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Ourania Makrygenni, Louise Vanmairis, Sabrina Taourit, Franck Launay, Alain Shum Cheong Sing, et al.. Selective Formation of Epoxy limonene Catalyzed by Phosphonyl/Arsonyl Derivatives of Trivalent Polyoxotungstates at Low Temperature. *European Journal of Inorganic Chemistry*, 2019, 2020 (7), pp.605-612. 10.1002/ejic.201901152 . hal-02426557

HAL Id: hal-02426557

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Selective Formation of Epoxylimonene catalyzed by Phosphonyl/Arsonyl Derivatives of trivalent polyoxotungstates at low temperature.

Ourania Makrygenni,^[a] Louise Vanmairis,^[b] Sabrina Taourit,^[b] Franck Launay,^[c] Alain Shum Cheong Sing,^{*[b]} Anna Proust,^[a] H  l  ne G  rard,^[d] and Richard Villanneau^{*[a]}

Abstract: The catalytic performances of three organophosphonyl/arsonyl derivatives of POMs were evaluated for the epoxidation of limonene in acetonitrile, using aqueous H₂O₂ as the oxidant. All three W-based POMs catalysts operated without any additional transition-metal ions and displayed excellent conversion for the limonene to temperatures varying from 4 to 50  C. Furthermore, the use of B, α -[NaHAsW₉O₃₃{P(O)R}₂]³⁻ (R = ^tBu, -CH₂CH₂CO₂H) complexes led to the complete conversion of limonene to epoxy limonene at 4  C. The selectivity of the reaction was modulated by varying the reaction solvent, and it was found that allylic reactions were favoured in ethanol. The effect of the catalyst protonation was also investigated by DFT calculations, highlighting the role of protons in the epoxidation process.

Key Topic: Epoxidation with Polyoxometalates

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Introduction

The catalytic epoxidation of alkenes by hydrogen- or alkyl-peroxides is often considered as a model reaction in coordination chemistry. However, it is inexact and, in some respects, unfair to restrict this specific reaction to an "academic" tool for comparison of the catalysts performance. Indeed several chemical multi-steps processes of interest are based on the formation of organic epoxides, the most relevant being the formation of adipic acid from cyclohexene and of cyclic carbonates from the corresponding alkenes and CO₂. The rationale for the use of polyoxometalates (POMs) in this context is justified by the research for efficient catalysts working in the mildest conditions. Indeed, various examples of epoxidation reactions catalyzed by heteropolytungstates without any additional transition-metal ions¹ (including hybrid derivatives of POMs)^{2,3} have been reported in the literature.

Protocols with hydrogen- or alkyl-peroxides involving efficient W(VI)-peroxides catalytic species are generally characterized by negligible oxidant decomposition pathways and good to excellent selectivities.⁴ Some of us recently reported the promising catalytic performance of an organophosphonyl derivative of the arsenotungstate B, α -[AsW₉O₃₃]⁹⁻, namely B, α -[NaHAsW₉O₃₃{P(O)CH₂CH₂CO₂H}₂]³⁻ (figure 1).⁵ This particular compound has proved to be an efficient catalyst when used with H₂O₂ at room temperature.

Among the different substrates, terpenes are a natural class of compounds that can be found in large amounts in essential oils, used in food and beverage flavorings, perfumery and chemical industry. For instance, citrus essential oils (containing terpenes and, to a lesser extent, their oxygenated derivatives: alcohols, aldehydes, esters, ketones and a few amount of epoxides) are obtained from the peel with an overall yield in the range of 0.1 to 1%, depending on the raw material. However, in the case of perfumery, terpenes such as limonene and β -pinene play a minor role in the aromas despite their presence in highest concentration. In this regard, processes for the sustainable production of higher value (oxygenated) terpenes derivatives from these citrus by-products are now explored and established.⁶ In this study, we focused on the catalytic transformation of the most abundant compounds, limonene preferentially, into the mono-epoxides. Indeed, the limonene oxides can be used as platform molecules for the synthesis of multifunctional bio-based polycarbonates polymers.⁷

In the present manuscript, a complete study of the catalytic performances of three different POM hybrids, namely B, α -[NaHAsW₉O₃₃{P(O)R}₂]³⁻ (R = ^tBu (1), -CH₂CH₂CO₂H (2)) and A, α -[NaHPW₉O₃₄{As(O)*p*-C₆H₄NH₂]₂]³⁻ (3) anions, toward the epoxidation reaction of alkenes (cyclooctene, cyclohexene and limonene) is presented. Furthermore, DFT calculations were performed to address the localization of the catalytically active site, highlighting the role of the proton in the epoxidation process. In addition, the behavior of different catalysts toward the terpenic part of an essential oil deriving from sweet orange from Portugal was verified.

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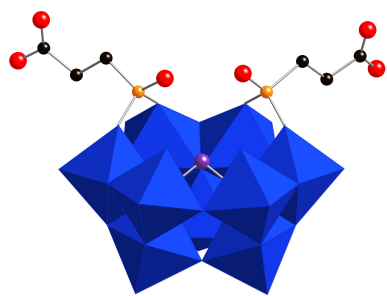


Figure 1. Structural representation of the B,α -[AsW₉O₃₃{P(O)CH₂CH₂CO₂H}₂]⁵⁻ anion (2). WO₆ polyhedra are shown in blue. The As, P, O and C atoms are shown respectively in prune, orange, red and grey.

Results and Discussion

1. Catalytic epoxidation of alkenes with (*n*-Bu₄N)₃[NaHASW₉O₃₃{P(O)R}₂]

Heteropolytungstates that do not contain any additional transition-metal ions,¹ including organophosphonyle derivatives of POMs,^{2,3} have been extensively tested as catalysts precursors in alkenes epoxidation reactions with H₂O₂ as oxidant.^{8,9,10} Among these examples, only the group of Bonchio used organophosphonyle derivatives of trivacant POMs. However, the experimental conditions were quite drastic since it required a relatively high reaction temperature with microwave assistance. As a result, the need for a less energy-consuming approach is worth being investigated.

