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Surface characterization of poly-2-vinylpyridine—A polymer for area selective deposition techniques

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Thin films of OH terminated poly-2-vinylpyridine (P2VP), a polymer with potential for infiltration mediated thin film deposition, area selective deposition (ASD), and small feature size development via block copolymer (BCP) self-assembly, have been studied with hard x-ray photoelectron spectroscopy. From the N 1s and C 1s core level spectra, accurate values for the binding energy positions of the species present in the films were obtained, providing clear evidence for signals associated with pyridine bonds. The aromatic ring on the P2VP side chain is clearly identified in the studied core levels. These observations allow for a complete understanding of the chemical environment of the polymer and provide evidence of the potential reactions that can occur with metal diffusion into P2VP. Transmission electron microscopy, attenuated total reflection infrared spectroscopy, and atomic force microscopy measurements reveal high quality films, and this work provides a reference base for this functional material in terms of its utility for ASD, BCP, and subsequent atomic layer deposition based polymer infiltration processes. *Published by the AVS.* <https://doi.org/10.1116/1.5115769>

I. INTRODUCTION

The limitations of traditional optical lithographic patterning is a well documented concern in the semiconductor industry.¹ The study of polymers as potential materials for use in area selective deposition (ASD) has become a major field of research for developing alternative nanoelectronic fabrication methods via the acceptance and blocking of infiltrating metals in a variety of different polymer types.^{2,3} This range includes polymers such as polystyrene (PS), poly(methyl methacrylate), poly(ethyleneoxide), and poly(vinylpyridine).^{4–8}

Poly-2-vinylpyridine (P2VP) is a leading material in ASD and block copolymer (BCP) research, being a pH-responsive component,^{9,10} while also demonstrating effective segregation when used in a block copolymer architecture with materials such as PS and polyisoprene,¹¹ a crucial aspect of BCP morphology.¹² P2VP has had promising experimental results, being reported to facilitate the fabrication of small and uniform features such as gold nanoparticles.¹³ PS-b-P2VP has been used to create small-device-like structures as reported by Morris and others.^{14,15} The presence of an unshared electron pair on the nitrogen atom of the pyridine ring can facilitate coordination bonding with various species such as metal-ligand complexes.^{16–18} These factors make P2VP an ideal

starting material for infiltration processes, such as atomic layer deposition (ALD), that yield key device elements such as metallic contacts and high-k dielectrics.

Although reported in ASD studies, in-depth photoemission analysis of P2VP films and other common BCP-relevant polymers is limited. The use of polymers for the fabrication of thin films highlights the need for more intricate studies of their surface properties to allow for a better understanding of the material chemistry of complex, chemistry rich, processes, such as ALD.¹⁹ Sequential infiltration synthesis ALD has been reported in the literature as a particularly successful method of infiltration for a range of polymers.^{20–22}

The present work focuses on a hard energy x-ray photoelectron spectroscopy (HAXPES) and subsequent angle resolved HAXPES (ARHAXPES) study of the polymer P2VP. The synchrotron radiation provides high resolution photoemission spectra, allowing for accurate determination of the chemistry of the polymer, while the ARHAXPES method characterizes the thickness and depth-resolved composition of the polymer.^{23,24}

The nitrogen atom in the pyridine ring of P2VP (see schematic in Fig. 1) has a localized lone pair of electrons that can interact with other atoms via sigma and pi-pi* interactions.²⁵

