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Capturing a Square Planar Gold(III) Complex Inside a Platinum Nanocage: A Combined Experimental and Theoretical Study

Emmanuel Puig,^a Christophe Desmarets,^{*a} Geoffrey Gontard^a, Marie Noelle Rager,^b Andrew L. Cooksy,^c and Hani Amouri,^{*a}

[a] Sorbonne Université-Faculté des Sciences et Ingénierie Campus Pierre et Marie Curie, CNRS, IPCM (UMR 8232), 4 place Jussieu, 75252 Paris cedex 05, France.

[b] Chimie ParisTech, PSL University, NMR Facility, F-75005 Paris, France.

[c] Department of Chemistry, San Diego State U., San Diego CA 921821030, USA

ABSTRACT: A novel synthetic procedure was set up to gain access to platinum coordination cages Pt₂L₄, which are less investigated compared to their palladium counterparts. This Pt₂L₄ nanocage exhibits an adequate cavity for guest encapsulation. Indeed the Au(III) metal complex [Au(bdt)₂]⁺ (bdt = benzene-1,2-dithiolate) was successfully captured inside the cavity, in contrast to the analogous palladium cage which failed to host the gold complex. This result represents a rare example where a metal complex with thioligands can be encapsulated in a coordination cage. Moreover it highlights the role of the metal center and the robustness of the platinum cage for host-guest chemistry. This discovery will inspire researchers in this area to pay more attention to Pt-cages. The host-guest system was fully characterized by NMR techniques and X-ray crystallographic analysis. Moreover the nature of the host-guest interaction in this unique example was investigated and rationalized by DFT computational studies.

■ INTRODUCTION

Host-guest chemistry of self-assembled coordination cages has attracted much attention in the last two decades.¹⁻⁶ These cages can now be obtained with different shapes and display cavities with controllable sizes.⁷⁻¹⁵ As a consequence, a wide range of applications have been investigated including molecular sensing, chiral separation, drug delivery and catalytic properties.¹⁶⁻²⁰ Among these cages, particular interest has been devoted to the formation of the M_2L_4 cage-type.²¹⁻²⁶ These architectures display high symmetry and are capable of encapsulating guests such as anions or organic molecules but to a lesser extent metal compounds.²⁷⁻³⁰ They are usually obtained by mixing metal precursors displaying four available coordination sites and bidentate ligands. While Pd_2L_4 cages have been intensively described, surprisingly so far only a few platinum analogues have been reported.^{31, 32}

Our group have developed a variety of metallacages of type $M_2(L^1)_4$ based on Co(II) and Cu(II) using a semi-rigid bidentate ligand where $L^1 = 1,3$ -bis(benzimidazol-1-ylmethyl)-2,5-dimethoxy-4,6-dimethylbenzene.³³⁻³⁶ We also demonstrated that these systems are capable of selectively encapsulating weakly coordinated anions such as BF_4 and PF_6 anions.³⁵ However in the presence of strongly coordinated anions the related coordination polymers were obtained instead, highlighting the lability of metal-ligand (L^1) interaction in these metallacages.³⁴

Subsequently, we shifted our efforts towards the design of another kind of M_2L_4 metallacage but using the more rigid pyridine-based ligands.³⁷ The latter are more strongly bound to the metal centers. Indeed we demonstrated that the metallacage Pd_2L_4 successfully accommodates anionic organometallic species $[Pt(NO_2)_4]^{2-}$ inside the cavity.²⁹ 1H -NMR studies and DFT calculations allowed us to rationalize the nature of non-covalent interactions occurring between the guest complex and the Pd_2L_4 cage. It is worth mentioning that encapsulating kinetically labile metal complexes inside self-assembled cages is a difficult task and only few examples have been reported.^{18, 27, 38}

■ RESULTS AND DISCUSSION

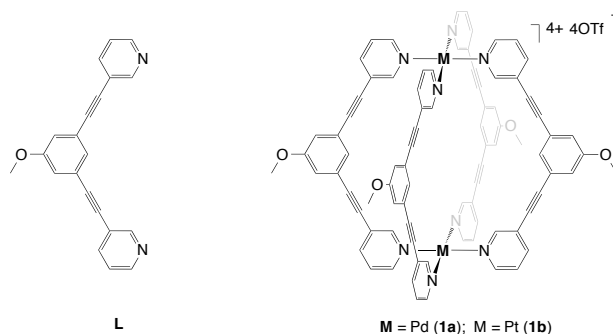
Pursuing our investigations in this area we sought again to accommodate metal complexes, but now displaying thio-ligands such as $[n-Bu_4N][Au(bdt)_2]$ (bdt = benzene-1,2-dithiolate). This type of molecule is highly challenging because it contains strongly coordinating sulfur centres and displays interesting magnetic and redox properties.^{39, 40} In this paper we report the synthesis of nanocages of the type M_2L_4 ($M = Pd, Pt$) (Figure

1) and our investigations towards the encapsulation of a square planar Au(III) guest molecule.

