation of **7** with Zn(OAc)₂ gave **8** in 83% yield. The metalation is conveniently followed by the changes in the absorption spectra. Actually, four bands (Q bands), responsible for the red-purple color in solution, are observed in the visible region (515, 549, 593 and 648 nm) of the absorption spectrum of free base porphyrin **7**. A very sharp and intense band (B band also called Soret band) is also present in the near UV region

[a] Dr. T. M. N. Trinh, Dr. I. Nierengarten, Dr. H. Ben Aziza, E. Meichsner, Dr. M. Holler, M. Chessé, Dr. J.-F. Nierengarten
Laboratoire de Chimie des Matériaux Moléculaires, Université de Strasbourg et CNRS (UMR 7509), Ecole Européenne de Chimie, Polymères et Matériaux
25 rue Becquerel, 67087 Strasbourg Cedex 2, France.
E-mail: nierengarten@unistra.fr

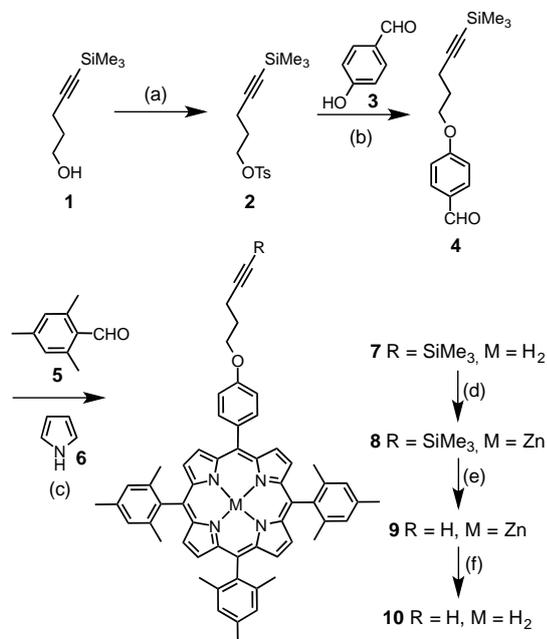
[b] Dr. H. Ben Aziza, Prof. R. Abidi
Laboratoire d'Applications de la Chimie aux Ressources et Substances Naturelles et l'Environnement, Faculté des Sciences de Bizerte, Université de Carthage
7021 Zarzouna Bizerte, Tunisia.

[c] Prof. E. Maisonhaute
Sorbonne Universités, UPMC Univ Paris 06, UMR 8235, Laboratoire Interfaces et Systèmes Electrochimiques, F-75005 Paris, France.
E-mail: emmanuel.maisonhaute@upmc.fr

[d] C. Bijani, Dr. Y. Coppel, Dr. B. Delavaux-Nicot
Laboratoire de Chimie de Coordination du CNRS, Université de Toulouse, UPS, INPT, 205 route de Narbonne, BP 44099, F-31077 Toulouse Cedex 4, France.
E-mail: beatrice.delavaux-nicot@lcc-toulouse.fr

excellent yields from 1,4-dialkoxybenzene and paraformaldehyde in the presence of BF₃.Et₂O.^[3] They also proposed the name *pillararene* for this class of macrocycles

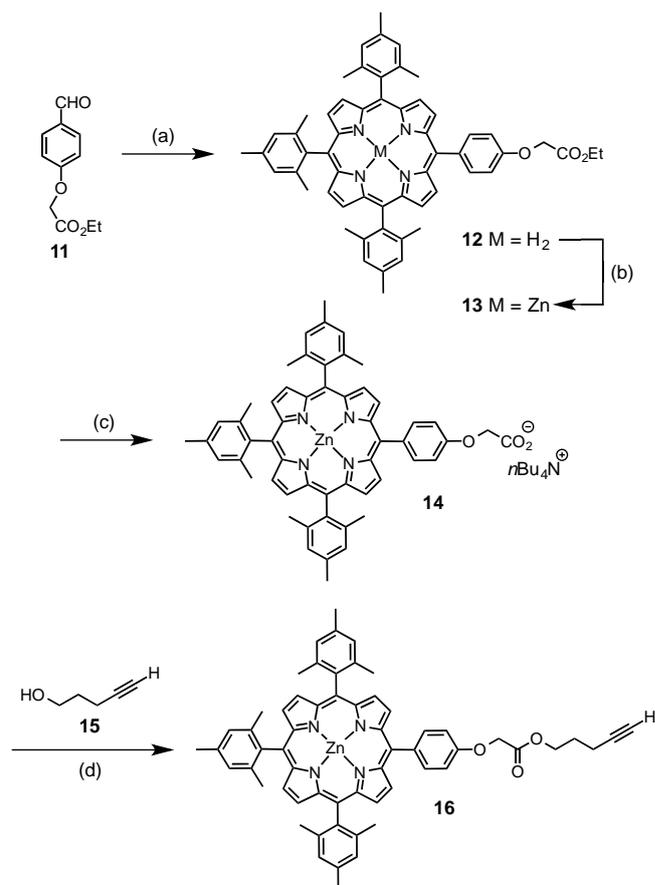
(419 nm). Upon metalation, the four Q bands collapse into two bands (550 and 587 nm) due to the higher symmetry of the porphyrin core when metallated (D_{4h} vs. D_{2h}), but the Soret band (422 nm) is barely affected.^[9] The changes in the visible region result also in a clear color change when going from **7** to **8**. Effectively solutions of free base porphyrin **7** are red-purple while solutions of Zn(II)-porphyrin **8** are pink-purple.



Scheme 1. Reagents and conditions: (a) *p*-TsCl, pyr, CH_2Cl_2 , rt (77%); (b) K_2CO_3 , DMF, 80°C (85%); (c) $\text{BF}_3\cdot\text{Et}_2\text{O}$, CHCl_3 , rt, then chloranil (15%); (d) $\text{Zn}(\text{OAc})_2\cdot 2\text{H}_2\text{O}$, $\text{CHCl}_3/\text{MeOH}$, Δ (83%); (e) TBAF, THF, 0°C (91%); (f) TFA, CH_2Cl_2 , rt (96%).

Reaction of trimethylsilyl (TMS)-protected alkyne **8** with tetra-*n*-butylammonium fluoride (TBAF) in THF afforded Zn(II)-porphyrin **9** bearing a terminal alkyne function. Inspection of the ^1H NMR spectrum of **9** indicates the disappearance of the diagnostic TMS signal observed at $\delta = 0.20$ ppm for **8** and the concomitant apparition of a triplet ($J = 3$ Hz) at $\delta = 2.04$ ppm corresponding to the resonance of the terminal alkyne proton. Moreover, in the IR spectrum of **9**, the C-H stretching band characteristic of the terminal alkyne function is observed at 3302 cm^{-1} . Finally, Zn(II)-porphyrin **9** was also demetallated by treatment with trifluoroacetic acid (TFA) to afford the corresponding free-base porphyrin (**10**) that was used as a reference compound for the photophysical studies.

The preparation of a second clickable Zn(II)-porphyrin building block having a slightly longer spacer between the porphyrin core and the terminal alkyne function was also achieved (Scheme 2). Aldehyde **11** was prepared from **3** and ethyl bromoacetate according to a previously reported method.^[10] A₃B porphyrin **12** was then obtained in 14% yield by reaction of **11** (1 equiv.) and **5** (3 equiv.) with pyrrole (**6**, 4 equiv.) under typical Lindsey conditions. Subsequent treatment with $\text{Zn}(\text{OAc})_2$ afforded the corresponding Zn(II)-porphyrin (**13**) in 99% yield.

