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**Electrosynthesis of polypyrrole nano/micro structures using an electrogenerated oriented polypyrrole nanowire array as framework.**

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**Abstract**

The purpose of this paper is to show that it is possible to increase the diameter and the length of the nanostructures of a framework formed of oriented polypyrrole nanowires that has been prepared by a templateless electrochemical method based on the use of a pyrrole solution containing a high concentration of weak-acid anion and a low concentration of non-acidic anion. The dimensions of the initial nanowires are increased by performing an additional electrosynthesis in a 'classical' monomer solution. Depending on the polarization time of this last synthesis (a few tens of seconds), wires with various diameters, from one hundred up to several hundred nanometers, are obtained. In addition to the variation of the nanowire size, these findings confirm, as outlined in the reaction mechanism we have proposed, that the base of the nanowires is surrounded by a thin non-conductive polymer *i.e.* by an overoxidized polypyrrole film. Actually this paper shows a proof-of-concept. Indeed one can imagine that the second polymeric electrodeposit could be performed using an organic monomer solution, using functionalized pyrrole monomer to fabricate a biosensor having large specific area, and/or using anions which could be drugs.

**Keywords**

Polypyrrole, nanostructures, nanowire, microstructure, electrochemical synthesis

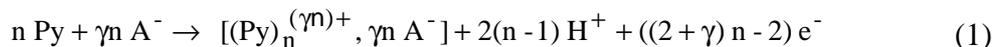
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## 1. Introduction

Polypyrrole (PPy) is one of the most widely used conducting polymers, because of its numerous advantages such as its biocompatibility, environmental stability or ease of preparation even under nanostructured form. Polypyrrole nanostructures are generally synthesized either by a chemical or by an electrochemical route using soft or hard templates [1-11] or sacrificial oxidative templates such as  $\text{MnO}_2$  [12-15]. For instance, PPy nanotubes can be synthesized by using a self-degraded methyl orange (MO) template method [16-19]. However, it has also been established that polypyrrole nanostructures can also be prepared without the use of any template [20-24]. Notably, in that respect, in our previous works [25-27], we brought to light that different polypyrrole nanostructures can be synthesized by using a one-step electrochemical synthesis, without using any template. Indeed, nanostructured polypyrrole films, which are superhydrophilic, are electrogenerated in the presence of (i) a high concentration of weak-acid anions conferring to upon the monomer solution a pH between 6 and 10, and (ii) a low concentration of non-acidic anions such as perchlorate ions. Notably, depending on the perchlorate concentration of a 0.15 M pyrrole solution containing 0.2 M monohydrogenophosphate, different PPy nanostructures can be synthesized. In the absence or in the presence of a very low concentration of perchlorate ions ( $< 10^{-4}$  M), an ultra-thin non-conductive overoxidized PPy (OPPy) film is obtained [28, 29]. This finding was notably determined in particular by XPS analyses [28] and EIS studies [29]. Notice that due to its compactness, this OPPy film prevents the diffusion to the electrode surface of large redox species such as ascorbic acid and dopamine [30] or  $\text{Fe}(\text{CN})_6^{4-}$  and  $\text{Ru}(\text{NH}_3)_6^{2+}$  [31] avoiding their electro-oxidation. ~~It is not the case for~~ On the contrary, small molecules such as  $\text{H}_2\text{O}_2$  and  $\text{H}_2\text{O}$  can diffuse across this OPPy film, and therefore be oxidized at the electrode surface [30, 31]. In the presence of a higher concentration of perchlorate ions ( $> 10^{-4}$  M), an oriented nanowire (NW) array or a network of more or less interconnected nanofibers is formed [26].

The oxidation of pyrrole (Py) leads to the formation of polypyrrole under its oxidized form, doped with anions ( $A^-$ ), according to the following reaction [32, 33]:



where  $\gamma$  stands for the doping level of the polymer, it is generally in the range of 0.25 to 0.33 [34, 35].

In the presence of weak-acid anions, the protons released during Py oxidation (reaction (1)) are captured by these anions:



The mechanism that we have previously proposed in order to explain the formation of a nanostructured polypyrrole film is summarized in Figure 1. It is based on the variation of the interfacial concentration of anions. Indeed, ~~during pyrrole oxidation~~ during pyrrole oxidation protons are released during pyrrole oxidation, and collected by the weak-acid anions present in the solution (reaction (2)), which results in a drastic decrease, or elimination, of anions at the electrode/solution interface. As pyrrole oxidation requires the presence of anions (reaction (1)), which are no longer available, this reaction cannot occur anymore. Instead, water oxidation takes place, leading to the formation of hydroxyl radicals. These radicals can either react with the already formed polypyrrole film, resulting in its overoxidation [36], or with themselves, leading to the formation of  $H_2O_2$  molecules. These molecules are subsequently oxidized into  $O_2$  molecules that form nanobubbles inside the polymer, protecting it against the hydroxyl radical action and therefore preventing locally its overoxidation allowing the conservation of some conductive zones (step 2 in Fig. 1). After bubbles evolution, the electrooxidation of the monomers takes place at these zones, leading to the formation of nanorods/nanowires (step ③ in Figure 1). Obviously, the diameter of these PPy nanostructures depends on the size of the  $O_2$  bubbles. Notice that this process differs from the

one involving gas bubbles as template for PPy microstructure electrosynthesis [37-39]. Therefore, the nature of the anions present in the monomer solution and the release of protons during Py electropolymerization are the two main parameters involved in polypyrrole nanostructure formation. Under potentiostatic conditions, using a Pt anode, the diameter of the PPy nanowires is about 80 nm and their length depends on the electrode polarization time [25]. It has been shown that PPy nanostructures are also electrogenerated in the presence of a low monomer concentration and a high non-acidic anion concentration [26].