Our first approach was to examine whether our Pd_2L_4 cage available in hand, ($L = 1,3$ -bis(pyridin-3-ylethynyl)-5-methoxybenzene) is adequate to host the gold(III) complex since we demonstrated that such a metallacage was able to accommodate $[PtCl_4]^{2-}$ species inside the cavity.³⁷ We reasoned that recognition could be enhanced through electrostatic interactions between the positively charged Pd(II) cations and the anionic nature of the guest but also by hydrogen bonding provided by the *endohedral* hydrogens of ligand L pointing towards the interior of the cavity. The host-guest studies were carried out in $CD_3CN/DMSO-d_6$ (4:1) mixture due to solubility limitations. Thus, when one equivalent of $[n-Bu_4N][Au(bdt)_2]$ was added to a solution of cage $[Pd_2(L)_4][OTf]_4$ (**1a**), the 1H NMR spectrum showed the formation of multiple peaks relative to the starting material **1a** and the appearance of the free ligand L and unidentified materials (Figure S1). Further addition of another equivalent of $[Au(bdt)_2]^-$ to the reaction mixture provided L in equal amount to the metallacage **1a**, while addition of a third equivalent led to the decomposition of the metallacage and formation of the free ligand L almost quantitatively along with free $[Au(bdt)_2]^-$ as well. We reasoned that the analogous Pt_2L_4 cage should display stronger Pt-ligand interaction and hence should be more appropriate to study this type of host-guest interaction with the challenging Au(III) guest $[n-Bu_4N][Au(bdt)_2]$.

Our first step was to find a rational synthetic approach to this kind of platinum nanocage Pt_2L_4 . Gratifyingly after several attempts we found that treatment of ligand L with freshly prepared $[Pt(EtCN)_4][OTf]_2$ precursor in acetonitrile solution at $80^\circ C$ for several days afforded the desired nanocage complex $[Pt_2L_4][OTf]_4$ (**1b**) in 80% yield (figure 1).

Figure 1. Schematic representation of the ligand L and the metallacages $[M_2L_4][OTf]_4$. $M = Pd$ (**1a**); $M = Pt$ (**1b**).



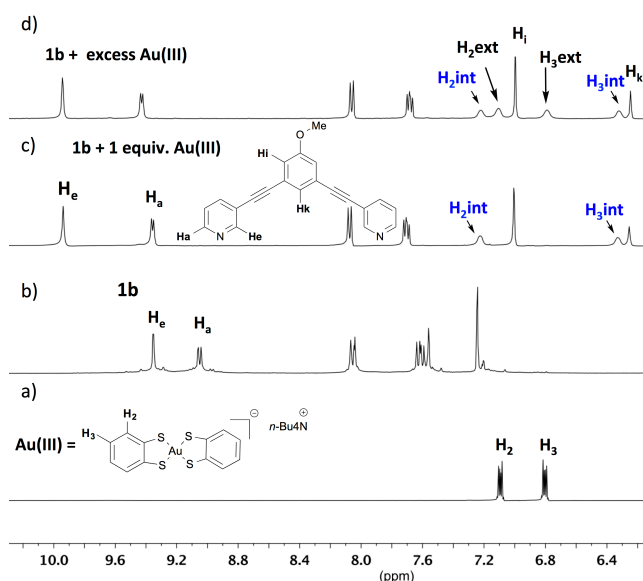
The infrared spectrum of **1b** showed the presence of triflate anions bands at 1256 cm^{-1} and displayed the alkyne stretching band at 2219 cm^{-1} . Moreover the integrity of the nanocage $[\text{Pt}_2\text{L}_4][\text{OTf}]_4$ (**1b**) was demonstrated by electrospray spectrometry, in which the $[\text{Pt}_2\text{L}_4]^{4+}$ fragment with varying numbers of triflate counterions was observed. The ^1H NMR spectrum of **1b** was recorded in CD_3CN due to solubility limitations. The spectrum shows that the coordinated ligands in **1b** display a symmetric pattern close to the free ligand L. A downfield shift relative to free ligand was observed, particularly for the H_a and H_c protons (Figure S2). Full characterization of **1b** is given in the supporting information.