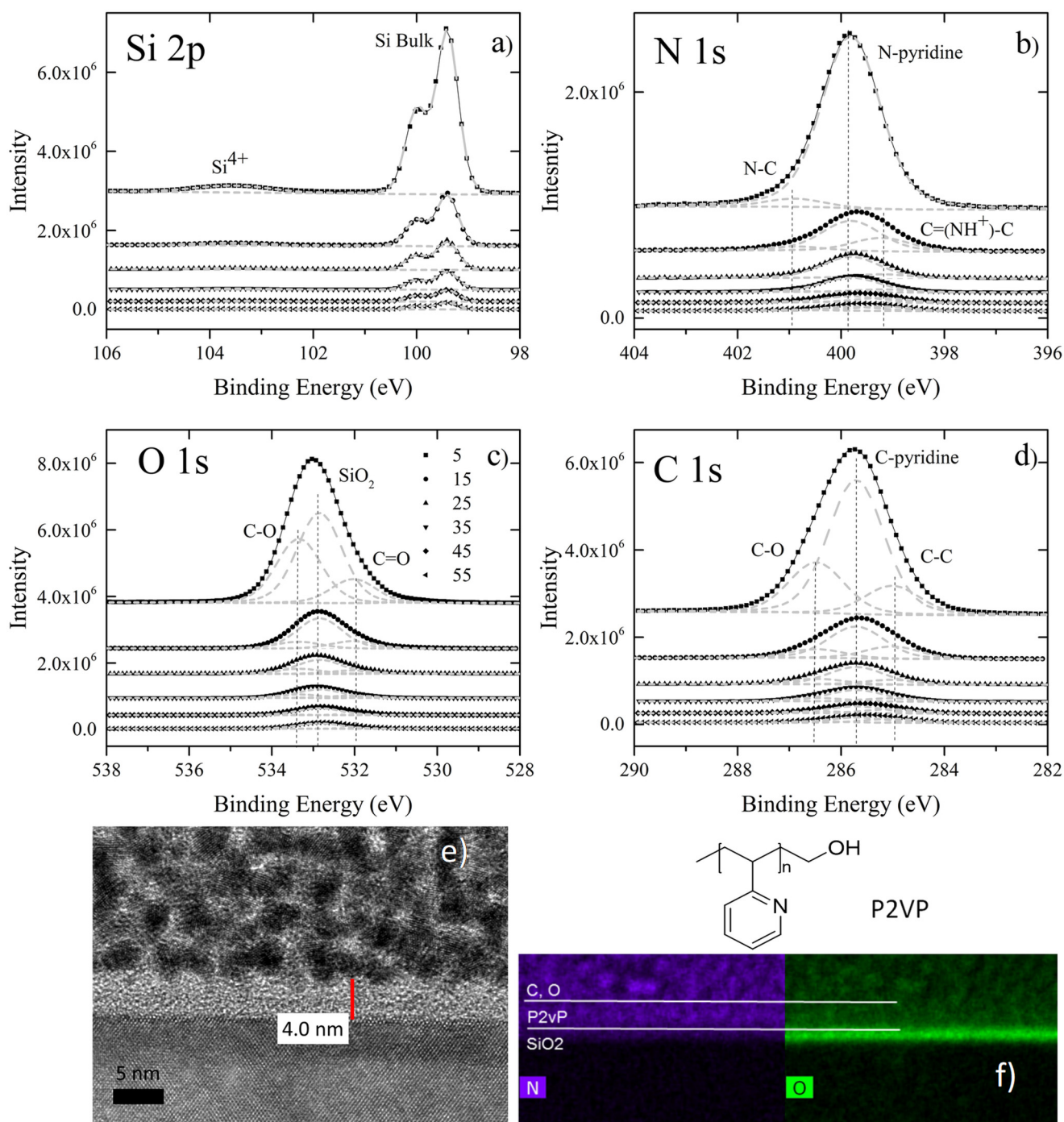


FIG. 1. (a) Si 2p, (b) N 1s, (c) O 1s, and (d) C 1s core level AR HAXPES spectra of a C/P2VP/SiO₂/Si sample. (e) TEM cross-sectional image of P2VP film. (f) EDX mapping by TEM of P2VP film for N and O signals.

In the case of N-pyridine-metal complexes, the electron lone pair on the nitrogen is involved in the bonding interaction. Our previous work presented evidence for Cu metal ion infiltration and coordination with pyridine via liquid phase salt solution processing.²⁶ The role of the electron lone pair is of critical importance in facilitating the coordination bond formation with the infiltrating metal, and this is not thoroughly explored in photoemission reports to date.²⁷ In this letter, we report HAXPES measurements including the ARHAXPES dependence of the C 1s and N 1s signals originating from the P2VP film.