The oriented PPy nanowire array could be used as framework for the growth of larger and longer nano/micro structures. Therefore, a second electrosynthesis could be done, using for example functionalized monomer in order to prepare a biosensor [40-46], or using organic solvent containing anions which could be drugs [47-50].

The aim of the present work is to show that it is possible to vary the diameter and the length of oriented polypyrrole nanowires by performing an additional electrosynthesis in a classical pyrrole solution. Moreover, it ~~allows to confirm~~ provides confirmation that the nanowire base is surrounded by an overoxidized polypyrrole (non-conductive polymer) layer. These experiments also allow us to determine whether or not the polypyrrole nanowires are conductive over their entire surface (top and sides). Therefore, the strategy developed in the present work involves two steps. The first one ~~requires~~ is the electrogeneration of oriented nanowires, which requires the use of a monomer solution containing a high concentration of weak-acid anions and a low concentration of perchlorates. The second step ~~is based on~~ implies an anodic polarization in a classical monomer solution, without any weak-acid anions.

## 2. Experimental

The polypyrrole films were synthesized using preliminarily distilled pyrrole. All the solutions were prepared with bi-distilled water. Py,  $K_2HPO_4$ ,  $KCH_3COO$ ,  $K_2CO_3$  and  $LiClO_4$  were purchased from Aldrich.

The electrochemical experiments were performed in a classical three-electrode electrochemical cell. A platinum foil was used as counter electrode and a saturated calomel electrode (SCE) was used as the reference electrode. A double junction was used in order to avoid chloride diffusion into the Py solution. The working electrode was a Pt electrode ( $S = 0.07 \text{ cm}^2$ ) for electrochemical studies and Au/mica substrate for SEM observations. The preparation of the substrates was the following: a thin gold film ( $\sim 80 \text{ nm}$ ) was deposited under low pressure ( $10^{-4} \text{ Pa}$ ) by thermal evaporation on a mica substrate.

All the electrosyntheses were performed under potentiostatic conditions, in  $0.15 \text{ M Py} + 0.2 \text{ M K}_2\text{HPO}_4 + 10^{-3} \text{ M LiClO}_4$  for the first synthesis and in  $0.15 \text{ M Py} + 0.2 \text{ M LiClO}_4$  for the second one, during given times named  $t_1$  and  $t_2$ , respectively. The pH of the first monomer solution is 8.9.

The superhydrophilic character of the PPy films was checked by a drop test experiment: a small pure water droplet is deposited onto the film and the spreading of the drop is analyzed. When the film is nanostructured the drop spreads out whereas it does not for classical cauliflower-like films.

As far as the electropolymerization of Py is concerned, an Autolab PGSTAT30 potentiostat (Ecochemie) controlled with the GPES software was employed. The film morphology was examined under a field emission gun scanning electron microscope (FEG-SEM), Ultra55 Zeiss, operating at 5 kV.

### 3. Results and discussion

Firstly, the framework composed of oriented PPy nanowires was synthesized by performing an electropolymerization of pyrrole monomers under potentiostatic conditions at  $0.75 \text{ V/SCE}$  in a solution composed of  $0.15 \text{ M pyrrole}$ ,  $0.2 \text{ M K}_2\text{HPO}_4$  and  $10^{-3} \text{ M LiClO}_4$  (see the SEM micrograph in Fig. 1). The average diameter of the nanowires is about  $80 \text{ nm}$ . Notice that the framework can also be obtained using acetate or carbonate instead of

monohydrogenophosphate. Indeed by replacing  $\text{HPO}_4^{2-}$  by  $\text{CO}_3^{2-}$  or  $\text{CH}_3\text{COO}^-$ , we have obtained the same results *i.e.* nanostructured PPy films (see Table 1). Then, another polypyrrole deposit was performed onto these nanostructures. This second electrosynthesis was conducted in a monomer solution containing only perchlorates as anions, solution which generally leads to the formation of a PPy film with cauliflower-like structure.

Figures 2A and 2B (curves (a)) show the anodic current *versus* time responses for the first and second synthesis, respectively (5 experiments:  $t_1 = 200$  s,  $t_2 = 30, 60, 90, 180$  or  $300$  s). For comparison, the curve recorded at a bare Pt electrode, in the experimental conditions employed for the second synthesis, is also depicted in Fig. 2B (curve (b)). The shape of the  $I(t)$  curves and the total anodic charge (Fig. 3) clearly depend on the experimental conditions. After the second electrosynthesis, if the deposition times are longer than 90 seconds, the PPy films are no more superhydrophilic *i.e.* no more nanostructured. As previously stated, curves (a) and (b) in Fig. 2B correspond to the two-step and to the one-step 300 second synthesis, respectively. The shape of these  $I(t)$  responses differs regarding the 90 first seconds, which is consistent with the fact that the substrates are not the same (NW-PPy/Pt and bare Pt). After this period (90 s), (a) and (b) curves are parallel in good agreement with the cauliflower structure obtained when the second synthesis lasts more than 90 s (at 0.75V/SCE) (see below).