We then investigated the host-guest properties of the metallacage **1b** towards the metal complex guest $[n\text{-Bu}_4\text{N}][\text{Au}(\text{bdt})_2]$ under similar conditions to those used for the palladium cage $[\text{Pd}_2\text{L}_4][\text{OTf}]_4$ (**1a**). It is worth mentioning that examples where a square metal complex is encapsulated within a metallacage, especially in the case of a platinum capsule, are limited^{27, 28} and more scarce when the complex displays thioligands because they are also able to bind the metal centre.⁴¹ When the host molecule $[\text{Pt}_2\text{L}_4][\text{OTf}]_4$ (**1b**) was treated by one equivalent of the guest complex $[n\text{-Bu}_4\text{N}][\text{Au}(\text{bdt})_2]$ in a mixture of $\text{CD}_3\text{CN}/\text{DMSO-}d_6$ (4:1), pronounced chemical shift changes were observed for the anionic complex and the cage as well, suggesting the encapsulation of the guest. For instance the signals attributed to hydrogens H_c and H_a for **1b** moved from δ 9.34 ppm and δ 9.06 ppm to δ 9.92 ppm and δ 9.39 ppm while the NMR resonances of the aromatic protons H_2 and H_3 for the guest changed from δ 7.09 ppm and δ 6.78 ppm to δ 7.21 ppm and δ 6.31 ppm (Figure 2c, Figure S3). Addition of another equivalent of guest molecule led to two novel signals at δ 7.09 ppm and δ 6.78 ppm, which correspond to the signals of free Au(III) guest (Figure 2d). No changes in the ^1H NMR spectrum occurred in the presence of an excess of guest molecules. Moreover upon leaving the solution to stand over time no evolution occurred, suggesting that this host-guest system is kinetically and thermodynamically stable. These results contradict completely those obtained for the analogous and more labile Pd_2L_4 metallacage **1a**, highlighting the robust nature of our platinum cage **1b**.

We then attempted to isolate the host-guest assembly. Thus, treatment of **1b** with excess of $[n\text{-Bu}_4\text{N}][\text{Au}(\text{bdt})_2]$ in CH_3CN , afforded a green precipitate which was isolated and characterized by NMR spectroscopy in DMSO (^1H -, ^{13}C -, 1D and 2D-NMR experiments) as $[\text{Au}(\text{bdt})_2\text{CPT}_2\text{L}_4][\text{Au}(\text{bdt})_2]_3$ (**2**). Moreover, the existence of assembly **2** in solution was confirmed by

electrospray mass spectrometry. The +3 and +2 charge state of the $\{[\text{Au}(\text{bdt})_2\text{CPT}_2\text{L}_4]\}$ and $\{[\text{Au}(\text{bdt})_2\text{CPT}_2\text{L}_4][\text{Au}(\text{bdt})_2]\}$ fragments appear respectively at $m/z = 701.76338$ and at $m/z = 1291.10670$ and were verified by comparison of the observed and theoretical isotopic patterns (Figure

Figure 2. ^1H NMR host-guest studies with metallacage **1b** in the presence of $[n\text{-Bu}_4\text{N}][\text{Au}(\text{bdt})_2]$ in $\text{CD}_3\text{CN}/\text{DMSO-}d_6$ (4:1) a) free $[n\text{-Bu}_4\text{N}][\text{Au}(\text{bdt})_2]$ guest b) the cage complex **1b** c) 1 eq. of $[n\text{-Bu}_4\text{N}][\text{Au}(\text{bdt})_2]$ d) excess of $[n\text{-Bu}_4\text{N}][\text{Au}(\text{bdt})_2]$. (Hint = H interior Hext = H exterior).