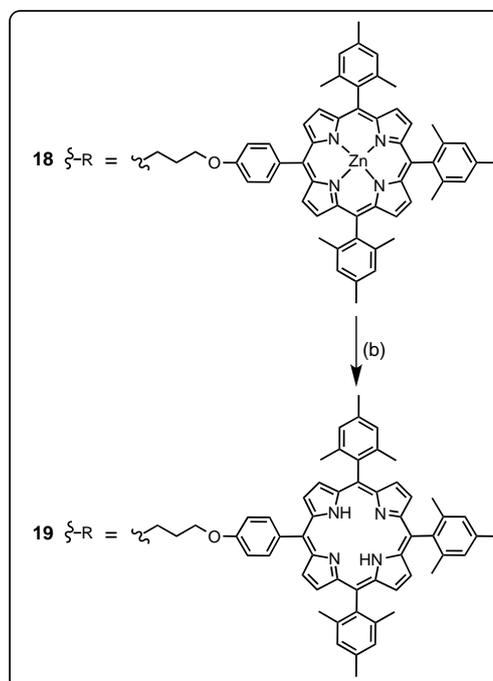
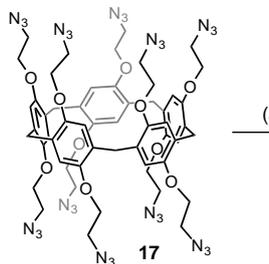


Scheme 2. Reagents and conditions: (a) **5**, **6**, $\text{BF}_3\cdot\text{Et}_2\text{O}$, CHCl_3 , rt, then chloranil (14%); (b) $\text{Zn}(\text{OAc})_2\cdot 2\text{H}_2\text{O}$, $\text{CHCl}_3/\text{MeOH}$, Δ (99%); (c) K_2CO_3 , H_2O , THF, Δ then $n\text{Bu}_4\text{NOH}$ (85%); (d) DCC, DMAP, HOBT, CH_2Cl_2 , 0 to 25°C (95%).

Hydrolysis of the ester function of **13** was achieved by treatment with K_2CO_3 in a mixture THF and water. The resulting carboxylate is poorly soluble in CH_2Cl_2 as its potassium salt. In order to obtain a good solubility in organic solvents, tetrabutylammonium hydroxide was added to the reaction mixture and carboxylate **14** was conveniently extracted with CH_2Cl_2 as its tetrabutylammonium salt. Finally, reaction of **14** with alcohol **15** under esterification conditions using *N,N*-dicyclohexylcarbodiimide (DCC), 4-(dimethylamino)pyridine (DMAP) and 1-hydroxybenzotriazole (HOBT) afforded clickable Zn(II)-porphyrin **16** in 95% yield.

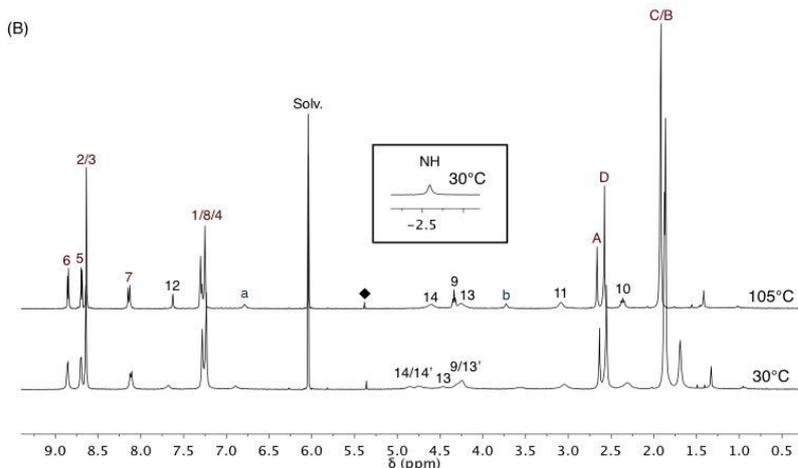
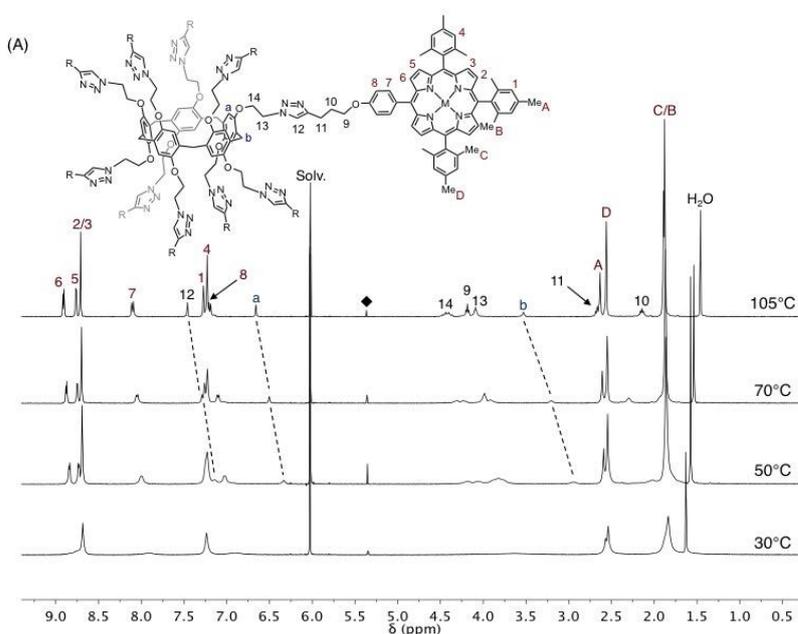
The grafting of the Zn(II)-porphyrin building blocks (**9** and **16**) onto the pillar[5]arene scaffold was achieved under CuAAC conditions (Scheme 3). Pillar[5]arene building block **17** bearing ten peripheral azide functions was prepared according to a previously reported procedure.^[7a] The reaction of **17** with terminal alkyne **9** was achieved under the typical CuAAC conditions developed for the functionalization of multi-azide cores^[11-12] ($\text{CuSO}_4\cdot 5\text{H}_2\text{O}$, sodium ascorbate, $\text{CH}_2\text{Cl}_2/\text{H}_2\text{O}$). It can be noted that it was difficult to follow the progress of the reaction by simple thin layer chromatography (TLC) analysis. Indeed, IR spectroscopy was a precious complementary tool to indicate the total consumption of the azide residues by monitoring the disappearance of the diagnostic azide band at 2089 cm^{-1} . In this particular case, completion of the reaction was clearly observed after 48 h. After work-up and purification, compound **18** was

obtained in 96% yield. The ten peripheral Zn(II)-porphyrin moieties were then demetallated by treatment of **18** with TFA to



Scheme 3. Preparation of porphyrin-substituted pillar[5]arene derivatives. *Reagents and conditions:* (a) **9** or **16**, CuSO₄·5H₂O, sodium ascorbate, CH₂Cl₂, H₂O, 25°C [from **9**: **18** (96%); from **16**: **20** (66%)]; (b) TFA, CH₂Cl₂, rt [from **18**: **19** (92%); from **20**: **21** (97%)].

The reaction conditions used for the preparation of **18** were then applied to the second clickable Zn(II)-porphyrin derivative (**16**). A mixture of **17** (1 equiv.), **16** (12 equiv.), CuSO₄·5H₂O (0.4 equiv.), sodium ascorbate (0.3 equiv.) in CH₂Cl₂/H₂O was vigorously stirred at room temperature for 6 days. After work-up and purification by column chromatography on SiO₂ followed by gel-permeation chromatography (Biobeads SX-1, CH₂Cl₂), pillar[5]arene derivative **20** was obtained in 66% yield. Finally, treatment of **20**



with TFA afforded the corresponding free base porphyrin-substituted pillar[5]arene derivative (**21**) in 97% yield.

Characterization of the porphyrin-substituted pillar[5]arene derivatives.