II. EXPERIMENT

P2VP films of 0.2 wt.% were prepared by dissolving OH terminated poly-2-vinylpyridine (P2VP-OH) polymer in tetrahydrofuran and subsequently spin coating at 3000 rpm for 30 s onto Si substrates following the process used in Ref. 28. P2VP-OH is a hydroxyl chain terminated polymer brushlike system that can be coordinated to a silicon oxide surface through a condensation process.^{29,30}

HAXPES measurements were undertaken at SOLEIL Synchrotron at the GALAXIES beamline. Spectra were

acquired at pressures of 1×10^{-9} mbar with photon beam energies of 3 keV. A Si(111) monochromator was used for the 3 keV beam. The angles chosen for ARHAXPES were in the range of 85° – 25° , in intervals of 10° (with normal emission being 90°). The sample position at every angle was calibrated to ensure that the photoelectron counts were at their maximum value. The fitting of the photoemission data was performed by AAnalyzer[®], a program that allows for the simultaneous peak fitting of data acquired at all takeoff angles.³¹ For all of our studies, the use of simultaneous methods is of high importance.³² Calculations of the thickness and chemical composition of each layer were obtained using the multilayer method (MLM).^{33–35}

High resolution transmission electron microscopy (TEM) experiments to obtain the thickness of the P2VP-OH layer were performed on an FEI OSIRIS TEM, with sample preparation performed in an FEI Helios Nanolab 450 S. The sample was capped with Pt via a gas injection system.

Grazing-angle attenuated total reflection infrared (ATR IR) spectroscopy was performed using a Harrick VariGATR accessory on a Nicolet iS50 spectrometer (unpolarized, angle of incidence is 65° , 32 scans, resolution 8 cm^{-1}). Atomic force microscopy (AFM) was used to analyze the P2VP-OH surface roughness on a Bruker Dimension 3100 AFM.

III. RESULTS

The measured ARHAXPES spectra of the P2VP-OH film are shown in Fig. 1, with all spectra being shifted so that the Si⁰ peak on the Si 2*p* signal [Fig. 1(a)] occurs at 99.4 eV binding energy.³⁵ Si⁴⁺ was observed at 103.4 eV. Figure 1(b) shows the N 1*s* fit, where C–N=C bonds (N-Pyridine) are the main components of the peak, occurring at 399.8 eV, while N–C at 400.9 eV was also included in the fit.^{36,37} Another signal at 399.2 eV was associated with pyridine–hydrogen interaction (C=NH⁺–C bonds). Similar binding energy components in the N 1*s* signal have been reported in the literature but have so far remained uncategorized.^{38,39}

The O 1*s* signal was fitted with four signals [Fig. 1(c)]. The most prominent signal at 532.8 eV is associated with silicon oxide, which is located at the interface between P2VP-OH and the silicon wafer. The other three signals were associated with oxygen in polymer chain bonds, residual oxygen, and adsorbed surface oxygen at 530.6, 532, and 533.3 eV, respectively.

The C 1*s* signal in Fig. 1(d) was fitted with five signals, one of them at 285 eV being associated with C–C bonds within the polymer chain as reported by Briggs *et al.*⁴⁰ The signal with the highest intensity, observed at 285.7 eV is characteristic of pyridinic carbon.^{41,42} The C–O signal at 286.5 eV was associated with oxygen bonding to the polymer chain and suboxides at the interfaces. The signals for C=O and O–C=O at 287.6 and 288.7 eV were attributed to surface and residual oxygen.^{43,44} The peak fitting parameters for the signals associated with P2VP-OH are displayed in Table I.

By using AR HAXPES, the thicknesses of the different layers present in the sample were obtained. For each core level signal (N 1*s*, Si 2*p*, C 1*s*, O 1*s*), the attenuation length

TABLE I. Binding energy and peak fit parameters for the P2VP film.