The anodic charges ( $Q$ ) *versus* the electrode polarization time are plotted in Figure 3. In this figure, the black open squares indicate the PPy films which are superhydrophilic *i.e.* which are nanostructured. With regard to the second synthesis (green squares), the  $Q(t)$  relationship is linear. One can notice that when the polarization lasts 300 s, the charge passed during the second synthesis is about 20% higher than those recorded in the same conditions but using a bare Pt electrode, 215 mC and 177 mC respectively, in good agreement with the initial real surface of the electrode (3D / 2D).

Some of these polypyrrole films have been observed by SEM. In Figure 4 are reported micrographs obtained for a Au/mica substrate first polarized during 600 s and then only a part of its surface has undergone a second polarization of 50 seconds in the 'classical' Py solution. Micrograph A shows the film after the first polarization whereas micrographs B-D show the film after the second polarization. Notice that micrographs A and B were done at the same magnitude. From these micrographs, it is clear that the growth of the polymer occurs on the whole surface of the nanowires, since their diameter increases from 80 nm to 200 nm and their length increases from 500 nm [25] up to 1500 nm. Besides, the SEM micrographs presented in Figure 5 show different kinds of (nano)structures depending on the second synthesis duration, the first synthesis duration being 200 s. When the second polarization duration increases, the polypyrrole nanowire diameter increases too. Finally, for 300 seconds of polymerization, the PPy film has a cauliflower-like structure (see micrographs E and F in Fig. 5) because during the electrode polarization the diameter of the nanowires has so increased that their coalescence has occurred.

The diameter and the length of the polypyrrole nanostructures obtained after the second synthesis *versus* the polymerization duration are plotted in Figure 6. The nanowire diameter increases with the second polarization time, leading to a decrease of the space between the nanowires until they coalesce and form a 2D film that is no more superhydrophilic, confirming

that the PPy film is no longer nanostructured. Before coalescence, the growth rate in terms of fiber length is about  $17 \text{ nm s}^{-1}$ , which is almost 20 times faster than the rate determined with the Py solution used for the framework synthesis (first synthesis) [25]. This finding could be explained by the fact that the Py electropolymerization is limited by the concentration of the anions present at the electrode/solution interface [51]. Finally, a series of PPy framework was prepared by varying the polarization durations ( $t_1$ ), leading to different anodic charges ( $Q_1$ ). A second polarization, lasting a given time ( $t_2 = 10, 15, 20, 30, 40$  or  $50$  seconds), was then

performed in the ‘classical’ Py solution (solution containing only perchlorates as anions). The anodic charges after the second synthesis ( $Q_2$ ) vs. anodic charges after the first synthesis ( $Q_1$ ) are plotted in Fig. 7A. As it can be noticed from this figure, for a given  $t_2$ , the relationship between  $Q_1$  and  $Q_2$  is linear and does not pass through the origin:

$$Q_2 = \alpha + \beta Q_1 \quad (3)$$

Whatever  $t_1$  is, the first synthesis leads to PPy nanowire arrays having the same NW density and the same NW diameter, as these parameters only depend on the nature of the electrode, the applied potential, and the Py solution composition. Therefore, the variation of  $t_1$  *i.e.*  $Q_1$  only leads to the variation of the PPy nanowire length [25]. As the relationship between the charge passed during the second synthesis and the one passed during the first synthesis (which determines the length of the wires) is linear, one can conclude that the Py polymerization takes place on the entire surface of the nanowires and not only on their top. In this latter case,  $Q_2$  should be independent of the wire length *i.e.* of  $Q_1$ . This conclusion is in good agreement with the SEM observations (Figs. 4, 5, 6).

$\alpha$  and  $\beta$  parameters of equation (3) *versus* the polarization time of the second synthesis are depicted in Figures 7B and 7C, respectively. These relationships are linear.  $Q_2$  corresponds to Py oxidation that leads to the generation of PPy at a place that can be divided into three zones, named a, b, and c in Figure 8.

$$Q_2 = Q_a + Q_b + Q_c \quad (4)$$

As it can be noticed from Figure 8, zone a corresponds to the enlargement of the nanowires. Therefore,  $Q_a$  is function of  $Q_1$ . Zones b and c correspond to the elongation of the nanowires and the enlargement of the nanowires formed during the second synthesis, respectively. Therefore both  $Q_b$  and  $Q_c$  are independent from  $Q_1$ , which sets the length of the nanowire.  $\alpha$  is proportional to  $t_2$  as evidenced by Fig. 7B, in good agreement with Fig. 6.

Consequently, taking into account that  $\alpha = Q_b + Q_c$  and  $\beta = a t_2$ , the combination of equations (3) and (4) gives

$$Q_2 = Q_b + Q_c + a t_2 Q_1 \quad (5)$$

All these findings allow us to confirm that the PPy nanowires of the framework (first synthesis) are conductive on their entire surface *i.e.* on their top and their sides, as it is possible to deposit another polypyrrole film all around the nanostructures whose diameter increases with the polarization time of the second synthesis (Fig. 6). Moreover, these findings confirm that the layer which surrounds the nanowire base is non-conductive *i.e.* it is an overoxidized PPy film, as shown in the mechanism proposed in our previous work regarding the formation of the different polypyrrole nanostructures (see Fig. 1).