The ^1H -NMR spectrum displayed two series of resonances for the Au(III) complex : two doublets of doublets at δ 7.08 ppm and δ 6.80 ppm corresponding to H_2 and H_3 protons of free $\text{Au}(\text{bdt})_2$ molecules and two other doublets of doublets at δ 7.13 ppm and δ 6.34 ppm corresponding to H_2 and H_3 protons of encapsulated Au(III) complex, respectively in 3/1 ratio. 2D NOESY experiment showed correlation between H_c proton of Pt_2L_4 host and H_2 proton from $[\text{Au}(\text{bdt})_2]$ guest, proving the encapsulation of the latter. In addition H_k protons which point to the interior of the cage moved up-field (from δ 7.56 ppm to δ 6.22 ppm) and are perpendicular to the Au(III) guest since no NOE correlation was observed between H_k and H_2 or H_3 . Finally, NOESY spectra showed exchange between the encapsulated Au(III) complex and the free $[\text{Au}(\text{bdt})_2]$ -molecule is occurring (Figure S5). Interestingly the ^{195}Pt -H couplings are initially visible for complex **1b** but disappear just after guest addition confirming this dynamic exchange. As shown in Figure S4 in the supporting information, both free cage and complexed-cage protons were seen in the spectrum, during sequential

guest additions, indicating a slow exchange of the host-guest complex.⁴²

$\text{Bu}_4\text{N}[\text{Au}(\text{bdt})_2]$ provided crystals adequate for X-ray structural determination. The solid state structure was identified stoichiometrically as $[\text{Au}(\text{bdt})_2\text{C}(\text{Pt}_2\text{L}_4)][(\text{Au}(\text{bdt})_2)_{2.5}(\text{OTf})_{0.5}]$ (**3**) i.e. a 1:1 ratio of $[\text{Au}(\text{bdt})_2\text{C}(\text{Pt}_2\text{L}_4)[\text{Au}(\text{bdt})_2]_3$ and $[(\text{Au}(\text{bdt})_2\text{C}(\text{Pt}_2\text{L}_4))_2(\text{Au}(\text{bdt})_2)(\text{OTf})]$. In all cages the X-ray structure showed the presence of one encapsulated molecule of the anionic $[\text{Au}(\text{bdt})_2]^-$ complex inside the cavity of Pt_2L_4 cage and the presence of free $[\text{Au}(\text{bdt})_2]^-$ in accord with the $^1\text{H-NMR}$ solution studies (Figure 4). The Pt—N bond lengths lie in the range of 2.01-2.03 Å. The Pt---Pt distance is 11.880(2) Å and the average distance between two facing phenyl rings is around 10.7 Å so that the cavity size is comparable to that observed for cage **1a**.³⁷ The anionic $[\text{Au}(\text{bdt})_2]^-$ guest is located inside the cavity and adopts a perpendicular orientation relative to planes containing the square platinum centers. The three metal centers Pt---Au---Pt are not perfectly aligned, and the Pt---Au distances are 5.962(1) and 5.999(1) Å, precluding any direct metal-metal interaction. However, the guest interacts with the frame of the cage via hydrogen bonding interactions between each sulfur atom of the $[\text{Au}(\text{bdt})_2]^-$ and two *ortho* aromatic protons from two different coordinated pyridines. Indeed, the C---S and H---S distances lie respectively in the range of 3.65-3.89 Å and 2.76-2.99 Å, with C-H---S bond angles of ca. 139 to 171°. These values are comparable to those reported for complexes displaying C-H---S bonding interactions.⁴³ On the other hand the structure reveals the presence of other anionic molecules $[\text{Au}(\text{bdt})_2]^-$ located outside the assembly, but stacking perfectly over the platinum centres of the cage (Figure 3). The Au---Pt distances are 4.626 (1) and 4.698 (1) Å, larger than the sum of van der Waals radii (3.73 Å) precluding any metal—metal interaction.⁴⁴

Figure 3. Experimental and theoretical isotope distribution of a) $\{[\text{Au}(\text{bdt})_2\text{C}(\text{Pt}_2\text{L}_4)]\}$ and b) $\{[\text{Au}(\text{bdt})_2\text{C}(\text{Pt}_2\text{L}_4)[\text{Au}(\text{bdt})_2]\}$