Core level peak	$h\nu$ (eV)	Gaussian (eV)	Lorentzian (eV)	Chemical environment	Binding energy (eV)
C 1 <i>s</i>	3000	1.11	0.27	C–C	285.0
				C-Pyridine	285.7
N 1 <i>s</i>	3000	1.13	0.51	C=NH ⁺ –C	399.2
				N-Pyridine	399.8

of electrons, cross section, areal density, and lattice constant were considered. The assumed composition of the material for theoretical modeling was a layered C/P2VP/SiO₂/Si structure. The thickness calculations were performed using the MLM with the XPS intensity of each species that depends on the takeoff angle.⁴⁵ The bandgap and density of P2VP-OH used for the calculations was 2.25 eV and 1.257 g/cm^3 , respectively, in accordance with the literature.^{46,47}

The ARHAXPES results are shown in Fig. 2. P2VP and SiO₂ dependences were obtained with the C-Pyridine and Si⁴⁺ signals, respectively. The uncertainty of the experimental data represents the variation in signal intensity across individual scans. The proposed structure of the analyzed material is presented in the inset of Fig. 2(a), featuring a C layer above the P2VP-OH film, which itself is above an SiO₂ layer

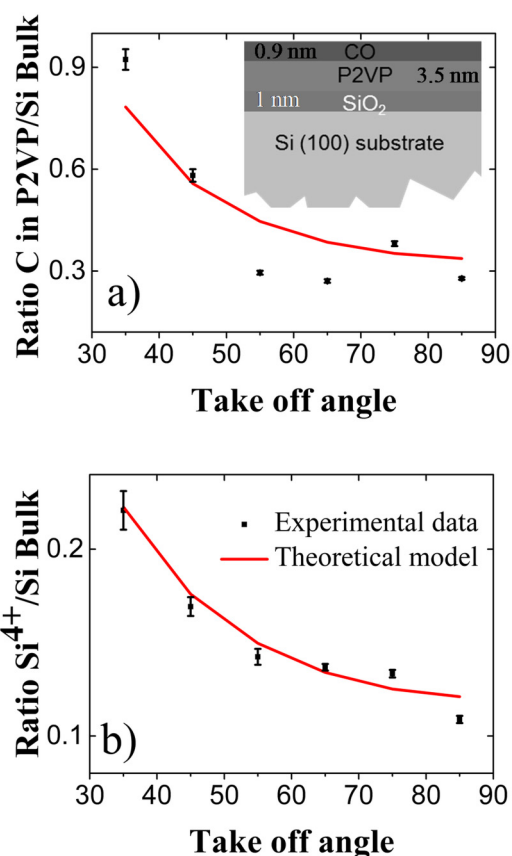


Fig. 2. Angle resolved HAXPES dependence: (a) carbon in P2VP layer and (b) Si⁴⁺ in silicon oxide layers. The thickness and chemical composition calculation of each chemical species were considered in the sample.

on top of the Si bulk. The intensity of the signals associated with the films was divided by the silicon bulk intensity. This was then plotted against the takeoff angle. Comparing the raw data with the theoretical model (MLM), stoichiometry and thickness of the thin layers were obtained.

The thicknesses determined were 1.00 ± 0.03 , 3.54 ± 0.06 , and 0.92 ± 0.01 nm for silicon oxide, P2VP, and adventitious carbon layers, respectively. The error was determined by considering two scenarios: when the intensity ratio of the films over the bulk was at maximum and when it was at minimum. TEM images were correlated to photoemission results as shown in Fig. 1(e), showing a thin film of around 4 nm on top of a silicon dioxide layer. Energy-dispersive x-ray spectroscopy (EDX) mapping by TEM provides evidence that nitrogen is present in this P2VP region [see Fig. 1(f)]. The stoichiometry of the layers was consistent for each layer according to the chemical composition.

Grazing-angle ATR IR spectroscopy allowed the acquisition of IR spectra even from the ultrathin P2VP-OH film used in this work, after subtracting the silicon background signal. The resulting spectrum is shown in Fig. 3(a). The bands at 2927 and 2854 cm^{-1} correspond to the vibrational mode of the aliphatic polymer backbone. Deformations in the free backbone lead to the signals at 1473 and 1436 cm^{-1} . The stretching modes of the pyridine ring are located at 1592 and 1570 cm^{-1} with an additional ring deformation band at 987 cm^{-1} .⁴⁸ The feature at 787 cm^{-1} corresponds to out-of-plane

deformations of the hydrogen atoms of the pyridine ring. It is shifted to slightly higher frequencies compared to 1,2-substituted aromatic rings due to the higher electronegativity of the nitrogen atom. The signals at 1233 and 1153 cm^{-1} correspond to C—O—C and C—OH vibrations, respectively, resulting from the terminating hydroxy groups of the polymer. These data indicate that partial condensation of the hydroxy group with the hydroxy functionalized silicon substrate may occur during film deposition. This hypothesis agrees with the observation of a visible band at 891 cm^{-1} corresponding to a Si—C bond vibration.⁴⁹

