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Design of effective catalysts based on ZnLaZrSi oxide systems for obtaining 1,3-butadiene from aqueous ethanol

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KEYWORDS ethanol-water mixture, 1,3-butadiene, ZnLaZrSi oxide system, silica support.

ABSTRACT

ZnLaZrSi oxide systems prepared with a silica component of the different nature have been studied in 1,3-butadiene production from aqueous ethanol. The following silica materials were used: KSKG, A-175, A-380, SBA-15, MCM-41, MCM-48, MCF and dealuminated BEA zeolites. The characteristics of the porous structure of the silica support, such as porosity, pore size distribution, specific and external surface areas, were found not to be critical parameters for achieving a high 1,3-butadiene yield during the EtOH-H₂O mixture conversion in the presence of ZnLaZrSi oxide catalysts. On the contrary, the quantity and strength of Lewis acid sites, which in turn differ depending on the choice of silica material, have a significant impact on 1,3-butadiene selectivity and yield. The highest values of the selectivity of 1,3-butadiene formation (up to 68 %) and yield as well as stability toward deactivation in the presence of H₂O were achieved over ZnLaZr-KSKG, ZnLaZr-SBA-15 and ZnLa-Zr₁SiBEA (with mononuclear isolated tetrahedral Zr(IV) species). The productivity of ZnLa-Zr₁SiBEA catalyst accounts for 0.324 g_{1,3-BD}·g_{cat}⁻¹·h⁻¹ (T = 648K, WHSV = 2.88 h⁻¹, 80 vol% EtOH in water as an EtOH source). The main reason for the decrease in 1,3-butadiene yield in the presence of H₂O in the reaction mixture was shown to be a deactivation of acetaldehyde condensation sites on the catalyst surface, while the rate of acetaldehyde formation decreases slightly. According to ¹H-¹³C CP/MAS NMR spectroscopic results, the use of aqueous ethanol as the feed for ETB-process is very advantage to prevent the carburization of the catalysts.

INTRODUCTION

Large investments in the development of ethanol production technology from renewable raw materials contribute to a significant reduction in the costs for manufacture, and accordingly make it quite feasible to partially replace industrially important processes based on petroleum with eco-friendlier processes based on bioethanol in the near future.¹⁻³ A number of publications⁴⁻⁶ are also provided justification (primarily from an environmental point of view) the starting the production of such an industrially important substrate as 1,3-butadiene (1,3-BD) from ethanol (EtOH). In particular, the project for bio-butadiene production has been initiated in 2013.⁷ Development of active and highly selective catalysts of the ethanol-to-butadiene (ETB) technology is one of the main tasks for the process implementation in industry.⁸

A modern approach to the process implementation of both 1,3-BD and ethylene production from ethanol involves the use of cheaper ethanol-water mixtures as raw materials that can be obtained by flash separation of fermentation ethanol, instead of expensive rectified or 100 vol% ethanol. For instance, the 50 vol% ethanol can be obtained by pervaporation of fermentation broth through a membrane. Micheline company has patented the technology of the 1,3-BD production from 80 vol% ethanol.⁹ Moreover, the presence of over 5-10 % water in the feed has been shown to significantly affect activity and selectivity of the ETB process catalysts.¹⁵⁻¹⁹ Thus, development of the catalysts able to convert ethanol-aqueous mixtures containing 50-80 vol% ethanol into 1,3-BD is a very important task which is insufficiently studied now.

Based on the studies of various catalytic systems presented in the literature, activity and selectivity of catalyst in the ETB process can be concluded to equally depend on both catalyst composition and its preparation method. Most of the known catalysts contain silicon dioxide. It has been suggested that active sites of the key step of ETB process, aldol condensation, are acid-

base pairs formed in contact of catalytically active metal cation and silicon oxide, usually depicted as M-O-Si.^{8,20}

