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Lithium borates from the glass to the melt: a temperature-induced structural transformation viewed from the boron and oxygen atoms

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Abstract : A multi-edge study of the local structure of lithium borate glasses and melts has been carried out using X-ray Raman scattering (XRS) as a function of temperature. Thanks to a wide range of compositions, from pure B₂O₃ up to the metaborate composition, we are able to finely interpret in detail the modifications of the local environment of both the boron and oxygen atoms in terms of boron coordination number, formation of non-bridging oxygens (NBOs) and polymerization's degree of the borate framework as a function of temperature and composition. A temperature induced ⁴B to ³B conversion is observed above the glass transition temperature (T_g) from the glass to the melt from the triborate composition up to the metaborate composition. Two distinct melt structures are reported: a well polymerized borate network – with few NBOs - below the triborate composition, and a depolymerized borate network above the diborate composition with a rapid increase of the number of NBOs when Li₂O is added. These two structurally distinct melts allow explaining the two dynamical regimes observed for lithium ion diffusion.

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1. Introduction

Ionic transport, diffusivity and viscosity are fundamental properties of melts, whose implications are at the origin of very important processes in material and Earth sciences. The modification of glass structure with temperature is therefore of major interest for the understanding of relaxation phenomena, but also to better understand the solid-liquid transition at the local scale. The high-temperature liquid is a well-defined phase in internal thermodynamic equilibrium, with constant and fast exchanges between atomic configurations through structural relaxation and diffusion processes. Unlike alkali silicate glasses whose properties and structure are mainly governed by the proportion of non-bridging oxygen (NBO), the behavior of alkali borate glasses is more complex.¹ The higher level of intricacy originates from the existence of two distinct coordination states of boron atoms: the planar units BO_3 and the negatively charged tetrahedral units BO_4^- , the combination of these two building blocks allowing a wide variety of tessellations of the tridimensional space.^{2,3}

Alkali borate glasses show an increasing fragility as the alkali oxide content increases.⁴⁻⁶ This fragile behaviour, characteristic of strong glass-formers, indicates a rapid raise in the configurational entropy above the glass transition temperature T_g , contrary to the strong B_2O_3 liquid. The link between structural properties and the entropy increase is not well understood, partly due to the lack of data on the structural reorganization when going from the glass to the liquid. At room temperature, the structure of alkali borate glasses is well described in the literature in terms of short- and medium-range order.^{1,3,7-12} A temperature induced ^{10}B to ^{11}B conversion has also been predicted by molecular dynamics and thermodynamic models^{13,14} and evidenced using *in situ* high-temperature sample environments coupled to Raman spectroscopy,¹⁵⁻¹⁷ X-ray or neutron diffraction^{15,18-21} or ^{11}B NMR.^{22,23} However, experimental data of the $\text{Li}_2\text{O}-\text{B}_2\text{O}_3$ system are rather sparse and mostly concern a few compositions in the range of 20-40 mol%. It is then worthwhile to study the structural properties of a wide range of compositions as a function of temperature from the glass to the melt.

Central to these problems is the bridging or non-bridging role of oxygen, the archetypal ligand in crystalline and glassy oxides. Many techniques have been used to probe the oxygen environment and in the last few years X-ray Raman scattering (XRS) spectroscopy has emerged as a powerful and versatile local probe of the electronic structure of low-Z elements under a wide range of experimental conditions.²⁴⁻²⁹ XRS spectra are sensitive to a multitude of electronic processes related to inelastic scattering and core electron excitations. These latter ones correspond mainly to the absorption edges of the various elements present in the sample and located in the probed energy-loss window.^{25,30,31} As for the conventional X-ray Absorption Near Edge Spectroscopy (XANES), XRS provides the opportunity to measure several absorption edges, such as the B and O K-edges, almost simultaneously, by simply selecting the energy loss range of interest. In the case of lithium borate glasses, the simultaneous measurement of both B and O K-edges is a powerful characterization method to get structural information on the local environment of every atom constituting the glassy network.

In this paper, we present a multi-edge study of the local structure of lithium borate glasses as a function of temperature - from room temperature up to 800°C. To reach the liquid state, an

aerodynamic levitation device coupled to a laser heating system was inserted into the sample chamber of a XRS spectrometer, allowing to monitor both the boron coordination number and the formation of NBO atoms as a function of temperature and glass composition.

2. Materials and Methods

a. Synthesis

Lithium borate glasses were synthesized by melting weighed mixtures of lithium carbonate (Li_2CO_3 - 99.99% - Alfa Aesar) and boric acid (H_3BO_3 - 99.5% - Sigma Aldrich) in a Pt/Rh crucible at 1000°C during 20 minutes. The crucible was then quenched in water. Five different glasses were prepared with the following molar compositions: $x\text{Li}_2\text{O}-(1-x)\text{B}_2\text{O}_3$, where $x = 0.5, 0.33, 0.25, 0.1$ and 0 . The five glasses will be denoted as LB ($\text{Li}_2\text{O}-\text{B}_2\text{O}_3$), LB2 ($\text{Li}_2\text{O}-2\text{B}_2\text{O}_3$), LB3 ($\text{Li}_2\text{O}-3\text{B}_2\text{O}_3$), LB9 ($\text{Li}_2\text{O}-9\text{B}_2\text{O}_3$) and v- B_2O_3 , respectively. Their respective glass transition temperatures are equal to 420°C , 500°C , 487°C , 343°C and 260°C .³²