#### 4. Conclusions

It is possible to use oriented polypyrrole nanowires obtained by a one-step templateless electrochemical method as framework to perform a second synthesis in a pyrrole solution that does not contain a high concentration of weak-acid anions. Depending on the polarization duration of this synthesis, the diameter of the nanowires varies from a few tens of nanometers up to several hundred nanometers, and after a threshold of polarization time cauliflower-like structures are obtained. These findings are important because they allow us to confirm the global mechanism we proposed to achieve the electrogeneration of PPy nanostructure without the use of any template, just by using a monomer solution containing a high concentration of weak-acidic anions having  $pK_a > 6$ . Finally, this paper shows a proof-of-concept, indeed one can imagine that the second electrodeposit could be performed using functionalized pyrrole monomer to fabricate electrochemical biosensors with large specific area, for example or using an organic solvent, and/or using anions which could be drugs.

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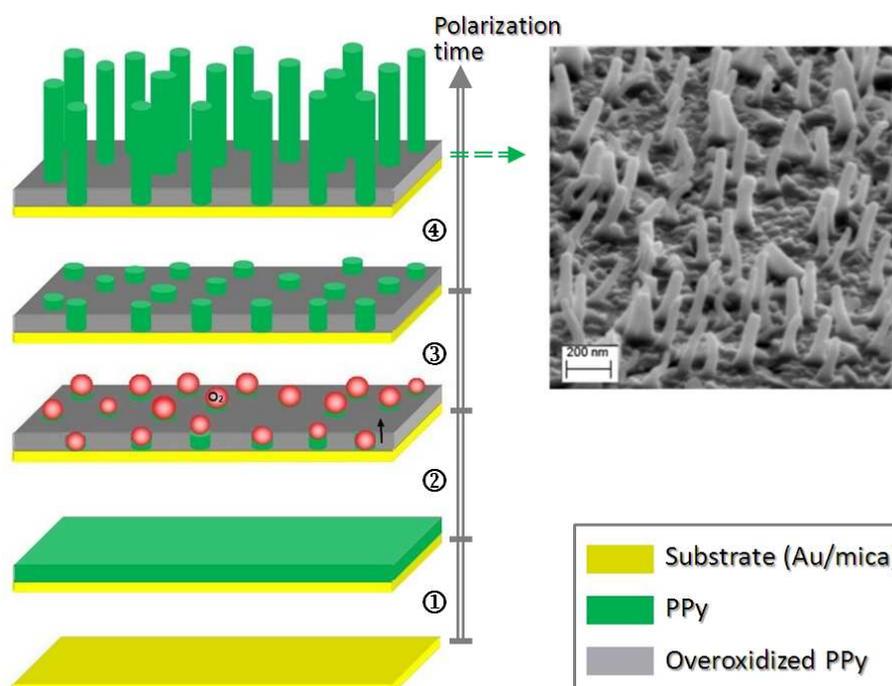
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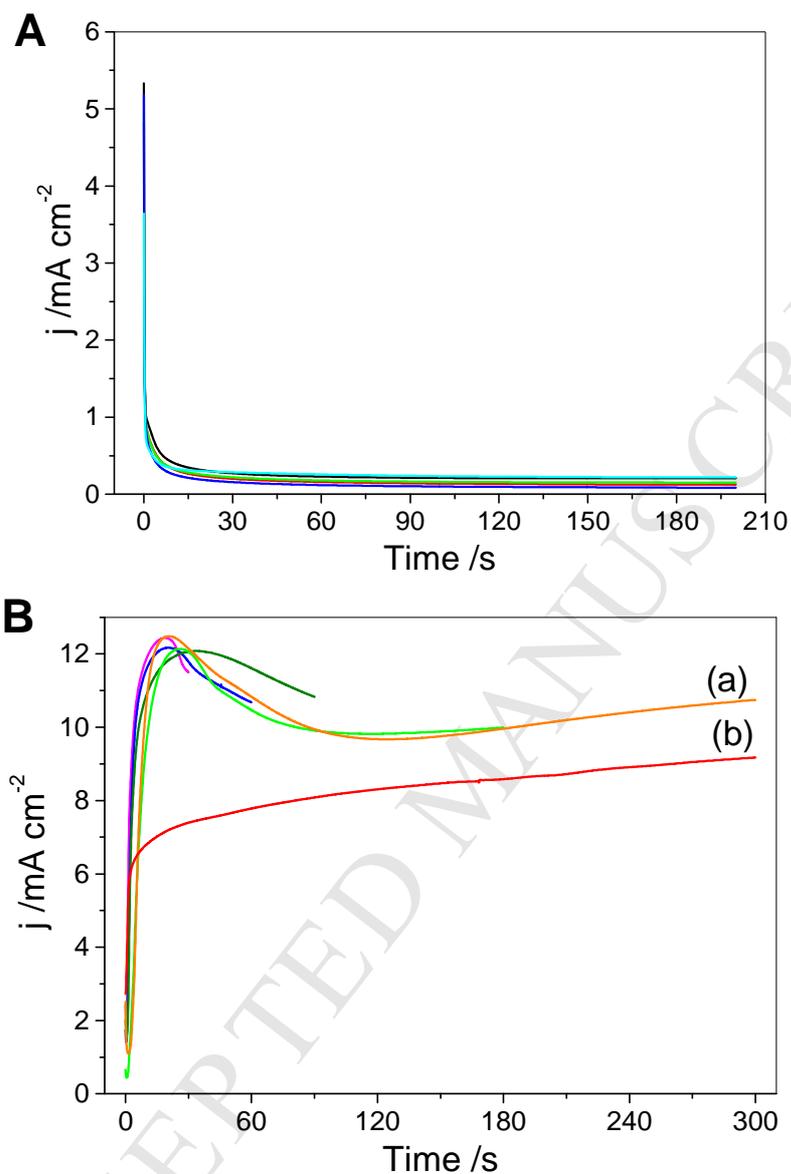
**Table 1:** First synthesis : influence of the nature of the weak-acid anion. PPy films obtained by polarization at 0.78 V/SCE for 600 s in 0.15 M Py aqueous solution containing 0.2 M of weak-acid anion +  $10^{-3}$  M perchlorate or containing 0.2 M perchlorate (no weak-acid anion).

