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Evolution of interstellar organic compounds under asteroidal hydrothermal conditions

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A B S T R A C T

Carbonaceous chondrites (CC) contain a diversity of organic compounds. No definitive evidence for a genetic relationship between these complex organic molecules and the simple organic molecules detected in the interstellar medium (ISM) has yet been reported. One of the many difficulties arises from the transformations of organic compounds during accretion and hydrothermal alteration on asteroids. Here, we report results of hydrothermal alteration experiments conducted on a common constituent of interstellar ice analogs, Hexamethylenetetramine (HMT – C₆H₁₂N₄). We submitted HMT to asteroidal hydrothermal conditions at 150 °C, for various durations (up to 31 days) and under alkaline pH. Organic products were characterized by gas chromatography mass spectrometry, infrared spectroscopy and synchrotron-based X-ray absorption near edge structure spectroscopy. Results show that, within a few days, HMT has evolved into (1) a very diverse suite of soluble compounds dominated by N-bearing aromatic compounds (> 150 species after 31 days), including for instance formamide, pyridine, pyrrole and their polymers (2) an aromatic and N-rich insoluble material that forms after only 7 days of experiment and then remains stable through time. The reaction pathways leading to the soluble compounds likely include HMT dissociation, formose and Maillard-type reactions, e.g. reactions of sugar derivatives with amines. The present study demonstrates that, if interstellar organic compounds such as HMT had been accreted by chondrite parent bodies, they would have undergone chemical transformations during hydrothermal alteration, potentially leading to the formation of high molecular weight insoluble organic molecules. Some of the diversity of soluble and insoluble organic compounds found in CC may thus result from asteroidal hydrothermal alteration.

Keywords:

Hexamethylenetetramine
Hydrothermal alteration
Carbonaceous chondrites
Interstellar organic compounds
IOM
GC-MS
FTIR
XANES

1. Introduction

Chondrite parent bodies accreted a diversity of materials, including icy grains, minerals and organic molecules (Aléon, 2010; Boogert et al., 2015; Herbst and van Dishoeck, 2009; Remusat, 2016; Remusat et al., 2009, 2006; Scott and Krot, 2007). Carbonaceous chondrites (CC) contain up to 4 wt% of organic matter (OM), dominated by an insoluble organic matter fraction (IOM – 75 to 90 wt%) associated with a minor fraction of soluble organic matter (SOM – 10–25 wt%). While the SOM is composed of a diversity of small organic molecules (such as carboxylic acids, amino acids, nucleobases, aliphatic and polycyclic aromatic hydrocarbons, sugars,...), the IOM consists in high molecular weight molecules rich in hetero-elements (N, S, O) made of small aromatic units and

short aliphatic chains (Alexander et al., 2007; Derenne and Robert, 2010; Orthous-Daunay et al., 2013; Remusat et al., 2005b, 2005a; Vinogradoff et al., 2017).

The origin of chondrite OM remains the subject of intense debates, in particular because this mixture of compounds, likely coming from different sources, may have undergone several episodes of evolution (Aléon, 2010; Alexander et al., 2017, 2007; Le Guillou et al., 2014; Remusat et al., 2010; Yabuta et al., 2007). Organic molecules may form through different processes in various astrophysical environments including molecular clouds, circumstellar envelopes and protoplanetary disks (Boogert et al., 2015; Caselli and Ceccarelli, 2012; Herbst and van Dishoeck, 2009). Pre-accretion organic synthesis may have occurred through plasma-induced condensation reactions at elevated temperatures (up to 1000 K) (Kuga et al., 2015; Morgan et al., 1991; Sagan and Khare, 1979; Saito and Kimura, 2009), through gas-grain reactions (Fischer-Tropsch-type process for instance) at moderated temperatures (< 600 K)

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(Anders et al., 1973; Llorca and Casanova, 2000; Nuth et al., 2008) or through irradiation-driven and thermal-driven reactions within ices, usually on silicate grains, at low temperature (< 280 K) (Allamandola et al., 1988; Bernstein et al., 1995; Ciesla and Sandford, 2012; Laurent et al., 2015; Muñoz Caro and Schutte, 2003; Theulé et al., 2013; Vinogradoff et al., 2013).

In addition to such diversity of possible sources for OM, a number of mineralogical studies of CC have pointed out that most parent bodies underwent hydrothermal alteration and metamorphism (Brearley, 2006; Keil, 2000; Mccsween, 1979). While the impact of hydrothermal processes on mineral phases has been well identified and constrained (e.g. formation of phyllosilicates, carbonates and sulfides), it remains a subject of investigations for OM (Alexander et al., 2014, 2007; Burton et al., 2012; Cody et al., 2011; Cody and Alexander, 2005; Glavin et al., 2010; Le Guillou et al., 2014; Martins and Sephton, 2009; Orthous-Daunay et al., 2013; Remusat et al., 2010; Vinogradoff et al., 2017; Yabuta et al., 2017). The present study aims at experimentally constrain the impact of hydrothermal alteration on a molecule that is typically formed in interstellar ice analogs.

Ice experiments are conducted to simulate the evolution processes of the interstellar ice-mantled dust from early molecular cloud phase to planetary disk phase (Allamandola et al., 1988; Bernstein et al., 1995; Muñoz Caro and Schutte, 2003). They are classically performed by irradiation (usually at 10K) with ultraviolet light (UV) or charged particles, and/or by warming of ice mixtures composed of water, methanol and ammonia. When warmed up to room temperature, a refractory organic residue forms. The main product is hexamethylenetetramine molecule (HMT, $C_6H_{12}N_4$) which can represent up to ~50 wt% of the residue as determined by infrared spectroscopy (Bernstein et al., 1995; Cottin et al., 2001; Muñoz Caro and Schutte, 2003; Vinogradoff et al., 2015, 2013, 2011). Of note, irradiation itself is not required to form HMT. The detailed mechanism of HMT formation shows that the molecule is actually formed during the heating of the ice analog from three molecules: formaldehyde, ammonia and formic acid (Vinogradoff et al., 2012), which are molecules commonly observed in interstellar ices of molecular clouds and young stellar objects (Boogert et al., 2008; Gibb et al., 2004). Although HMT forms at relatively high temperature in laboratory experiments (280 K), it has been suggested that on longer timescales, it may form at temperatures as low as 200 K in protoplanetary or cometary environments (Vinogradoff et al., 2013). HMT has not been observed in astrophysical environments because it does not have rotational mode for gas phase detection and because its IR bands are overlapped with silicate bands in solid phase (Bowey et al., 1998).

There are several scenarios that could account for the accretion of HMT in chondrites. First, if HMT precursors-bearing ices were heated above 200 K in the protoplanetary disk, HMT might form (Vinogradoff et al., 2015, 2013). Above 200 K, the water is sublimated but HMT can remain stable up to 400–500 K (Vinogradoff et al., 2012), absorbed onto silicate grains or as individual solid grains for instance. HMT molecules could have then survived in the protoplanetary disk and may have been accreted on asteroids. Another possibility is that ice mixtures containing water and the HMT precursors were directly co-accreted in asteroids (Le Guillou and Brearley, 2014). Over the course of the warming of the asteroids, at temperature below the water melting point (< 250–270 K), reactions similar to the one observed in annealing ice experiments could occur and lead to the formation of HMT, at the expense of formaldehyde and ammonia. Given that up to 15 wt% of water is present in chondrites, if HMT precursors are indeed associated with water ice, it is possible that a significant amount of HMT could be formed in the very early stages of the asteroids life. In any case, the refractory organic residues of interstellar ice experiments, including HMT, are likely to be present in

primitive bodies (Caselli and Ceccarelli, 2012; Ciesla and Sandford, 2012; Danger et al., 2016; Gudipati et al., 2015).

