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Protonated ions as systemic trapping agents for noble gases: From electronic structure to radiative association

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The deficiencies of argon, krypton, and xenon observed in the atmosphere of Titan as well as anticipated in some comets might be related to a scenario of sequestration by H_3^+ in the gas phase at the early evolution of the solar nebula. The chemical process implied is a radiative association, evaluated as rather efficient in the case of H_3^+ , especially for krypton and xenon. This mechanism of chemical trapping might not be limited to H_3^+ only, considering that the protonated ions produced in the destruction of H_3^+ by its main competitors present in the primitive nebula, i.e., H_2O , CO , and N_2 , might also give stable complexes with the noble gases. However the effective efficiency of such processes is still to be proven. Here, the reactivity of the noble gases Ar, Kr, and Xe, with all protonated ions issued from H_2O , CO , and N_2 , expected to be present in the nebula with reasonably high abundances, has been studied with quantum simulation method dynamics included. All of them give stable complexes and the rate coefficients of their radiative associations range from 10^{-16} to 10^{-19} $\text{cm}^3 \text{s}^{-1}$, which is reasonable for such reactions and has to be compared to the rates of 10^{-16} to 10^{-18} $\text{cm}^3 \text{s}^{-1}$, obtained with H_3^+ . We can consider this process as universal for all protonated ions which, if present in the primitive nebula as astrophysical models predict, should act as sequestration agents for all three noble gases with increasing efficiency from Ar to Xe. *Published by AIP Publishing.* <https://doi.org/10.1063/1.4994630>

I. INTRODUCTION

The problem of the noble gases' deficiency observed in Titan's atmosphere¹ has led to tentative explanations through several models, none of them fully satisfactory. Theoretical quantum chemistry, with high level *ab initio* simulations, demonstrated that, at least from a thermodynamic point of view, H_3^+ is able to act as a trap for noble gases by forming complexes in the gas phase.^{2,3} The coupling of these results with those of an elaborate model giving the abundance profile of H_3^+ in the conditions of the cooling proto-planetary disks⁴⁻⁶ led to the conclusion that H_3^+ might be abundant enough in the outer protosolar nebula to trap, at least Xe and Kr, prior to their condensation epochs, implying that their abundances should be solar in Saturn's current atmosphere and below the observational limit in Titan. The same scenario predicts that comets agglomerated from ices formed at high heliocentric distances might also be depleted in Kr and Xe.

In the astrophysical models, the crucial point is the accurate determination of the abundance of H_3^+ , which implies evaluating not only its production but also its destruction along the cooling evolution of the proto-planetary disk. That is why a complete modeling should take into account the molecules

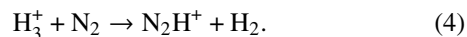
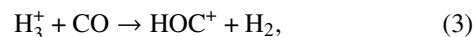
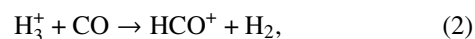
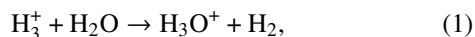
able to react with H_3^+ during this period and the evolution of their abundances with time.

From a chemical point of view, when we looked at the main destructive agent for H_3^+ , i.e., H_2O , we found that the result of the destructive reaction, i.e., H_3O^+ , could also act as a trapping agent for the noble gases, leading to chemically stable complexes⁶ via radiative associations with the noble gases. Then we have to consider the possibility of a similar behavior for all protonated ions, among which those issued from the other main destructors of H_3^+ present in the environment. Nevertheless, to be conclusive, it implies performing a complete quantitative study of these radiative associations and in particular to evaluate reliable rate constants of these reactions which belong to a category usually considered not to be very efficient.

In the solar nebula, the H_3^+ formed can be destroyed by the reactions with the most abundant molecules of the gas phase, namely, H_2O , CO , and N_2 .⁶ In a steady state, formation and destruction rates balance each other, but when the solar nebula cools down, the destructive molecules do not remain in the gas phase (the upper limit being a local temperature of 150 K), being trapped themselves as clathrates or simply condensing, and the balance has to be adjusted. However, if these molecules are reacting with H_3^+ to give ions with a significant capability for trapping the noble gases, most adjustments of the density of H_3^+ could be avoided, each H_3^+ destroyed being replaced

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by another trapping ion. This can be argued safely as long as the trapping capabilities are more or less equivalent. In these reactions, H_3^+ plays the role of a universal proton donor; the reactions implied in such a scheme are

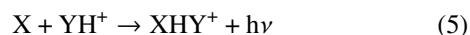


All these ions are well-known and ubiquitous in astrophysical objects. Looking at the reported original detections in the interstellar medium (ISM) and circumstellar envelopes, we can cite⁷ HCO^+ first seen in 1970 and confirmed five years later,⁸ HN_2^+ ,^{9,10} and HOC^+ .^{11,12} We can also mention further detections of HCO^+ ,¹³ HN_2^+ ,¹⁴ and HOC^+ .¹⁵ The detection of H_3O^+ has also been reported in the ISM.^{16,17} Moreover, all these reactions [Eqs. (1)–(4)] are known to be thermodynamically (exothermic) and dynamically efficient (<http://www.rate99.co.uk/>).