Compounds **18-21** were characterized by ^1H and ^{13}C NMR, IR and UV/vis spectroscopies. For all the compounds, the proposed structures were also confirmed by MALDI-TOF mass spectrometry. For both free-base porphyrin derivatives (**19** and **21**), the ^1H NMR spectra recorded at room temperature revealed a single set of signals for all the peripheral porphyrin moieties thus showing that they are all equivalent. This is in perfect agreement with the 5-fold symmetrical structures of compounds **19** and **21**. In addition to the signals corresponding to the 10 equivalent peripheral moieties, the resonances arising from the pillar[5]arene core are clearly observed for both **19** and **21**. Actually, the ten aromatic protons of the macrocyclic core are equivalent and give rise to a singlet at $\delta = 6.81$ for **19** and 6.69 ppm for **21**. Similarly, a singlet is observed for the methylene bridging moieties of the pillar[5]arene core ($\delta = 3.45$ and 3.34 ppm for **19** and **21**, respectively). Whereas well resolved signals are observed in the ^1H NMR spectrum of the free base porphyrin derivatives (**19** and **21**) recorded at room temperature, the spectra of the corresponding Zn(II)-porphyrin analogues (**18** and **20**) are broad under the same conditions. This is shown in Figure 1A for compound **18**. Variable NMR studies revealed however a perfectly reversible narrowing of the signals by increasing the temperature. Indeed, the ^1H NMR spectra of **18** and **20** recorded at high temperature show the expected set of signals (Figures 1A and S13b). These observations suggest a fast dynamic exchange between different conformers on the NMR timescale at high temperatures. Importantly the dynamic motions at the origin of the broadening of the spectra of **18** and **20** at room temperature are specific to the presence of the Zn(II) centers. Effectively, as shown in Figure 1B, the ^1H NMR spectra of free-base porphyrin **19** reveals only minor changes as a function of the temperature.

Figure 1. ^1H NMR spectra (400 MHz, $\text{CDCl}_2/\text{CDCl}_2$) of compounds **18** (A) and **19** (B) recorded at different temperatures (\blacklozenge : CH_2Cl_2).

In this particular case, limited molecular motions of the peripheral groups at low temperature results in a slight broadening of the signals corresponding to the protons located close to the core moiety of the molecule. At higher temperature, the increased thermal agitation of the peripheral substituents contributes to sharpen these signals ($\text{H}_{\text{a-b}}$ and H_{12} , see Figure 1B). Moreover, agitation also contributes to simplify the pattern observed for the signals of some specific protons. For example, the two CH_2 groups of the ethylene linker are both diastereotopic ($\text{H-13/13}'$ and $\text{H-14/14}'$, see Figure 1) and accordingly give rise to four sets of signals at 30°C. In contrast, only two signals are observed at 105°C for the same protons (H-13 and H-14 , see Figure 1B) and both methylene subunits appear as enantiotopic under these conditions. As already observed for glycopillar[5]arenes,^[5a] the thermal agitation of the peripheral substituents prevents the transfer of the chiral information between the peripheral subunits and the central core.

In the case of **18** and **20**, the dramatic changes observed in the ^1H NMR spectra as a function of the temperature cannot be simply explained by an increased thermal agitation of the

peripheral substituents. The Zn(II) centers play obviously a key role in the dynamic motions contributing to the broadening of the spectra at low temperature. Actually, the presence of 1,2,3-triazole moieties in **18** or **20** could result in intramolecular interactions with neighboring Zn(II)-porphyrin subunits as schematically depicted in Figure 2. Aromatic compounds with a donating nitrogen moiety are well known ligands of Zn(II)-porphyrins.^[13-15] Pyridine and imidazole derivatives are typical examples. Even if 1,2,3-triazoles are weaker ligands owing to their lower basicity, their binding to Zn(II)-porphyrins has been already documented in the literature.^[14]

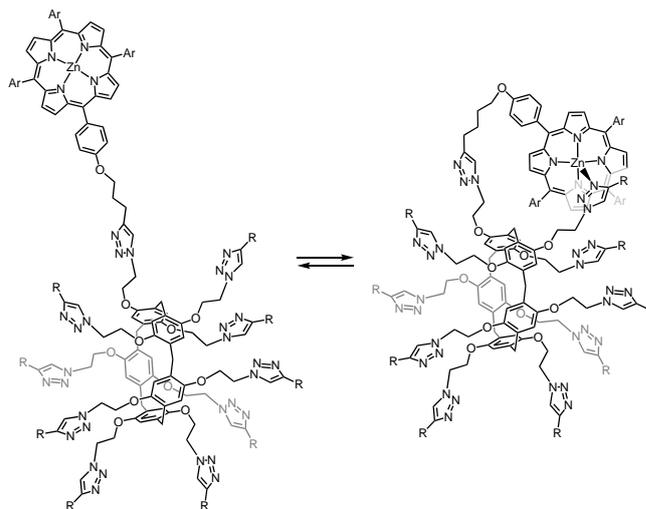


Figure 2. Schematic representation of the dynamic conformational changes resulting from the intramolecular coordination of a 1,2,3-triazole moiety to a Zn(II)-porphyrin subunit of compound **18**.

The binding constant for the intramolecular equilibrium between coordinated and uncoordinated Zn(II)-porphyrins in **18** is temperature dependent. At high temperature, the dynamic equilibrium is displaced in favor of uncomplexed species and the conformation of **18** must be similar to that of its analogous free-base analogue (**19**). Indeed, as shown in Figure 1, the ^1H NMR spectra recorded at 105°C for **18** and **19** are very similar. In contrast, at low temperature, association is favored thus generating important conformational changes in **18**. The coordination of the Zn(II)-porphyrin moieties implies a folding of the molecule and aromatic porphyrin rings are located closer to the core moiety of the molecule. This view is clearly supported by the dramatic shielding observed for the signals of some protons by decreasing the temperature (H-a , H-b and H-12 , see Figure 1A). Indeed, these chemical shift changes result from the ring current effect of the aromatic porphyrin moieties that are spatially close to H-a , H-b or H-12 when intramolecular coordination of Zn(II) centers with 1,2,3-triazole groups occurs. ^1H NMR spectra of **18** were also recorded at low temperature (from room temperature to -70°C). A continuous broadening was observed thus showing that the coordination/decoordination process remains faster than the NMR timescale even at -70°C and/or that compound **18** does not adopt a preferential folded conformation at -70°C. A similar behavior has been also observed for compound **20** with longer spacer subunits. In both cases, this interpretation was further supported by the self-diffusion coefficients (D) of compounds **18-21** deduced from

DOSY experiments (see Table S1). Comparison of the free-base porphyrin derivatives with their Zn(II) coordinated analogues revealed an increase in their D value (13(±5)% for **19** and 9(±6)% for **21**) corresponding to a respective volume reduction of 32(±12)% (**18** vs. **19**) and 22(±15)% (**20** vs. **21**). These observations showing a reduced average volume for the Zn(II)-porphyrin derivatives is fully consistent with a dynamic intramolecular coordination process leading to a folded conformation.

The occurrence of intramolecular interactions of 1,2,3-triazole groups with neighboring Zn(II)-porphyrin subunits in **18** and **20** was also supported by their absorption spectra recorded in CH₂Cl₂ at room temperature (Figures 3, S11g and S13g).