1.1. Study of catalysts reactivity toward cyclooctene including previously reported results. In this context, we recently synthesized and used a family of heteropolytungstates-based hybrids for their covalent immobilization onto the surface of various silica-based supports.¹¹ The performances of these catalysts were first evaluated in the epoxidation of cyclooctene and cyclohexene with aqueous H₂O₂ as a model reaction, in homogeneous conditions or after their anchoring onto the support.⁵ In the course of this study, B,α -[NaHASW₉O₃₃{P(O)R}₂]³⁻ (R = -C(CH₃)₃ = ^tBu (1), -CH₂CH₂CO₂H (2)) and A,α -[NaHPW₉O₃₄{P(O)*p*-C₆H₄NH₂]₂]³⁻ (3) complexes demonstrated increased catalytic performances in homogeneous conditions. As one example, we have shown that the carboxylic acid-containing B,α -[NaHASW₉O₃₃{P(O)CH₂CH₂CO₂H}₂]³⁻ (2) derivative led to almost complete conversion of cyclooctene after 6 h at room temperature (73% for cyclohexene after 24 h, see figure 2). The calculated turnover frequency (TOF) was ca. 45 h⁻¹ when using a molar ratio of catalyst/alkene/H₂O₂ equal to 1/250/250. This TOF reached 300 h⁻¹ at 50°C with an excess of H₂O₂ (3 eq.) in regards to the substrate. This catalyst was also covalently grafted onto a {NH₂}-functionalized mesoporous SBA-15 silica support. In these heterogeneous conditions, interesting conversions for both alkenes were maintained and, last but not least, a better epoxide selectivity for the cyclohexene, compared to the homogeneous conditions, was observed.⁵

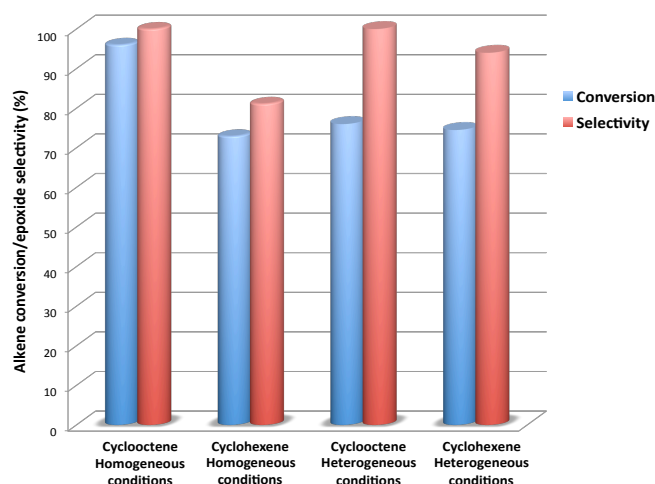


Figure 2. Conversion (in blue) and epoxide selectivity (in red) in the epoxidation of cyclooctene and cyclohexene with aqueous H₂O₂ using (*n*-Bu₄N)₃NaH[AsW₉O₃₃{P(O)CH₂CH₂CO₂H}₂]³⁻ (2) in homogeneous and heterogeneous conditions, in acetonitrile, at room temperature. The molar ratios catalyst/alkene/H₂O₂ were equal to 1/250/250 after 24h.

For purpose of comparison, the catalytic performances of the corresponding *tert*-butyl derivative, first identified as (*n*-Bu₄N)₃[NaHASW₉O₃₃{P(O)^tBu}₂], were also evaluated in the homogeneous epoxidation of cyclooctene (Table 1).⁵ In this first study, the behavior of both complexes (with R = -CH₂CH₂CO₂H and ^tBu) was different since for the latter the conversion was found very low (7.5%) after one day at room temperature. With this catalyst, it is noteworthy that an increase of the temperature up to 50°C allowed the recovery of a similar catalytic efficiency to that of the carboxylic acid analogue.

We thus carried out additional investigations to determine the origin of such reactivity differences. Samples of the [AsW₉O₃₃{P(O)^tBu}₂]⁵⁻ anion were prepared with different degrees of protonation. In the first place, the mono-protonated (*n*-Bu₄N)₃[NaHASW₉O₃₃{P(O)^tBu}₂]⁴⁻ (1) species was synthesized (see preparation in SI) and the catalytic behavior of this new sample was found equivalent to that of other compounds (cyclooctene conversion = 100%, at 20°C after 24h, see Table 1). In a second experiment, a sample of the protonated (*n*-Bu₄N)₃[NaHASW₉O₃₃{P(O)^tBu}₂]⁴⁻ (1) was treated with one equivalent of a methanolic solution (1 mol.l⁻¹) of *n*-Bu₄NOH, leading to the *in situ* formation of the [NaASW₉O₃₃{P(O)^tBu}₂]⁴⁻ anion. Using this sample for the cyclooctene epoxidation led to a dramatic decrease of the catalytic performances, attaining values obtained in our previous publication.⁵ It is, thus, possible to assign the previously reported results to the non-protonated (*n*-Bu₄N)₃[NaASW₉O₃₃{P(O)^tBu}] compound (see Table 1). Such synergetic process may be supported by previous studies by the group of Bonchio. It was shown that the catalyst efficiency was modulated through the protonation state of POMs (in their case γ-[SiW₁₀O₃₆]⁸⁻).¹² In addition, the groups of Kholdeeva and Poblet have recently emphasized the role of protons in heterolytic activation of H₂O₂ in the case of Nb-substituted polyoxometalates,¹³ as it can be observed for mesoporous niobium silicates¹⁴ or Zr-Based Metal-organic frameworks.¹⁵

Table 1. Comparison of the cyclooctene conversion (%) in the epoxidation of cyclooctene with aqueous H₂O₂ (30%) after 24h using the different

catalysts. Catalyst: 24 μmol . Acetonitrile: 20 mL. Ratio catalyst/cyclooctene/ H_2O_2 = 1/250/250.

Catalysts	T(°C)	(%)	ref
$(^n\text{-Bu}_4\text{N})_3[\text{NaHAsW}_9\text{O}_{33}\{\text{P}(\text{O})\text{CH}_2\text{CH}_2\text{CO}_2\text{H}\}_2]$ (2)	20	96	5
$(^n\text{-Bu}_4\text{N})_3[\text{NaHPW}_9\text{O}_{34}\{\text{P}(\text{O})\text{-C}_6\text{H}_4\text{NH}_2\}_2]$ (3)	20	97	5
$(^n\text{-Bu}_4\text{N})_4[\text{NaAsW}_9\text{O}_{33}\{\text{P}(\text{O})^i\text{Bu}\}_2]$	20	7.5	5
$(^n\text{-Bu}_4\text{N})_4[\text{NaAsW}_9\text{O}_{33}\{\text{P}(\text{O})^i\text{Bu}\}_2]$	50	100	5
$(^n\text{-Bu}_4\text{N})_3[\text{NaHAsW}_9\text{O}_{33}\{\text{P}(\text{O})^i\text{Bu}\}_2]$ (1)	20	100	This work

1.2. Theoretical insight in the exact localization of the proton in $(^n\text{-Bu}_4\text{N})_3[\text{NaHAsW}_9\text{O}_{33}\{\text{P}(\text{O})\text{R}\}_2]$

Having in mind these experimental results, we were interested in determining the nature of the catalytic site and especially the role of protons in the reactivity of the $(^n\text{-Bu}_4\text{N})_3[\text{NaHAsW}_9\text{O}_{33}\{\text{P}(\text{O})\text{R}\}_2]$ compounds (R = $-\text{CH}_2\text{CH}_2\text{CO}_2\text{H}$, ^iBu). A DFT study was thus carried out (see experimental part for computational details) aiming at analyzing the energetic and structural factors associated to the protonation of the phosphonyle polyoxoanions. This was first carried out on a slightly simplified model of the anionic part of the compound, namely $[\text{NaHAsW}_9\text{O}_{33}\{\text{P}(\text{O})\text{CH}_3\}_2]^{3-}$. Working on such highly charged species can be justified by the non-coordinating nature of the counter ions ($^n\text{-Bu}_4\text{N}^+$).