Effect of the used silica materials (along with catalyst composition) on activity of the ETB process catalysts is also shown in some papers. Using $\text{ZrO}_2\text{-ZnO/SiO}_2$ and $\text{ZrO}_2\text{-ZnO-La}_2\text{O}_3\text{/SiO}_2$ systems, Jones et al.^{13,21} have shown that a pore size of SiO_2 used for the catalysts preparation affects their activity in the conversion of ethanol and ethanol-acetaldehyde (EtOH-AA) mixture into 1,3-BD. Sushkevich et al.²² and Dagle et al.²³ have observed a relationship between butadiene selectivity and concentration of Lewis acid sites for Ag/ ZrO_2 -silica systems that depends on the nature of the used silica component. Pomalaza et al.^{24,25} have recently shown that ZnTa oxide catalyst based on mesoporous silica TUD-1 is more active in ETB process than similar system based on SiO_2 or dealuminated BEA zeolite. Chae et al.²⁶ have studied the effect of ordered mesoporous silica (OMS) on activity of $\text{Ta}_2\text{O}_5\text{/OMS}$ in 1,3-BD production from EtOH-AA mixture. The pore size and particle size of OMS samples are shown to be more important than mesopore structure such as pore size and pore shape to optimize catalytic performance of Ta_2O_5 -containing catalysts. Gao et al.²⁷ have come to a similar conclusion during studying $\text{ZrO}_2\text{-SiO}_2$ systems prepared using various silica materials in EtOH-AA mixture conversion into 1,3-BD. Cheong et al.²⁸ have shown that high-performance ZrO_2 -containing catalysts of 1,3-BD production from EtOH-AA mixture can be obtained using mesocellular siliceous foam (MCF) as a silica component. The use of zeolite as a matrix for the formation of M-O-Si active sites also allows preparing highly selective catalysts for the conversion of both EtOH and EtOH-AA mixture.²⁹⁻³⁵

In view of the foregoing considerations, the type (nature) of a silica component is assumed to have a decisive influence on activity of silica-based catalysts in 1,3-BD production from EtOH-

H₂O mixture as raw materials. As previously shown, catalytic systems containing zinc, lanthanum, zirconium and siliceous oxides exhibit quite high activity in the conversion of both ethanol and ethanol-water mixtures into 1,3-BD.^{12,36,37} Thus, we have used ZnLaZrSi oxide systems to study the effect of the silica component type on their catalytic properties in 1,3-BD production from EtOH-H₂O mixture. Since there is no univocal opinion in the literature concerning the best silica support, a series of the samples based on commercially available silica (KSKG, A-175, A-380), ordered mesoporous silica (SBA-15, MCM-41, MCM-48), mesocellular siliceous foam (MCF) and dealuminated BEA zeolites was investigated.

MATERIALS AND METHODS

Catalyst preparation

KSKG industrial silica gel, aerosils A-175 and A-380, ordered mesoporous silica SBA-15, MCM-41 and MCM-48, mesocellular siliceous foam (MCF) and dealuminated BEA zeolites were used for the preparation of ZnLaZrSi oxide samples. KSKG industrial silica gel (silica gel granular with large pores, specific surface area is 283 m² g⁻¹) was pre-treated with dilute nitric acid, washed with double distilled water and calcined at 773 K. Dealumination of a parent tetraethylammonium BEA zeolite (Si/Al = 17) was carried out by its treatment in a 13 mol·L⁻¹ HNO₃ solution at 353 K for 4 h. The resulting dealuminated BEA zeolite with Si/Al ratio of 1000 was marked as SiBEA.

Introduction of zirconium (6 wt % relative to ZrO₂) into the silica samples (namely pre-treated KSKG, Aerosils A-175 and A-380, ordered mesoporous silica SBA-15, MCM-41, MCM-48, mesocellular siliceous foam and SiBEA) was carried out by wet kneading (fraction < 0.1 mm) with analytical grade ZrO(NO₃)₂·2 H₂O, followed by drying (393 K, 2 h) and calcination (773 K,

3 h). Introduction of zirconium into SiBEA was also carried out by stirring for 3 h at 353 K in 200 mL of isopropanol solution containing $1 \cdot 10^{-3} \text{ mol} \cdot \text{L}^{-1}$ of $\text{Zr}(\text{NO}_3)_4 \cdot 4 \text{ H}_2\text{O}$ (Sigma Aldrich, 99.9 %). The obtained suspensions (pH = 6.8) were stirred in an evaporator under vacuum of a water pump in air at 353 K for 1 h until isopropanol was evaporated. The resulting solids, washed three times in distilled water and dried in air at 353 K for 24 h, were finally calcined at 773 K in flowing air for 3 h. The samples contained 1 and 3 wt % of Zr were labelled as Zr_1SiBEA and Zr_3SiBEA .