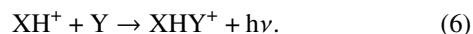
In the solar system, the H_3^+ emission has been recorded towards the auroral zone of Jupiter, soon followed by its detection in giant planets Saturn and Uranus,¹⁸ whereas it took ten more years to succeed in the detection of H_3^+ and its isotopologs^{19,20} in the ISM.

The role of H_3^+ itself, as a possible trapping agent for noble gases, has already been fully investigated,²¹ but one has to be aware that till now this scenario has never been quantified for any other trapping agent beyond a static description of the trapping processes.

In a general way, two radiative association processes would lead to the XHY^+ complex (with $\text{X} = \text{Ar}, \text{Kr}, \text{Xe}$ and $\text{Y} = \text{H}_2\text{O}, \text{CO}, \text{N}_2$), namely,



and



Because the concentrations of the XH^+ species are so low (no detection has been reported apart from that of $^{36}\text{ArH}^+$ in a supernova remnant known as the Crab nebula),²² the formation of the XHY^+ complexes via [Eq. (6)] is assumed to be negligible and therefore the only route to be considered is the radiative association [Eq. (5)].

II. COMPUTATIONAL STRATEGY

To reduce the length of the story, only a brief overview of the key theoretical points of the static and dynamic aspects of this research is presented (see Appendices for more details). In the present study, we employ a fully *ab initio* approach, from the calculations of the reactive potential surfaces to the determination of the corresponding radiative association rate constants.

Concerning the thermodynamic part, all treatments belong to the *ab initio* coupled cluster area, i.e., CCSD and CCSDT associated with a correlation consistent basis set cc-pVQZ and its augmented versions (aug-) of increasing flexibility in the

valence shell. The calculations were performed with the G09 version of the Gaussian package.²³

- (1) The first task is to determine the structure of the complexes. All stationary points were optimized and the stability conditions were verified by a vibrational analysis.
- (2) In order to get reliable geometries and stabilities, the potential energy surface around the minima had to be reconstructed to correct the bias introduced by the artificial stabilization of the energy of such weakly bound complexes due to the fact that the interacting fragments borrow each other's basis sets to lower their energies. This artifact is known as the basis set superposition error (BSSE). It was corrected here using the counterpoise method,²⁴ for all points of the re-optimization of the geometrical parameters implied in the description of the two entities interacting in the complex. The difficulty is to determine the fragments to be considered when several separations are possible, a situation we have already faced with the complexes between H_3^+ and the noble gases. Usually we can refer to those of the lowest energy dissociation. However, when it was unclear which separation to use, we relied on a topological analysis (see Sec. IV and Appendix A) of the complex to get the partitioning of the physical space into well-defined interacting fragments.

For the dynamic part, the radiative association calculations are done in a pseudo-diatom approximation (see Appendix B). We relied on the previous computational experiments²¹ performed for the rate constants of the radiative associations of argon with H_3^+ .

III. STRUCTURE AND STABILITY OF THE PROTONATED IONS

Taking H_3^+ as a reference for the capture of noble gases X, we report some preceding results for the structure (Fig. 1) and binding energies (Table I) of the X-H_3^+ complexes.²

For all X-H_3^+ complexes, the most stable geometry is that of C_{2v} symmetry in which X points to an apex of the H_3^+ triangle (Fig. 1); this occurs at all levels of theory.

The convergence of the computational methods used allows us to consider the energetic calculations reliable by 1%–2% when reaching the CCSDT/aug-cc-pVQZ level.

A. The structures of protonated ions with noble gases $[\text{X-ion-H}]^+$ ($\text{X} = \text{Ar}, \text{Kr}, \text{and Xe}$)

All complexes, whatever the noble gas, are structurally similar (Fig. 2). The geometries presented in Table II show that the BSSE corrections, which could lead to changes of 1%–2% in the most influenced bond lengths, cannot be escaped.



FIG. 1. The stable structure of $\text{X} \cdots \text{H}_3^+$ for $\text{X} = \text{Ar}, \text{Kr}, \text{and Xe}$.

TABLE I. Stabilities of the $X \cdots H_3^+$ complexes (kcal/mol), including harmonic zero point energies^a and BSSE correction, related to reactions R1 and R2.^b Basis sets: /Q = cc-pVQZ; /aug-Q = aug-cc-pVQZ.

Complex Reactions	Ar \cdots	H ₃ ⁺	Kr \cdots	H ₃ ⁺	Xe \cdots	H ₃ ⁺
	R1	R2	R1	R2	R1	R2
CCSD(T)/Q	7.27	16.01	13.03	8.60	23.22	2.93
CCSD(T)/aug-Q	7.35	16.16	13.50	8.54	23.64	2.78

^aCalculated at level CCSD(T)/Q.

^bR1 = $X \cdots H_3^+ \rightarrow X + H_3^+$ and R2 = $X \cdots H_3^+ \rightarrow XH^+ + H_2$.

This is especially true if spectroscopic data are to be calculated in order to help with the detection of such species. It should be emphasized that contrary to *a posteriori* corrections of the energy of the equilibrium structures, the optimization procedure reported here are carried out including an *a priori* correction of the BSSE artifact for all the geometrical parameters defining the molecular structures.