Here, we use HMT as a representative, simple molecule, resulting from interstellar ice evolution. We conducted experiments for different durations (up to 31 days) under hydrothermal conditions typical of those that existed on asteroids, i.e. at 150 °C (423 K) under alkaline pH (Brearley, 2006). Both soluble and insoluble experimental products were characterized by gas-chromatography coupled to mass spectrometry (GC-MS), Fourier transform infrared spectroscopy (FTIR) and synchrotron-based X-ray absorption near edge structure (XANES) spectroscopy. Altogether, the present study illustrates the diversity of molecules that can derive from a simple soluble organic molecule under hydrothermal conditions, including insoluble macromolecules. We discuss the significance of the present results for the origin of chondritic OM.

2. Methods

2.1. Hydrothermal experiments and products extraction

Hydrothermal experiments were carried out within Parr® PTFE reactors maintained at 150 °C ~ 5 bars (saturation vapor pressure) for different durations (2, 7, 20 and 31 days). HMT powder (> 99% – Sigma Aldrich) was mixed with bi-distilled water, previously mixed with KOH to adjust the pH at 10, to obtain solutions at 0.7 M (mol L^{-1}). To note, this concentration is below the saturation point of HMT in water at room temperature (6 mol L^{-1}). Bi-distilled water solution at pH 10 was degassed with argon before experiments. 2 mL of these solutions were introduced into 12 mL PTFE reactors. PTFE reactors were closed under argon atmosphere (> 99.999%; Air Liquide, ALPHAGAZ 1) in a glove box (< 0.5 ppm O_2).

At the end of each experiment, solutions were centrifuged in Eppendorf bio-pur tubes (2 mL) for 6 min at 12,000 rpm. The pH at the end of each experiments was around 10. These solutions were directly analyzed by XANES spectroscopy while liquid-liquid extractions were performed using 0.4 mL of dichloromethane (DCM) on 0.4 mL aliquots to perform GC-MS analyses (Fig. 1). Of note, no organic compounds were observed after 31 day-long blank experiments carried out with pure bi-distilled water.

An insoluble, dark brown solid material was recovered for experiments longer than 7 days. This IOM was washed twice with water, methanol, and DCM/methanol (1:1 vol/vol), as well as with a HF/HCl solution (3:1 vol/vol, 12 M/3 M) for 24 h. The IOM were then dried in an oven for 4 days at 60 °C under primary vacuum before bulk elemental composition analyses and infrared and XANES spectroscopy experiments (Fig. 1).

2.2. Elemental analysis

The hydrogen, nitrogen and carbon contents of the IOM were determined using the Thermo-Fisher Flash 2000 CHNS-O analyzer operating at IStEP (Paris, France). A mass of 0.7 mg of IOM was combusted under oxygen/helium flux at 960 °C. N_2 , CO_2 , and H_2O released by combustion were separated by a chromatography column and quantified using a thermal conductivity detector. Sulphanilamide ($C_6H_8N_2O_2S$) was used as a standard.

2.3. Gas chromatography–mass spectrometry

Gas Chromatography–Mass Spectrometry (GC-MS) analyses were carried out using the Agilent Technologies 6890N gas chromatograph coupled with an Agilent Technologies 5973 network mass spectrometer operating at METIS (Paris, France). The capillary column is a RTX-5Si/MS (30 m × 0.25 mm, 0.5 μm film) coated with chemically bound Restek (low-polarity phase, suitable for semi-volatile, hydrocarbon, amine, phenol compounds). The temperature

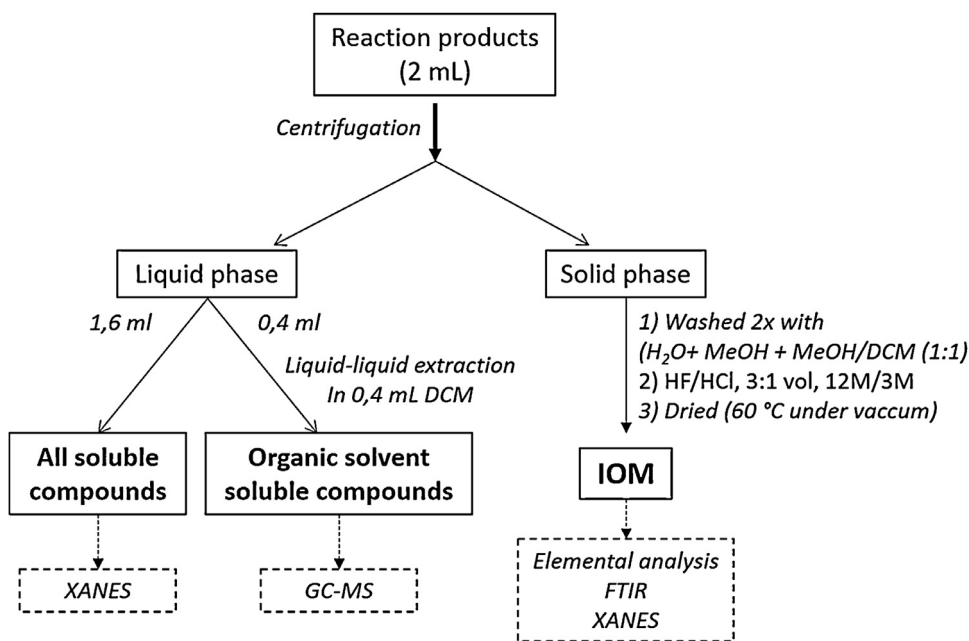


Fig. 1. Schematic representation of the preparation of experimental residues for GC-MS, IR and XANES analyses.

injection was 80 °C in splitless mode and the GC oven program started at 50 °C–320 °C at 4 °C/min with He as carrier gas. The solvent delay was as short as possible (3 min) and no derivatization was performed. Measurements were performed using an electron energy of 70 eV, an ion source temperature of 220 °C and a scanning rate of 2.24 scan s⁻¹ from 35 to 700 *m/z*. Products were identified using the NIST database. Data processing was done using the OpenChrom software, namely subtraction of the solvent background (DCM) taken at 3.8 min followed by an automatic peak detection using the mass selective detector (MSD) first derivative process (signal to noise ratios > 3) and then each peak is integrated using the MSD integrator option. This procedure allowed quantifying a molecular size index in our experiments that corresponds to the ratio between the integrated area of all peaks appearing between 25 and 54 min (hereafter called “heavy compounds”, HC) and the integrated area of all peaks appearing between 5 and 25 min (hereafter called “light compounds”, LC).

2.4. FTIR

FTIR analyses were performed using the Vertex 70 spectrometer (Bruker) operating at the MNHN (Paris, France). Less than 0.5 mg of powdered sample was pressed onto the diamond crystal surface using a single reflection in the attenuate total resonance (ATR) device, ATR-quest (Specac). The absorption of the infrared beam by the samples was measured from 370 to 4000 cm⁻¹ with a spectral resolution of 4 cm⁻¹. Spectra were accumulated 96 times (vs. 128 for the background to ensure a good signal to noise ratio). Each IR spectrum was corrected with a baseline generated by the Bruker Opus software, integrated and normalized to its total area.

2.5. XANES spectroscopy

2.5.1. XANES data acquisition procedure

XANES data were collected on the 10ID-1 STXM beamline (Kaznatcheev et al., 2007) at the Canadian Light Source (CLS) and on the HERMES STXM beamline (Belkhou et al., 2015; Swaraj et al., 2017) at the synchrotron SOLEIL. At CLS, a 100 nm thick titanium filter is used to remove the contribution of second order light

while at SOLEIL, beamline optical elements are exposed to a continuous flow of pure O₂ to remove carbon contamination. Microscope chambers were pumped down to less than 100 mTorr after sample insertion. Energy calibration is done before measurements using the well-resolved 3p Rydberg peak of gaseous CO₂ at 294.96 eV for the C K-edge and using the 1s→π* photo-absorption resonance of gaseous N₂ at 400.8 eV for the N K-edge.