Spheroidal samples were prepared by melting small pieces of the as-synthesized glasses in a levitation apparatus to obtain homogeneous samples of approximately 2mm in diameter. The principle of the levitation chamber is described in detail elsewhere.³³ For this experiment, the spherical sample was placed in the convergent-divergent nozzle that directed an oxygen gas jet supporting the sample. An additional regulated helium flow was also used to reduce the x-ray absorption in the chamber. Then the sample was melted using two 125 W CO_2 laser beams focused onto the sample by means of spherical mirrors. The specimen temperature was measured by an optical pyrometer operating at a wavelength of $1.625\ \mu\text{m}$ at the spot on the surface illuminated by the x-ray beam.

b. XRS Spectrometers

i. ID16 Beamline - ESRF

The XRS experiments on LB2, LB3 and LB glasses were carried out at the ID16 Inelastic X-ray beamline of the European Synchrotron Radiation Facility (ESRF).³⁴ The inelastic X-ray scattering experiments previously possible on port ID16 of the ESRF are now carried out on port ID20. The incident beam was monochromatized with Si(111) and Si(220) monochromators and the beam focused to $115\ (\text{H}) \times 40\ (\text{V})\ \mu\text{m}^2$. The scattered radiation was analyzed with nine bent Si(660) analyzers ($R = 1\ \text{m}$) and detected with a photon-counting MAXIPIX bidimensional pixel detector.³⁵ The spectra were collected by scanning the incident energy while keeping the analyzers energy to $9.6885\ \text{keV}$, so that the energy transfer was scanned around both the B K -edge at $194\ \text{eV}$ and the O K -edge at $535\ \text{eV}$. The total energy resolution of the setup was $0.7\text{--}0.8\ \text{eV}$ as measured by fitting a Gaussian profile to the quasi-elastic scattering peak. All the measurements were made at an average scattering angle of 30° in 2θ . Under these conditions all 9 analyzers are in a low exchanged momentum q configuration ($q \sim 2.54\ \text{\AA}^{-1}$) and the corresponding spectra can be summed. An accumulation time of at least 2 h was

necessary to achieve good statistics. A background subtraction has been achieved and each spectrum has been normalized to the area. The normalization range was taken between 185 and 218 eV for the B K-edge and between 530 and 555 eV for the O K-edge.

ii. GALAXIES Beamline - SOLEIL

The measurements of ν -B₂O₃ and LB9 glasses were performed at the GALAXIES beamline at the SOLEIL Synchrotron.^{36,37} The incident energy was monochromatized with a Si(111) double – crystal monochromator and the beam focused to ~ 100 (H) x 30 (V) μm^2 . The scattered radiation was analyzed with 4 spherical Si(660) analyzers (R = 1m) and detected with a silicon drift diode (SDD). The spectra were collected by keeping the analyzers fixed at their quasi-backscattering energy of 9.720 keV (Bragg angle $\sim 86^\circ$) while scanning the incident X-ray energy. For example, the O K-edge at 535 eV was measured by varying the incident X-ray energy around 10.255 keV. The total energy resolution of the setup was ~ 1.5 eV. All the measurements were made at an average scattering angle of 40° in 2θ . As in the previous case, under these conditions all 4 analyzers are in a low transfer momentum q configuration ($q \sim 3.4 +0.4/-0.6 \text{ \AA}^{-1}$) and the corresponding spectra can be summed. An accumulation time of at least 5h was necessary to achieve good statistics. After background subtraction each spectrum has been normalized to the area. The normalization range was taken between 185 and 218 eV for the B K-edge and between 530 and 555 eV for the O K-edge.

3. Results

The five glassy samples – accompanied with the L7B3 composition published elsewhere⁷, span an extended concentration range of Li₂O, starting from pure vitreous ν -B₂O₃ up to 70 mol.% of Li₂O in the case of L7B3. All the glasses were measured by XRS from room temperature up to 800-900°C depending on the Li₂O content - the different temperature points are shown in Fig. 1. The highest temperature achieved for each composition was decided on the basis of lithium volatilization, which leads to a rapid reduction of the sphere diameter. All the measurements showed here were performed at a constant sphere diameter.

For each temperature, both the B and O K-edges XRS spectra were measured almost simultaneously. (Fig. 2) As previously described,^{4,7,38-40} the B K-edge spectrum, which consists of an intense π^* peak at 194 eV and a broader σ^* peak located around 203 eV, is mainly informative of the nature of the boron coordination polyhedron. On the other hand, the O K-edge XRS spectrum, represented by a narrow peak at 536 eV and a broader one around 545 eV, is mainly indicative of the connectivity between the boron polyhedra.