In a general way, the $[\text{NaAsW}_9\text{O}_{33}\{\text{P}(\text{O})\text{R}\}_2]^{4-}$ anions are obtained through the grafting of two phosphonyle groups onto four oxygen atoms from the lacuna of the $\text{B}_{\alpha}\text{-}\{\text{AsW}_9\text{O}_{33}\}^{9-}$ unit. Consequently, the two free O atoms of the lacuna are strongly nucleophilic. Furthermore, the two O atoms from the $\{\text{RP}=\text{O}\}$ functions present nucleophilic properties, as shown before.^{11,16} For this reason, we have considered that the Na^+ cation was located inside the lacuna and linked to these four O atoms, a location which was found as a minimum (Figure 3) by DFT calculations (see Computational details). Despite the absence of X-ray diffraction structural determination available for the family of compounds $(^n\text{-Bu}_4\text{N})_3[\text{NaHAsW}_9\text{O}_{33}\{\text{P}(\text{O})\text{R}\}_2]$,¹⁷ the coordination of Na^+ is in line with the partial structure determined by X-Ray crystallography obtained on $(^n\text{-Bu}_4\text{N})_3[\text{NaHPW}_9\text{O}_{34}\{\text{As}(\text{O})\text{-C}_6\text{H}_4\text{NH}_2\}_2]$ (3).¹⁶ In contrast, in the absence of experimental elements, the position of the H^+ counterion had to be determined computationally. This was carried out through geometry optimization of the regioisomers differing from the proton position (see figure S1 for complete results). The isomers where the proton is localized on one of the two nucleophilic O atoms that delimit the lacuna (Figure 3) are the most stable. This is in accordance with the experimental results obtained by the group of Mizuno in the case of other lacunary POMs such as $[\gamma\text{-}\{\text{SiW}_{10}\text{O}_{32}(\text{H}_2\text{O})_2\}_2(\mu\text{-O})_2]^{4-}$.¹⁸ The presence of the proton is shown to significantly lengthen the $\text{Na}\dots\text{O}(\text{H})\text{W}$ distance, compared to when the proton is absent (Figure 3). In conclusion, it is probably more accurate to consider one $\{\text{W}=\text{O}\}$ group that delimit the lacuna of the $\{\text{XW}_9\}$ scaffold as a $\{\text{W}-\text{OH}\}$ group. Further insight within the formation of the catalytically active site was sought through examination of the interaction of H_2O_2 with the lacuna.

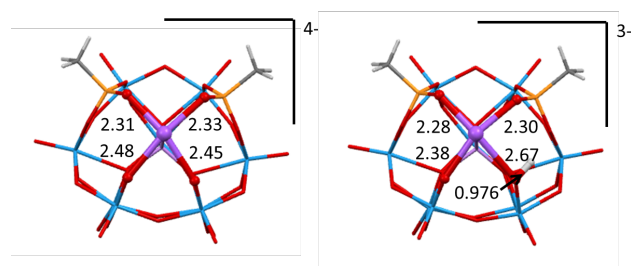


Figure 3: Optimized structure of the $[\text{NaAsW}_9\text{O}_{33}\{\text{P}(\text{O})\text{CH}_3\}_2]^{4-}$ (left) and $[\text{NaHAsW}_9\text{O}_{33}\{\text{P}(\text{O})\text{CH}_2\}_2]^{3-}$ (right) anions showing the location of the Na^+ cation (violet) inside the lacuna of the anion and of the proton (white) on one of nucleophilic WO. All the atoms - W (in blue), O (in red), As (in purple), P (in orange), C (in grey), H (in white) - are represented in stick mode, except for the lacuna (ball and sticks). Distances within the ball and stick part are in Å.

1.3. Theoretical examination of the interaction of H_2O_2 with $(^n\text{-Bu}_4\text{N})_3[\text{NaHAsW}_9\text{O}_{33}\{\text{P}(\text{O})\text{CH}_3\}\{\text{P}(\text{O})\text{CH}_2\text{CH}_2\text{COOH}\}]$

Regarding the literature and the impact of acetic acid in epoxidation processes,¹⁹ it was also interesting to determine whether the presence of carboxylic acid functions found on the hybrid of POM could affect the chemical process. To carry out this theoretical study, one CH_3 group of $[\text{NaAsW}_9\text{O}_{33}\{\text{P}(\text{O})\text{CH}_3\}_2]^{4-}$ was replaced by a $\text{CH}_2\text{CH}_2\text{CO}_2\text{H}$ group. In this case, a similar structure was obtained for the protonated product $[\text{NaAsW}_9\text{O}_{32}(\text{OH})\{\text{P}(\text{O})\text{CH}_3\}\{\text{P}(\text{O})\text{CH}_2\text{CH}_2\text{COOH}\}]^{3-}$ (designed as $\{\text{W}-\text{OH}\}$, see figure 4, top/left and figure S2 for complete results), and no significant impact of the carboxylic acid function was observed at this stage. In particular, as only one carboxylic acid arm was added (see Figure 4), this simplification obviously led to a dissymmetry. However, we found that the coordination of the proton on one or the other $\{\text{W}=\text{O}\}$ group of the lacuna differed by less than $0.1 \text{ kcal}\cdot\text{mol}^{-1}$. Furthermore, the most stable structure resulting from the transfer of the carboxylic acid proton onto the second $\{\text{W}=\text{O}\}$ group of the lacuna was found to be more than $15 \text{ kcal}\cdot\text{mol}^{-1}$ higher in energy. Consequently, the possibility for transprotonation from the $\text{CH}_2\text{CH}_2\text{CO}_2\text{H}$ chain to the POM framework was definitively discarded.