For the present study, samples were deposited on Si₃N₄ windows (50 nm thick membrane). Image stacks were collected with energy increments of 1 eV over the 250–450 eV energy range to measure N/C values, following the procedure described in Alleon et al., (2015). Additional image stacks were collected with energy increments of 0.1 eV over the carbon (270–350 eV) and the nitrogen (390–420 eV) absorption ranges to gain insights on carbon and nitrogen speciation. Data were collected with a dwell time of one millisecond per pixel to prevent irradiation damages as recommended by Wang et al. (2009) and processed using the aXis2000 software (ver. 2.1n). The C- and N-XANES spectra shown here correspond to homogeneous areas of several tens of squared micrometers each.

2.5.2. XANES data deconvolution procedure

The same deconvolution procedure in three steps was applied to all C-XANES spectra: i) background subtraction, ii) normalization to the total carbon content, and iii) spectral fitting using Gaussian functions. Background subtraction consisted in the subtraction of a linear regression over the 270–282 eV energy range. Then, following Barré et al. (2016) and Alleon et al. (2017), C-XANES spectra were normalized to their area between 280 eV and 291.5 eV and N-XANES spectra were normalized to their area between 395 and 406.5, thereby ensuring chemical consistency (a spectrum showing a more prominent absorption at a given energy must have a less intense absorption at other energies).

After normalization, Gaussian functions with a given full-width at half maximum (0.6 eV) were used for C-XANES signal deconvolution. Their positions were fixed following Urquhart et al. (1999), Dhez et al. (2003) and De Gregorio et al. (2010) (see Table 1). The height of the Gaussian curves were used as a proxy of the abundance of the functional group absorption. Due to overlapping energies of several resonances, absorption peaks at the N K-edge

Table 1

Positions and peak attributions for the deconvolution of the C-XANES spectra, from [Urquhart et al. \(1999\)](#), [Dhez et al. \(2003\)](#) and [De Gregorio et al. \(2010\)](#).

X-Ray absorption (eV)	Functional groups
284.4	Quinones
284.7; 285; 285.4	Aromatics; olefins
285.8	Imines
286.2; 286.6; 286.8	Ketones; phenols; nitriles
287.6; 287.9	Aliphatics
288.2	Amides
288.55	Carboxyls; esters; acetals
289	Aldehydes
289.4	Hydroxyls
289.9	Aliphatics; carbamoyls
290.3	Carbonates

cannot be univocally assigned to given functional groups. Still, the N local coordination can be partly inferred based on the energy of absorption features: imine, nitrile and pyridinic N will generate absorption features below 400 eV, while absorption features above 400 eV will indicate the presence of amide, nitro and pyrrolic N ([Leinweber et al., 2007](#); [Myneni, 2002](#)).

Following previous studies ([Alleon et al., 2016b, 2017](#); [Bernard et al., 2012, 2015](#); [Le Guillou et al., 2013, 2014](#); [Vinogradoff et al., 2017](#)), semi-quantitative parameters were extracted from the XANES spectra: 1) the aromatic/olefin/imine index that corresponds to the relative content of Aromatic C, C=C and C=N (i.e.

the sum of the areas of the Gaussian functions used to deconvolve the absorption signal in the range 284.4–285.8 eV), 2) the ketone/phenol/nitrile index that corresponds to the relative content of C=O, Aromatic-OH and C≡N (i.e. the sum of the areas of the Gaussian functions used to deconvolve the absorption signal in the range 286.2–286.8 eV) and 3), the amide/carboxyl/ester/acetal index that corresponds to the relative content of N–C=O and COOH/COOR (i.e. the sum of the areas of the Gaussian functions used to deconvolve the absorption signal in the range 288.2–288.55 eV). These values carry a degree of uncertainty of $\pm 10\%$.

3. Results

3.1. Evolution of the soluble compounds

3.1.1. XANES data

XANES spectra of the starting and of the residual solutions were collected prior any extraction. The C-XANES spectrum ([Fig. 2b](#)) of the starting HMT solution (0.7 M) exhibits a broad absorption band centered at 290 eV, attributed to the $1s \rightarrow \sigma^*$ electronic transitions of amine (C–N) groups ([Myneni, 2002](#); [Solomon et al., 2009](#)). The N-XANES spectrum ([Fig. 2c](#)) of the starting HMT solution exhibits a broad absorption feature centered at 406.8 eV with two peaks at 399.2 and 401.2 eV. Given the HMT structure, these peaks might be assigned to tertiary amines functions (N–(CH₂)₃) involved in a cyclic structure. Although HMT exhibits an initial N/C value of 0.57, the mean N/C value of all the residual solutions is 0.27, thereby highlighting a chemical evolution of HMT during experiments.