Weak-acid anion	Py solution pH	Anodic charge (mC)	Superhydrophilic PPy film <sup>(1)</sup>	
HPO <sub>4</sub> <sup>2-</sup>	8.9	9.7	Yes	This work
CH <sub>3</sub> COO <sup>-</sup>	8.3	5.9	Yes	This work
CO <sub>3</sub> <sup>2-</sup>	11.1	2.8	Yes	This work
PO <sub>4</sub> <sup>3-</sup>	12.6	1.15	No <sup>(2)</sup>	Ref. 26
H <sub>2</sub> PO <sub>4</sub> <sup>-</sup>	4.8	350	No	Ref. 26
No ([ClO <sub>4</sub> <sup>-</sup> ] = 0.2M)	8.9 <sup>(3)</sup>	500 mC	No	Ref. 26

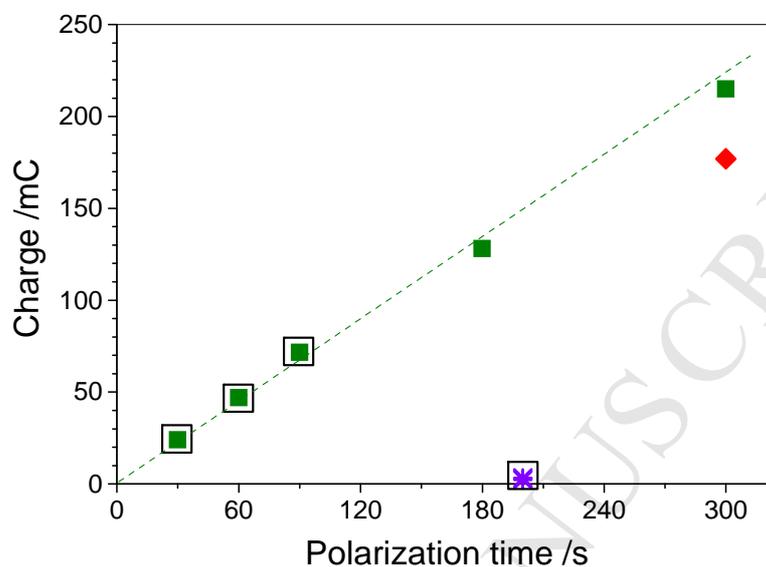
<sup>(1)</sup> Nanostructured film ; <sup>(2)</sup> Overoxidized PPy film ; <sup>(3)</sup> pH adjusted with KOH



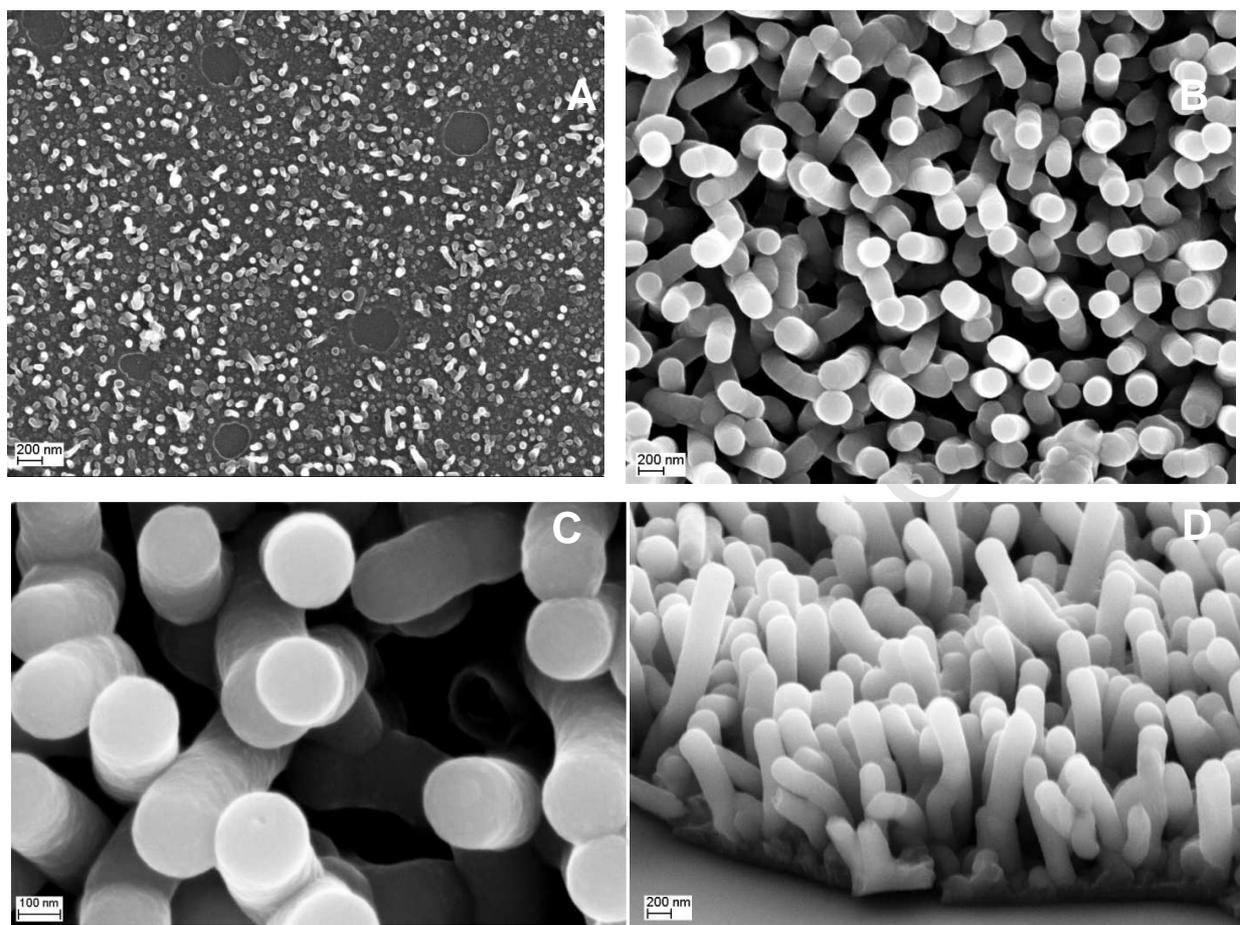
**Figure 1.** General mechanism of PPy nanowire electrogeneration. Left: Drawing showing the evolution of the polymer deposit with the anodic polarization time, in Py aqueous solution containing a high concentration of  $\text{K}_2\text{HPO}_4$  and a low concentration of  $\text{LiClO}_4$ . Right: SEM micrograph (side view,  $60^\circ$ ) of PPy nanowires electrogenerated at 0.75 V/SCE for 200 s in 0.15 M Py + 0.2 M  $\text{K}_2\text{HPO}_4$  +  $10^{-3}$  M  $\text{LiClO}_4$  aqueous solution. Steps ①, ③, ④: Py oxidation ; step ② : water oxidation leading to  $\text{OH}^\bullet$  and  $\text{O}_2$  formation.  $\text{O}_2$  locally protect the PPy film against the action of  $\text{OH}^\bullet$  which overoxidizes PPy. Steps ①, ②, ③ last around a few seconds.