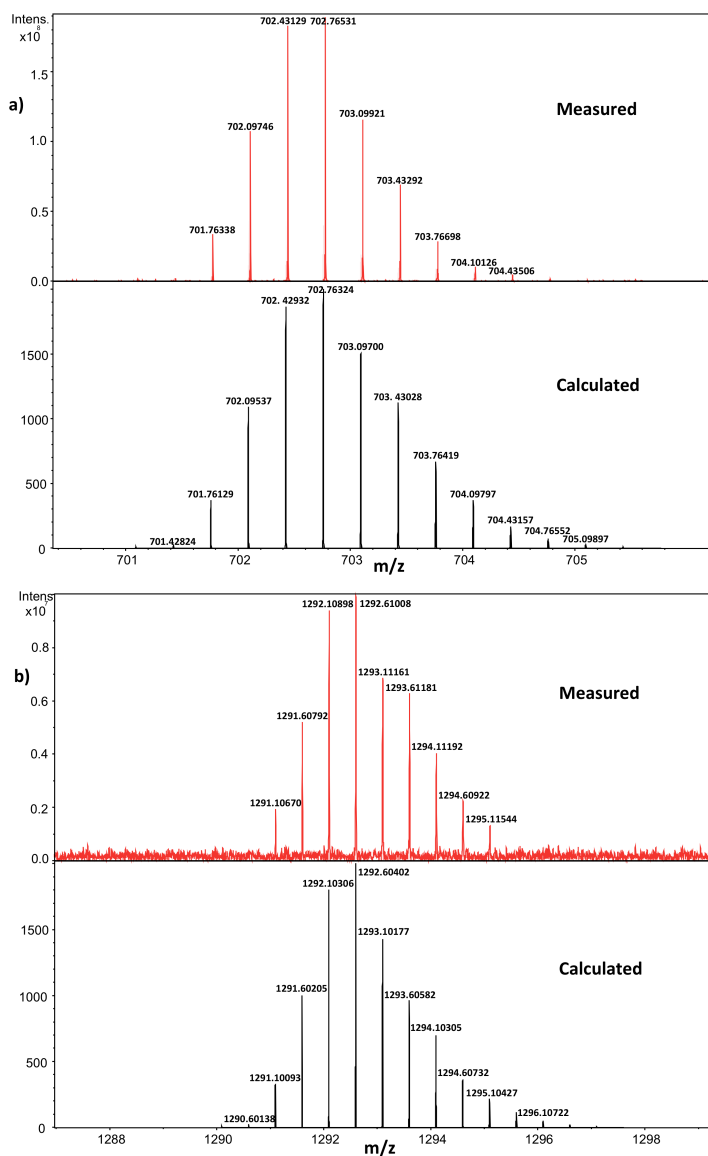
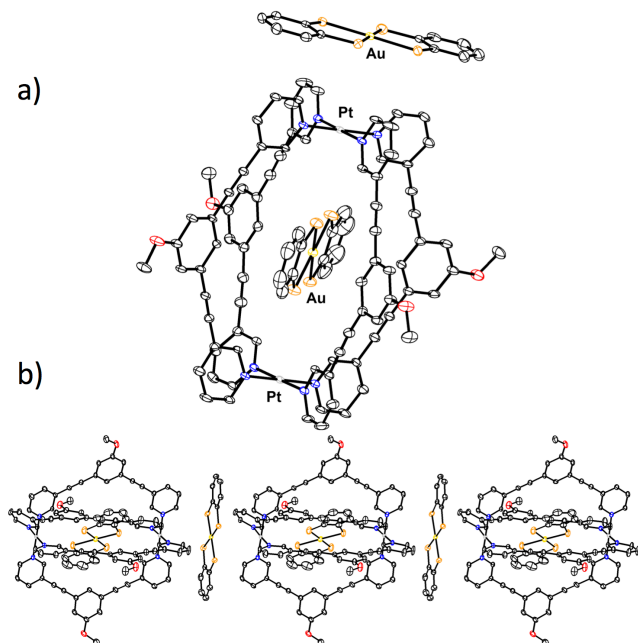


Figure 4. a) X-ray crystal structure of $[\text{Au}(\text{bdt})_2]\text{C}(\text{Pt}_2\text{L}_4)[(\text{Au}(\text{bdt})_2)]^{2+}$. For clarity purposes only the encapsulated Au(III) guest and one $\text{Au}(\text{bdt})_2$ outside the cavity perched over the Pt(II) centre are shown. b) 1D chain made of host-guest assemblies interacting with the outside $\text{Au}(\text{bdt})_2^{-1}$ species. All hydrogen atoms are omitted, thermal displacement ellipsoids are shown at 30% probability.



To shed light on the nature of the host-guest interactions in the supramolecular assembly, properties of the host-guest complex were calculated by density functional theory (DFT), using both the B3LYP-D3⁴⁵⁻⁴⁷ and M06⁴⁸ methods, and several basis sets for comparison.⁴⁹⁻⁵³ Additional details are provided in the Supplemental Information.