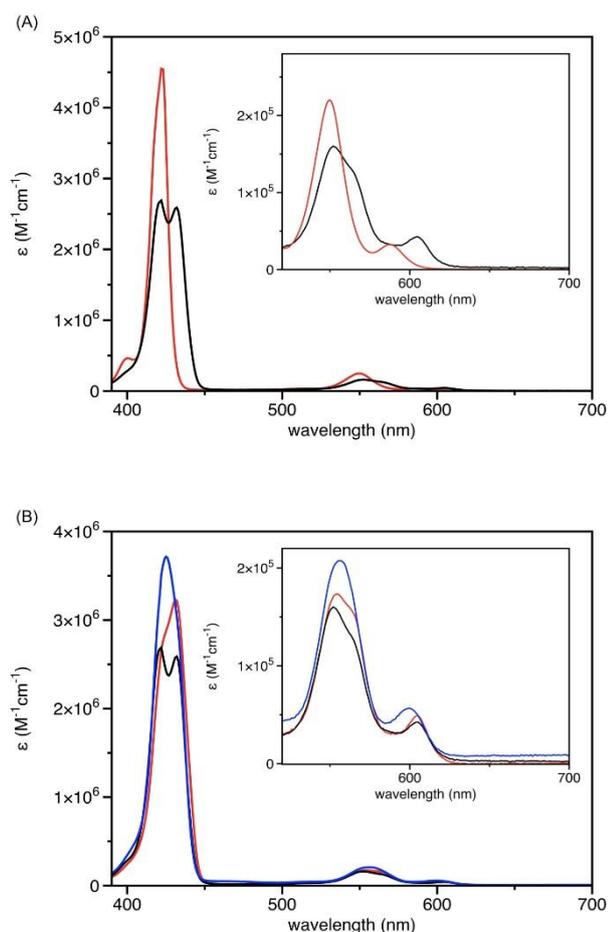


Figure 3. (A) Absorption spectra of compounds **18** (black) and **9** (red) recorded in CH₂Cl₂ at 25°C. (B) Absorption spectra of compounds **18** recorded in different solvents at 25°C (blue: CHCl₃, black: CH₂Cl₂, red: toluene).

In both cases, the absorption bands of the Zn(II)-porphyrin moieties are effectively broadened and red-shifted when compared to those of appropriate model compounds. This is a clear signature^[13-15] for the apical coordination of Zn(II)-porphyrin moieties in **18** and **20** by a nitrogen ligand. Moreover, no evidence of molecular aggregation (e.g. intermolecular coordination) could be detected within the range of concentrations used for the optical measurements (10⁻⁵ to 10⁻⁷ M) as the absorbance follows the Beer-Lambert law for both **18** and **20**. Therefore, the apical coordination of the Zn(II)-porphyrin moieties by 1,2,3-triazoles in **18** and **20** occurs in an

intramolecular manner. It can be also noted that the shape of the absorption spectra of both **18** and **20** is significantly affected by the nature of the solvent as shown in Figure 3B for dichloromethane, chloroform and toluene. Close inspection of the shape of the Soret band suggests a higher degree of association in the less polar solvent (toluene). In this solvent, the red-shifted absorption maxima of the Soret band that is typical for coordinated Zn(II)-porphyrins becomes effectively more pronounced. This observation is consistent with the solvent-dependence of the binding constant (K_A) reported in the literature for complexes resulting from the intermolecular association of N-ligands and Zn(II)-porphyrin derivatives.^[14] The K_A values are in general higher in less polar solvents.

The absorption spectra of the free-base porphyrin-substituted pillar[5]arene derivatives **19** and **21** are depicted in Figures 4. In contrast to what was observed for their Zn(II) analogues, the absorption spectra of **19** and **21** are only very slightly broadened when compared to model free-base porphyrin **10**. Indeed, their absorption spectra corresponds well to the sum of their constitutive subunits (10 x **10**) thus showing rather limited electronic interactions (if any) among the free-base porphyrin moieties in **19** and **21**.

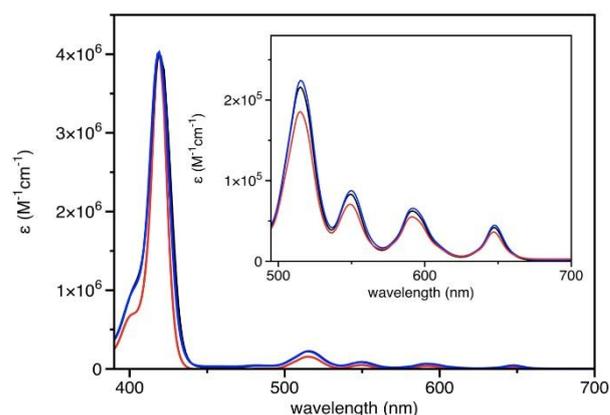


Figure 4. Absorption spectra of compounds **19** (blue), **21** (black) and **10** (red) recorded in CH₂Cl₂ at 25°C.

These observations further support our interpretation of the UV/vis data of the Zn(II) analogues **18** and **20**. The splitting of the Soret band and the overall broadening of the absorption spectra observed for **18** and **20** are not related to electronic interactions between their peripheral porphyrin subunits but result from an intramolecular coordination-driven conformational folding of the molecules. Indeed, the molecular motions resulting from the coordination/decoordination phenomena observed for compounds **18** and **20** mimic the blooming of a flower. As deduced from the variable temperature NMR studies, the folded conformation is favored at low temperature. By increasing the temperature, the intramolecular coordination is disfavored and the molecules adopt an open conformation like a flower that bloom.

Blooming of the molecular flowers with a chemical stimulus (input). In order to gain more control on the conformational changes observed for **18** and **20**, we became interested in using an external input to trigger the blooming of molecular flowers **18**

and **20**. For this purpose, 1-phenylimidazole (**22**) was selected as an external chemical input. Indeed, compound **22** is a stronger ligand for Zn(II)-porphyrins^[13-15] when compared to 1,2,3-triazoles and one may anticipate that a preferential coordination of the Zn(II)-porphyrin groups in **18** or **20** will occur in the presence of an excess of **22**. As shown in Figure 5, dramatic changes were evidenced in the UV/vis spectrum of **20** upon addition of an excess of imidazole derivative **22** (14 equiv.).

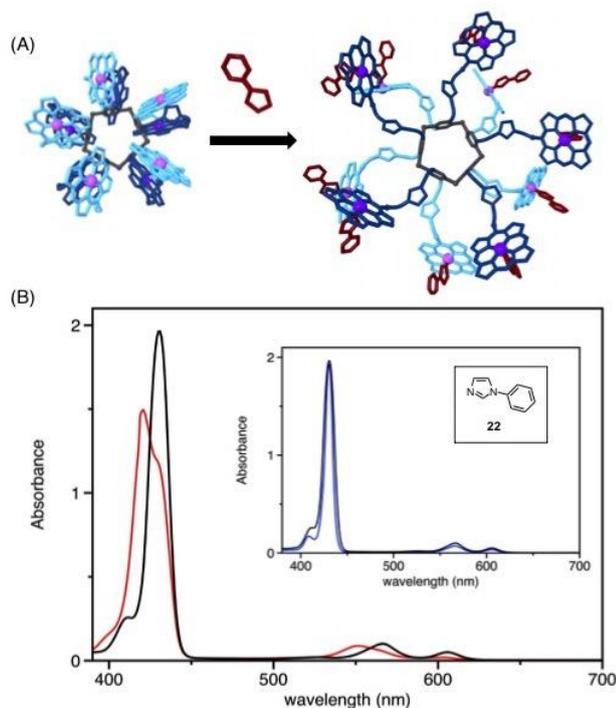


Figure 5. (A) Schematic representation of the blooming of molecular flower **20** upon addition of 1-phenylimidazole (**22**). (B) Absorption spectra of compound **20** before (red) and after (black) addition of 14 equiv. of **22** (CH_2Cl_2 , 25°C , $[\mathbf{20}] = 0.48 \mu\text{M}$); inset: comparison of the absorption spectra of pillar[5]arene **20** (black) and model porphyrin **16** (blue) in the presence of 14 equiv. of **22** (CH_2Cl_2 , 25°C , $[\mathbf{20}] = 0.48 \mu\text{M}$, $[\mathbf{16}] = 3.6 \mu\text{M}$).