Then, all Zr-Si oxide samples were sequentially impregnated to incipient wetness with aqueous solutions of chemically-pure-grade $\text{La}(\text{NO}_3)_3 \cdot 6\text{H}_2\text{O}$ and $\text{Zn}(\text{O}_2\text{CCH}_3)_2$ with intermediate drying at 393 K (2 h) and calcination at 773 K (3 h). The 2 wt % ZnO–3 wt % La-containing samples were prepared and marked as ZnLaZr–KSKG (A-175, A-380, SBA-15, MCM-41, MCM-48, MCF) and ZnLa– Zr_1SiBEA and ZnLa– Zr_3SiBEA .

Catalyst characterization

X-ray diffraction (XRD) patterns of the powder samples were recorded using a D8 Advance (Bruker AXS GmbH, Germany) diffractometer with monochromatized Cu $K\alpha$ radiation (nickel filter, $\lambda = 0.15418 \text{ nm}$).

Nuclear magnetic resonance (NMR) spectra were recorded using a Varian VNMRS 600 MHz spectrometer (14.1 Tesla 51mm bore Oxford superconducting magnet) operating at a ^{29}Si Larmor frequency of 119.229 MHz and using a 3.2 mm magic angle spinning (MAS) probe head. ^{29}Si MAS NMR spectra were recorded with cross-polarization (^1H – ^{29}Si CP/MAS NMR) and without it (^{29}Si MAS NMR). Chemical shifts of silicon were measured relative to tetramethylsilane (TMS). ^{29}Si MAS NMR spectra were obtained with rotors spinning with a frequency of 20 kHz,

using a single pulse acquisition and a recycle delay of 4 s. All measurements were carried out at room temperature.

The morphology of the samples was observed by means of scanning electron microscopy (SEM) using a Tescan MIRA 3 microscope (electron beam energy of 20 keV).

The nitrogen adsorption-desorption isotherms were measured at 77 K with the porous analyzer Sorptomatic-1990. Before measurements, the samples were evacuated at 473 K for 2 h. The specific surface area of the samples was determined by the Brunauer-Emmett-Teller (BET) method. External surface area and micropore volume of the investigated samples were estimated from *t*-plot method. The mesopore size of the samples were obtained by the Barrett-Joyner-Halenda (BJH) method.

Analysis of acid properties of the catalyst samples was performed by adsorption of pyridine followed by Fourier Transform Infrared (FTIR) spectroscopy using a Spectrum One FTIR spectrometer (Perkin Elmer, USA). Typically, 24 scans with a resolution of 1 cm^{-1} were coded to give one spectrum. The catalyst sample was pressed at $\sim 2\text{ ton} \cdot \text{cm}^{-2}$ into thin wafers of ca. $12\text{ mg} \cdot \text{cm}^{-2}$ and placed inside the IR cell. Before adsorption of pyridine, the cell was connected to a vacuum-adsorption apparatus allowing a residual pressure below 10^{-3} Pa . The samples were outgassed (10^{-3} Pa) at 693 K for 1 h. Then, pyridine was adsorbed at 423 K for 15 min. The spectra were recorded under ambient conditions after pyridine desorption at 423, 523 and 623 K for 30 min. All measured spectra were recalculated to a “normalized” wafer weight and surface.

Catalytic activity measurements

Catalytic activity tests were carried out in a fixed-bed flow quartz reactor with an inner diameter of 4 mm at 648–673 K and atmospheric pressure. Samples with grains of 0.25–0.5 mm were loaded into the reactor. Aqueous ethanol feed (96–50 vol% EtOH – 4–50 vol% H₂O) was

introduced to an evaporator via a syringe infusion pump, and argon with flow of $7 \text{ mL} \cdot \text{min}^{-1}$ was carried the feed into the reactor. Reaction was carried out at weight hourly space velocity (WHSV) by EtOH of 0.4–1.2 $\text{g}_{\text{EtOH}} \cdot \text{g}_{\text{cat}}^{-1} \cdot \text{h}^{-1}$. Before the reaction, the catalysts were annealed at 773 K under flowing argon for 1 h. The reagents and reaction products were analyzed on a gas chromatograph (KristalLyuks 4000M, MetaChrome) equipped with a TCD detector and a packed column (10 % NiSO₄ on coal, 3 m × 3 mm) for CO and CO₂, and a FID detector and a capillary column (HP-FFAP, 50 m × 0.32 mm) for organic compounds.