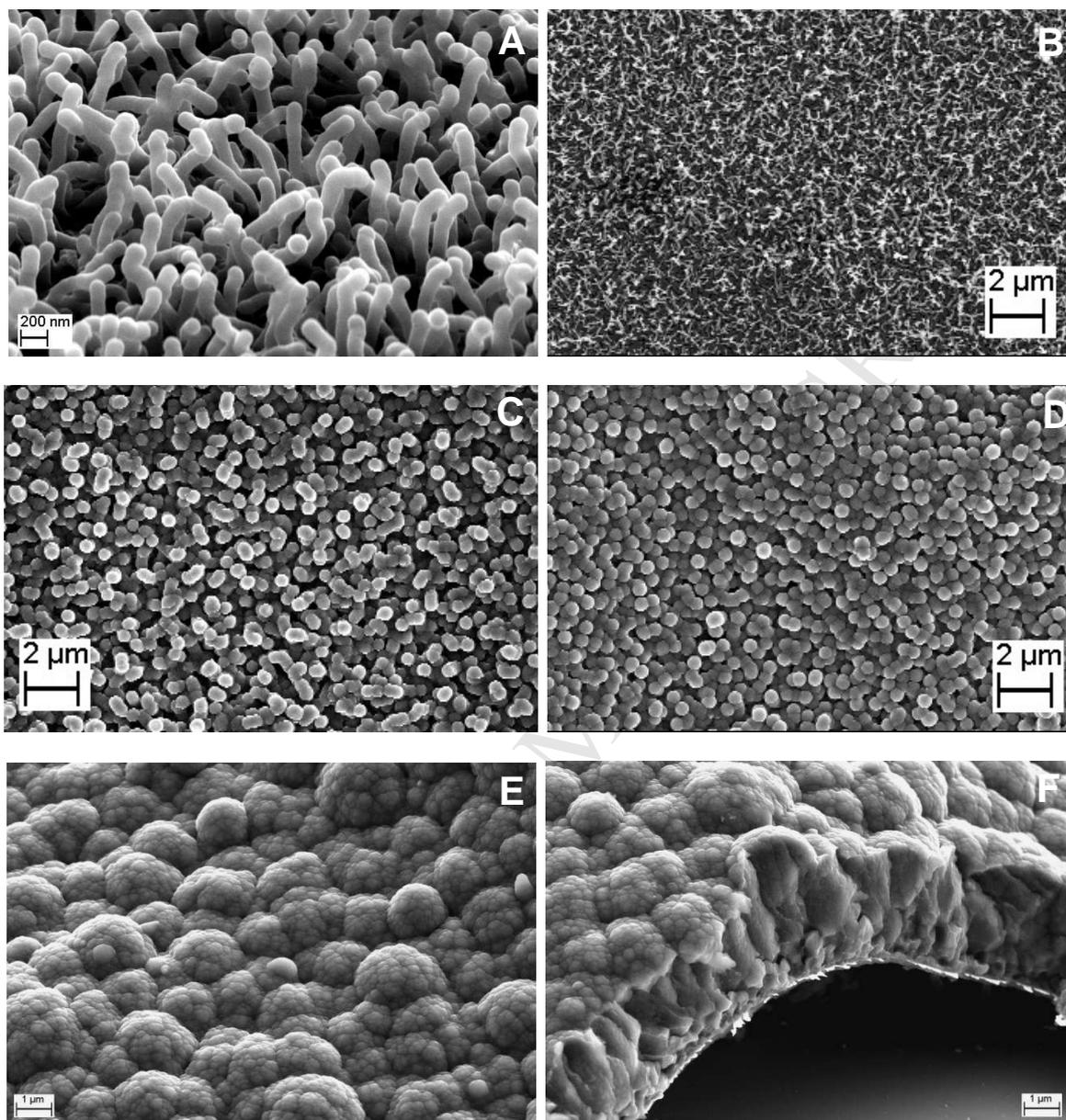
**Figure 2.** PPy electrosynthesis under potentiostatic conditions. Current density vs. time at a Pt electrode polarized at 0.75 V/SCE in 0.15 M Py + 0.2 M  $\text{K}_2\text{HPO}_4$  +  $10^{-3}$  M  $\text{LiClO}_4$  (A) and then in 0.15 M Py + 0.2 M  $\text{LiClO}_4$  (B curves (a)). B curve (b) bare Pt electrode polarized at 0.75 V/SCE in 0.15 M Py + 0.2 M  $\text{LiClO}_4$ . In all cases, the surface used to calculate J is  $0.07 \text{ cm}^2$ .