The Soret band is red-shifted and becomes narrow. A similar effect is also observed for the Q bands. These are diagnostic signatures for an effective apical coordination of **22** onto the Zn(II)-porphyrin subunits of **20**. Importantly, the absorption spectrum of **20** upon addition of **22** corresponds well to the one recorded for model porphyrin **16** in the presence of imidazole **22** (Figure 5B, inset) thus suggesting that the intramolecular coordination of Zn(II)-porphyrin moieties by 1,2,3-triazole subunits is not effective anymore. Coordination of the Zn(II)-porphyrins of compound **20** by imidazole ligand **22** is thus capable of preventing the coordination-induced folding in the multi-Zn(II)-porphyrin array. As schematically shown in Figure 5A, addition of ligand **22** to **20** allows for a complete unfolding of the compound thus mimicking the blooming of a flower.

In order to further evidence this blooming event upon addition of imidazole **22**, ^1H NMR binding studies were also carried out. For both **18** and **20**, dramatic changes were observed in the ^1H NMR spectra recorded after addition of imidazole **22**. This is shown in Figure 6 for compound **18**.

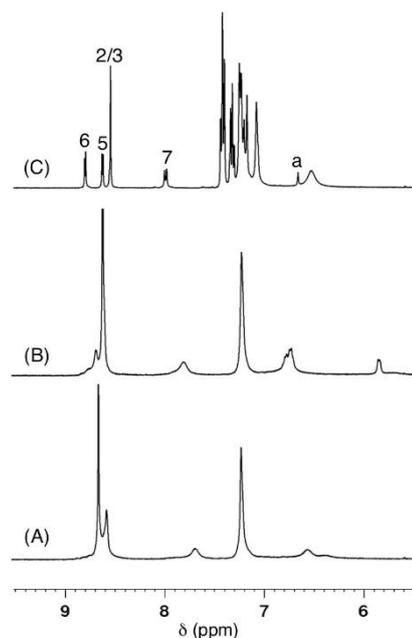


Figure 6. Aromatic region of the ^1H NMR spectra of pillar[5]arene **18** (400 MHz, CD_2Cl_2 , 25°C) recorded before (A) and after addition of 5 (B), and 45 (C) equivalents of imidazole **22** (for the numbering, see Fig. 1).

The broad signals of the pillar[5]arene derivative are sharpened by adding increasing amount of ligand **22**. This clearly shows that the dynamic folding at the origin of the broadening of the ^1H NMR spectrum of **18** is progressively prevented by coordination of imidazole ligands **22** to the Zn(II) centers of **18**. In other words, the intermolecular coordination of **22** to the Zn(II)-porphyrin subunits of **18** is strong enough to prevent the intramolecular coordination of the 1,2,3-triazole moieties in the multiporphyrin array. As a result, addition of an excess of **22** promotes the blooming of molecular flower **18**. A similar behavior was also evidenced for derivative **20** (Figure S16b). In both cases, the blooming event was also supported by DOSY experiments. Effectively, significant changes in the D values were observed upon addition of ligand **22** to CD_2Cl_2 solutions of **18** and **20** (see Table S2). In the presence of an excess of ligand **22**, the D values were decreased by $13.0(\pm 5)$ (**18**) and $14.0(\pm 6)\%$ (**20**). These changes in D values correspond to a coordination-induced volume expansion of $46(\pm 12)$ (**18**) and $49(\pm 15)\%$ (**20**). Interestingly, the D values deduced for **18** and **20** in the presence of an excess of **22** ($2.2(\pm 0.1)$ and $2.1(\pm 0.1) \times 10^{-10} \text{ m}^2 \cdot \text{s}^{-1}$ respectively) are very similar to those recorded for the corresponding free-base porphyrin derivatives **19** and **21** ($2.2(\pm 0.1)$ and $2.2(\pm 0.1) \times 10^{-10} \text{ m}^2 \cdot \text{s}^{-1}$, respectively) thus providing good evidence for similar unfolded structures.

Electrochemical properties. The electrochemical properties of pillar[5]arenes **18-21** were determined by cyclic voltammetry (CV) and Osteryoung Square Wave Voltammetry (OSWV).^[16] For the sake of comparison, electrochemical measurements have been also carried out with model Zn(II)-porphyrin derivatives **9** and **16**. All the experiments were performed at room temperature in CH_2Cl_2 solutions containing tetra *n*-butylammonium tetrafluoroborate (0.10 M) as supporting electrolyte and ferrocene (Fc) as internal reference, with a Pt wire as the working electrode and a saturated calomel electrode

(SCE) as a reference. Potential data for all of the compounds are collected in Table 1.

Table 1. Electrochemical data of compounds **9**, **16** and **18-21** determined by OSWV on a Pt working electrode in $\text{CH}_2\text{Cl}_2 + 0.1\text{M } n\text{-Bu}_4\text{NBF}_4$ at room temperature. Ferrocene is used as internal reference.^[a,b]

	E_3^{ox}	E_2^{ox}	E_1^{ox}	E_1^{Red}	E_2^{Red}	E_3^{Red}	E_4^{Red}
19		+1.66	+1.06	-1.23	-1.60	-	
21		+1.70	+1.06	-1.24	-1.60		
9		+1.18	+0.85	-1.40	-1.61	-1.82	
16		+1.18	+0.85	-1.41	-1.83		
18	+1.29	+0.92	+0.73	-1.41	-1.59	1.88	
20	+1.24	+0.89	+0.72	-1.43	-1.54	-1.65	-1.94

OSWVs were obtained using a sweep width of 20 mV, a frequency of 20 Hz, and a step potential of 5 mV. [b] Potential values in oxidation (E_{ox}) and in reduction (E_{red}) in Volt vs. SCE (Fc^+/ Fc is observed at $0.55 \text{ V} \pm 0.01 \text{ V}$ vs. SCE).

Free-base porphyrin-substituted pillar[5]arenes **19** and **21** display the diagnostic electrochemical signature of H_2 -porphyrin derivatives.^[17] The first one-electron transfer is reversible both in oxidation and in reduction, while the second one is at most only partially reversible as deduced from their CVs (see ESI). The redox potentials are similar for both compounds thus indicating that the length of the spacer has no significant influence on their electrochemical properties. Moreover the observed electrochemical processes are due to the simultaneous oxidation or reduction of ten independent free-porphyrin units in a similar environment. As reported for dendrimers bearing multiple peripheral redox centers, this indicates that electrostatic interactions between porphyrinic cations or anions are shielded by the supporting electrolyte.^[18]

The typical electrochemical features of Zn(II)-porphyrin systems are clearly observed for model compounds **9** and **16**. When compared to **19** and **21**, all the redox processes are cathodically shifted. In contrast, the electrochemical behavior of multi-Zn(II)-porphyrin arrays **18** and **20** is significantly different (see ESI). As a typical example, the CVs recorded for pillar[5]arene **20** and the corresponding model compound (**16**) are depicted in Figure 7.