B. The binding energies of protonated ions with noble gases [X-ion-H]⁺ (X = Ar, Kr, and Xe)

The evolutions of the complex stabilities within the series Ar \rightarrow Kr \rightarrow Xe present the same trend: as for the XH_3^+ complexes, the stability decreases in the R1 fragmentation from Ar to Xe, whereas it increases for R2 (see Table III).

Looking at the energetic balance, one has the following:

- (1) For $X-H_3O^+$, we found that, contrary to the reference XH_3^+ , all complexes prefer a common dissociative process leading to $X + H_3O^+$ whatever the noble gas. The point to be noted is that, when comparing the stability values in Tables III and I, it is obvious that each H_3O^+ ion issued from the destruction of H_3^+ by H_2O in the gas

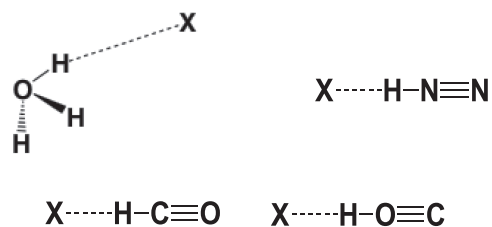


FIG. 2. Equilibrium structures of the X-protonated ions with X = Ar, Kr, and Xe.

phase should be able to play an almost equivalent role in the capture process of noble gases.

- (2) For $X-HCO^+$ and $X-HOC^+$, one has two different situations although they are isomers whose difference is only due to the attachment of the proton to one or the other extremity of the CO diatomic species. The case of $XHCO^+$ is very close to $X-H_3O^+$: for all X, R1 is the preferred dissociative channel with similar energies increasing from Ar to Xe. The case of $XHOC^+$ is different and closer to XH_3^+ . If the R1 dissociation is favored for Ar, it is the R2 dissociation that is favored for Xe. The case of Kr is very sensitive to the method used. Indeed, $X-HOC^+$ dissociates following the R1 route for Density Functional Theory (DFT) but following the R2 route for MP2 (not shown here). However, higher levels of theory, as CCSD and CCSDT, agree on this complex being intermediate between two cluster structures as shown below in the topological analysis of the electronic structure of this complex. Finally, confronting the values obtained for both complex families shows that their stabilities are equivalent to XH_3^+ reported in Table I (7.3, 8.5, 2.8 kcal), indicating that the trapping power of HCO^+

TABLE II. Raw and BSSE-optimized equilibrium geometries^a of the $X \cdots ion^+$ complexes (X = Ar, Kr, and Xe) and rotational constants in GHz at the CCSD(T)/cc-pVQZ level of theory and dipole moments in Debye at CCSD.

Complex	Raw-optimized geometry				BSSE-optimized geometry				A	B	C	μ
	X-H _X	H _X -O	H _O -O	H _O -O-H _O	X-H _X	H _X -O	H _O -O	H _O -O-H _O				
Ar $\cdots H_3O^+$	2.053	0.991	0.974	111.21	2.068	0.990	0.974	111.29	319.6	4.175	4.140	8.3
Kr $\cdots H_3O^+$	2.151	0.997	0.974	111.16	2.187	0.995	0.974	111.22	318.2	3.215	3.195	10.2
Xe $\cdots H_3O^+$	2.293	1.005	0.974	110.87	2.340	1.002	0.974	110.95	318.3	2.721	2.706	10.9
	X-H _X	H _X -C	C-O		X-H _X	H _X -C	C-O					
Ar $\cdots HCO^+$	2.110	1.112	1.110		2.133	1.111	1.110			1.973	1.973	5.3
Kr $\cdots HCO^+$	2.199	1.120	1.110		2.239	1.118	1.110			1.468	1.468	8.0
Xe $\cdots HCO^+$	2.316	1.135	1.111		2.372	1.130	1.110			1.244	1.244	9.2
	X-H _X	H _X -O	O-C		X-H _X	H _X -O	O-C					
Ar $\cdots HOC^+$	1.695	1.073	1.153		1.707	1.071	1.153			2.724	2.724	3.3
Kr $\cdots HOC^+$	1.737	1.135	1.146		1.739	1.139	1.151			2.035	2.035	4.8
Xe $\cdots HOC^+$	1.648	1.579	1.142		1.642	1.611	1.142			1.516	1.516	5.0
	X-H _X	H _X -N	N-N		X-H _X	H _X -N	N-N					
Ar $\cdots HNN^+$	1.881	1.076	1.096		1.893	1.075	1.096			2.383	2.383	4.2
Kr $\cdots HNN^+$	1.950	1.099	1.096		1.991	1.092	1.096			1.764	1.764	6.4
Xe $\cdots HNN^+$	2.049	1.142	1.097		2.057	1.145	1.097			1.508	1.508	7.1

^aBond lengths in Å; angles in deg.

TABLE III. Stabilities of the $X \cdots H_3^+$ complexes (kcal/mol), including harmonic zero point energies^a and BSSE correction, related to reactions R1 and R2.^b Basis sets: /Q = cc-pVQZ; /aug-Q = aug-cc-pVQZ.