In a second step, the interaction of H_2O_2 with the $[\text{NaAsW}_9\text{O}_{32}(\text{OH})\{\text{P}(\text{O})\text{CH}_3\}\{\text{P}(\text{O})\text{CH}_2\text{CH}_2\text{COOH}\}]^{3-}$ model compound was examined. Thus, the formation of a hydroperoxide site $\{\text{W}-\text{OOH}\}$ (associated to a free water molecule), instead of the $\{\text{W}-\text{OH}\}$ group, was found to be quasi athermic. Moreover, it can be stabilized by more than $10 \text{ kcal}\cdot\text{mol}^{-1}$ through coordination of water to the Na^+ center (figure 4, top/right).

On the other hand, the formation of a peracid function on the carboxylic acidic arm was slightly less favored than the formation of the hydroperoxide ($+1.4 \text{ kcal}\cdot\text{mol}^{-1}$, $\{\text{CO}-\text{OOH}\}$ figure 4, bottom/left). This preference did not change when H_2O was bound to Na^+ (total: $+2.7 \text{ kcal}\cdot\text{mol}^{-1}$, see Table S3).

Finally, the formation of a peroxy-hydroxo pair (designed as $\{\text{W}-\text{OO}\}$ figure 4, bottom/right, and figure S4 for complete studies) after proton transfer from the hydroperoxide to the other $\{\text{W}=\text{O}\}$ site of the lacuna, was also slightly disfavored ($+2.5$ or $+4.2 \text{ kcal}\cdot\text{mol}^{-1}$, without or with a water molecule coordinated on the Na^+ cations respectively). Despite these small energy differences, and considering:

- i) the catalytic activities of the protonated and unprotonated form of the $[\text{NaAsW}_9\text{O}_{33}\{\text{P}(\text{O})\text{R}\}_2]^{4-}$ anions on the one hand,

- ii) the catalytic activities of the different $[\text{NaHAsW}_9\text{O}_{33}\{\text{P}(\text{O})\text{R}\}_2]^{3-}$ anions on the other hand, it can be suggested that the true catalytic center for the epoxidation reaction is the hydroperoxylated $\{\text{W}-\text{OOH}\}$ form, even in the case where $\text{R} = -\text{CH}_2\text{CH}_2\text{COOH}$.

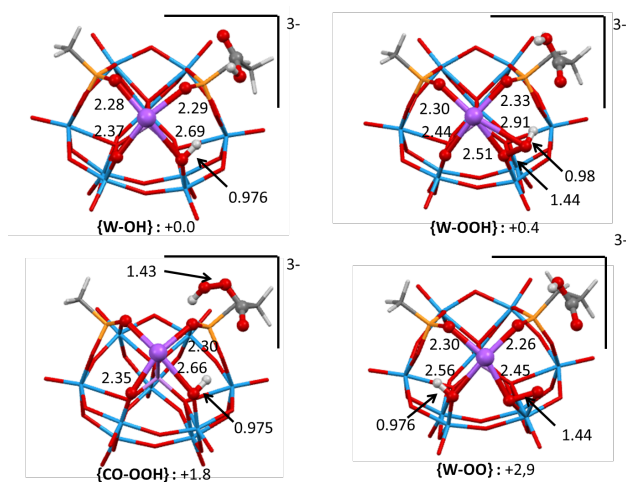


Figure 4: Optimized structure of the $[\text{NaAsW}_9\text{O}_{32}(\text{OH})\{\text{P}(\text{O})\text{CH}_3\}\{\text{P}(\text{O})\text{CH}_2\text{CH}_2\text{COOH}\}]^{3-}$ anion ($\{\text{W}-\text{OH}\}$, top/left) and of its potential products of reaction with H_2O_2 : $[\text{NaAsW}_9\text{O}_{32}(\text{OOH})\{\text{P}(\text{O})\text{CH}_3\}\{\text{P}(\text{O})\text{CH}_2\text{CH}_2\text{COOH}\}]^{3-}$ ($\{\text{W}-\text{OOH}\}$, top/right), $[\text{NaAsW}_9\text{O}_{32}(\text{OH})\{\text{P}(\text{O})\text{CH}_3\}\{\text{P}(\text{O})\text{CH}_2\text{CH}_2\text{COOOH}\}]^{3-}$ ($\{\text{CO}-\text{OOH}\}$, bottom/left) and $[\text{NaAsW}_9\text{O}_{31}(\text{OH})(\text{OO})\{\text{P}(\text{O})\text{CH}_3\}\{\text{P}(\text{O})\text{CH}_2\text{CH}_2\text{COOH}\}]^{3-}$ ($\{\text{W}-\text{OO}\}$, bottom/right). Colour codes are similar to Figure 3. Distances within the ball and stick part are in Å (not given if equal to the $\{\text{W}-\text{OH}\}$ distance). Energies of the reaction products + H_2O are computed in $\text{kcal}\cdot\text{mol}^{-1}$ with respect to $\{\text{W}-\text{OH}\} + \text{H}_2\text{O}$.

2. Oxidation of limonene in the presence of the ($n\text{-Bu}_4\text{N}$) $_3$ [NaHAsW $_9$ O $_{33}$ {P(O)(R)} $_2$] (compounds 1 and 2) or ($n\text{-Bu}_4\text{N}$) $_3$ [NaHPW $_9$ O $_{34}$ {As(O)p-C $_6$ H $_4$ NH $_2$ }] $_2$] (3) catalysts. In the present work, the catalytic oxidation of bio-sourced molecules that are contained in essential oils of citrus has been also investigated. Indeed, citrus fruit is considered as one of the highest values fruits crops because of its good resistance during transport and storage phases.²⁰ As a result, the essential oils obtained from the peel contain valuable components such as terpenes and, to a lesser extent, their oxygenated derivatives such as alcohols, aldehydes, esters, ketones and a few amounts of oxides. Part of the essential oil of citrus fruit, in particular unsaturated terpenic hydrocarbons such as limonene is undesirable for the cosmetics and perfumes industries. Limonene has been indeed classified as an allergen since the European Directive (2003/15/CE).

Furthermore, efficient and easy recovery processes of the terpenic fraction allowing the utilization of sustainable liquid-liquid extraction, in particular using ethanol/water mixtures, now exist.²¹ Our concern was then to devise sustainable production of higher value oxygenated derivatives of terpenes. In this regard, this study focused on the catalytic transformation of the most abundant compound, the limonene, preferentially into its corresponding mono-epoxides.