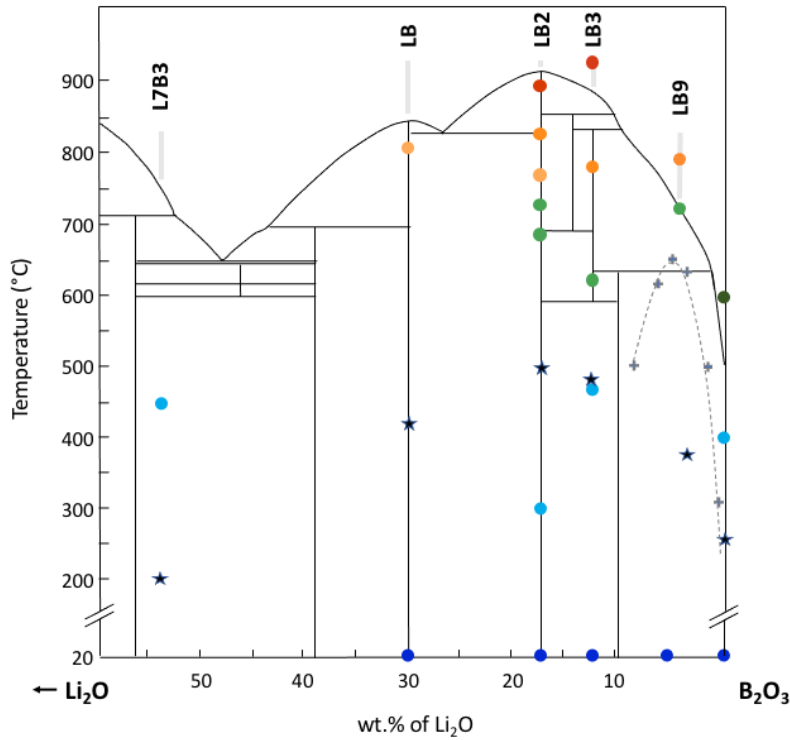


Fig. 1: Phase diagram of the $\text{Li}_2\text{O}:\text{B}_2\text{O}_3$ binary system. The colored filled circles are indicative of the measured temperatures for each glass composition: dark blue (room temperature), cyan (300-500°C), green (600-700°C), orange (~ 800°C) and red (~ 900°C). The L7B3 composition at 450°C was measured previously.⁷ The black stars are indicative of the glass transition temperatures.³² The dashed grey line shows the subliquidus miscibility dome described by Shaw and Uhlmann⁴¹

3.1 Temperature dependence of the B K-edge

Except for $\nu\text{-B}_2\text{O}_3$ and LB9, all the studied glasses show significant spectral transformations of the B K-edge as a function of temperature. (Fig. 2) When the temperature rises, the main modifications reside in the general increase in intensity of the π^* peak and the disappearance of the peak around 200 eV, which are both indicators of a temperature induced $^{[4]}\text{B}$ to $^{[3]}\text{B}$ conversion. This conversion takes place gradually above T_g to finally reach some kind of *equilibrium* around 800-900°C. Slight fluctuations of the intensity of the π^* peak caused by subtraction of the Compton scattering background are not significant enough to affect the conclusions. $\nu\text{-B}_2\text{O}_3$ and LB9 do not show marked modifications by increasing temperature: only a slight shift of the π^* peak accompanied by a slight reduction of the intensity is observed with the temperature increase. A similar shift (around 0.1-0.2eV) is also observed for the metaborate composition (LB). These subtle modifications are caused by thermal fluctuations similarly to what has been experimentally observed at the Al K-edge of corundum by X-ray absorption near-edge spectroscopy.⁴²