The gas-phase host-guest complexation energies were calculated to be quite similar among the different basis sets and DFT methods, suggesting that these methods are adequate for estimating the energetics. The interior gold guest is predicted to be more strongly bound than the exterior guest, by 12–18 kcal/mol in the absence of dispersion and by 42–47 kcal/mol when dispersion is included. This makes sense, given that the interior guest has more near-atom interactions available and also the possibility of π -stacking interactions between the central methoxyphenyl ring of each arm and the phenyl rings of the guest. Furthermore, the overall binding energy of the two guests with the host is greater than the sum of the individual binding energies, indicating that each guest activates the host towards accepting a second guest. The donor-acceptor interaction energies, including charge-transfer contributions, were estimated by NBO analysis. For both the interior and exterior guests, the dominant contribution arises from the guest acting as a donor to the organic component of the host cage. The Pt atoms are too distant from the guest molecules for any metal-metal host-guest interaction to account for more than about a quarter of the total binding energy. Looking specifically for Pt-Au interactions, we find that in the case of the exterior guest, most of the energy contribution involving the neighbouring Pt atom is ascribed to direct Pt-Au interaction, whereas for the

interior guest the Pt atoms interact roughly equally with the Au atom and with the associated S atoms. In either case, however, Pt-Au bonding appears to be a small contribution to the stabilization of these guests. Instead, most of the donor-acceptor stabilization is provided by the nearest-neighbour interactions, especially the electropositive hydrogen atoms on the host aryl groups as the acceptors and the guest sulfur atoms and aryl carbons as donors. This is consistent in part with the NBO orbital analysis, which is further discussed in the SI. Unfortunately, a full-scale energy decomposition analysis is prohibitive for a system of this size. However, the present DFT calculations strongly indicate that both guest molecules are stabilized by a combination of dispersion, donor-acceptor interaction, and electrostatics. The interior guest appears to be stabilized largely by the dispersion interaction, whereas electrostatic interactions may play a larger role in stabilizing the exterior guest. Overall, the interior guest is predicted to be more strongly bound to the cage by roughly 40–50 kcal/mol.

■ CONCLUDING REMARKS

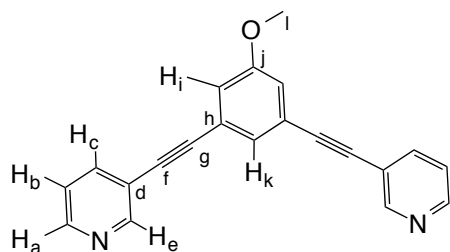
In this work we have investigated the encapsulation of square planar gold complex $[n\text{-Bu}_4\text{N}][\text{Au}(\text{bdt})_2]$ containing thioligands using metallacages of type M_2L_4 $\text{M} = \text{Pd}$ (**1a**), $\text{M} = \text{Pt}$ (**1b**). The latter was prepared using a novel synthetic procedure. While the Pd_2L_4 cage decomposed in the presence of the guest molecule $[\text{Au}(\text{bdt})_2]^-$, in stark contrast the platinum cage successfully accommodated an anionic square planar gold(III) complex as demonstrated by NMR techniques (2D NOESY). This result constitutes a rare example in the literature. Furthermore, single crystal X-ray diffraction studies confirm the presence of the gold guest inside the cavity stabilised through hydrogen-bonding with the ligands of the cage Pt_2L_4 . The X-ray structure also revealed the presence of a weak interaction between the platinum centre of the assembly and another anionic gold complex located outside the cavity. DFT calculations have been advanced to rationalize the nature of host-guest interactions in this rare and unique assembly

■ EXPERIMENTAL SECTION

General synthetic methods

All solvents used were reagent grade or better. Commercially available reagents were used as received. The ligand **L**, the solvated platinum building block $[\text{Pt}(\text{EtCN})_4][\text{OTf}]_2$, and the gold-bis(dithiolene) complex were prepared according to published methods.^{40, 54, 55} All experimental manipulations were carried out under argon using Schlenk techniques. IR spectra were recorded on a Bruker

Tensor 27 (Bruker Corp., Rheinstetten, Germany) equipped with a Harrick ATR. Elemental analyses were performed at the Microanalytical Laboratory of the Sorbonne-universit . Positive ESI mass spectra were obtained using a triple quadrupole mass spectrometer (Quattro II Micromass, Waters, UK). Automatic data acquisition was processed using the software Masslynx V4.0. NMR experiments were carried out on a Bruker Avance II 300 MHz, or Bruker Neo 500 MHz spectrometers operating at 300 K with chemical shifts referenced to residual solvent peaks. Chemical shifts are reported as parts per million (ppm) and coupling constants (J) in hertz (Hz). Standard abbreviations indicating multiplicity were used as follows: m = multiplet; t = triplet; d = doublet; s = singlet and b = broad.