However, the oxidation of terpenes is particularly versatile. Various compounds can be obtained depending on the catalysts and/or the reaction conditions.²² In the case of

limonene, the monoxide forms (*cis*-1,2-epoxylimonene and *trans*-1,2-epoxylimonene) are considered as valuable platform compounds, but alcohols (carveol), ketones (carvone), diols, aldehydes and over-oxidized forms (such as 1,2,8,9-diepoxyimonene) are systematically obtained as well (see figure 5).²³

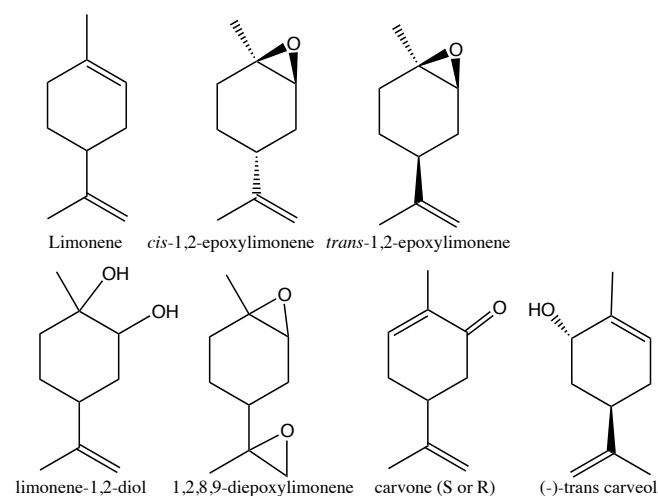


Figure 5. Limonene and most frequent oxidized forms of limonene.

2.1. Selectivity of the epoxidation of pure R(+)-limonene

Prior to the transformation of the citrus essential oil, we first performed an extended study on a commercial sample of R(+)-limonene, in order to check the catalytic performance of the ($n\text{-Bu}_4\text{N}$) $_3$ [NaHAsW $_9$ O $_{33}$ {P(O)R} $_2$] $^{3-}$ ($\text{R} = -\text{C}(\text{CH}_3)_3$ (1) or $-\text{CH}_2\text{CH}_2\text{CO}_2\text{H}$ (2)) and ($n\text{-Bu}_4\text{N}$) $_3$ [NaHPW $_9$ O $_{34}$ {As(O)p-C $_6$ H $_4$ NH $_2$ }] $_2$] (3) derivatives in the presence of the most abundant component of the oil. Due to slightly lower reactivity of limonene compared to simple cycloalkenes, all these catalytic tests were performed in acetonitrile with a two-fold excess of H_2O_2 with regard to the substrate (molar ratio of catalyst/alkene/ H_2O_2 equal to 1/250/500). The effect of the temperature was also studied and tests were carried out at 4, 25 and 50°C. In all these experiments, the decomposition of H_2O_2 in the presence of the catalysts was not observed, and the mass balances for the organic derivatives were systematically determined using an internal standard (tetradecane).

In acetonitrile, all three hybrids of POMs led to high conversions of limonene at 25°C, especially for ($n\text{-Bu}_4\text{N}$) $_3$ [NaHAsW $_9$ O $_{33}$ {P(O) t Bu}] $_2$] (1) and ($n\text{-Bu}_4\text{N}$) $_3$ [NaHAsW $_9$ O $_{33}$ {P(O)CH $_2$ CH $_2$ CO $_2$ H}] $_2$] (2), for which the cyclooctene conversions were respectively 95% and 100 % after 24h. The aniline derivative ($n\text{-Bu}_4\text{N}$) $_3$ [NaHPW $_9$ O $_{34}$ {As(O)p-C $_6$ H $_4$ NH $_2$ }] $_2$] (3) displayed a more modest catalytic activity (conversion = 77 % after 24h) than the two other POM catalysts. Furthermore, for this specific compound, prolonging the reaction up to 48h did not increase significantly the limonene conversion (81 %), probably due to the deactivation/decomposition of the catalyst. This result is not particularly surprising since the $\{\text{A},\alpha\text{-PW}_9\}$ scaffold is generally considered to be less stable than the $\{\text{B},\alpha\text{-X}^{\text{III}}\text{W}_9\}$ (X = As, Sb) ones.²⁴

The catalytic results at 4, 25 and 50°C with $B_{\alpha}\text{-[NaHASW}_9\text{O}_{33}\{\text{P(O)CH}_2\text{CH}_2\text{CO}_2\text{H}\}_2\}^{3-}$ (**2**) are reported in figure 6 (left).

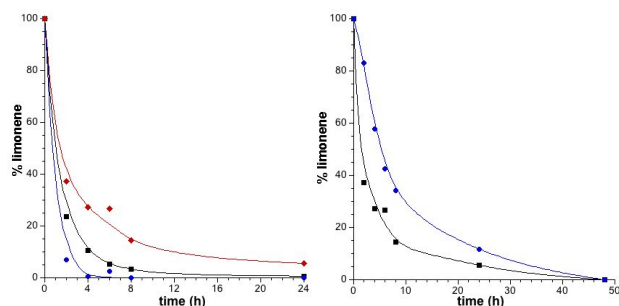


Figure 6. Left: time profiles of the limonene content in the presence of $\text{H}_2\text{O}_2/\text{[NaHASW}_9\text{O}_{33}\{\text{P(O)CH}_2\text{CH}_2\text{CO}_2\text{H}\}_2\}^{3-}$ (**2**) at 4°C (red), 25°C (black) and 50°C (blue). Right: time profiles of the limonene content at 4°C in the presence of $[\text{NaHASW}_9\text{O}_{33}\{\text{(PO)}^t\text{Bu}\}_2]^{3-}$ (**1** in blue) and $[\text{NaHASW}_9\text{O}_{33}\{\text{P(O)CH}_2\text{CH}_2\text{CO}_2\text{H}\}_2]^{3-}$ (**2**, in black).

With this catalyst, the reaction with H_2O_2 was completed after 8h at 25°C (TOF = 95 h^{-1}) whereas it was completed after 4h at 50°C (TOF = 116 h^{-1}). Furthermore, even at a low temperature (4°C) both catalysts were still active (limonene conversion of 88% after 24h for $(^n\text{-Bu}_4\text{N})_3[\text{NaHASW}_9\text{O}_{33}\{\text{P(O)}^t\text{Bu}\}_2]^{3-}$ (**1**) and 95 % for $(^n\text{-Bu}_4\text{N})_3[\text{NaHASW}_9\text{O}_{33}\{\text{P(O)CH}_2\text{CH}_2\text{CO}_2\text{H}\}_2]^{3-}$ (**2**) and led to complete limonene conversions after two days (Fig. 6, right).