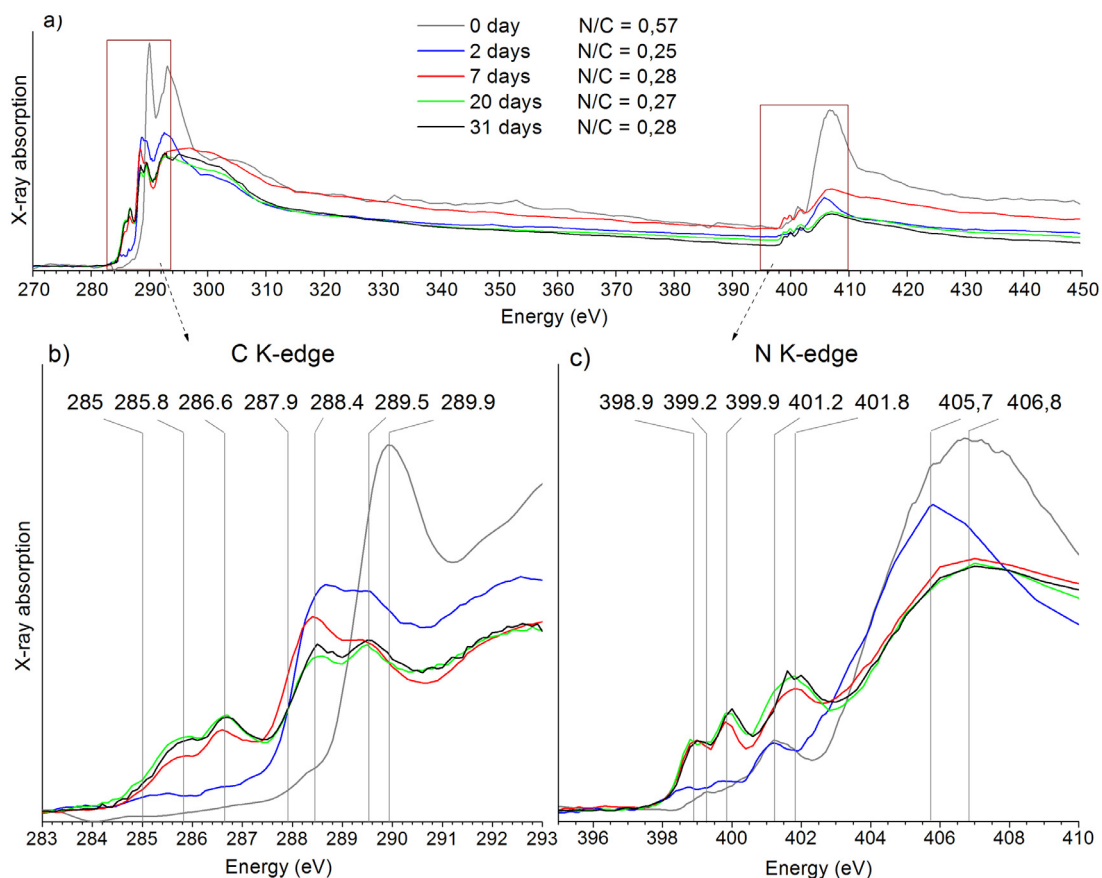


Fig. 2. XANES spectra of residual solutions. (a) XANES spectra of the starting HMT solution (0.7 M) and of the solutions after 2, 7, 20, 31 day-long hydrothermal experiments performed at 150 °C (spectra are normalized to the carbon content). (b,c) C-XANES spectra normalized to the carbon content and N-XANES spectra normalized to the nitrogen content. Attribution of absorption features: 285 eV: aromatics/olefins; 285.8 eV: imines; 286.2–286.8: ketones/phenols/nitriles; 287.2–287.9 eV: aliphatics; 288.2–288.7 eV: amides/carboxyls/esters/acetals; 289.5 eV: hydroxyls; 289.9 eV: amines/aliphatics/carbamoyls; 398.8–399.9 eV: imines/nitriles/pyridinic N; 401–401.5: amides; 401.5–402.5 eV: pyrroles.

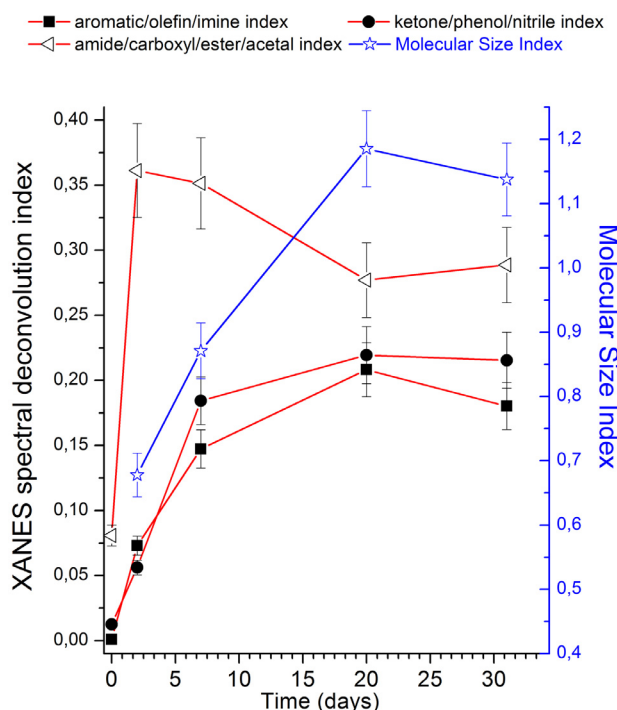


Fig. 3. Evolution of the aromatic/olefin/imine index, the ketone/phenol/nitrile index, the amide/carboxyl/ester/acetal index and the molecular size index of the solutions as a function of experimental duration.

After 2 days at 150 °C, the residual solution exhibits a C-XANES spectrum with two main peaks at 288.4 and 289.5 eV, attributed to the presence of amide (N=C=O), acetal (O-C-O), ester (O=C-O-R) and/or carboxyl (O=C-OH) groups and to the presence of hydroxyl groups (C-OH), respectively (Lessard et al., 2007; Robin et al., 2015; Rouchon and Bernard, 2015). Two additional, less intense absorption features can be observed at 285.8 eV, corresponding to $1s \rightarrow \pi^*$ electronic transitions of imine (C=N) groups, and at 286.6 eV attributed to $1s \rightarrow \pi^*$ electronic transitions of ketone (C=O), phenol (Ar-OH) and/or nitrile (C≡N) groups (Dhez et al., 2003; Hitchcock et al., 2005). The N-XANES spectrum of this residual solution exhibits two peaks at 398.9 and 399.9 eV, likely corresponding to

$1s \rightarrow \pi^*$ electronic transitions of imine (C=N), nitrile (C≡N) and/or pyrrole/pyrimidine groups and a peak at 401.2 eV, related to the contribution of $1s \rightarrow \pi^*$ transitions in amine and/or amide groups (Alleon et al., 2017, 2016a; Leinweber et al., 2007; Myneni, 2002).

With increasing experimental duration (from 2 to 20 days), the relative concentrations of aromatic/imine carbons and ketone/phenol/nitrile carbons significantly increase (peaks at 285.8 and 286.6 eV) while the relative concentrations of amide/carboxyl/ester/acetal and hydroxyl carbons slightly decrease (peaks at 288.4 and 289.5 eV) in the residual solution (Figs. 2 and 3). Consistently, the contribution of the broad absorption feature at 406.8 eV decreases while the contributions of the imine/nitrile/pyrrole/pyrimidine groups (peaks at 398.9 and 399.9 eV) and of the amide/amine/pyrazole groups (peaks at 401.2 and 401.8 eV) increase. After 20 days, the mean molecular composition of the organic compounds in solution stops evolving (Fig. 3).

3.1.2. GC-MS data

GC-MS data were collected on the DCM liquid-liquid extracts. The gas chromatogram of the DCM extract (Fig. 4) of the starting HMT solution (0.7 M) only displays the HMT peak at 20.73 min (m/z 140). This peak is still observed after 2 and 7 days (very small peak) of hydrothermal alteration but no more after 20 days, indicating that the total degradation of HMT takes between 7 and 20 days at 150 °C (Fig. 4). Concomitant with the degradation of HMT, the number of soluble organic compounds (number of peaks in the chromatogram) increases with increasing experimental duration (Fig. 4).

Light compounds -LC- with m/z up to 136 can be identified between 5 and 25 min and correspond mainly to three families of nitrogen-rich molecules (Fig. 4, Table 2): pyridine derivative molecules (6-membered ring molecules - $C_5H_5N \pm$ alkyl substituents), imidazole/pyrazole derivative molecules (5-membered ring molecules - $C_3H_4N_2 \pm$ alkyl substituents), and pyrazine derivative molecules (6-membered ring molecules - $C_4H_4N_2 \pm$ alkyl substituents). Oxygen-bearing products are also observed, such as N-N-dimethyl formamide (C_3H_7NO - first peak at 5.37 min) that form during the first 2 days of hydrothermal alteration, and pyrazolidinones and imidazolidinones (i.e. pyrazoles or imidazoles with an additional carbonyl group) that require longer times to form (they are observed after 20 days) (Table 2).

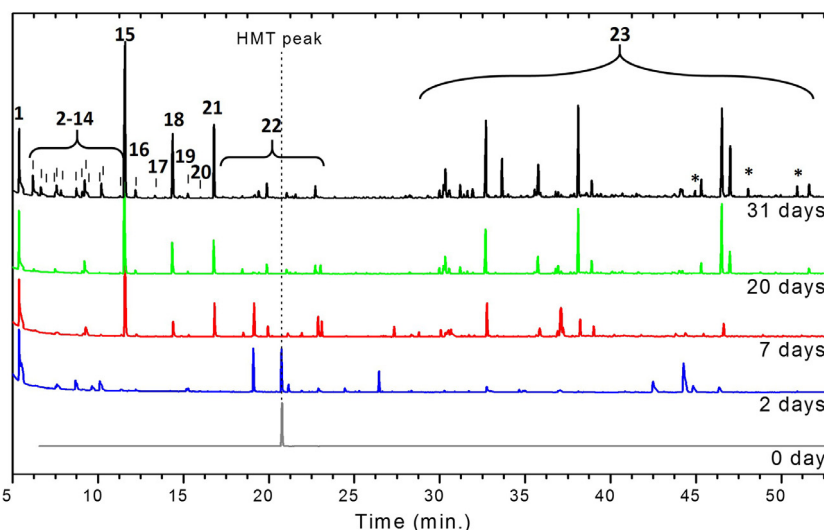


Fig. 4. Gas chromatograms of the starting HMT solution (0.7 M) and of the DCM extracts after hydrothermal alteration for 2, 7, 20 and 31 days normalized to the N,N formamide (1). Peaks between 5–25 min of the 31 days extract have been assigned to known molecules (Table 2). Stars (*) indicate contaminations from the GC-MS column.