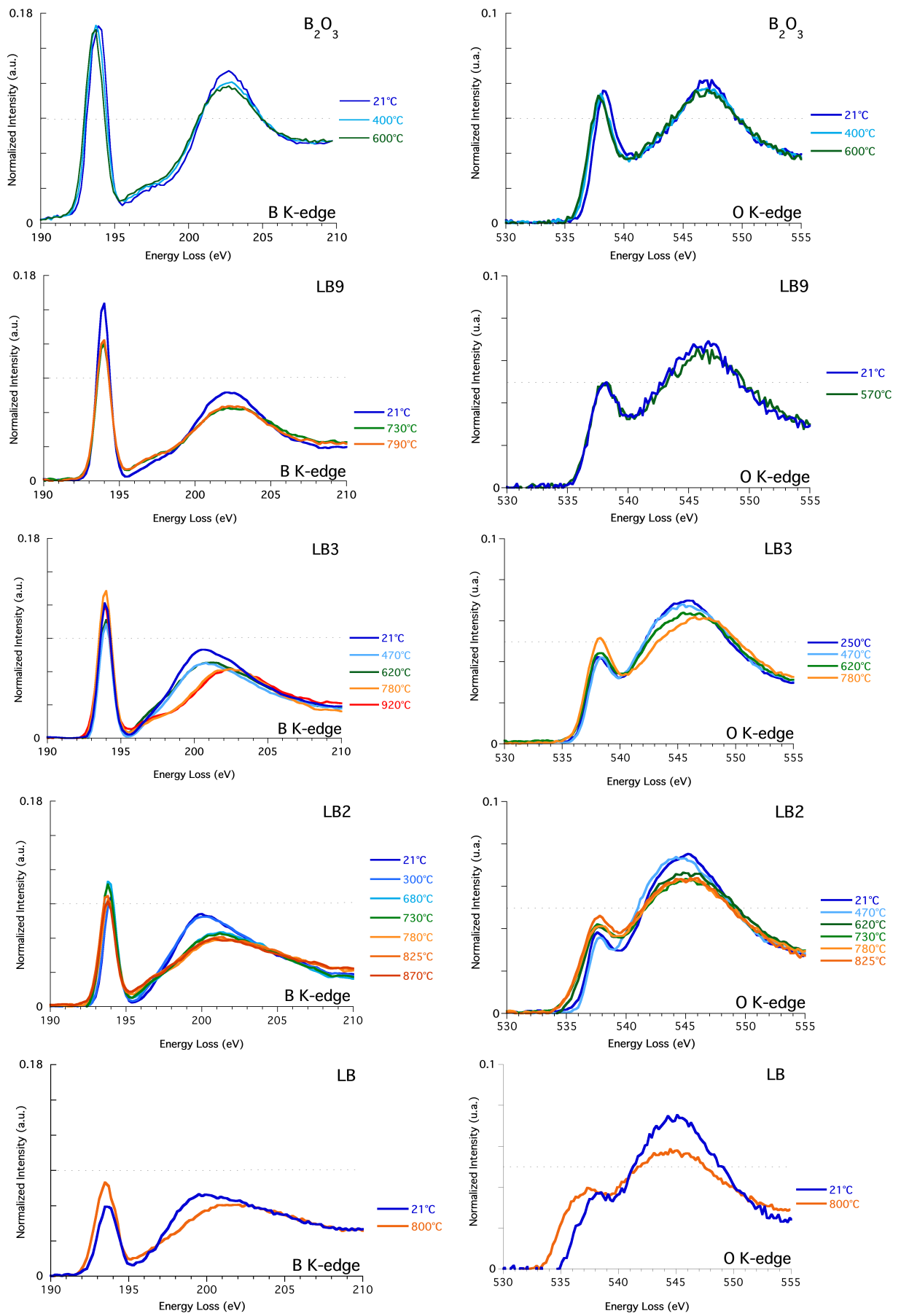


Fig. 2: *in situ* B K-edge (left) and O K-edge (right) XRS spectra of v - B_2O_3 , LB9, LB3, LB2 and LB glasses as a function of temperature. All the spectra were measured in the levitation chamber with the exception of the B K-edge of LB9 at room temperature, which has been measured in a standard sample holder.

3.2 Temperature dependence of the O K-edge

With the exception of a global broadening of the peaks related to an increase of the topological disorder when temperature is increased, the O K-edge XRS spectra do not show significant changes at low Li₂O content, *i.e.* for v-B₂O₃ and LB9 glasses. (Fig. 2) This is in line with the observations for the B K-edge spectra. The O K-edge spectra of LB3 are instead temperature dependent, as shown by the shift of the σ^* peak at 543eV and the intensity increase of the π^* peak around 537eV. From LB3 to LB, the spectral modifications are more and more pronounced, indicative of an increasing ^[4]B to ^[3]B conversion.³⁸ Additionally, an extra peak arises at low energy around 535eV, whose intensity increases with Li₂O content. This spectral feature is related to the formation of NBOs in the melt as a consequence of the ^[4]B to ^[3]B conversion.^{7,38}

4. Discussion

4.1 Temperature-induced structural modifications on the borate network

Depending on the stoichiometry, both the short- and medium range orders of alkali borate glasses may be affected by temperature. The monitoring of the proportion of three-fold coordinated boron atoms using XRS is made possible since the π^* peak of the B K-edge is solely related to triangular BO₃ units. By using a data treatment procedure previously described,⁷ it is possible to follow the proportion of ^[3]B as a function of temperature in order to quantify the temperature-induced structural modifications (Fig. 3)¹⁵ With the exception of LB9 and L7B3, all the glasses show a significant temperature induced ^[4]B to ^[3]B conversion. The proportion of converted boron atoms is about 15-20% in the temperature range under study, which is in line with Raman or NMR experimental data on lithium borates.^{15,16,43,44} The average boron coordination number $n_{BO}(x,T)$ is temperature dependent for intermediate compositions, *i.e.* LB3, LB2 and LB. In the case of LB2, $n_{BO}(x,T)$ decreases from 3.5 at RT to about 3.3-3.4 around 800°C, *i.e.* $\frac{\Delta n_{BO}}{\Delta T} \approx -3 \cdot 10^{-4} K^{-1}$, which agrees with the predictions of thermodynamic models and with the very few neutron diffraction experiments available at elevated temperature.⁴⁵ This decrease of the average B-O coordination number from the glass to the melt takes place only above the glass transition temperature. Below T_g, nothing really happens at the local scale, *i.e.* some thermal relaxations of defects/constraints on the medium range order may happen without significantly modifying the local boron environment.

Following the conversion reaction $BO_4 \Leftrightarrow BO_3 + NBO$, each BO₄ unit that is converted into a planar BO₃ unit involves the formation of at least one NBO in order to maintain charge neutrality of the system. Even if the exact quantification of NBOs in borate using XRS is difficult, the intensity of the peak at 535eV in the O K-edge spectrum is a good indicator of the concentration of NBOs in the glass.³⁸ By plotting the intensity of the peak at 535 eV versus temperature for different Li₂O concentrations (Fig. 3), we clearly show the correlation between the ^[4]B to ^[3]B conversion and the formation of NBOs, and that the formation of NBOs is enhanced when the Li₂O content is increased. NBOs start to appear around 25-30mol% Li₂O and keep increasing up to 50mol%. Between 25 and

50mol% Li_2O , the glass structure is mainly composed by triborate and pentaborate groups, which are gradually replaced by the di-triborate and eventually by the diborate superstructural units both characterized by the presence of a $^{[4]}\text{B}-\text{O}-^{[4]}\text{B}$ bound unit which is energetically less favourable.^{2,7,44} These diborate units seem to be easily breakable upon heating moving on to the formation of $\text{B}\text{O}_2\text{O}^-$ and/or $\text{B}\text{O}\text{O}_2^{2-}$ units in the melt, as suggested by MD simulations of Varsamis et al,⁴⁶ where O refers to a bridging oxygen.