To our knowledge these are the first examples of POM-based catalytic systems that are efficient at such low temperature. Even if both catalysts are hardly as efficient as the classical iron(II) and iron (III) non porphyrinic systems (such as (poly)pyridyl, salen, benzylamine or heterocycle-based complexes)¹⁹ that operate at 0°C, the present systems have the great advantage of not using acetic acid as a co-factor in the catalytic processes.

The selectivity in epoxides was also found strongly dependent on the temperature. Table 2 presents the yields for the different products in the oxidation of limonene at 25°C for the three catalysts.

Table 2. Yields (%) for the reaction products in the oxidation of limonene with aqueous H_2O_2 in the presence of the 3 catalysts at 25°C. Catalyst: 24 μmol . Acetonitrile: 20 mL. Ratio catalyst/limonene/ H_2O_2 = 1/250/500.

Catalysts	1,2-epoxy-limonene		diol	diepoxy	others
	trans	cis			
$[\text{NaHASW}_9\text{O}_{33}\{\text{P(O)CH}_2\text{CH}_2\text{CO}_2\text{H}\}_2]^{3-}$	76	7.5	4.5	3	9
$[\text{NaHASW}_9\text{O}_{33}\{\text{P(O)}^t\text{Bu}\}_2]^{3-}$	47.5	6.0	7.5	13.5	20.5
$[\text{NaHPW}_9\text{O}_{34}\{\text{As(O)}p\text{-C}_6\text{H}_4\text{NH}_2\}_2]^{3-}$	46.5	16.5	2.5	0.5	11.0

At this temperature, the *trans*-1,2-epoxylimonene was always the major product of the reaction, but the selectivity did not exceed 76 % for the most selective catalyst $(^n\text{-Bu}_4\text{N})_3[\text{NaHASW}_9\text{O}_{33}\{\text{P(O)CH}_2\text{CH}_2\text{CO}_2\text{H}\}_2]^{3-}$ in these conditions. Taking into account the *cis*-epoxide, the diepoxylimonene and the limonene-1,2-diol (which results from the opening of the oxirane ring) the sum of the selectivities into the monoepoxides and their derivatives are fairly good: 91% for $(^n\text{-Bu}_4\text{N})_3[\text{NaHASW}_9\text{O}_{33}\{\text{P(O)CH}_2\text{CH}_2\text{CO}_2\text{H}\}_2]^{3-}$ (**2**), 78% for $(^n\text{-Bu}_4\text{N})_3[\text{NaHASW}_9\text{O}_{33}\{\text{P(O)}^t\text{Bu}\}_2]^{3-}$ (**1**) and 86% for $(^n\text{-Bu}_4\text{N})_3[\text{NaHPW}_9\text{O}_{34}\{\text{As(O)}p\text{-C}_6\text{H}_4\text{NH}_2\}_2]^{3-}$ (**3**). However at this temperature, and *a fortiori* at the highest temperature tested

(50°C), the formation of products that came from allylic reactions (especially carvone and carveol) cannot be avoided. Moreover, decreasing the temperature down to 4°C for the two $B_{\alpha}\text{-[NaHASW}_9\text{O}_{33}\{\text{P(O)R}\}_2\}^{3-}$ catalysts led to excellent epoxide selectivities:

- for R = $-\text{C}(\text{CH}_3)_3$ (**1**), the conversion at 24h was 88% (100 % after 48h) with a selectivity for the *trans*-1,2-epoxylimonene of 100 % (TOF = 9 h^{-1}).

- for R = $-\text{CH}_2\text{CH}_2\text{CO}_2\text{H}$ (**2**), the conversion at 24h was 95% (TOF = 10 h^{-1}) with a final selectivity of 94% for the *trans*-1,2-epoxylimonene and 6% for the *cis*-epoxide.

With both catalysts, no products from allylic reactions were detected. The selective formation of epoxides/diols was expected for POMs that do not contain any additional transition-metal ions. Due to the presence of d^0 centers, these compounds are indeed good candidates for the (selective) electrophilic attack on double bond,²⁵ rather than the alpha allylic oxidation or hydroxylation, that can lead respectively to the formation of carvone and carveol derivatives. In this regard, they combined at once very good conversion and selectivity, better than most of the classical epoxidation catalysts, such as heterogeneous silylated Ti-MCM-41 materials, which led to TOF up to 60 h^{-1} but at high temperature (90°C) (limonene conversion = 89 % after 24h, and 1,2-epoxide selectivity = 78%).²⁶ Comparison with the peroxotungstates complexes described in the literature (including the classical Ishii–Venturello system) is also in favor of the $B_{\alpha}\text{-[NaHASW}_9\text{O}_{33}\{\text{P(O)R}\}_2\}^{3-}$ anions. Indeed, such systems generally displayed lower catalytic performances at room temperature (for the limonene).²⁷

Moreover, even if complete conversion and/or very good selectivity can be obtained with the Ishii–Venturello catalyst (at room temperature, limonene conversion = 94 % after 1,5h, TOF = 125 h^{-1} , limonene epoxide selectivity = 97%), this catalyst generally worked in harmful chlorinated solvents.²⁸ In order to avoid such solvents, alternative strategies were proposed by certain groups, such as the use of eco-friendly bi/triphasic processes²⁹ or the immobilization of the Ishii–Venturello catalyst onto macroreticular amberlite support.³⁰ In this last study, this immobilization allowed the authors the use of CH_3CN , but resulted in a lower catalytic efficiency: at T = 38°C, limonene conversion = 84 % after 24%, and excellent selectivity for the monoepoxide = 93%.

We also verified the reactivity of the $(^n\text{-Bu}_4\text{N})_3[\text{NaHASW}_9\text{O}_{33}\{\text{(PO)}^t\text{Bu}\}_2]^{3-}$ (**1**) at 4°C, in three other solvents: tetrahydrofuran (THF), ethyl acetate (AcOEt) and ethanol. In all cases (see Table 3), the reaction was slower than in acetonitrile but the limonene conversions after 2 days were rather satisfying (95 % in THF, 93 in AcOEt and 98 % in ethanol). Concerning the other catalysts and the fact that they are much less soluble in these three solvents, it was not possible to evaluate their catalytic performances.

Table 3. Limonene conversion (%) and yields (%) for the reaction products in the oxidation of limonene with aqueous H_2O_2 in the presence of $(^n\text{-Bu}_4\text{N})_3[\text{NaHASW}_9\text{O}_{33}\{\text{P(O)}^t\text{Bu}\}_2]^{3-}$ (**1**) at 4°C in THF, AcOEt and ethanol after 48h. Ratio catalyst/limonene/ H_2O_2 : 1/250/500.