Table 2

Nature of the organic molecules identified using GC-MS within the DCM extract of the HMT solution (0.7M) after 31 days of hydrothermal alteration at 150 °C. The retention time, the molecular ion (m/z), the m/z of the main fragments (sorted by decreasing intensity - m/z) and their possible assignments are reported for each "light compound" observed.

Number	Retention time	[M +] m/z	Main fragments* m/z	Possible assignment	
				Name	Formula
1	5.37	73	44,42	N-N-dimethyl formamide	C ₃ H ₇ NO
2	6.18	98	97,42,56	C2-methyl, 2-pyrazoline/imidazoline,	C ₅ H ₁₀ N ₂
3	6.65	96	95,81,68,54	C2-alkyl imidazole or pyrazole	C ₅ H ₈ N ₂
4	6.94	83	56	1-methyl-1H-1,2,4-triazole	C ₃ H ₅ N ₃
5	7.48	93	92,66,65,39	Pyridine, 3-methyl	C ₆ H ₇ N
6	7.56	96	98,42,68,54	C2-alkyl imidazole or pyrazole	C ₅ H ₈ N ₂
7	7.82	112	97,111,42,56	2-Pyrazoline, 1,3,4-trimethyl-	C ₆ H ₁₂ N ₂
8	8.73	112	97,111,83,42	2-Pyrazoline, 3-ethyl,1-methyl	C ₆ H ₁₂ N ₂
9	9.07	108	42,40,39	Pyrazine, 2,6-dimethyl	C ₆ H ₈ N ₂
10	9.22	82	54,42,81,40	1H-imidazole, 1-methyl	C ₄ H ₆ N ₂
11	9.39	110	95,109,42	C3-alkyl imidazole or pyrazole	C ₆ H ₁₀ N ₂
12	10.1	107	106,79,66,92	C2-alkyl pyridine	C ₇ H ₉ N
13	10.18	110	109,68,95,42	C3-alkyl imidazole or pyrazole	C ₆ H ₁₀ N ₂
14	11.28	107	106,79,92,77	C2-alkyl pyridine	C ₇ H ₉ N
15	11.55	96	95,68,42,81	Imidazole-1,4-dimethyl	C ₅ H ₈ N ₂
16	12.18	122	42,81,109	Pyrazine, trimethyl	C ₇ H ₁₀ N ₂
17	13.33	99	98,44,42	Pyrrolidinone, -methyl	C ₅ H ₉ NO
18	14.36	96	95,68,42,56	C2-alkyl imidazole or pyrazole	C ₅ H ₈ N ₂
19	15.25	136	54,42,39	Pyrazine tetramethyl	C ₆ H ₁₂ N ₂
20	15.99	114	44,113,42	2-imidazolidinone, 1,3-dimethyl	C ₅ H ₁₀ N ₂ O
21	16.79	110	109,95,68,42	C3-alkyl imidazole or pyrazole	C ₆ H ₁₀ N ₂
22	17.2-22	110, 124	-	C3/C4-alkyl imidazole or pyrazole	C ₆ H ₁₀ N ₂
23	25-54	-	-	Heavy compounds (HC)	C ₇ H ₁₂ N ₂

Heavy compounds -HC- with m/z comprised between 136 and 254 can be observed at retention times between 25 and 55 min (Fig. 4). Although the exact nature of these compounds is not readily assigned, they exhibit characteristic fragments of 6/5-membered ring pyridine/imidazole/pyrazine. They thus could be further derivatives of the pyridine/imidazole/pyrazine LC (with additional alkyl or CO₂ or NH₂ substituents) or pyridine/imidazole/pyrazine polycyclic compounds. Of note, the abundance of HC increases with increasing hydrothermal alteration duration (Figs. 3 and 4). In fact, the molecular size index (i.e. the HL/LC ratio) increases with experimental duration up to 20 days (Fig. 3).

3.2. Formation of an insoluble material

A dark brown solid material, representing less than 1% wt of the initial HMT, forms during experiments longer than 7 days. This material being insoluble in water, MeOH and DCM is logically called IOM hereafter. Elemental analyses conducted on the IOM from the 31 day-long experiment reveals that it contains 53.4 ± 0.2 wt% of C, 5.3 ± 0.01 wt% of H, 18.7 ± 0.1 wt% of N and 22.6 ± 0.2 wt% of O, i.e. its formula can be written C₁₀₀H₁₂₀N₃₀O₃₁.

The IOMs from the 7, 20 and 31 day-long experiments are spectroscopically very similar (Figs. 5 and 6), thereby highlighting a similar chemical composition. Their IR spectra display stretching vibrations of NH and OH bonds (3600–3000 cm⁻¹), C–H bonds (2850–2950 cm⁻¹), C=C or C=N bonds (1550–

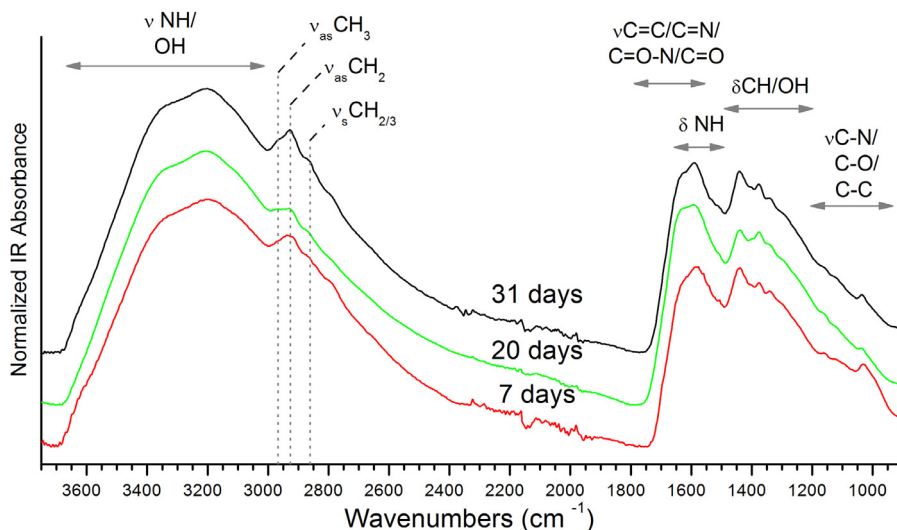


Fig. 5. Infrared spectra of the IOM formed during the 7, 20 and 31 day-long hydrothermal experiments performed at 150 °C using the starting HMT solution (0.7 M) (ν : stretching bands, as: asymmetric mode, s: symmetric mode, δ : bending bands).