Catalytic activity of the catalysts was characterized by EtOH conversion (X , %), selectivity to products (S_i , %), 1,3-BD yield ($Y_{1,3\text{-BD}}$, %) and 1,3-BD productivity ($P_{1,3\text{-BD}}$, $\text{g}_{1,3\text{-BD}} \cdot \text{g}_{\text{cat}}^{-1} \cdot \text{h}^{-1}$):

$$X = \frac{n_{\text{EtOH}}^0 - n_{\text{EtOH}}}{n_{\text{EtOH}}^0} \cdot 100\%,$$

$$S_i = \frac{n_i}{(n_{\text{EtOH}}^0 - n_{\text{EtOH}})} \cdot 100\%,$$

$$Y_i = \frac{X \cdot S_i}{100},$$

$$P_{1,3\text{-BD}} = \frac{Y_{1,3\text{-BD}} \cdot \text{WHSV} \cdot 0.587}{100},$$

where n_{EtOH}^0 is the initial amount of C_1 moles of EtOH; n_{EtOH} and n_i are the amount of C_1 moles of unreacted EtOH and product i in the stream of the reaction products, respectively; 0.587 is the maximum possible amount of 1,3-BD (g) that can be produced from 1 g of EtOH from a stoichiometric equation: $2 \text{ C}_2\text{H}_5\text{OH} \rightarrow \text{C}_4\text{H}_6 + \text{H}_2 + 2 \text{ H}_2\text{O}$.

The formation rates of 1,3-butadiene, acetaldehyde (AA) and ethylene+diethyl ether (DEE) were calculated per unit mass of catalyst (r_m , $\text{mmol} \cdot \text{g}_{\text{cat}}^{-1} \cdot \text{h}^{-1}$), per unit surface area (r_s , $\mu\text{mol} \cdot \text{m}^2_{\text{cat}}^{-1} \cdot \text{h}^{-1}$) and per unit external surface area ($r_{\text{extern.}}$, $\mu\text{mol} \cdot \text{m}^2_{\text{cat}}^{-1} \cdot \text{h}^{-1}$):

$$r_{m_i} = \frac{X \cdot S_i \cdot \text{WHSV}}{M_{\text{EtOH}} \cdot 10},$$

$$r_{s_i} = \frac{1000 \cdot r_{m_i}}{S_{\text{BET}}},$$

$$r_{\text{extern},i} = \frac{1000 \cdot r_{m_i}}{S_{\text{ext.}}},$$

where M_{EtOH} is molar mass of ethanol ($\text{g} \cdot \text{mol}^{-1}$); S_{BET} is specific surface area of the samples determined by the Brunauer-Emmett-Teller method ($\text{m}^2 \cdot \text{g}^{-1}$); $S_{\text{ext.}}$ is external surface area of the samples ($\text{m}^2 \cdot \text{g}^{-1}$).

RESULTS AND DISCUSSION

Structural and textural characteristics of the catalysts

XRD patterns (Figure 1) of the prepared catalysts indicate amorphous structure of the supported silica materials typical of the initial silica mesoporous molecular sieves (MMS). The low-angle reflexes characteristic for the MMS supports are absent in the XRD patterns due to the filling of the support pores with the active phase resulting in an increase the number of atoms per a unit cell. The only low-intense reflection present in these XRD patterns can be attributed to monoclinic ZrO_2 phase. XRD patterns of the zeolite supported samples demonstrate highly crystalline structure characteristic of BEA zeolite with main well-defined reflections at $2\theta = 7.7^\circ$; 22.5° .