Solvent	Conv. (%)	1,2-epoxy-limonene		diol	diepoxy	others
		trans	cis			
THF	95	44.0	0	5.0	46.0	0
AcOEt	93	24.3	0	55.7	8.7	0
Ethanol	98	0	0	33.4	24.7	40.3

The catalytic performances of $(^n\text{-Bu}_4\text{N})_3[\text{NaHASW}_9\text{O}_{33}\{\text{P(O)}^t\text{Bu}\}_2]^{3-}$ (**1**) were dramatically

different in the three solvents: i) the *trans*-epoxylimonene was never the major constituent of the final mixture and the formation of the *cis*-isomer was never observed; ii) in THF, the epoxidation reaction was really efficient, even for the terminal alkene functions, with the major constituents being the 1,2,8,9-diepoxyimonene and the *trans*-1,2-epoxylimonene; iii) the epoxides hydrolysis was favored in AcOEt (55.7 % of diols); iv) surprisingly, allylic products were formed in large amounts (40.3 %) in ethanol, while the monoepoxides were not detected. This is an interesting result since it is possible to effectively guide the course of the reaction by simply modifying the temperature and/or the solvent.

2.2. Epoxidation of the terpene part of sweet orange essential oil at 4°C with $(n\text{-Bu}_4\text{N})_3[\text{NaHAsW}_9\text{O}_{33}\{\text{P}(\text{O})^t\text{Bu}\}_2]$ (1).

We finally investigated the behavior of the $(n\text{-Bu}_4\text{N})_3[\text{NaHAsW}_9\text{O}_{33}\{\text{P}(\text{O})^t\text{Bu}\}_2]$ catalyst (1) toward the terpene part (see deterpenation protocol in Experimental Section) of an essential oil of sweet orange from Portugal. Limonene is the major component of this fraction (90%), but other compounds (especially monoterpenes such as α -pinene, β -myrcene, sabinene, 3-carene...) are also present.³¹ The oxidation of the pure essential oil was performed in conditions identical to those used for the limonene, with a catalyst/alkene/ H_2O_2 ratio equal to 1/250/500 in acetonitrile at 4°C, in order to optimize the potential formation of the different epoxides. We found that the oxidation of limonene in the essential oil was somewhat slower than starting from pure limonene in the course of time (conversion at 24 h equal to 71% for the essential oil, 88% in the case of pure limonene, see figure 7). Besides the remaining limonene, the main constituents of the mixture after oxidation were the *trans*-1,2-epoxylimonene (52%), the corresponding *cis*- isomer (2%) and the compounds initially present in the mixture.

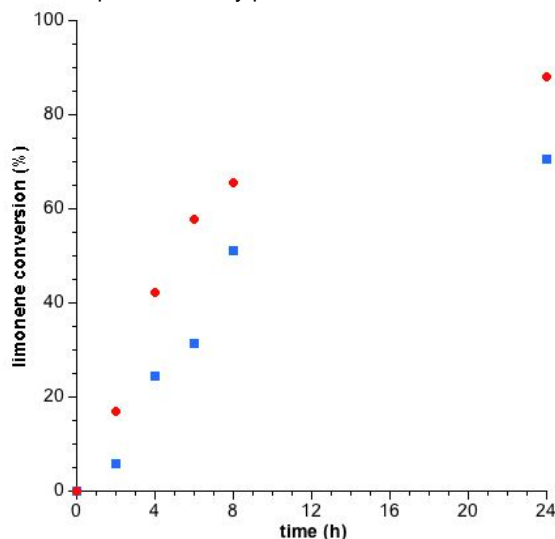


Figure 7. Time profile for limonene conversion starting from pure limonene (red dots) or sweet orange essential oil (blue square) with aqueous H_2O_2 in the presence of $(n\text{-Bu}_4\text{N})_3[\text{NaHAsW}_9\text{O}_{33}\{\text{P}(\text{O})^t\text{Bu}\}_2]$ (1) at 4°C in acetonitrile.

No trace of oxidation products of the other initial components of the essential oil could be detected, excepted for α -pinene, for which a weak amount of the corresponding epoxide was detected (less than 1%). In particular β -pinene, the other main constituent of the essential oil, did not react

with H_2O_2 in the presence of the catalyst. Thus, non-oxidized derivatives (remaining limonene, β -pinene...), α -allylic oxidation and α -allylic hydroxylation products represent 45 % of the final mixture.

Conclusions

In the present study, the performances of organophosphonyl/arsonyl derivatives of POMs, *i.e.* $(n\text{-Bu}_4\text{N})_3[\text{NaHAsW}_9\text{O}_{33}\{\text{P}(\text{O})\text{R}\}_2]$ ($\text{R} = {}^t\text{BuPO}$ (1) and $\text{CH}_2\text{CH}_2\text{COOH}$ (2)) and $(n\text{-Bu}_4\text{N})_3[\text{NaHPW}_9\text{O}_{34}\{\text{As}(\text{O})p\text{-C}_6\text{H}_4\text{NH}_2\}_2]$ (3), as terpenes epoxidation catalysts were explored. These complexes (to a lesser extent for $(n\text{-Bu}_4\text{N})_3[\text{NaHPW}_9\text{O}_{34}\{\text{As}(\text{O})p\text{-C}_6\text{H}_4\text{NH}_2\}_2]$) showed a high efficiency for the conversion of limonene whatever the temperature used (from 4 to 50°C). Indeed, $(n\text{-Bu}_4\text{N})_3[\text{NaHAsW}_9\text{O}_{33}\{\text{P}(\text{O})\text{R}\}_2]$ ($\text{R} = \text{CH}_2\text{CH}_2\text{COOH}$ and ${}^t\text{Bu}$) allowed the complete conversion of pure limonene at 50°C in a few hours in CH_3CN (ca. 12h at room temperature). As expected, the selectivity of this epoxidation reaction decreased by increasing the temperature. At 4°C in acetonitrile, the complete conversion of the limonene into the *trans*-epoxylimonene as the sole reaction product was observed after 48h for the most selective catalyst $(n\text{-Bu}_4\text{N})_3[\text{NaHAsW}_9\text{O}_{33}\{\text{P}(\text{O})^t\text{Bu}\}_2]$ (1). Moreover, varying the solvent allowed modulating the selectivity of the catalytic process, leading to the formation of products from allylic reactions (carvone and carveol) in ethanol.