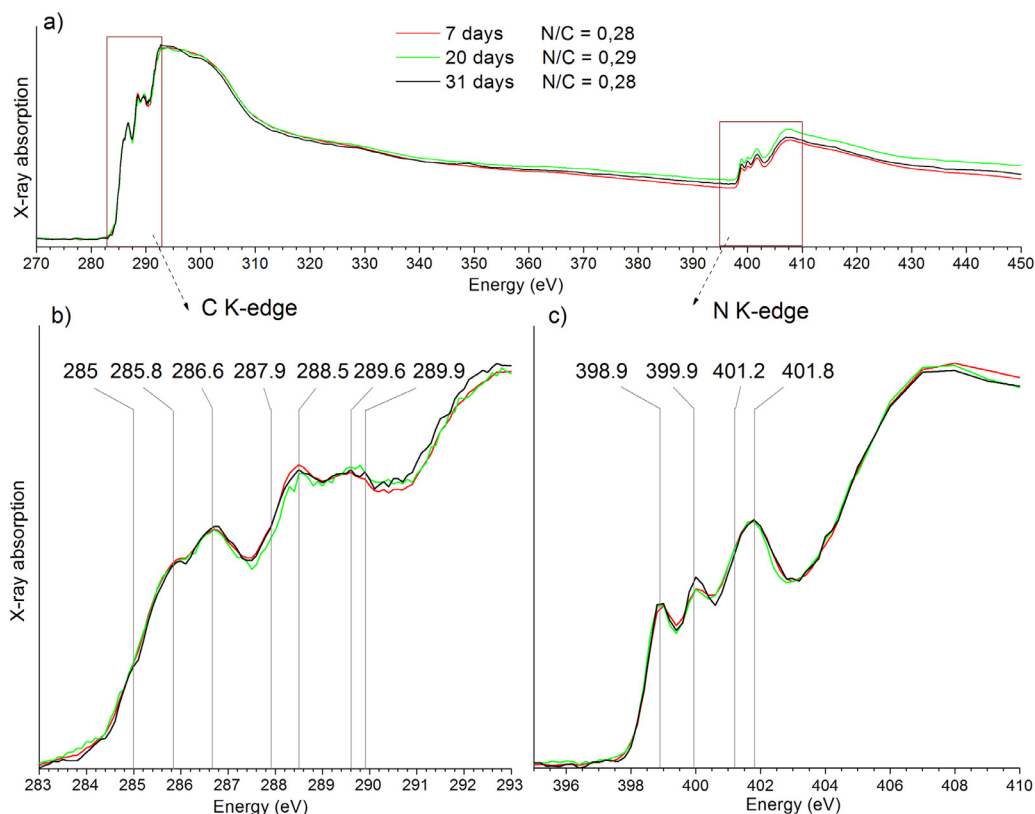


Fig. 6. XANES spectra of IOM. (a) XANES spectra of the IOM formed after 7, 20, 31 day-long hydrothermal experiments performed at 150 °C (spectra are normalized to the carbon content). (b,c) C-XANES spectra normalized to the carbon content and N-XANES spectra normalized to the nitrogen content. Attribution of absorption features: 285 eV: aromatics/olefins; 285.8 eV: imines; 286.2–286.8: ketones/phenols/nitriles; 287.2–287.9 eV: aliphatics; 288.2–288.7 eV: amides/carboxyls/esters/acetal; 289.5 eV: hydroxyls; 289.9 eV: amines/aliphatics/carbamoyls; 398.8–399.9 eV: imines/nitriles/pyridinic N; 401–401.5 eV: amides; 401.5–402.5 eV: pyrroles.

1620 cm^{-1}), and C=O/C=O–N bonds (1650–1750 cm^{-1}), and bending vibrations of NH bonds (1600–1500 cm^{-1}), C–H or OH bonds (below 1500 cm^{-1}), together with stretching vibrations of C–O and C–N bonds (900–1200 cm^{-1}) (Coates, 2006; Colthup, 1950). The C-XANES spectra of these IOMs (Fig. 6b) reveal the presence of aromatic or olefin (285 eV), imine (285.8 eV), ketone/phenol/nitrile (286.6 eV), carboxyl/ester/acetal (288.5 eV) and hydroxyl (289.6 eV) groups. Peaks at 398.9 and 399.9 eV observed in the N-XANES spectra support the presence of imine/nitrile and/or pyrrole/pyrimidine compounds, while the peaks at 401.2 and 401.8 eV indicate the presence of amide/pyrazole compounds. N/C values of 0.28 are estimated for these IOMs based on XANES data.

4. Discussion

The present hydrothermal alteration experiments show that, within a few days, HMT may evolve into (1) a very diverse suite of soluble compounds, including N–N-dimethyl formamide and N-rich heterocyclic compounds (> 150 species after 31 days), and (2) a polymeric N-rich IOM that forms after only 7 days of experiment (Fig. 7).

4.1. Formation of a diverse suite of new soluble compounds

HMT is formed from formaldehyde and ammonia and also decomposes, in water at temperature > 70 °C, into six molecules of formaldehyde and four of ammonia (Blažević et al., 1979; Meissner et al., 1954). The exact kinetics of HMT decomposition in water remains unknown. Our data show that HMT does not persist much

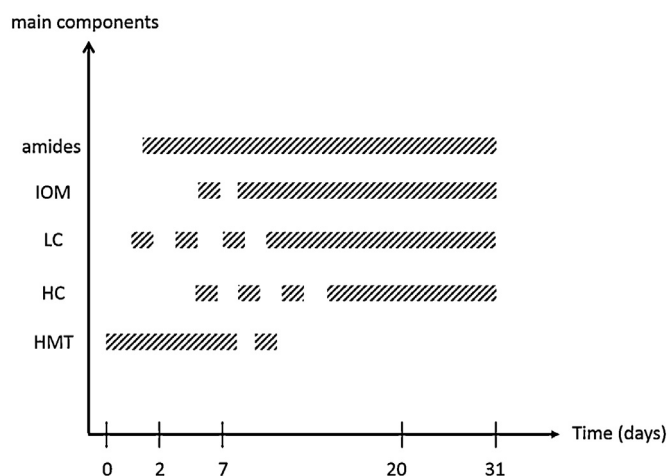


Fig. 7. Schematic representation of the detection of HMT, LC, HC, IOM and amides (mainly N,N-dimethylformamide) in experimental residues as a function of experimental duration.

longer than 7 days at 150 °C in alkaline solution, suggesting that HMT decomposition is instantaneous on asteroidal time scales.

In parallel, the present results evidence that the products of HMT decomposition can react to form a diversity of organic compounds under hydrothermal conditions typical of asteroids. In particular, numerous N-rich compounds form (> 150 species after 31 days, with m/z comprised between 59 and 254), likely through chain reactions (Fig. 8). Note that the N/C of the reaction products