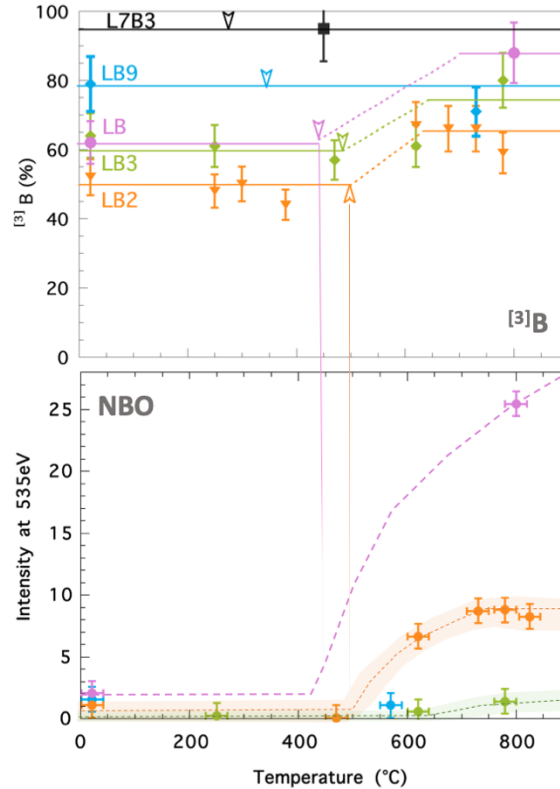


Fig. 3: Evolution of the proportion of $^{[3]}\text{B}$ (top) and the intensity at 535eV (bottom) as a function of temperature and Li_2O content: LB9 (cyan), LB3 (green), LB2 (orange), LB (pink) and L7B3 from reference⁷ (black). The plain and dashed colored lines are a guide for the eyes, and the arrows are indicative of the glass transition temperatures of the different glasses.¹³

4.2 Melt structure: existence of two distinct melt structures

The present study gives access to the structure of the melt for a wide range of Li_2O concentrations. By plotting the XRS spectra of both B and O K-edges for the five compositions at about 800°C, it is clear that the melt structure depends on the stoichiometry of the glass (Fig. 4). Two sets of melts compositions can be distinguished:

- a well connected borate network is observed in the melts of $\nu\text{-B}_2\text{O}_3$, LB9 and LB3, characterized by a strong and sharp π^* peak at the B K-edge and an almost zero intensity at 535 eV at the O K-edge spectra.
- a more and more depolymerized borate network beyond the LB3 composition with the appearance of numerous NBOs in the melt, as revealed by the intensity increase at 535 eV in the O K-edge spectra when adding Li_2O .

These observations indicate that the melt structure goes from a highly polymerized borate network to a more open and depolymerized structure with an increasing Li_2O content. The transition around 25 mol% Li_2O between these two melt structures has to be put in parallel with the persistence of the boroxol rings in the melt up to 20-25 mol%, as observed by Raman spectroscopy.^{2,16,17,47,48} The presence of boroxol rings in the liquid state would help to stabilize the connectivity of the borate framework for low-alkali glasses. These two melt structures (polymerized vs. depolymerized) could be at the origin of the subliquidus miscibility dome suggested by Shaw and Uhlmann⁴¹ and calculated by Kim and Sanders⁴⁹ which extends approximately from 2-18 mol% in the case of $\text{Li}_2\text{O}-\text{B}_2\text{O}_3$. For a better insight, the subliquidus miscibility gap has been reproduced in figure 1. The existence of these two melt structures would explain (i) the non-existence of crystalline phases in this concentration range, and (ii) the coexistence of two sub-domains in the glass with an “alkali-rich” domain and a pure B_2O_3 domain containing the boroxol rings.⁷

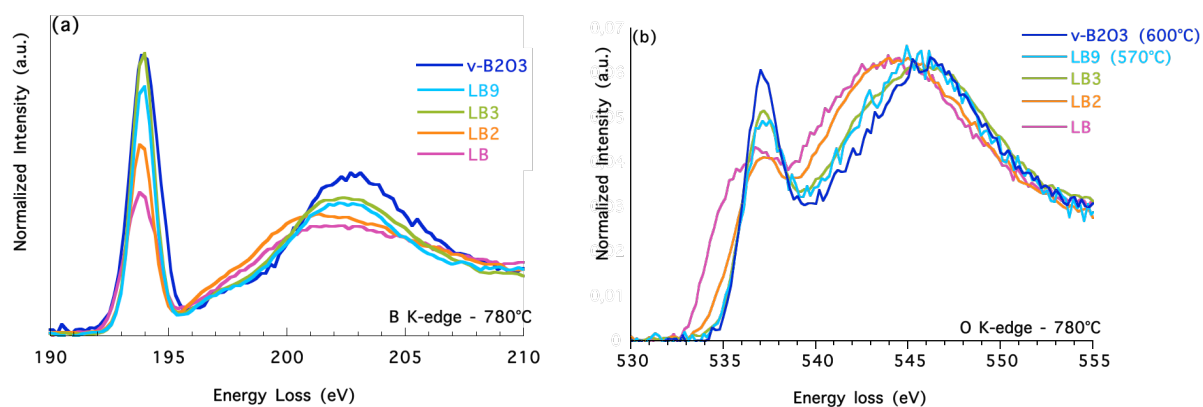


Fig. 4: B K-edge (Left) and O K-edge (Right) XRS spectra of LB, LB2, and LB3 glasses measured at 780°C, LB9 at 570°C and v- B_2O_3 at 600°C.