Complex	[Ar-ion-H] ⁺		[Kr-ion-H] ⁺		[Xe-ion-H] ⁺	
	R1	R2	R1	R2	R1	R2
[ion-H] ⁺ = H ₃ O ⁺						
CCSD(T)/Q	4.24	78.47	5.80	67.19	7.85	52.53
CCSD(T)/aug-Q	4.43	77.55	6.00	65.84	8.12	50.76
[ion-H] ⁺ = HCO ⁺						
CCSD(T)/Q	4.31	54.54	5.81	43.21	7.99	28.66
CCSD(T)/aug-Q	4.53	54.78	5.73	42.71	8.26	28.04
[ion-H] ⁺ = HOC ⁺						
CCSD(T)/Q	10.82	23.70	15.16	15.21	23.60	6.92
CCSD(T)/aug-Q	10.90	23.86	15.47	15.16	24.59	7.07
[ion-H] ⁺ = HNN ⁺						
CCSD(T)/Q	7.74	34.37	10.32	24.11	14.26	11.33
CCSD(T)/aug-Q	7.92	34.27	10.27	23.35	15.47	11.34

^aCalculated at level CCSD(T)/Q.

^bR1 = $X \cdots H_3^+ \rightarrow X + H_3^+$ and R2 = $X \cdots H_3^+ \rightarrow XH^+ + H_2$.

and HOC⁺ could be at least as efficient as the one of XH₃⁺.

- (3) For HNN⁺ the situation is closer to that of HOC⁺ but with an inversion of the preference for the dissociative reaction only for xenon: Ar and Kr prefer to dissociate through the R1 mechanism, whereas xenon dissociates through the R2 route. The trend is the same, that is, an increase of the binding energies relative to R1 dissociation and a decrease for those relative to R2 when going from Ar to Xe, with a crossing of the values for Xe (although the energy difference between two routes is much smaller than for HOC⁺). According to the DFT method (not shown), XeHNN⁺ dissociates through R1, whereas for all the other post HF methods, it prefers R2. Similar to KrHOC⁺, this shows that high level methods are needed to describe these rather weakly bonded complexes.

IV. BONDING PICTURES IN THE NOBLE GAS COMPLEXES

Even though, most of the time, the fragments interacting in the complexes can be identified from their preferred dissociation mechanism, a better understanding of bonding pictures can be obtained using both the topological analysis of the electron localization function (ELF)^{25–27} and the quantum theory of atoms in molecules (QTAIM).^{28,29} As these methodologies have been widely discussed in the literature, they are only recalled in Appendix A. Introduced³⁰ by Bader and Stephens and later revisited³¹ by Fradera *et al.*, the delocalization index (DI) is a measure of the electron-sharing between two atoms and can be compared to other bond order indices. The relative percentage of the two resonance structures Ia and Ib, respectively, [X + ion⁺] and [X-H⁺ + neutral], has been evaluated using the DI of protonated bonds. For example, in

the case of X–H₃O⁺ complexes, Ia and Ib are calculated as follows:

$$I_a = 100 \left(\frac{DI_{OH_x}}{DI_{OH_x} + DI_{XH_x}} \right) \quad \text{and} \quad I_b = 100 - I_a.$$

H_X is here the hydrogen located in front of the noble gas. In this work, these results have also provided a rationale about the choice of the partition of the systems into fragments to be used for BSSE corrections, especially when there is no straightforward clue, such as for KrHOC⁺.

The ELF topological analysis applied previously to X_n(H₃⁺) (X = Ne, Ar, and Kr; n = 1,3) complexes has shown the crucial role of the charge transfer from the noble gases toward the H₃⁺ cation as a driving force in the bonding along the series. It was also consistent with the increase in the atomic polarizabilities from neon to krypton.³ In this work, a net positive charge q(X) also discloses a charge transfer from the noble gas towards the cation (H₃O⁺, HCO⁺, or HN₂⁺). However, the magnitude of q(X) remains small (<0.2 e), except for XeOC⁺ (0.45 e).

The ELF localization domains are displayed in Fig. 3. The ELF population analysis for all X–H₃O⁺, X–HCO⁺, X–HOC⁺, and X–HN₂⁺ complexes is reported in Tables IV–VII, respectively.

As observed in Fig. 3 and Table IV, the ELF topology of X–H₃O⁺ complexes shows two well-separated domains. The first one corresponds to the neutral X noble gas [q(X) remaining always lower than 0.1 electron]; it splits into a core basin C(X) and valence basins V(X). The second domain corresponds to H₃O⁺ where one O–H bond viewed as the protonated V(H_X, O) topological basin is stretched toward the noble gas. The population of this latter basin was found to be lower than 1.5 electrons, whereas the other V(O, H) populations remain close to 2 electrons. It should be noted that no covalent character was identified for the X–H₃O⁺ interaction since there are no disynaptic basins located between X and H. It is fully consistent with a large weight of around 75% for Ia in all these complexes. Thus, it gives a better understanding of the fragmentation used for the BSSE corrections for which the R2 route is energetically preferred.