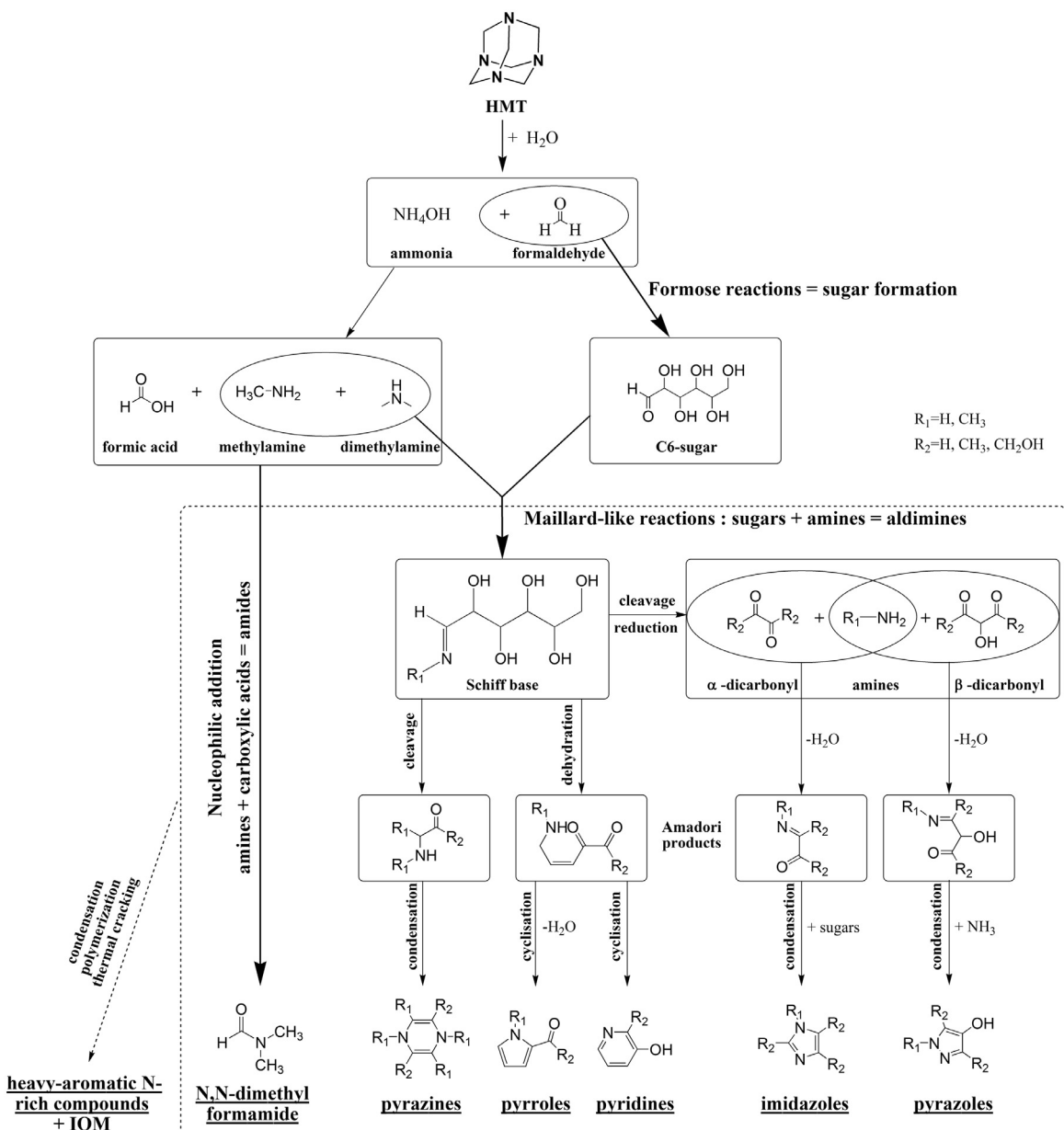


Fig. 8. Chemical pathways for the transformation of HMT during hydrothermal alteration experiments at 150 °C, pH 10, based on both existing literature data and the present results. With increasing time, the HMT is degraded and the number and the complexity of produced compounds increases. Successive Formose and Maillard-like reactions might have led to the formation of the LC, HC and IOM products.

is twice as low as that of HMT, suggesting that nitrogen has been lost to the gas phase as N₂ or NH₃ during experiments.

Formaldehyde may condense and form sugar derivatives via the Formose reaction (Breslow, 1959; Butlerov, 1861). The Formose reaction generally refers to the formation of aldose and ketose sugars through the Aldol condensation of glycolaldehyde with formaldehyde (Breslow, 1959; Shapiro, 1987). Hydrothermal Formose reactions may lead to the formation of three to six carbon sugar molecules in a few minutes at 200 °C (Kopetzki and Antonietti, 2011) and to some N-heterocycle compounds in the presence of ammonia (Kort, 1970).

At temperatures higher than 100 °C, an equimolar solution of formaldehyde and ammonia may also lead in a few hours to the formation of amine compounds, such as methylamine (CH₃NH₂), dimethylamine (CH₃)₂NH and trimethylamine (CH₃)₃N (Jones and Wheatley, 1918; Werner, 1917). Noted that this reaction also leads

to the formation of amino acids in acid hydrolysis conditions (Fox and Windsor, 1970; Hulett et al., 1971) and in alkaline hydrolysis conditions with the presence of glycolaldehyde (Kebukawa et al., 2017). In the present experiments, the formation of N-N-dimethyl formamide likely results from nucleophilic addition reactions between amine compounds and formic acid (Jursic and Zdravkovski, 1993; Mitchell and Reid, 1931), i.e. between products of formaldehyde oxidation (Fig. 8). Of note, the N-methyl formamide and the formamide compounds should also be formed, but cannot be detected without derivation by GC-MS.

With increasing experimental duration and with the temperature (150 °C), these amine compounds likely react with sugars through Amadori rearrangements and Maillard-like reactions (Ames, 1992; Nursten, 2005). In fact, although Maillard reactions generally involve amino acids and sugars (Maillard, 1912), the cyclic aldimines produced during the present experiments

(pyridines, imidazoles and pyrazines with many alkyl substitutions) are common Maillard reaction products, as are soluble O-rich compounds of high molecular weight, such as Schiff bases and di-carbonyl compounds (Ames, 1992; Nursten, 2005). Maillard-like reactions likely occur in the present experiments after the formose reaction because of the large quantity of available amines (Ellis, 1959; Namiki, 1988; Yamashita and Naraoka, 2014).

The continual increase of the molecular size index with increasing experimental duration highlights the production of high molecular weight soluble compounds (i.e. HC compounds). However, the chemical pathways that led to their formation remain unclear. These compounds could be seen as soluble melanoidins that constitute typical end products of Maillard-like reactions and, as such, may have formed via multiple condensation/polymerization reactions between low molecular weight intermediate products (Fig. 8).

After 20 days, the soluble compounds only slightly evolve with increasing experimental duration as indicated by GC-MS and XANES data, highlighting that a metastable thermodynamic equilibrium is likely achieved (Helgeson et al., 1998). Basic extrapolations suggest that such a mixture of soluble N-rich low and high molecular weight compounds may persist under asteroidal conditions for much longer durations.

4.2. From soluble to insoluble organic matter

A solid, N-rich IOM ($C_{100}H_{120}N_{30}O_{31}$) forms after 7 days of hydrothermal alteration (Figs. 5 and 6). This IOM can be seen as complex aromatic molecules containing N-rich (imines with many alkyl substitutions) and O-rich groups (ketones, phenols, carboxyls and hydroxyls) as shown by XANES and IR data. Its molecular structure does not vary with increasing experimental duration. Note that the N/C of this IOM is similar to that of the soluble compounds (~ 0.28) and twice as low as the initial HMT value.

The formation of IOM from soluble compounds has been documented to occur during (secondary) thermal cracking of soluble asphaltene at relatively low temperature ($\sim 150^\circ\text{C}$) in terrestrial settings (Bernard and Horsfield, 2014; Hill et al., 2003; Michels et al., 1996). These complex cracking reactions eventually lead to the formation of gases and of insoluble pyrobitumen residues that constitute as stable end products that may accumulate in a closed system (Behar et al., 1992; Bernard and Horsfield, 2014; Ungerer et al., 1988).

Complex organic materials, likely insoluble, may alternatively form through condensation, dehydration and cyclization reactions involving sugars produced from formaldehyde via the Formose reaction (Cody et al., 2011; Kebukawa et al., 2013; Kort, 1970; Weber, 2001). Such polymerization can occur in a few days at temperatures as low as 363 K if glycolaldehyde is added (with $\text{Ca}(\text{OH})_2$ as catalyst) (Kebukawa et al., 2013; Kopetzki and Antonietti, 2011). The presence of ammonia enhances the IOM yield significantly at temperatures below 150°C (Kebukawa et al., 2013; Kort, 1970). In such case, N-rich heterocyclic compounds are formed and incorporated into the polymeric structure (Kebukawa et al., 2013; Kort, 1970). Kinetic studies have shown that the half-life for such organic solid formation at 0°C is $\sim 2 \times 10^4$ years in the absence of ammonia vs only 70 years in the presence of ammonia (Kebukawa and Cody, 2015).