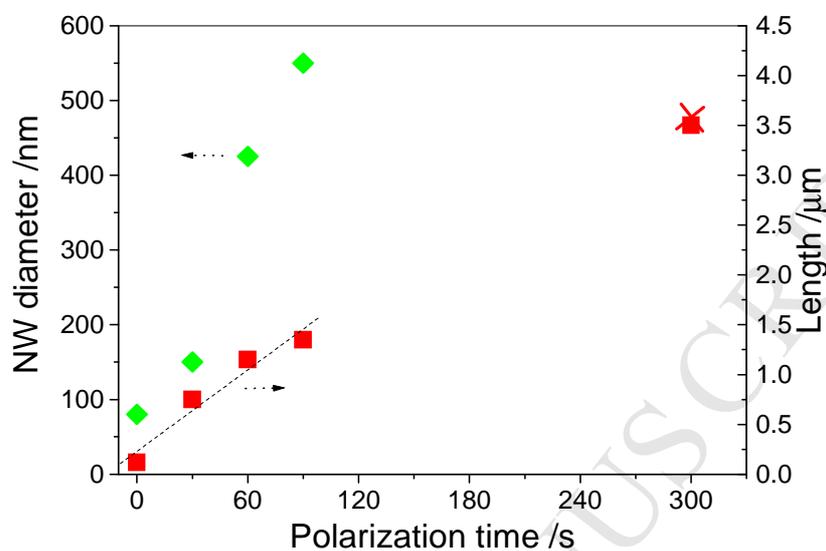
**Figure 3.** Anodic charge vs. polarization time, (★) first syntheses in 0.15 M Py + 0.2 M  $\text{K}_2\text{HPO}_4$  +  $10^{-3}$  M  $\text{LiClO}_4$  (the  $j(t)$  curves are shown in Fig. 2A ;  $Q \approx 1$  mC) ; (■) second syntheses in 0.15 M Py + 0.2 M  $\text{LiClO}_4$  (curves (a) in Fig. 2B) ; (◆) synthesis in 0.15 M Py + 0.2 M  $\text{LiClO}_4$  (curve (b) in Fig. 2B) (at a bare Pt electrode). The black open squares indicate the superhydrophilic PPy films *i.e.* the nanostructured films.