The topological analysis of XHCO⁺ complexes (Table V) shows a very similar pattern to that of H₃O⁺ complexes, i.e., two well-separated domains around the noble gas and the HCO⁺ ion. Previously, the population of the protonated V(H_X, C) bonding basin remains lower than 2 electrons for all noble gases. The I_a and I_b weights have been found to be quite similar for both XH₃O⁺ and XHCO⁺ complexes. This confirms again the validity of the fragmentation into the noble gas and HCO⁺ moieties used for the BSSE corrections.

Except for ArHOC⁺ showing a very similar topology to other Ar species (I_a > I_b), the bonding picture of XHOC⁺ complexes differs from previous compounds (see Table VI). First, the charge q(X) increases drastically from 0.1 e (Ar) to 0.45 e (Xe). This is in agreement with the increase of the weight of [X-H]⁺ (I_b) in Kr and Xe complexes (see Tables V and VI). The ELF topology of KrHOC⁺ undergoes a clear modification since the V(H_X, O) basin splits into non-bonding basins V(H) and V(O). Thus, the weights of I_a and I_b configurations become very similar in agreement with the practically

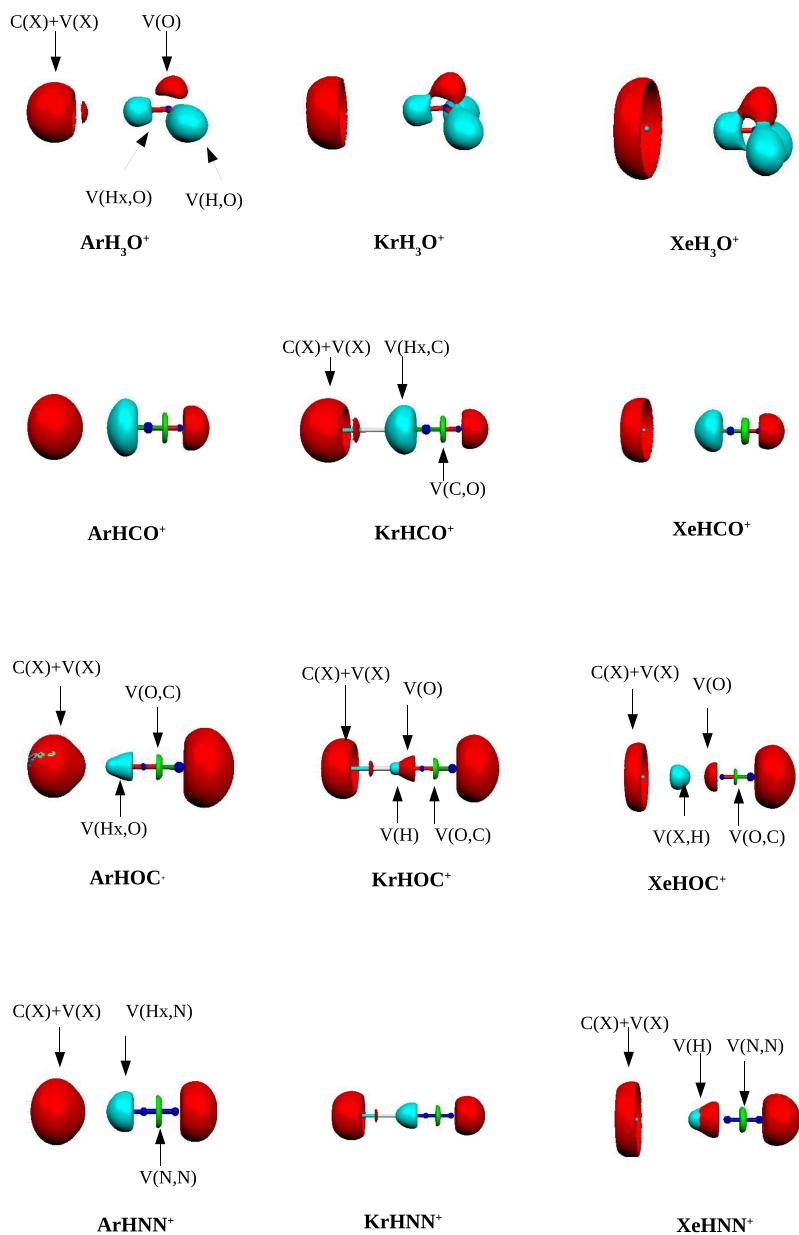


FIG. 3. View of the electron localization function (ELF = 0.85) isosurfaces (top to bottom) for $X\text{-H}_3\text{O}^+$, $X\text{-HCO}^+$, $X\text{-HOC}^+$, and $X\text{-HNN}^+$. Color code: magenta: core basins, red: non-bonding basins, green: bonding basins, light blue: protonated basins.

equal binding energies of 15.5 and 15.2 kcal/mol, respectively, obtained from the dissociation R1 and R2 routes. For XeHOC^+ , the $V(\text{H}_X, \text{O})$ basin disappears in favor of a new disynaptic protonated $V(\text{H}, \text{X})$ basin. Consequently, the weight of I_b becomes larger than I_a in agreement with the binding energies, respectively, obtained from R1 (24.6 kcal/mol) and R2 (7.1 kcal/mol) routes.