AFM measurements were also performed on the analyzed P2VP-OH samples to assess the physical surface profile of the spun-on polymer film [Fig. 3(b)]. The average root-mean-square surface roughness of the P2VP-OH according to AFM analysis was approximately 0.21 nm, which indicates the presence of a very smooth film and is similar to values obtained for the Si substrate.

IV. CONCLUSION

In summary, we have investigated P2VP-OH with a 3 keV HAXPES photon beam, resulting in a nondestructive and high resolution chemical analysis of the film. The P2VP-OH signals were deconvoluted in a robust and highly detailed process. The ARHAXPES analysis of the polymer indicates that pyridine N will facilitate metal incorporation into the film, with N 1s showing a signal corresponding to H reaction with the N-pyridine lone pair, which has not yet been reported, to the best of our knowledge. Our results suggest that a change in the binding energy corresponds with the lone pair interactions. ARHAXPES and TEM confirm the presence of a uniform thin film of the polymer. The IR spectroscopy indicates the condensation incorporation of the P2VP-OH into the substrate. The work in this letter should act as a reference for future work on this polymer in ASD and BCP infiltration studies.

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¹R. A. Segalman, *Mat. Sci. Eng. R.* **48**, 191 (2005).

²X. Jiang and S. F. Bent, *J. Electrochem. Soc.* **154**, D648 (2007).

³R. Chen, H. Kim, P. C. McIntyre, and S. F. Bent, *Appl. Phys. Lett.* **84**, 4017 (2004).

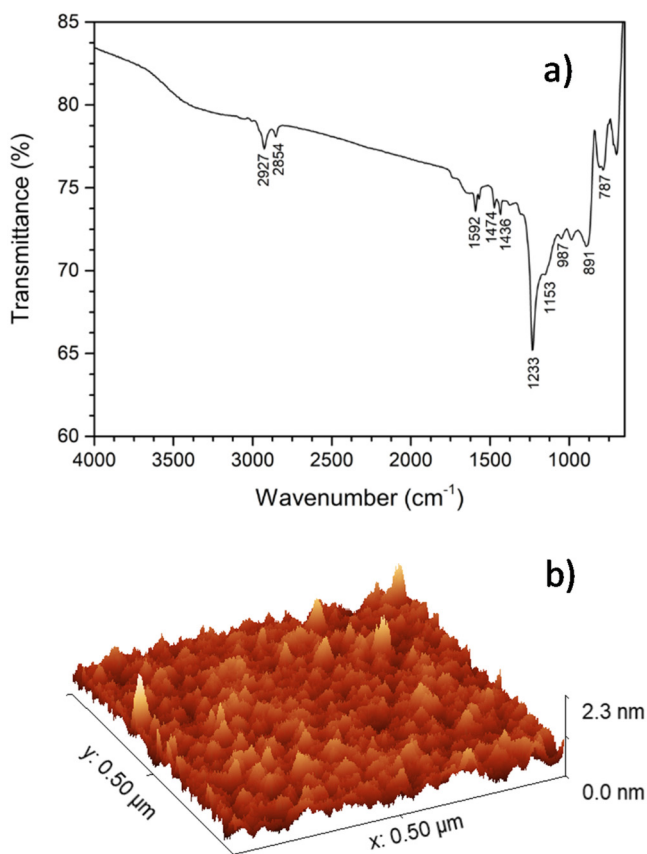


Fig. 3. (a) Grazing-angle ATR IR spectrum of a thin P2VP film. (b) AFM 3D height image of P2VP surface.