4.3 Medium-range order (MRO) in the melt

Alkali borate glasses are known to show medium range order, which can be classified into well-defined structural motifs known as superstructural units. In order to acquire information about a possible MRO organization in the melt, four crystals with well-defined superstructural units were used as reference samples: c- $\text{Li}_2\text{B}_4\text{O}_7$ for diborate units ($\text{B}_4\text{O}_5\text{O}_4^{2-}$), c- LiBO_2 for metaborate chains (BO_2^-)_∞, c- $\text{Li}_6\text{B}_4\text{O}_9$ for metaborate chains of four-members ($\text{B}_4\text{O}_6\text{O}_3^{6-}$) and c- Li_3BO_3 for orthoborate units (BO_3^{3-}).² The O K-edge XRS spectrum of LB2 at 800°C accompanied with those of c- $\text{Li}_2\text{B}_4\text{O}_7$ and c- LiBO_2 are shown in Figure 5. Even if the thermal vibrations broaden the spectral features, it is clear that the O K-edge of the LB2 melt shows similarities with c- $\text{Li}_2\text{B}_4\text{O}_7$, particularly when the first 5 eV of the edge are considered. The main difference resides in the formation of a few percent of NBOs which gives an additional contribution at 535 eV and their appearance coincides with the disappearance of some of the BO_4 units. This means that the polymerized borate framework is partly preserved in the LB2 melt. On the contrary, the O K-edge of LB melt is intermediate between c- $\text{Li}_6\text{B}_4\text{O}_9$ and c- Li_3BO_3 , (Fig. 5) suggesting that the polymerized structure of the LB glass disappears almost completely at high-temperature to finally be constituted by a mixture of metaborate chains of four BO_3 members ($\text{B}_4\text{O}_6\text{O}_3^{6-}$), orthoborate units (BO_3^{3-}) and probably pyroborate units ($\text{B}_2\text{O}_4^{4-}$).

Consequently, the structure of LB melt has to show NBO's rich domains, in which lithium ions are free to move. These two examples highlight the existence of two distinct melt structures up to the diborate composition, *i.e.* 33 mol%. Figure 6 schematically summarizes the main temperature induced structural modifications in the $\text{Li}_2\text{O}-\text{B}_2\text{O}_3$ system going from the glass to the melt.

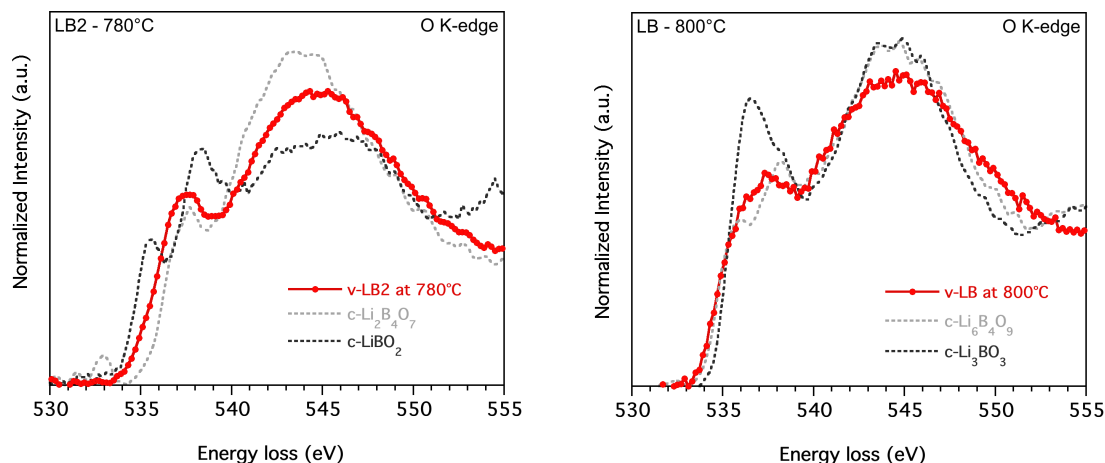


Fig. 5: Left: O K-edge XRS spectra of LB2 melt (red points) accompanied with two crystalline phases: $\text{Li}_2\text{B}_4\text{O}_7$ (grey) and LiBO_2 (black). Right: O K-edge XRS spectra of LB melt (red points) accompanied with two crystalline phases: $\text{Li}_6\text{B}_4\text{O}_9$ (grey) and Li_3BO_3 (black).