TABLE IV. ELF population analysis of $X\text{-H}_3\text{O}^+$ ($X = \text{Ar}, \text{Kr}, \text{and Xe}$). Populations of basins and QTAIM charges are in electrons.^a

$X\text{-H}_3\text{O}^+$	$C(X)+V(X)$	$V(\text{H}_X, \text{O})$	$q(X)$	I_a	I_b
$\text{Ar-H}_3\text{O}^+$	17.96	0.93	0.04	80	20
$\text{Kr-H}_3\text{O}^+$	35.94	1.29	0.06	76	24
$\text{Xe-H}_3\text{O}^+$	25.91	1.23	0.09	74	26
H_3O^+	...	1.86			

^aThe 28 innermost electrons of Xe are not treated explicitly but through an effective core potential.

Protonated N_2 , though isoelectronic with protonated CO isomers, offers another situation. For XHN_2^+ complexes (see Table VII), two well-separated domains are observed for Ar and Kr. The first domains contain the noble gas basins whereas the others are distributed over the HN_2^+ cation. Thus, the weight I_a remains larger than I_b . By contrast, the XeHN_2^+ bonding picture looks like KrHOC^+ since the protonated $V(\text{H}_X,$

TABLE V. ELF population analysis of $X\text{-HCO}^+$ ($X = \text{Ar}, \text{Kr}, \text{and Xe}$). Populations of basins and QTAIM charges are in electrons.^a

$X\text{-HCO}^+$	$C(X)+V(X)$	$V(\text{H}_X, \text{C})$	$q(X)$	I_a	I_b
Ar-HCO^+	17.96	2.36	0.04	86	14
Kr-HCO^+	35.93	2.38	0.07	82	18
Xe-HCO^+	25.90	2.39	0.10	79	21
HCO^+		2.21			

^aThe 28 innermost electrons of Xe are not treated explicitly but through an effective core potential.

TABLE VI. ELF population analysis of X–HOC⁺ (X = Ar, Kr, and Xe). Populations of basins and QTAIM charges are in electrons.^a

X-HOC ⁺	C(X)+V(X)	V(H _X ,O)	V(H _X)	V(O)	V(H _X , X)	q(X)	I _a	I _b
Ar–HOC ⁺	17.90	2.42	0.10	61	39
Kr–HOC ⁺	35.85	...	0.25	4.86	...	0.15	52	48
Xe–HOC ⁺	24.27	4.52	1.81	0.45	24	76
HOC ⁺		2.32						

^aThe 28 innermost electrons of Xe are not treated explicitly but through an effective core potential.

TABLE VII. ELF population analysis of X–HNN⁺ (X = Ar, Kr, and Xe). Populations of basins and QTAIM charges are in electrons.^a

X-HNN ⁺	C(X)+V(X)	V(H _X ,N)	V(H _X)	V(N)	q(X)	I _a	I _b
Ar–X–HNN ⁺	17.93	2.74	0.07	73	27
Kr–X–HNN ⁺	35.88	2.79	0.12	68	32
Xe–X–HNN ⁺	25.82	...	0.33	2.49	0.18	35	65
HN ₂ ⁺		2.60					

^aThe 28 innermost electrons of Xe are not treated explicitly but through an effective core potential.

N) basin splits into V(H_X) and V(N) valence basins. The weight I_a (35%) remains smaller than I_b (65%), again in accordance with the preferred dissociation mechanism (see Table III).

V. RATE CONSTANTS OF THE RADIATIVE ASSOCIATIONS

The formation of all complexes occurs without an activation barrier. The rate constants of these radiative associations are reported in Tables VIII–X (for Ar, Kr, and Xe, respectively) for temperatures in the range of 10–100 K, which includes typical temperature interval (30–100 K) in the protosolar nebula.³² All rates present similar variations with the temperature, i.e., a maximum around 60 K.

Regarding Ar complexes with the four different ions, the rates stay close to the value $0.5 \times 10^{-18} \text{ cm}^3 \text{ s}^{-1}$ obtained for H₃⁺, within one order of magnitude, with a small advantage to ArHOC⁺ whose higher binding energy gives a formation rate going up to $1 \times 10^{-18} \text{ cm}^3 \text{ s}^{-1}$. Likewise, the maximum rate constants of Kr complexes are in between 0.4×10^{-18} and

TABLE VIII. Rate constants for the radiative association of Ar with protonated ions in $10^{-18} \text{ cm}^3 \text{ s}^{-1}$ as a function of T (K).

T (K)	ArH ₃ ⁺	ArH ₃ O ⁺	ArHCO ⁺	ArHOC ⁺	ArHNN ⁺
10	0.33	0.16	0.09	0.68	0.33
20	0.43	0.22	0.11	0.94	0.45
30	0.49	0.26	0.13	1.11	0.52
40	0.52	0.28	0.14	1.21	0.57
50	0.53	0.29	0.15	1.25	0.58
60	0.53	0.29	0.14	1.25	0.58
70	0.52	0.28	0.14	1.22	0.56
80	0.50	0.27	0.13	1.18	0.54
90	0.48	0.26	0.13	1.13	0.52
100	0.47	0.25	0.12	1.08	0.49

TABLE IX. Rate constants for the radiative association of Kr with protonated ions in $10^{-17} \text{ cm}^3 \text{ s}^{-1}$ as a function of T (K).