- ⁴R. Lundy, S. P. Flynn, C. Cummins, S. M. Kelleher, M. N. Collins, E. Dalton, S. Daniels, M. A. Morris, and R. Enright, *Phys. Chem. Chem. Phys.* **19**, 2805 (2017).
- ⁵Q. Peng, Y.-C. Tseng, S. B. Darling, and J. W. Elam, *Adv. Mater.* **22**, 5129 (2010).
- ⁶D. Borah et al., *J. Phys. D Appl. Phys.* **44**, 174012 (2011).
- ⁷J. Varghese, T. Ghoshal, N. Deepak, C. O'Regan, R. W. Whatmore, M. A. Morris, and J. D. Holmes, *Chem. Mater.* **25**, 1458 (2013).
- ⁸S. Malynych, I. Luzinov, and G. Chumanov, *J. Phys. Chem. B* **106**, 1280 (2002).
- ⁹C. Corten, K. Kretschmer, and D. Kuckling, *Beilstein J. Org. Chem.* **6**, 756 (2010).
- ¹⁰U. Borchert, U. Lipprandt, M. Bilanz, A. Kimpfler, A. Rank, R. Peschka-Suss, R. Schubert, P. Lindner, and S. Forster, *Langmuir* **22**, 5843 (2006).
- ¹¹D. Cho, A. Noro, A. Takano, and Y. Matsushita, *Macromolecules* **38**, 3033 (2005).
- ¹²F. S. Bates and G. H. Fredrickson, *Phys. Today* **52**(2), 32 (1999).
- ¹³S. G. Jang, A. Khan, C. J. Hawker, and E. J. Kramer, *Macromolecules* **45**, 1553 (2012).
- ¹⁴P. Mokarian-Tabari, R. Sentharamaikannan, C. Glynn, T. W. Collins, C. Cummins, D. Nugent, C. O'Dwyer, and M. A. Morris, *Nano Lett.* **17**, 2973 (2017).
- ¹⁵C. Cummins and M. A. Morris, *Microelectron. Eng.* **195**, 74 (2018).
- ¹⁶N.-G. Kang, B.-G. Kang, H.-D. Koh, M. Changez, and J.-S. Lee, *React. Funct. Polym.* **69**, 470 (2009).
- ¹⁷S. Pal, "Pyridine: A useful ligand in transition metal complexes," in *Pyridine*, edited by P. P. Pandey (InTechOpen, London, 2018), pp. 57–74.
- ¹⁸U. Rafique, M. Mazhar, S. Ali, and F. A. Khwaja, *Synth. Met.* **78**, 73 (1996).
- ¹⁹Z. Zhang, T. Dwyer, S. M. Sirard, and J. G. Ekerdt, *J. Vac. Sci. Technol. A* **37**, 020905 (2019).
- ²⁰D. Berman, S. Guha, B. Lee, J. W. Elam, S. B. Darling, and E. V. Shevchenko, *ACS Nano*. **11**, 2521 (2017).
- ²¹J. W. Elam et al., *ECS Trans.* **69**, 147 (2015).
- ²²Y.-C. Tseng, Q. Peng, L. E. Ocola, D. A. Czaplewski, J. W. Elam, and S. B. Darling, *J. Mater. Chem.* **21**, 11722 (2011).
- ²³J.-P. Rueff et al., *J. Synchrotron Radiat.* **22**, 175 (2015).
- ²⁴A. Herrera-Gomez, F. S. Aguirre-Tostado, P. G. Mani-Gonzalez, M. Vazquez-Lepe, A. Sanchez-Martinez, O. Ceballos-Sanchez, R. M. Wallace, G. Conti, and Y. Uritsky, *J. Electron Spectrosc. Relat. Phenomena* **184**, 487 (2011).
- ²⁵T. Sierański, *J. Mol. Model.* **23** (2017).
- ²⁶M. Snelgrove et al., "Hard x-ray photoelectron spectroscopy study of copper formation by metal salt inclusion in a polymer film," *J. Phys. D. Appl. Phys.* (unpublished).
- ²⁷P. Papadopoulos, D. Peristeraki, G. Floudas, G. Koutalas, and N. Hadjichristidis, *Macromolecules* **37**, 8116 (2004).