Synthesis of [PtL₄][OTf]₄ (**1b**)

To an oven dried Schlenk flask equipped with a magnetic stirrer was added [Pt(EtCN)₄][OTf]₄ (126 mg, 0.176 mmol, 1 equiv.), 5-Methoxy-1,3-bis(pyridin-3-ylethynyl)benzene (L) (140 mg, 0.451 mmol, 2 equiv.) which were solubilized in argon-degassed MeCN (30 ml) freshly distilled over CaH₂. The solution was heated to reflux and maintained under argon with stirring for four days at 80°C. The solution was then allowed to cool at room temperature and concentrated under vacuum to 3 ml. The mixture was precipitated by slow addition of freshly distilled and degassed diethylether (20 ml). The white solid was filtered using a filtrating cannula, washed three times with diethylether (10 ml) and recrystallized from a mixture of CH₂Cl₂, DMSO/Et₂O. Dried under vacuum to afford the desired product as air-stable white solid. 170 mg (86 %).

¹H NMR (300 MHz, CD₃CN, 298K): δ 9.34 (s, 8H, H_c); 9.06 (dd, 8H, J=6.0, 0.5 Hz, H_a); 8.05 (dt, 8H, J=9.0, 0.5 Hz, H_c); 7.61 (ddd, 8H, J=9.0, 6.0, 0.5 Hz, H_b); 7.56 (t, 4H, J=1.3 Hz, H_k); 7.24 (d, 8H, J=1.3 Hz, H_i), 3.82 (s, 12H, OMe). IR (KBr): (ν, cm⁻¹) ν(C≡C) = 2219, ν(C-F) = 1256, ν(S=O) = 1231, ν(S-O) = 1029. ES-MS (m/z): [Pt₂L₄]⁴⁺: 407.59; found: 407.84, [Pt₂L₄(CF₃SO₃)₃]³⁺: 593.10; found: 593.84, [Pt₂L₄(CF₃SO₃)₂]²⁺: 964.13; found: 964.63; Anal. calcd for [Pt₂L₄][OTf]₄.DMSO.CH₂Cl₂: C 45.72, H 2.70, N 4.69; Found: C 45.90, H 1.93, N 4.71.

Synthesis of [(Au(bdt)₂CPt₂L₄)[Au(bdt)₂]₃ (**2**)

[*n*-Bu₄N][Au(bdt)₂] (40 mg, 0.056 mmol) was added to a solution of [PtL₄][OTf]₄ (31 mg, 0.014 mmol) in CH₃CN (5 mL). The solution was stirred at room temperature for 30 min, during which time a green precipitate formed. The solid was collected by filtration, washed with diethyl ether (3 × 5 mL), and dried under vacuum: 42 mg of green solid (87%); This metallacage was isolated and characterized as [Au(bdt)₂CPt₂(L₄)[Au(bdt)₂]₃. ¹H NMR (500 MHz, DMSO-d₆): δ 9.88 (s, 8H, H_c), 9.45 (d, J = 6.0 Hz, 8H, H_a), 8.23 (dt, J = 7.9 Hz, J = 1.4 Hz, 8H, H_c), 7.87 (dd, J = 7.9 Hz, J = 6.0 Hz, 8H, H_b), 7.13 (dd, J = 5.9 Hz, J = 3.1 Hz, 4H, H_{2int}), 7.08 (dd, J = 5.9 Hz, J = 3.2 Hz, 12H, H_{2ext}), 7.02 (d, J = 1.2 Hz, 8H, H_i), 6.80 (dd, J = 5.9 Hz, J = 3.2 Hz, 12H, H_{3ext}), 6.34 (dd, J = 5.9 Hz, J = 3.1 Hz, 4H, H_{3int}), 6.22 (t, J = 1.2 Hz, 4H, H_k), 3.72 (s, 12H, OMe). ¹³C-NMR (CDCl₃, 125 MHz): δ = 158.8 (C_j), 154.4 (C_c), 151.1 (C_a), 143.3 (C_c), 141.7 (C_{1ext}), 128.3 (C_{2int}), 128.1 (C_b), 128.0 (C_{2ext}), 126.9 (C_k), 123.8 (C_{3int}), 123.3 (C_{3ext} and C_d), 122.4 (C_h), 118.3 (C_i), 94.4 (C_g), 83.9 (C_f), 55.7 (C_i). ES-HRMS (m/z): [L₄Pt₂(Au(bdt)₂)₃]³⁺: 701.76129; found: 701.76338, [L₄Pt₂(Au(bdt)₂)₂]²⁺: 1291.10093; found: 1291.10670. All attempts to obtain crystals of this compound were unsuccessful.