T (K)	KrH ₃ ⁺	KrH ₃ O ⁺	KrHCO ⁺	KrHOC ⁺	KrHNN ⁺
10	1.75	0.019	0.010	0.26	0.12
20	2.43	0.026	0.019	0.35	0.17
30	2.91	0.031	0.021	0.42	0.20
40	3.23	0.034	0.024	0.47	0.21
50	3.42	0.035	0.026	0.50	0.22
60	3.50	0.035	0.026	0.51	0.22
70	3.50	0.034	0.024	0.52	0.21
80	3.44	0.033	0.023	0.51	0.21
90	3.35	0.021	0.021	0.51	0.20
100	3.24	0.020	0.029	0.49	0.19

TABLE X. Rate constants for the radiative association of Xe with protonated ions in $10^{-17} \text{ cm}^3 \text{ s}^{-1}$ as a function of T (K).

T (K)	XeH ₃ ⁺	XeH ₃ O ⁺	XeHCO ⁺	XeHOC ⁺	XeHNN ⁺
10	4.90	0.064	0.031	0.44	0.29
20	6.96	0.087	0.042	0.61	0.40
30	8.36	0.10	0.050	0.74	0.47
40	9.24	0.11	0.056	0.82	0.52
50	9.72	0.12	0.059	0.89	0.55
60	9.87	0.12	0.061	0.93	0.57
70	9.73	0.11	0.062	0.94	0.57
80	9.55	0.11	0.061	0.94	0.56
90	9.23	0.10	0.060	0.93	0.54
100	8.83	0.099	0.059	0.91	0.52

$5 \times 10^{-18} \text{ cm}^3 \text{ s}^{-1}$, in agreement with their relative binding energies. The efficiency of the reactions for Xe complexes increases again with the binding energy, starting from 0.6×10^{-18} up to almost $10 \times 10^{-18} \text{ cm}^3 \text{ s}^{-1}$.

Comparing these results with the rate constants of the H₃⁺ complexes, we observe that those of Ar complexes are in the same order of magnitude as ArH₃⁺, whereas the rates of Kr complexes are lower than KrH₃⁺, by almost one order of magnitude for KrN₂H⁺ and KrHOC⁺ and even two for KrH₃O⁺ and KrHCO⁺; the same behavior is observed for Xe complexes. However, it has to be noted that all the rates concerning Kr and Xe complexes are higher than those of Ar complexes and that even the smallest rates of Kr and Xe complexes can compare with those of ArH₃⁺. Thus, they all can be considered as quite efficient in terms of radiative association reactions.

It has also to be reminded that, since the approach applied treats the complexes as pseudo-diatomic systems, the accuracy of the rate constant calculations is lesser for the ions having linear geometry than for H₃⁺ that has a more sphere like symmetry.

VI. CONCLUDING REMARKS

An extensive study of the complexes issued from the trapping of the noble gases by the H₃⁺ ion had indicated that such complexes were far more stable than expected. The important result of the present study is that this property appears to generalize to other protonated ions, in particular ions issued

from common interstellar molecules as H_2O , CO , and N_2 which, through interaction with H_3^+ , give H_3O^+ , HCO^+ , HO^+ , and N_2H^+ . Apart from H_3O^+ , all of these complexes have a linear structure, the noble gas pointing to the hydrogen side. Although all these well-defined complexes are stabilized by long-range interactions, they do not belong to the same family. The topological analysis of the electron localization function shows that XH_3O^+ , XHCO^+ , ArHO^+ , ArN_2H^+ , and KrN_2H^+ complexes can be described as $(\text{X} \cdots \text{ion})$ systems, whereas XeHO^+ and XeN_2H^+ are better described as $(\text{XH}^+ \cdots \text{neutral})$ systems, the case of KrHO^+ remaining unclear. The partitioning so obtained is at the basis of a reliable evaluation of the BSSE corrections necessary to determine accurate geometries and rotational constants in view of the search for these species at radio wavelengths (in the laboratory or in space). However, the trapping ability of these ions had also to be investigated by evaluating their formation reactions rates. A fully *ab initio* approach has been applied in the framework of the one-dimensional reaction coordinate approximation. According to our results, the noble gas complexes with the ions obtained by the destructive reactions of H_3^+ have rate constants in the range of 10^{-19} to $10^{-17} \text{ cm}^3 \text{ s}^{-1}$ which can be considered fairly efficient values for radiative association reactions. Therefore, the destruction of H_3^+ also forms ions which have the capacity to trap noble gases as well.