A DFT study was performed in order to understand the role of protons in the $(n\text{-Bu}_4\text{N})_3[\text{NaHAsW}_9\text{O}_{33}\{\text{P}(\text{O})\text{R}\}_2]$ -catalyzed epoxidation by H_2O_2 . The exact localization of the protons associated to the anions in this family of complexes allowed the determination of the presence of a {W-OH} group at the lacuna of these hybrids of POMs. It is highly likely that the reaction of hydrogen peroxide with the POMs occurs at this {W-OH} group. However, no certainty on the mechanism can be drawn at this time on the basis of these sole thermodynamic considerations.

Finally, this study also highlighted that $(n\text{-Bu}_4\text{N})_3[\text{NaHAsW}_9\text{O}_{33}\{\text{P}(\text{O})^t\text{Bu}\}_2]$ (1) was a very specific catalyst for the oxidation of limonene in natural extracts. In a more general way, phosphonyl derivatives of trivalent POMs displayed good limonene conversion and high selectivity in the epoxylimonene formation. This class of catalysts may thus be considered as ideal candidates for the scale up of the selective oxidation of terpenes such as limonene, and its development as practical biorefinery, *i.e.* as a facility that processes agricultural material (such as crop waste) to commercially useful chemicals.³²

Experimental Section

Materials and Methods

The synthesis and the characterization of the $(n\text{-Bu}_4\text{N})_3[\text{NaHAsW}_9\text{O}_{33}\{\text{P}(\text{O})\text{R}\}_2]$ ($\text{R} = {}^t\text{Bu}$ (1) and $\text{CH}_2\text{CH}_2\text{COOH}$ (2)) and $(n\text{-Bu}_4\text{N})_3[\text{NaHPW}_9\text{O}_{34}\{\text{As}(\text{O})p\text{-C}_6\text{H}_4\text{NH}_2\}_2]$ (3) catalysts were previously described.¹⁶ Solvents and other reagents were obtained from various commercial sources and used as received. R-(+)-limonene (purity > 99%) stabilized with BHT was purchased from Fluka, tetradecane (purity >99%) from Siccapp emmop and the essential oil of sweet orange from Portugal was purchased

from COOPER – Coopération Pharmaceutique Française, Melun, France (lot N° 04050164/B). The samples were analyzed by gas-phase chromatography coupled to an electron impact mass spectrometer (70eV). The different components were separated using a non-polar capillary column (Supelco SPB5, 60 m*0,32 mm, stationary phase 0,25 µm), with helium as the gas vector (0.7 mL.min⁻¹). The injector temperature was fixed at 250°C in split mode (1/50). The oven was initially set at 60°C, increased until 250°C (4°C.min⁻¹) for 50 min, for a total duration of 97 min. The identification of the different components was performed by comparison of their calculated retention indexes and of the corresponding mass spectra with data from the literature³³ or available data banks (NIST, WILEY).

Computational Details. The calculations were carried out with the Gaussian 09 package of programs³⁴ using the B3PW91 functional.³⁵ Arsenic and tungsten were represented with the effective core potential SDDall³⁶ from the Stuttgart group and the associated basis set. The 6-31G** basis set was used for all the other atoms (H, C, O, Na and P). Full geometry optimizations were carried out using a polarizable continuum model (SM,³⁷ acetonitrile as a solvent) for solvation.

Catalytic studies.

The conditions used for the previously reported epoxidation of cyclooctene/cyclohexene are provided below.³⁸

The limonene oxidation experiments were realized with a molar catalyst/limonene/H₂O₂ ratio equal to 1/250/500 at three different temperatures (4°C, 25°C and 50°C).

In three vials, R-(+)-limonene (0.5 mL), H₂O₂ (0.7 mL of a 30%wt. aqueous solution), and tetradecane (0.45 mL, internal standard) were successively added in 10 mL of acetonitrile. The catalyst (ca. 40 mg) was then added and the vials were finally sealed and stirred at their corresponding temperatures. Samples (1.5 mL) of each solution were taken after 2 h, 4 h, 6 h, 8 h et 24 h and mixed with 1.5 mL of pentane. The organic layer that contains the limonene and the oxygenated derivatives was removed, dried with Na₂SO₄ and injected (0.08 µL) in GC/MS.

Deterpenation protocol²¹

The essential oil of sweet orange from Portugal (10 mL) was added to 30 mL of ethanol (70% vol). This mixture was degassed under N₂ for 10 min and was stirred for 4 days at 4° C. The suspension was then allowed to settle for 6 days. After separation, the terpene phase was mixed with water (1/1 volume) and solid NaCl was added in the mixture until saturation of the aqueous phase. After a few hours, the limpid organic phase was then separated using a separating funnel, dried with Na₂SO₄ and checked by GC-MS.

Acknowledgments.

This work was supported by the CNRS, Sorbonne Université and the Université de la Réunion. The authors also acknowledge support from the French Ministère de la Recherche et de l'Enseignement Supérieur for a PhD fellowship to Miss Ourania Makrygenni.

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- ³⁷ A. V. Marenich, C. J. Cramer, D. G. Truhlar, *J. Phys. Chem. B* **2009**, *113*, 6378-6396.
- ³⁸ Epoxidation of cyclooctene/cyclohexene: In homogeneous conditions, experiments of cyclooctene (or cyclohexene) epoxidation were performed in a 50 mL round-bottom flask equipped with a condenser and magnetic stirrer at room temperature or at 50°C. The catalyst (n-Bu₄N)₃NaH[NaHAsW₉O₃₃{P(O)CH₂CH₂COOH}₂] (80 mg, 24 μmol / 1 eq.), 20 mL of acetonitrile, 0.78 mL of cyclooctene or 0.60 mL of cyclohexene (6 mmol / 250 eq.), 0.9 mL of decane (as an internal standard) and 0.6 mL of H₂O₂ (30% / 6 mmol / 250 eq) were introduced successively in the flask. The resulting solutions were analysed by gas chromatography on a Delsi Nermag DN 200 GC apparatus equipped with a flame ionization detector and a Macherey-Nagel Optima-5 capillary column (length 30 m, internal diameter 0.32 mm, thickness 1 μm). In heterogeneous conditions, experiments of cyclooctene/cyclohexene epoxidation were performed in three 5 mL flasks (for analysis at 3, 6 and 24 h) under stirring at room temperature. The solid catalyst (n-Bu₄N)₃NaH[NaHAsW₉O₃₃{P(O)CH₂CH₂COOH}₂]@SBA-NH₂ (40 mg), - corresponding to 1.2.10⁻³ mmol/1 eq. of POM - 1 mL of acetonitrile, 39 μL of cyclooctene (250 eq.) or 30 μL of cyclohexene (250 eq.), 45 μL of decane and 30 μL of H₂O₂ (30%, 250 eq.) were introduced successively in each flask. The supernatant of the resulting suspensions was analyzed by gas chromatography as described above.