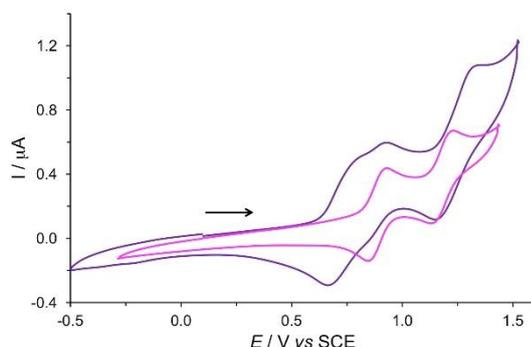
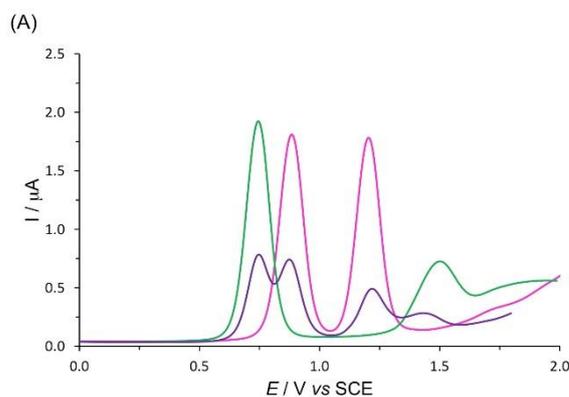


Figure 7. Cyclic voltammograms of compounds **16** (pink) and **20** (purple) on a Pt electrode in $\text{CH}_2\text{Cl}_2 + 0.1 \text{ M } [n\text{Bu}_4\text{N}][\text{BF}_4]$ at a scan rate of 200 mV s^{-1} (toward anodic potentials).

Three characteristic oxidation waves are clearly detected for compounds **18** and **20** in OSWVs and CVs. The first one ($E_{1\text{ox}}$) is cathodically shifted when compared to model Zn(II)-porphyrins **9** and **16** (E_1^{ox}), while the second one observed around 0.90 V (E_2^{ox}) is rather close to the oxidation potentials of the model compounds. The third one (E_3^{ox}) conversely is slightly anodically shifted. Similarly, the reduction of compounds **18** and **20** also involves an additional process when compared to the corresponding model Zn(II)-porphyrins (see ESI). In light of the NMR and UV-Vis spectroscopic data, it is believed that the origin of the supplementary oxidation and reduction processes observed for **18** and **20** must be related to the existence of intramolecular 1,2,3-triazole-porphyrin interactions in these compounds. Indeed, these interacting Zn(II)-porphyrin units are more electron-rich and should be at the same time easier to oxidize and more difficult to reduce. In order to confirm this hypothesis, further electrochemical investigations were carried out in the presence of imidazole ligand **22**. The OSWVs recorded for model compound **16** in the presence of **22** are shown in Figure 8A (anodic scan, for the cathodic scan, see the ESI). Upon addition of 1 equiv. of imidazole **22**, the oxidation waves of both coordinated and uncoordinated Zn(II)-porphyrins are simultaneously observed. Saturation is reached upon addition of 2 equiv. and only the electrochemical signature of [(**16**)(**22**)] is then detected. When compared to **16**, the first oxidation wave is cathodically shifted in [(**16**)(**22**)] while the second one is anodically shifted. This indicates that the radical cation of the Zn-porphyrin is stabilized upon complexation by **22**. In contrast, while the potential of the first reduction wave is significantly affected by the presence of **22**, the second reduction is always observed at the same redox potential. This latter observation suggests a decoordination of **22** upon the first reduction.



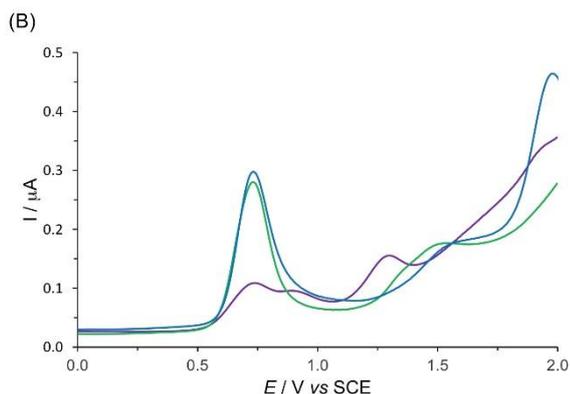


Figure 8. (A) OSWVs (anodic scan) of compound **16** before (pink) and after successive additions of **22** [1 (purple) and 2 (green) equiv.] and (B) compound **18** before (purple) and after successive additions of **22** [5 equiv. (green) and 10 equiv. (blue)] on a Pt electrode in CH_2Cl_2 + 0.1 M $[\text{nBu}_4\text{N}][\text{BF}_4]$ at room temperature (frequency 20 Hz, amplitude 20 mV, step potential 5 mV).

All these observations nicely confirm that the peculiar electrochemical behavior observed for the multiporphyrinic arrays **18** and **20** results effectively from the partial intramolecular coordination of Zn(II)-porphyrin moieties with 1,2,3-triazoles. In both cases, the relative intensity of the two first oxidation waves shows that at least 5 out of the 10 porphyrins are complexed by a triazole unit under these conditions (CH_2Cl_2 , room temperature), but not all of them. We also noticed that for both **18** and **20**, the ratio between the two first oxidation peaks is unchanged while switching the scan rate between 0.1 and 1 Vs^{-1} . This indicates that the dynamic interconversion process between coordinated and uncoordinated porphyrins does not occur within the voltammetric timescale.^[19] Moreover, as decoordination seems to occur upon the first reduction, an electrochemical input appears thus capable of triggering the blooming of molecular flowers **18** and **20**. This hypothesis is further reinforced by the only partial reversibility of the first reduction wave of **18** and **20** in agreement with a square scheme mechanism.^[20]

Finally, CVs and OSWVs of compound **18** and **20** were also recorded in the presence of imidazole **22**. As a typical example, the OSWVs recorded for **18** upon addition of **22** are depicted in Figure 8B. The addition of ligand **22** to polytopic receptor **18** is responsible for a simplification of the voltammogram and only the typical electrochemical signature of coordinated Zn(II)-porphyrin subunits is observed. The observed electrochemical processes are now arising from the simultaneous oxidation of the ten independent peripheral porphyrin moieties that are all in a similar state. We also notice that only 5 equiv. are necessary to switch ensure a total coordination of the ten porphyrin units, in agreement with the CVs of **18** and **20** without **22**.

Conclusions

Pillar[5]arene derivatives bearing peripheral porphyrin subunits have been efficiently prepared from a clickable pillar[5]arene building block (**17**) and Zn(II)-porphyrin derivatives bearing a terminal alkyne function (**9** and **16**). The characterization of the resulting deca-Zn(II)-porphyrin arrays (**18** and **20**) has been complicated owing to an intramolecular complexation of the

peripheral Zn(II)-porphyrin moieties by 1,2,3-triazole subunits. As a result of this intramolecular coordination, the ^1H NMR spectra of **18** and **20** are broad at room temperature. Variable temperature NMR studies have shown that the conformational equilibrium resulting from the coordination-driven molecular folding is dynamic and well-resolved ^1H NMR spectra have been recorded at high temperature. Actually, the thermodynamic equilibrium is displaced in favor of uncoordinated species at high temperature and molecules **18** and **20** adopt an open conformation similar to the one of their free-base porphyrin analogues (**19** and **21**) under these conditions. In contrast, coordination of the Zn(II)-porphyrin moieties by 1,2,3-triazole groups is favored at low temperature and thus the molecules adopt a folded conformation. Finally, we have shown that the coordination-driven folding of **18** and **20** can be controlled by an external chemical stimulus. Specifically, addition of an imidazole derivative (**22**) to solution of **18** or **20** prevents the intramolecular coordination at the origin of the folding. As a result, unfolding of molecules **18** and **20** occurs. This view is further supported by a variation of the diffusion coefficient as measured by DOSY and by the electrochemical signature. The resulting molecular motions triggered by the addition of the imidazole ligand mimics the blooming of a flower.

Experimental Section

Synthesis.

The preparation and characterization (Figures S1-14) of all the new compounds are described in the ESI.

Electrochemistry.