X-ray crystal structure of [Au(bdt)₂CPt₂L₄][Au(bdt)₂]_{2.5}(OTf)_{0.5} (**3**)

Diffusion of diethyl ether into acetonitrile/DMSO (4:1) solution of **1b** in the presence of excess [*n*-Bu₄N][Au(bdt)₂] provided suitable crystals. A single crystal was selected, mounted and transferred into a cold nitrogen gas stream. Intensity data was collected with a Bruker Kappa-APEX2 system using fine-focus sealed tube Mo-Kα radiation. Unit-cell parameters determination, data collection strategy, integration and absorption correction were carried out with the Bruker APEX2 suite of programs. The structure was solved with SHELXT-2014⁵⁶ and refined anisotropically by full-matrix least-squares methods with SHELXL-2014⁵⁶ using the WinGX suite.⁵⁷ Significant electron density remains near gold and platinum atoms despite our repeated attempts at improving absorption correction. The crystal contains severely disordered and partially missing molecules of volatile solvent (diethyl ether, dimethyl sulfoxide and/or acetonitrile). PLATON SQUEEZE⁵⁸ was used for the final refinement cycles to account for the resulting voids. The structure was deposited at the Cambridge Crystallographic Data Centre with number CCDC 1841449 and can be obtained free of charge via www.ccdc.cam.ac.uk.

Crystal data for **3**

Yellow plate, C_{134.5} H₁₀₄ Au_{3.5} F_{1.5} N₈ O_{7.5} Pt₂ S_{14.5}, triclinic P -1, a = 12.8265(9) Å, b = 21.1589(15) Å, c

= 27.9767(21) Å, $\alpha = 107.374(3)^\circ$, $\beta = 94.244(3)^\circ$, $\gamma = 98.877(2)^\circ$, $V = 7101.2(9) \text{ \AA}^3$, $Z = 4$, $T = 200(1) \text{ K}$, $\lambda = 0.71073 \text{ \AA}$, $\mu = 5.834 \text{ mm}^{-1}$, min / max transmission = 0.24 / 0.90, θ range = 0.77° to 30.55° , 209447 reflections measured, 43422 independent ($R_{\text{int}} = 0.0384$), 30862 observed [$I > 2\sigma(I)$], completeness = 0.997, 1661 parameters, 334 restraints, final R indices R_1 [$I > 2\sigma(I)$] = 0.0643 and wR_2 (all data) = 0.1750, GOF on $F^2 = 1.054$, largest difference peak / hole = 7.82 / -6.15 e.Å⁻³.

ASSOCIATED CONTENT

Supporting Information

The Supporting Information is available free of charge on the ACS Publications website. CCDC reference number 1841449 contains crystallographic data for complexes **3** presented in this paper. These data can be obtained free of charge from the Cambridge Crystallographic Data Centre via www.ccdc.cam.ac.uk/data_request/cif.

Additional figures and tables (PDF)

■ AUTHOR INFORMATION

Corresponding Author

*E-mail: hani.amouri@sorbonne-universite.fr. Phone: +33-1-44273083.

*E-mail: Christophe.desmarests@sorbonne-universite.fr.

Author Contributions

The manuscript was written through contributions of all authors. All authors have given approval to the final version of the manuscript.

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SYNOPSIS TOC. A novel platinum-based metallacage Pt_2L_4 has been designed and self-assembled. This nanocage displays a cavity for guest encapsulation. Indeed the Au(III) metal complex $[\text{Au}(\text{bdt})_2]^-$ (bdt = benzene-1,2-dithiolate) was successfully captured inside the cavity of the Pt_2L_4 and in contrast to that observed for the analogous Pd_2L_4 metallacage.