From an astrochemical point of view, these results support the hypothesis that the heavy noble gases could be sequestered by protonated species (H_3^+ or/and daughter molecules) in the gas phase of the protosolar nebula. These properties would then prevent the noble gases from condensation or incorporation in the ices formed at large heliocentric distances during the cooling of the protosolar nebula. The trapping ability of the products formed by the destruction of H_3^+ ions remains preserved since this mechanism leads to the same number of protonated trapping agents at least as efficient as the parent ion H_3^+ . These processes could explain the noble gas depletion in the planetesimals which, later on, agglomerated to form Titan.⁴

They are also relevant to the determination of the origin of comets since a larger noble gas deficiency in a comet may imply that this body agglomerated from grains formed at larger heliocentric distances, where the trapping efficiency by protonated species increases. The *in situ* measurements³³ in comet 67P/Churyumov-Gerasimenko made by the ROSINA mass spectrometer aboard the *Rosetta* spacecraft give $^{36}\text{Ar}/\text{H}_2\text{O} \sim 0.1\text{--}2.3 \times 10^{-5}$. This value, compared with $^{36}\text{Ar}/\text{H}_2\text{O} \sim 5.9 \times 10^{-3}$ in the protosolar nebula,³⁴ indicates a significant though moderate depletion, as expected from our model. The future measurement of Kr and Xe abundances in long period comets, such as C/1995 O1 (Hale-Bopp), will be needed to assess the validity of our model. On the other hand, similar determinations in short period comets, such as comet 67P/Churyumov-Gerasimenko, may not display values representative of those acquired at their formation epoch in the protosolar nebula. This alteration could result from the progressive devolatilization of these comets during their multiple orbits with perihelions close to the Sun.

Moreover, the mechanism proposed here is valid in the ISM where H_3^+ should be even more abundant than in the protosolar nebula. However, it has to be reminded that the H_3^+ abundance in the ISM is poorly constrained compared to those of HCO^+ and the other ions; this results from the fact that the H_3^+ abundance must be determined at infrared wavelengths, a difficult task to perform compared to the case of the other ions whose abundances can be fairly easily quantified in the radio domain.

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APPENDIX A: QUANTUM CHEMICAL TOPOLOGY

The methodology lies on the analysis of the gradient field of a scalar function by applying the theory of dynamical systems. The latter provides a partitioning of the molecular space into non-overlapping volumes (basins) localized around maxima of the scalar function and separated by zero flux surfaces. In the *atoms in molecular theory* (QTAIM), the function is the electron density. Each atom of a molecule corresponds to a basin and the integration of the electron density over the basin volume provides an atomic population. The atomic charge of a topological atom Ω , $q(\Omega)$, is calculated by subtracting the atomic population from the atomic number, $Z(\Omega)$. If effective core potentials (ECPs) are used for some atoms (case of Xe species), the calculation of their atomic charge involves Z_{eff} , the charge of the inner-core. If the scalar function is the electron localization function (ELF) of Becke and Edgecombe, atomic and non-atomic basins are found. The topology of the function can be related to chemical concepts issued from the Lewis theory. The core basins ($Z > 2$) are located at nuclear centers, whereas the valence basins are located in the remaining space: monosynaptic basins, labeled $V(X)$, usually correspond to the lone-pair regions of X atoms while disynaptic basins, labeled $V(X, Y)$, correspond to two-center X–Y bonds. Disynaptic basins $V(H, Y)$ are termed as protonated basins. The integration of the total electron density over the basin volume provides a covalent bond population for $V(X, Y)$ basins while lone-pair populations can be obtained for $V(X)$ basins.

APPENDIX B: QUANTUM DYNAMIC PROCEDURE

We used a complete *ab initio* method^{35–37} to determine rate constants, in which they are calculated by averaging over the cross sections with a Maxwellian velocity distribution at temperature T,

$$k(T) = \langle v, \sigma(v) \rangle,$$

as proposed³⁸ by Bacchus-Montabonel *et al.* The quantum expression of the cross section at a given relative collision energy E of the colliding fragments is given by

$$\sigma(E) = \sum_J \sigma_J(E) \quad (\text{B1})$$

with

$$\sigma_J(E) = \sum_{v'J'} \frac{64 \pi^5}{3} \frac{g}{c^3} \frac{g}{2\mu E} \nu_{E,v'J'}^3 S_{JJ'} M_{EJ,v'J'}^2, \quad (\text{B2})$$

where μ is the reduced mass, g is a parameter accounting for the degeneracy of the electronic state (here $g=1$), $S_{JJ'}$ are the Höln-London coefficients characteristic of the $J J'$ pair of rotational levels, and $M_{EJ,v'J'}^2$ is the free-bound transition moment matrix element of the dipole moment $D(\mathbf{R})$ between the energy normalized function f_{EJ} of the continuum and the bound rovibrational function $\phi_{v'J'}$ for the same potential. In this expression, $\nu_{E,v'J'}$ is the emitted photon frequency corresponding to the energy difference between the continuum level of energy E and the rovibrational bound state of energy $E_{v'J'}$ is given by

$$E = h\nu_{v'J'} = E - E_{v'J'}.$$

For diatomic molecules, the rovibrational bound states and corresponding free-bound dipole transition moments can be evaluated so that rate constants may be calculated with a quantum chemical treatment. This approach may be extended for larger molecules in the framework of the one-dimensional reaction coordinate approximation in which the complex can be divided into two fragments and potential energy is optimized for each distance between the fragments. Therefore, in this work, we considered the formation of noble gas complexes to occur along the [ion ··· noble gas] reaction coordinate, in its electronic ground state potential energy curve, as no excited states could be involved in the process. Although this model cannot take into account all degrees of freedom of the complex, it is shown to give reliable results even in the case of two molecular fragments (CH_3^+ and H_2).³⁸

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