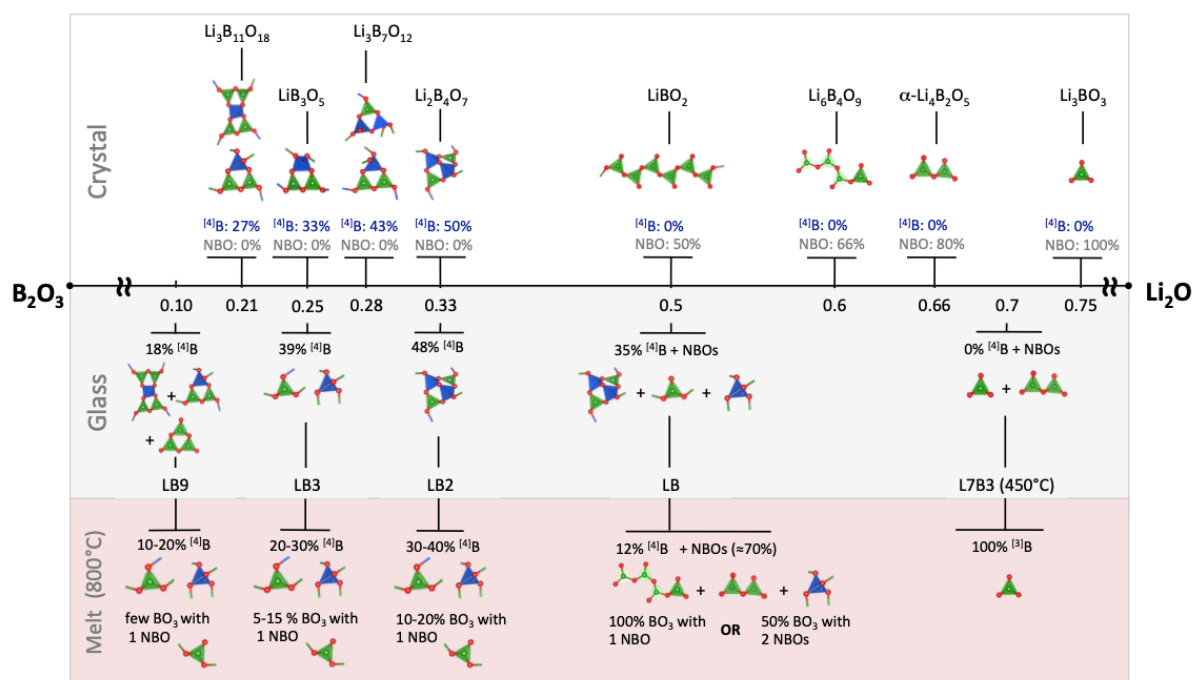


Fig. 6: Summary of the superstructural units found in crystalline phases (top), in glasses at room temperature (grey-shaded area) and in the melt (red-shaded area) for the $\text{Li}_2\text{O}-\text{B}_2\text{O}_3$ system.

4.4 Melt structure vs density at 800°C

Density is related to the atomic spatial distribution, *i.e.* the stacking of atoms, which will depend on the short-range order, but also on the medium-range organization via the composition in superstructural units. By plotting the structural characteristics of the melt at 800°C, *i.e.* the intensity of the peak at 535eV relative to the presence of NBOs and the proportion of ^{13}B , and the variation of the

density at 800°C⁵⁰ as a function of Li₂O content, (Fig. 7) three domains emerge from the composition plot :

Domain 1 (0-30 mol% Li₂O): a gradual ^[3]B → ^[4]B conversion is observed in the melt when lithium oxide is added. With the introduction of Li⁺ cations, the electroneutrality is fulfilled by the formation of negatively charged (BO₄)⁻ tetrahedral units - no or very few NBOs are observed in the 0 – 30 mol% Li₂O concentration range. The borate network is fully polymerized on the overall concentration range, but the 3D connectivity increases with the formation of the BO₄ units and the gradual disassembling of the boroxol rings into triborate and then ditriborate units.¹⁶ As a consequence, the density of the melt increases.

Domain 2 (30-45 mol% Li₂O): between 30 and 45mol%, there is a competition between the ^[3]B → ^[4]B conversion and the appearance of NBOs. Both processes have opposite effects on the melt density: the ^[3]B → ^[4]B conversion when Li₂O is added tends towards a densification of the melt whereas the formation of NBOs leads to a depolymerization of the borate network and consequently to an increase of the melt volume. A change of slop of density at 800°C as a function of concentration is observed in this concentration domain, where the vitrification enthalpy reaches its maximum value.⁵¹

Domain 3 (45-70 mol% Li₂O): from 45 up to 70mol%, the NBO content increases drastically supplanting the ^[3]B → ^[4]B conversion. A progressive depolymerization of the borate network is observed leading to a decrease of the melt density.

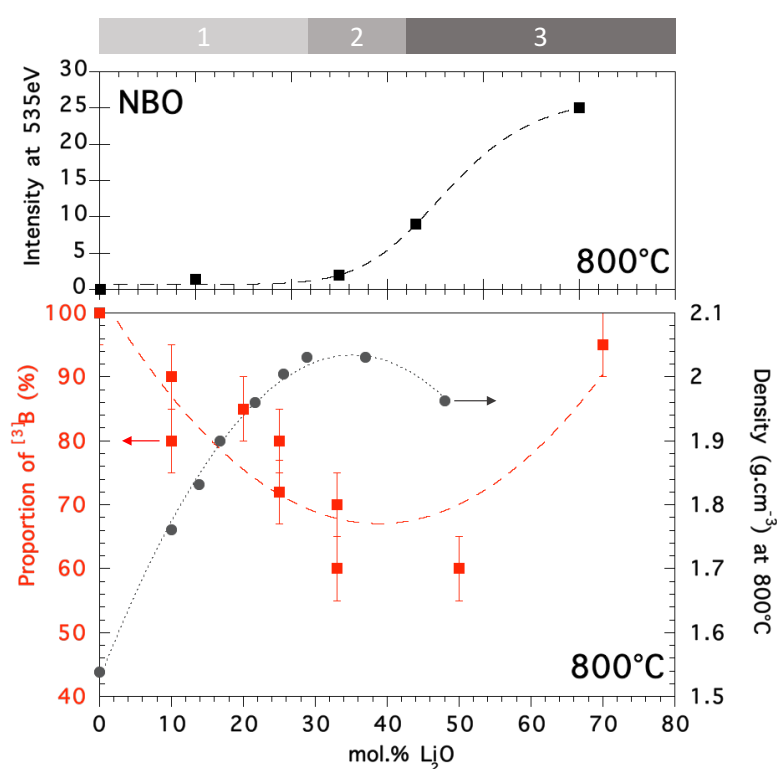


Fig. 7: Evolution of the intensity of the peak at 535 eV related with NBOs (black colored squares), the proportion of ^{13}B (red colored squares) and the density⁵⁰ at 800°C (grey colored circles) as a function of Li_2O content. The dashed lines are a guide for the eyes.