The electrochemical properties of compounds **9**, **16**, **18-21** were determined by cyclic voltammetry (CV) and Osteryoung Square Wave Voltammetry (OSWV) (See experimental selected curves in SI). The cyclic voltammetric measurements were carried out with a potentiostat Autolab PGSTAT100. Experiments were performed at room temperature in a homemade airtight three-electrode cell connected to a vacuum/argon line. The reference electrode consisted of a saturated calomel electrode (SCE) separated from the solution by a bridge compartment. The counter electrode was a platinum wire of ca. 1 cm^2 apparent surface. The working electrode was a Pt microdisk (0.5 mm diameter). The supporting electrolyte $[\text{nBu}_4\text{N}][\text{BF}_4]$ (Fluka, 99% electrochemical grade) was used as received and simply degassed under argon. Dichloromethane was freshly distilled over CaH_2 prior to use. The solutions used during the electrochemical studies were typically 2×10^{-4} or 3×10^{-3} M in compound and 0.1 M in supporting electrolyte. Before each measurement, the solutions were degassed by bubbling Ar and the working electrode was polished with a polishing machine (Presi P230). Under these experimental conditions, Fc^+/Fc is observed at $+0.55 \pm 0.01$ V vs. SCE. OSWVs were obtained using an amplitude of 20 mV, a frequency of 20 Hz, and a step potential of 5 mV. For addition experiments on **16** and **18**, a CH_2Cl_2 solution of 0.5 mL of compound **22** was prepared under argon, and the adjusted volume was syringed into the stirred electrochemical mixture under argon atmosphere.

Acknowledgements

Financial support by the International Center for Frontier Research in Chemistry, the LabEx "Chimie des Systèmes Complexes", the CNRS and the University of Strasbourg is gratefully acknowledged. TMNT thanks

the University of Strasbourg (IdEx 2013) and HBA the Tunisian Ministry of Higher Education and Scientific Research and Technology for their fellowships. We thank Dr. A. Saquet (LCC) for electrochemical measurements, Drs. M. Schmitt and E. Wasielewski (UMR 7509) for high-field NMR measurements and Dr. J.-M. Strub (ECPM) for the mass spectra. We further thank Drs. R. Caballero and M. Vartanian for the initial preparation of compound **16**.

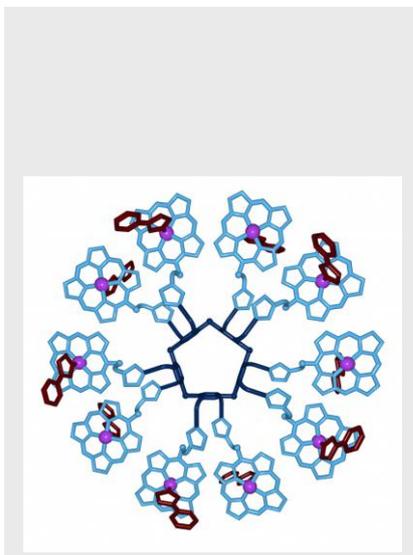
Keywords: Pillararene • Porphyrin • CuAAC • Molecular machine • Molecular folding