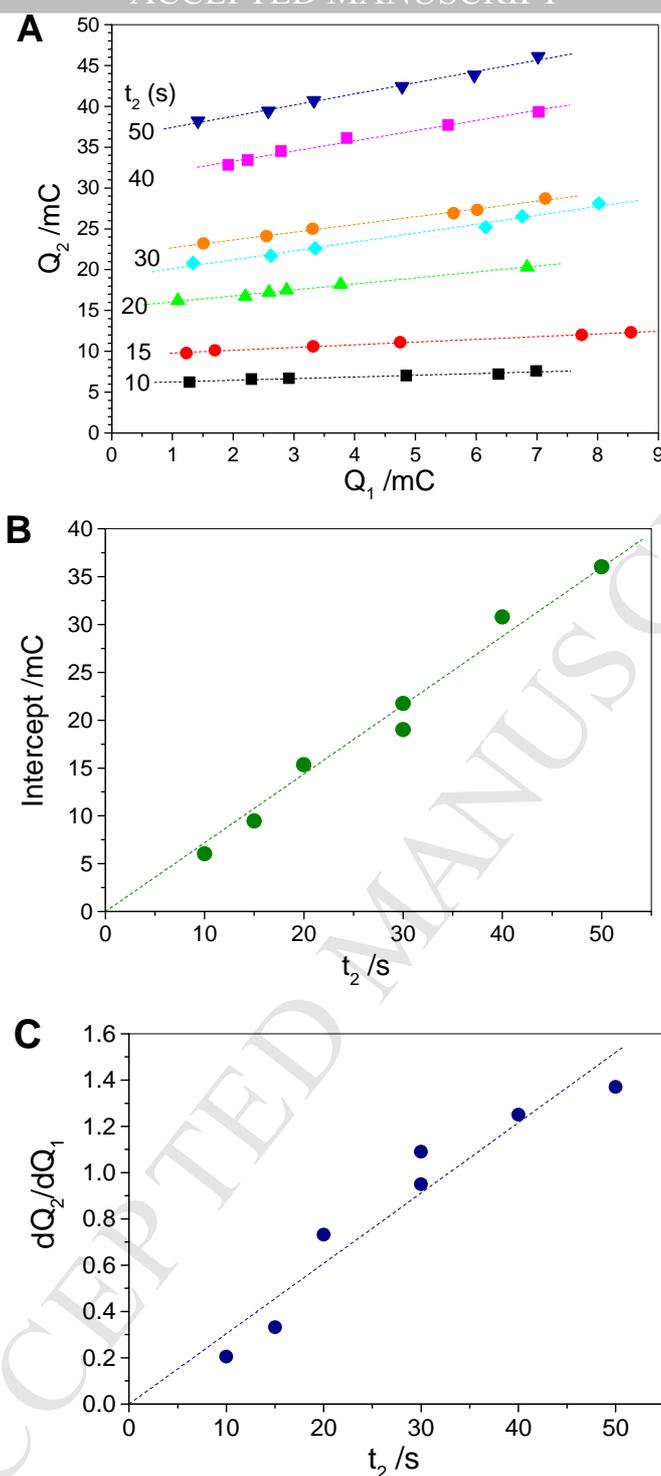
**Figure 4.** SEM micrographs of PPy. A) after a first synthesis at 0.75 V/SCE for 600 s in 0.15 M Py + 0.2 M  $\text{K}_2\text{HPO}_4$  +  $10^{-3}$  M  $\text{LiClO}_4$  aqueous solution. B-D) after a second synthesis at 0.75 V/SCE for 50 s in 0.15 M Py + 0.2 M  $\text{LiClO}_4$  aqueous solution. A,B,C top view, D side view ( $60^\circ$ ). Scale bar: 200 nm (A), (B), (D) ; 100 nm (C).



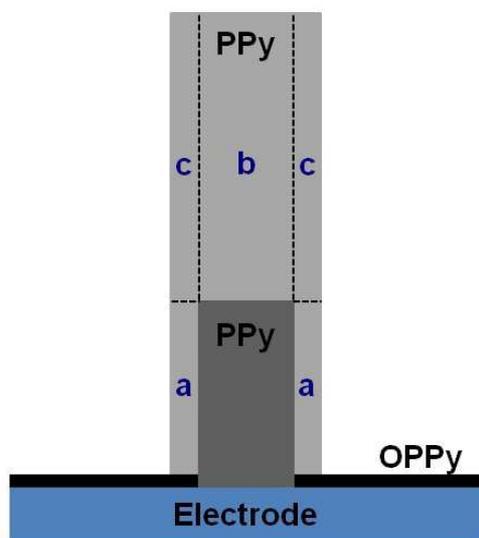
**Figure 5.** SEM micrographs of PPy films electrogenerated at 0.75 V/SCE in two steps. First step was performed in 0.15 M Py + 0.2 M  $\text{K}_2\text{HPO}_4$  +  $10^{-3}$  M  $\text{LiClO}_4$  solution for 200 s ; second step in 0.15M Py + 0.2 M  $\text{LiClO}_4$  solution for (A,B) 30 s, (C) 60 s, (D) 90 s, (E,F) 300 s. Scale bars: (A) 200 nm, (B-D) 2  $\mu\text{m}$ , (E, F) 1  $\mu\text{m}$ .



**Figure 6.** PPy nanowire (NW) characteristics after the second synthesis at 0.75 V/SCE in 0.15 M Py + 0.2 LiClO<sub>4</sub> at a NW-PPy/Pt electrode, the NW-PPy film being formed at 0.75 V/SCE for 200 s in 0.15 M Py + 0.2 M K<sub>2</sub>HPO<sub>4</sub> + 10<sup>-3</sup> M LiClO<sub>4</sub>. Left: NW diameter (◆), right: NW length (■) or film thickness (✱) vs. polarization time. Some SEM micrographs are shown in Fig. 5.



**Figure 7.** **A)** Anodic charge of the second synthesis performed during various times ( $t_2$ ) in a classical Py solution (0.2 M LiClO<sub>4</sub> + 0.15 M Py) vs. anodic charge of the first synthesis performed in 0.2 M K<sub>2</sub>HPO<sub>4</sub> + 10<sup>-3</sup> M LiClO<sub>4</sub> + 0.15 M Py solution, the polarization time of the first synthesis varying from 30 to 180 s (variation of the nanowire length). Electrode: Pt (0.07 cm<sup>2</sup>). Applied potential: 0.75 V/SCE. **B** and **C**) Y-intercept and slope of the relationship  $Q_2 = f(Q_1)$  (graph A) vs. the polarization time of the second synthesis, respectively.



**Figure 8.** Drawing showing one PPy nanowire electrogenerated at 0.75 V/SCE in two steps, the first synthesis being performed in 0.2 M  $\text{K}_2\text{HPO}_4 + 10^{-3}$  M  $\text{LiClO}_4 + 0.15$  M Py solution and the second one in a ‘classical’ Py solution (0.2 M  $\text{LiClO}_4 + 0.15$  M Py). Black: ultra-thin film of overoxidized PPy; dark grey: PPy nanowire after the first synthesis; light grey: PPy generated during the second synthesis.