4.5 Melt structure vs Li diffusion

The structural description of the melt in the section 4.4 is fully corroborated by the constant decrease of the activation energy of lithium diffusion above 20-30 mol% observed by high-temperature pulsed field gradient NMR.⁵² This decrease is a direct consequence of the reduction of dimensionality of the borate framework with the appearance of NBOs in the melts for Li_2O content above 45 mol%.⁵³ Both the diffusion coefficient of the lithium ions⁵² and the intensity of the NBO's peak at 535eV measured at 800°C are fully correlated, as shown on the Figure 8. The disappearance of BO_4 units and the formation of NBOs are correlated with a change in the structural role of Li. In the presence of BO_4 , Li^+ cations play a charge compensating role around the negatively charged BO_4 tetrahedra while Li^+ cations are modified when associated with NBOs. This change in the Li structural role results in modifications of the local Li environment, as observed by neutron diffraction.¹¹

The three structural domains described above are in fair agreement with the dynamical behavior of Li cations, *i.e.* a relatively slow translational diffusion up to 30 mol% corresponding to almost trapped ions in borate cages, followed by a steadily increase of the translational properties above 30 mol%. The role of NBOs in the mobility of Li cations is then fully pointed out, and is indicative of the presence of NBO rich-domains, whose size increases from 30 mol% up to 50 mol%, and where Li ions are free to move. This behavior seems to remain upon quenching, since the ionic conductivity of the glass is enhanced above 30-40 mol%.⁵⁴

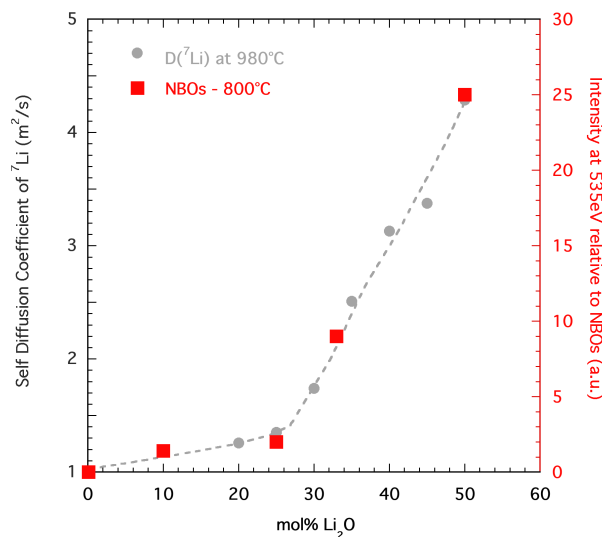


Fig. 8: Comparison between the intensity of the peak at 535eV at 800°C, which is indicative of the proportion of NBOs (red squares) and the self-diffusion coefficient of ^7Li measured by NMR⁵² at 975°C (grey circles) as a function of Li_2O content.

5. Conclusions

This exhaustive study on the temperature behavior of lithium borate glasses and melts allowed us to monitor simultaneously the coordination number of boron atoms and the BO/NBO ratio as a function of both temperature and composition. Low lithium content glasses (v-B₂O₃ and LB9) show almost no modification of their local structure with heating, probably because of the reminiscence of boroxol rings that help to stabilize the melt structure. A temperature induced ⁴B → ³B conversion is observed above T_g from the glass to the melt from the triborate composition up to the metaborate composition. The non-monotonic evolution of the melt density of the melt at 800°C can be easily explained in terms of ⁴B → ³B conversion and formation of NBOs.

Two distinct melt structures were evidenced: a well polymerized borate network – with few NBOs - below the triborate composition, and a depolymerized borate network above the diborate composition with a rapid increase of the number of NBOs when Li₂O is added. These two structurally distinct melts allow for explaining the two dynamical regimes observed for lithium ion diffusion,⁵² and could be at the origin of the increasing fragility as the alkali oxide content increases. This study highlights also the major role of oxygen, and especially NBOs in the structural properties of the melt and its consequence on the diffusivity of atoms.

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Supporting Information

XRS data treatment is fully described in the electronic supporting information for both B and O K-edges.

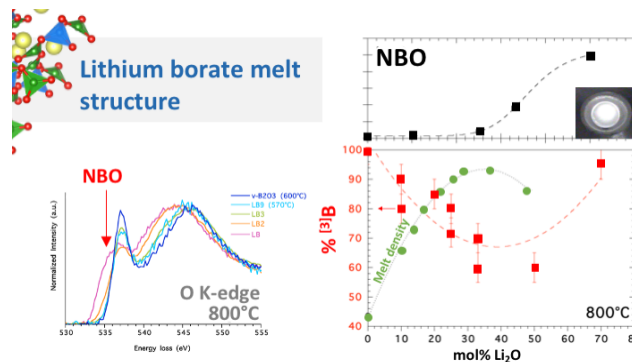
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High-temperature X-ray Raman Spectroscopy has been used to determine the melt structure of five compositions of the $\text{Li}_2\text{O}-\text{B}_2\text{O}_3$ system. A temperature induced ^{4}B to ^{3}B conversion is observed above the glass transition temperature, and two distinct melt structures were evidenced for compositions below and above the diborate composition. These two structurally distinct melts allow explaining the two dynamical regimes observed for lithium ion diffusion.