- [1] For reviews on pillar[*n*]arenes, see: (a) P. J. Cragg, K. Sharma, *Chem. Soc. Rev.* **2012**, *41*, 597-607. (b) M. Xue, Y. Yang, X. Chi, Z. Zhang, F. Huang, *Acc. Chem. Res.* **2012**, *45*, 1294-1308. (c) T. Ogoshi, *J. Incl. Phenom. Macrocycl. Chem.* **2012**, *72*, 247-262. (d) T. Ogoshi, T.-a. Yamagishi, *Eur. J. Org. Chem.* **2013**, 2961-2975. (e) D. Cao, H. Meier, *Asian J. Org. Chem.* **2014**, *3*, 244-262. (f) T. Ogoshi, T. Yamagishi, *Chem. Commun.* **2014**, *50*, 4776-4787. (g) N. L. Strutt, H. Zhang, S. T. Schneebeli, J. F. Stoddart, *Acc. Chem. Res.* **2014**, *47*, 2631-2642. (h) K. Yang, Y. Pei, Z. Pei, *Chem. Commun.* **2016**, *52*, 9316-9326. (i) T. Ogoshi, T.-a. Yamagishi, Y. Nakamoto, *Chem. Rev.* **2016**, *116*, 7937-8002. (j) T. Kakuta, T. Yamagashi, T. Ogoshi, *Chem. Commun.* **2017**, DOI: 10.1039/C7CC01833A.
- [2] G. W. Gribble, C. F. Nutaitis, *Tetrahedron Lett.* **1985**, *26*, 6023-6026.
- [3] T. Ogoshi, S. Kanai, S. Fujinami, T.-a. Yamagishi, Y. Nakamoto, *J. Am. Chem. Soc.* **2008**, *130*, 5022-5023.
- [4] (a) T. Ogoshi, T. Aoki, K. Kitajima, S. Fujinami, T.-a. Yamagishi, Y. Nakamoto, *J. Org. Chem.* **2011**, *76*, 328-331. (b) D. Cao, Y. Kou, J. Liang, Z. Chen, L. Wang, H. Meier, *Angew. Chem. Int. Ed.* **2009**, *48*, 9721-9723. (c) Y. Kou, H. Tao, D. Cao, Z. Fu, D. Schollmeyer, H. Meier, *Eur. J. Org. Chem.* **2010**, 6464-6470. (d) Y. Ma, Z. Zhang, X. Ji, C. Han, J. He, Z. Abliz, W. Chen, F. Huang, *Eur. J. Org. Chem.* **2011**, 5331-5335. (e) T. Boinski, A. Szumma, *Tetrahedron* **2012**, *68*, 9419-9422. (f) M. Holler, N. Allenbach, J. Sonet, J.-F. Nierengarten, *Chem. Commun.* **2012**, *48*, 2576-2578.
- [5] (a) I. Nierengarten, K. Buffet, M. Holler, S. P. Vincent, J.-F. Nierengarten, *Tetrahedron Lett.* **2013**, *54*, 2398-2402. (b) G. Yu, Y. Ma, C. Han, Y. Yao, G. Tang, Z. Mao, C. Gao, F. Huang, *J. Am. Chem. Soc.* **2013**, *135*, 10310-10313. (c) K. Buffet, I. Nierengarten, N. Galanos, E. Gillon, M. Holler, A. Imberty, S. E. Matthews, S. Vidal, S. P. Vincent, J.-F. Nierengarten, *Chem. Eur. J.* **2016**, *22*, 2955-2963. (d) S. P. Vincent, K. Buffet, I. Nierengarten, A. Imberty, J.-F. Nierengarten, *Chem. Eur. J.* **2016**, *22*, 88-92. (e) N. Galanos, E. Gillon, A. Imberty, S. E. Matthews, S. Vidal, *Org. Biomol. Chem.* **2016**, *14*, 3476-3481.
- [6] (a) I. Nierengarten, M. Nothisen, D. Sigwalt, T. Biellmann, M. Holler, J.-S. Remy, J.-F. Nierengarten, *Chem. Eur. J.* **2013**, *19*, 17552-17558. (b) K. Yang, Y. Chang, J. Wen, Y. Lu, Y. Pei, S. Cao, F. Wang, Z. Pei, *Chem. Mater.* **2016**, *28*, 1990-1993.
- [7] (a) I. Nierengarten, S. Guerra, M. Holler, J.-F. Nierengarten, R. Deschenaux, *Chem. Commun.* **2012**, *48*, 8072-8074. (b) I. Nierengarten, S. Guerra, M. Holler, L. Karmazin-Brelot, J. Barbera, R. Deschenaux, J.-F. Nierengarten, *Eur. J. Org. Chem.* **2013**, 3675-3684. (c) S. Pan, M. Ni, B. Mu, Q. Li, X.-Y. Hu, C. Lin, D. Chen, L. Wang, *Adv. Funct. Mater.* **2015**, *25*, 3571-3580. (d) I. Nierengarten, S. Guerra, H. Ben Aziza, M. Holler, R. Abidi, J. Barberá, R. Deschenaux, J.-F. Nierengarten, *Chem. Eur. J.* **2016**, *22*, 6185-6189.
- [8] (a) B. J. Littler, Y. Ciringhi, J. S. Lindsey, *J. Org. Chem.* **1999**, *64*, 2864-2872. (b) G. R. Geier III, B. J. Littler, J. S. Lindsey, *J. Chem. Soc. Perkin Trans. 2* **2001**, 701-711. (c) R. W. Wagner, D. S. Lawrence, J. S. Lindsey, *Tetrahedron Lett.* **1987**, *28*, 3069-3070. (d) J. S. Lindsey, R. W. Wagner, *J. Org. Chem.* **1989**, *54*, 828-836.
- [9] (a) X. Huang, K. Nakanishi, N. Berova, *Chirality* **2000**, *12*, 237-255. (b) T. Hashimoto, Y.-C. Choe, H. Nakano, K. Hirao, *J. Phys. Chem. A* **1999**, *103*, 1894-1904. (c) M. Gouterman, *J. Chem. Phys.* **1959**, *30*, 1139-1161.
- [10] J. Bhaumik, Z. Yao, K. Eszter Borbas, M. Taniguchi, J. S. Lindsey, *J. Org. Chem.* **2006**, *71*, 8807-8817.
- [11] H. C. Kolb, M. G. Finn, K. B. Sharpless, *Angew. Chem. Int. Ed.* **2001**, *40*, 2004-2021.
- [12] (a) I. Nierengarten, J.-F. Nierengarten, *Chem. Rec.* **2015**, *15*, 31-51. (b) I. Nierengarten, J.-F. Nierengarten, *Chem Asian J.* **2014**, *9*, 1436-1444.
- [13] A. Matea-Alonso, C. Sooambar, M. Prato, *C. R. Chimie* **2006**, *9*, 944-951 and references therein.
- [14] (a) N. T. Nguyen, G. M. Mamardashvili, O. M. Kulikova, I. G. Scheblykin, N. Z. Mamardashvili, W. Dehaen, *RSC Adv.* **2014**, *4*, 19703-19709. (b) C. Maeda, P. Kim, S. Cho, J. K. Park, J. M. Lim, D. Kim, J. Vura-Weis, M. R. Wasielewski, H. Shinokubo, A. Osuka, *Chem. Eur. J.* **2010**, *16*, 5052-5061. (c) C. Maeda, S. Yamaguchi, C. Ikeda, H. Shinokubo, A. Osuka, *Org. Lett.* **2008**, *10*, 549-552. (d) J.-P. Collin, F. Durola, V. Heitz, F. Reviriego, J.-P. Sauvage, Y. Trolez, *Angew. Chem. Int. Ed.* **2010**, *49*, 10172-10175.
- [15] (a) N. Armaroli, F. Diederich, L. Echegoyen, T. Habicher, L. Flamigni, G. Marconi, J.-F. Nierengarten, *New J. Chem.* **1999**, *23*, 77-83. (b) W.-S. Li, K. S. Kim, D.-L. Jiang, H. Tanaka, T. Kawai, J. H. Kwon, D. Kim, T. Aida, *J. Am. Chem. Soc.* **2006**, *128*, 10527-10532. (c) K. Yoosaf, J. Iehl, I. Nierengarten, M. Hmadeh, A.-M. Albrecht-Gary, J.-F. Nierengarten, N. Armaroli, *Chem. Eur. J.* **2014**, *20*, 223-231.
- [16] R. G. Compton, C. E. Banks, *Understanding Voltammetry*, Imperial College Press, London, **2011**.
- [17] (a) K.M. Kadish, L.R. Shiue, R.K. Rhodes, L.A. Bottomley, *Inorg. Chem.* **1981**, *20*, 1274-1277. (b) K.M. Kadish, R.K. Rhodes, *Inorg. Chem.* **1981**, *20*, 2961-2966. (c) C.-L. Lin, M.-Y. Fang, S.-H. Cheng, *J. Electroanal. Chem.* **2002**, *531*, 155-162. (d) F. D'Souza, P.M. Smith, L. Rogers, M.E. Zandler, E. Melvin, D.-M. S. Islam, Y. Araki, O. Ito, *Inorg. Chem.* **2006**, *45*, 5057-5065. (e) C.-Y. Huang, C.-Y. Hsu, L.-Y. Yang, C.-J. Lee, T.-F. Yang, C.-C. Hsu, K., C.-H. Ke, Y. O. Su, *Eur. J. Inorg. Chem.* **2012**, 1038-1047. (f) H. C. Cheng, P. P.Y. Chen, Su, Y. O., *Dalton Trans.* **2014**, *43*, 1424-1433.
- [18] (a) C. Amatore, Y. Bouret, E. Maisonhaute, J. I. Goldsmith, H. D. Abruna, *ChemPhysChem.* **2001**, *2*, 130-134. (b) C. Amatore, Y. Bouret, E. Maisonhaute, J. I. Goldsmith, H. D. Abruna, *Chem. Eur. J.* **2001**, *7*, 2206-2226. (c) C. Amatore, F. Grun, E. Maisonhaute, *Angew. Chem. Int. Ed.* **2003**, *42*, 4944-4947. (d) J. A. Camerano, M. A. Casado, U. Hahn, J. F. Nierengarten, E. Maisonhaute, C. Amatore, *New J. Chem.* **2007**, *31*, 1395-1399. (e) U. Hahn, E. Maisonhaute, C. Amatore, J. F. Nierengarten, *Angew. Chem. Int. Ed.* **2007**, *46*, 951-954. (f) P. Fortgang, E. Maisonhaute, C. Amatore, B. Delavaux-Nicot, J. Iehl, J. F. Nierengarten, *Angew. Chem. Int. Ed.* **2011**, *50*, 2364-2367. (g) P. Fortgang, M. Urbani, M. Holler, J.-F. Nierengarten, A. Moreau, B. Delavaux-Nicot, E. Maisonhaute, *Electroanalysis* **2015**, *27*, 1010-1016. (h) X. S. Zhou, B. W. Mao, C. Amatore, R. G. Compton, J. L. Marignier, M. Mostafavi, J. F. Nierengarten, E. Maisonhaute, *Chem. Commun.* **2016**, *52*, 251-263.
- [19] J.-M. Saveant, *Elements Molecular and Biomolecular Electrochemistry*, John Wiley and Son, Hoboken, New Jersey, **2006**.
- [20] C. Amatore, D. Genovese, E. Maisonhaute, N. Raouafi, B. Schöllhorn, *Angew. Chem. Int. Ed.* **2008**, *47*, 5211-5214.

Entry for the Table of Contents

FULL PAPER

Molecular flowers. Deca-Zn(II)-porphyrin arrays constructed on a pillar[5]arene scaffold adopt a folded conformation due to intramolecular complexation of the Zn(II)-porphyrin moieties by 1,2,3-triazole subunits. An external chemical stimulus can trigger the unfolding of these compounds thus mimicking the blooming of a flower.



Thi Minh Nguyet Trinh, Iwona Nierengarten, Haifa Ben Aziza, Eric Meichsner, Michel Holler, Matthieu Chessé, Rym Abidi, Christian Bijani, Yannick Coppel, Emmanuel Maisonhaute, Béatrice Delavaux-Nicot,* and Jean-François Nierengarten**

Coordination-driven folding in multi-Zn(II)porphyrin arrays constructed on a pillar[5]arene scaffold