- ²⁸R. Lundy et al., "Selective area deposition infiltration strategies and mechanisms in block copolymer and polymer brush films," ACS Omega (unpublished).
- ²⁹J. J. Lebrun and H. Porte, "Polysiloxanes," in *Comprehensive Polymer Science and Supplements*, edited by G. Allen and J. C. Bevington (Pergamon, New York, 1989), pp. 593–609.
- ³⁰C. J. Brinker, *J. Non Cryst. Solids* **100**, 31 (1988).
- ³¹A. Herrera-Gomez, M. Bravo-Sanchez, O. Ceballos-Sanchez, and M. O. Vazquez-Lepe, *Surf. Interface Anal.* **46**, 897 (2014).
- ³²J. Muñoz-Flores and A. Herrera-Gomez, *J. Electron Spectrosc. Relat. Phenomena* **184**, 533 (2012).
- ³³A. Herrera-Gomez, Internal Report (2008), available at <http://www.gro.cinvestav.mx/~aherrera/reportesInternos/arxpsAnalysisSharpInterfaces.pdf>.
- ³⁴O. Ceballos-Sanchez, A. Sanchez-Martinez, M. O. Vazquez-Lepe, T. Duong, R. Arroyave, F. Espinosa-Magana, and A. Herrera-Gomez, *J. Appl. Phys.* **112**, 053527 (2012).
- ³⁵A. Herrera-Gomez, Y. Sun, F.-S. Aguirre-Tostado, C. Hwang, P.-G. Mani-Gonzalez, E. Flint, F. Espinosa-Magaña, and R. M. Wallace, *Anal. Sci.* **26**, 267 (2010).
- ³⁶T. Susi, T. Pichler, and P. Ayala, *Beilstein J. Nanotech.* **6**, 177 (2015).
- ³⁷Y. Liu et al., *Adv. Funct. Mater.* **21**, 986 (2011).
- ³⁸A. Mohtasebi, T. Chowdhury, L. H. H. Hsu, M. C. Biesinger, and P. Kruse, *J. Phys. Chem. C* **120**, 29248 (2016).
- ³⁹M. T. Greiner, M. Festin, and P. Kruse, *J. Phys. Chem. C* **112**, 18991 (2008).
- ⁴⁰G. Beamson and D. Briggs, *J. Electron Spectrosc.* **62**, 220 (1993).
- ⁴¹G. Gabka, P. Bujak, K. Giedyk, K. Kotwica, A. Ostrowski, K. Malinowska, W. Lisowski, J. W. Sobczak, and A. Pron, *Phys. Chem. Chem. Phys.* **16**, 23082 (2014).
- ⁴²M. Barber, J. A. Connor, M. F. Guest, I. H. Hillier, and M. Schwarz, *J. Chem. Soc. Faraday Trans. 2* **69**, 551 (1973).
- ⁴³G. Calderón-Ayala et al., *Carbon Lett.* **21**, 93 (2017).
- ⁴⁴H.-W. Tien, Y.-L. Huang, S.-Y. Yang, J.-Y. Wang, and C.-C. M. Ma, *Carbon* **49**, 1550 (2011).
- ⁴⁵A. Herrera-Gomez, F. S. Aguirre-Tostado, Y. Sun, R. Contreras-Guerrero, R. M. Wallace, Y. Hisao, and E. Flint, *Surf. Interface Anal.* **39**, 904 (2007).
- ⁴⁶S.-W. Kuo, C.-H. Wu, and F.-C. Chang, *Macromolecules* **37**, 192 (2004).
- ⁴⁷S. U. D. Khan, B. Ahmed, S. K. Raghuvanshi, and M. A. Wahab, *Indian J. Pure Appl. Phys.* **52**, 192 (2014), available at <http://op.niscair.res.in/index.php/IJPAP/article/view/1767/285>.
- ⁴⁸L. C. Cesteros, J. R. Isasi, and I. Katime, *Macromolecules* **26**, 7256 (1993).
- ⁴⁹A. Grill, *Annu. Rev. Mater. Res.* **39**, 49 (2009).