Alternatively, in the presence of amines, IOM may directly form through condensation/polymerization of Maillard-like products such as the soluble melanoidins discussed in the previous section. These reactions which require reactive molecules, such as unsaturated imine/carbonyl compounds (Benzing-Purdie et al., 1985), lead to the formation of polymeric, insoluble melanoidins with a high molecular weight (Namiki, 1988; Nursten, 2005). The final composition of such Maillard insoluble melanoidins depends on

the amine/sugar ratio of the solution, as well as on temperature and pH (Cämmerer and Kroh, 1995).

The IOM that forms in the present experiments is insoluble in all organic solvents, as are chondritic IOM and terrestrial pyrobitumen, which might not be the case for the Kebukawa IOM and the Maillard insoluble melanoidins. Compared to these IOMs, the present IOM is significantly richer in nitrogen but contains less oxygen: it exhibits a N/C of 0.28 and a O/C of 0.32 (vs 0.1 and 0.4 for the Kebukawa IOM or up to 0.15 and up to 0.6 for the Maillard melanoidins, respectively). The amount of IOM that forms is also very different (less than 1 wt% of the starting material vs. 15 or 30 wt% for the Kebukawa IOM or the Maillard insoluble melanoidins respectively). These differences are likely related to the large initial concentration of nitrogen in our solution (initial N/C of 0.57 vs 0.1 or less than 0.15 for the Kebukawa IOM or the Maillard insoluble melanoidins respectively) which has led to different chemical pathways.

Elucidating the exact pathway of IOM formation in the present experiments remains a difficult task, especially as several routes may occur simultaneously. The fact that the molecular structure of the IOM that forms in the present experiments does not significantly vary with increasing experimental duration highlights that its precursors are already formed before the end of the first week of hydrothermal alteration. Interestingly, the IOM and the HC form concomitantly (Fig. 7), which suggests that both the IOM and the HC are products of reactions involving LC or that the IOM results from reactions involving LC and the firstly formed HC.

4.3. Implications for the origin of chondritic OM

A longstanding question in cosmochemistry is whether the chondritic organic compounds are formed in the ISM, in the protoplanetary disk or within asteroids during hydrothermal alteration (Aléon, 2010; Alexander et al., 2014, 2007; Glavin et al., 2010; Le Guillou et al., 2014; Remusat et al., 2010, 2009; Vinogradoff et al., 2017). A number of interstellar and nebular processes (ion-molecule reactions, gas-grain reactions, irradiation, heating) may lead to the formation of both soluble and insoluble OM (Danger et al., 2016, 2013; Herbst and van Dishoeck, 2009; Kuga et al., 2015; Laurent et al., 2015; Marcellus et al., 2017). Here, we investigated the impact of asteroidal hydrothermal conditions on a simple interstellar molecule such as HMT (which constitutes up to 50 wt% of ice analog residues – Cottin et al., 2001; Muñoz Caro and Schutte, 2003; Vinogradoff et al., 2015, 2013, 2012) assuming that it had been accreted by asteroids (Caselli and Ceccarelli, 2012; Gudipati et al., 2015).

At 150°C in hydrothermal conditions, HMT reacts (instantaneously on asteroidal timescales, i.e. 10^6 years – Fujiya et al., 2012) to form a very diverse suite of N-rich soluble compounds (> 150 species), as evidenced in the present study. The fact that CC contain amides and quite similar nitrogen-bearing aromatic molecules (Martins et al., 2008; Stoks and Schwartz, 1981; Yamashita and Naraoka, 2014) thus suggests that asteroids possibly accreted interstellar organic compounds such as HMT. Actually, more than thousands of different organic species have been identified in ice analog residues (Danger et al., 2016, 2013; Marcellus et al., 2015; Nuevo et al., 2008). Asteroidal hydrothermal alteration reactions involving interstellar organic compounds could be at the origin of a large part of the inventory of soluble chondritic molecules ($> 10,000$ different soluble molecules – Schmitt-Kopplin et al., 2010).

Although the chondritic IOM contains pyrrole and alkylpyridine groups (Pizzarello and Williams, 2012; Remusat et al., 2005b), it is not as N-rich as the IOM that forms in the present experiments ($C_{100}H_{60-80}O_{16-18}N_3S_{2-7}$ (Derenne and Robert, 2010) vs $C_{100}H_{120}O_{31}N_{30}$). In fact, in terms of N/C and XANES signatures, the insoluble organic products of the present experiments are closer to

organic materials of ultracarbonaceous micrometeorites (UCAMM – N/C of ~0.15–0.2 – Dartois et al., 2013; Yabuta et al., 2017). Yabuta et al. (2017) have recently interpreted the N-rich organic compounds from one UCAMM as products of incipient aqueous alteration reactions that likely occurred within water coming from locally melted cometary ice grains. These organic materials would thus be highly primitive and may contain the precursor of CC organic compounds (Yabuta et al., 2017). Altogether, the results of the present study thus shed a new experimental light on the comet-asteroid continuum.

Importantly, the N/C values of the soluble and insoluble products of the present experiments are lower than that of HMT, suggesting that significant amount of NH₃ and/or N₂ is released during hydrothermal alteration. Pizzarello and Williams (2012) detected up to 10 μg mg⁻¹ of NH₃ that was likely IOM-bound, i.e. trapped within the IOM porosity, indicating that NH₃ was present in the system during alteration. Still, a fraction of these gases could have been lost through degassing of the asteroids (Young et al., 2003; Alexander et al., 2010; Le Guillou et al., 2015), thereby having impacted the formation of the soluble and insoluble compounds (a loss of NH₃ would have decreased the pH of the system while a loss of H₂ would have increased its fO₂). Asteroidal hydrothermal alteration of organic compounds likely occurred in a continuously changing environment.

Lastly, the insoluble/soluble ratio of the products of the present experiments is significantly lower than that of chondrites. According to Kebukawa et al. (2013) and Kebukawa and Cody (2015), the presence of ammonia significantly enhances the formation of IOM from formaldehyde via the formose reaction at 150 °C. Given the large N concentration of the starting material of the present experiments (N/C ~0.57 vs ~0.1 for Kebukawa experiments) and the production of ammonia, the yield of IOM from HMT should thus have been quite high, at least if similar reaction pathways have occurred. The low IOM yield observed here (< 1 wt% vs ~15 wt% for Kebukawa experiments) suggests that, rather than the ammonia concentration, this is the addition of glycolaldehyde and Ca(OH)₂ that enhances the Formose reaction and produces such high IOM yield in the Kebukawa experiments (Kebukawa et al., 2013; Kebukawa and Cody, 2015). This illustrates the need to properly identify the parameters that impact hydrothermal alteration reactions in addition to the temperature, the pH, the nature of the starting materials and the reaction pathways, including any accreted organic or inorganic phase. In this respect, the exact role played by inorganic phases remains to be investigated. For instance, phyllosilicates, oxides and/or carbonates have been observed in close association with organic compounds in CC (Le Guillou et al., 2014; Le Guillou and Brearley, 2014; Zega et al., 2010). Because these minerals may exert an influence on the oxidation state of the system, their presence may strongly impact organic reactions (e.g., McCollom, 2013; Seewald, 2001).

5. Concluding remarks

The present study highlights the pertinence of performing laboratory experiments to improve the understanding of organic matter evolution in the context of asteroidal hydrothermal alteration, thereby shedding new light on the possible origin of chondritic organic compounds. The present experiments demonstrate that, if exposed to asteroidal conditions, a single molecule such as HMT will degrade and lead to the formation of hundreds of soluble compounds together with an insoluble organic material. These hydrothermal processes can be seen as instantaneous on asteroidal timescales. A step forward could be to investigate precisely the stability of organic hydrothermal products exposed to dry heating conditions, which likely existed in most asteroidal environments (Bonal et al., 2016; Remusat et al., 2016).

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