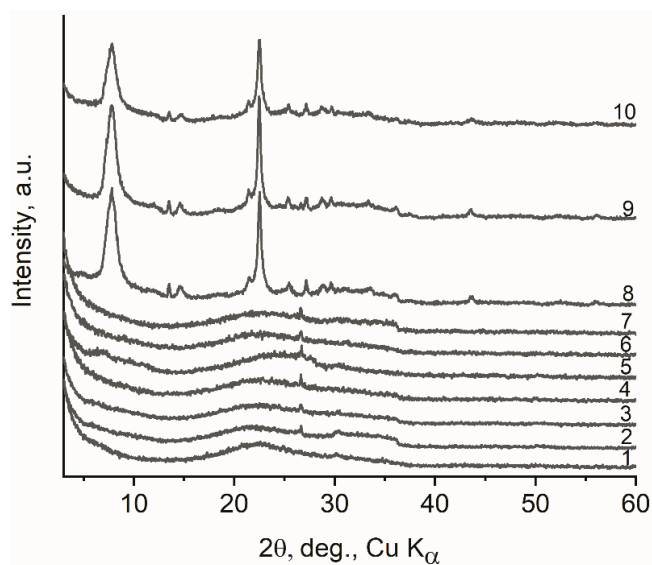


Figure 1. XRD patterns of the prepared catalysts: 1 – ZnLaZr–KSKG; 2 – ZnLaZr–A-175; 3 – ZnLaZr–A-380; 4 – ZnLaZr–SBA-15; 5 – ZnLaZr–MCM-41; 6 – ZnLaZr–MCM-48; 7 – ZnLaZr–MCF; 8 – ZnLaZr–SiBEA; 9 – ZnLa–Zr₁SiBEA; 10 – ZnLa–Zr₃SiBEA.

²⁹Si MAS NMR spectrum of ZnLa–Zr₃SiBEA zeolite presented in Figure S1 shows strong NMR signals in the range of –115 to –111 ppm. These signals correspond to Si atoms in Si(OSi)₄ species in non-equivalent T crystallographic sites (T-atoms) of BEA (9 sites)³⁸, which are not well resolved. At about –102 ppm there is a weak, hardly observable ²⁹Si NMR signal of (SiO)₃Si(OH) species. Its assignment is confirmed by the ¹H–²⁹Si CP/MAS NMR measurement (Figure S1 B), which significantly enhances the signal of ²⁹Si nuclei close to protons with respect to the other ²⁹Si signals. The intensity of the peak at about –102 ppm for ZnLa–Zr₃SiBEA is very low. It confirms the presence of low amount of (SiO)₃Si(OH) species formed in the vacant T-atom sites upon dealumination and consumed upon the reaction of these groups with incorporated zirconium cations into SiBEA zeolite, as reported earlier for Nb cations.³⁰

Table 1. Structural characteristics (nitrogen ad(de)sorption, 77 K) of the prepared catalysts.

Sample	$V_{mi}^{a)}$	D_{mi}	E_0	V_{me}	$S_{me}^{a)}$	$D_{me.max}$	$S_{ext.}^{a)}$	S_{BET}	$V_{\Sigma}^{c)}$	$ \Delta\mu_o $
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	cm ³ /g	nm	kJ/mol	cm ³ /g	m ² /g	nm	m ² /g	m ² /g	cm ³ /g	kJ/mol
ZnLaZr-KSKG	0.02	–	–	0.63	140	10.1±1.9	80	260	0.82	11.0
ZnLaZr-A-175	0.01	–	–	–	–	~40 (5-100)	131 ^{d)}	160	0.80	10.5
ZnLaZr-A-380	*	*	–	1.05	170 ^{d)}	28±9.0	–	205	1.08	10.1
ZnLaZr-SBA-15	0.03	–	–	0.64	185	7.4±0.4	80	340	0.81	10.9
ZnLaZr-MCM-41	*	*	–	0.33	500	2.2	40	615	0.46	10.8
ZnLaZr-MCM-48	0.01	–	–	0.50	255	6.5±0.4	85	380	0.69	11.6
ZnLaZr-MCF	0.01	–	–	0.89	190	8.8±1.5	130	355	1.15	11.5
ZnLaZr-SiBEA	0.22 ^{b)}	0.90	3.9	0.37	–	~50	95	570	0.59	–
ZnLa-Zr ₁ SiBEA	0.18 ^{b)}	0.96	4.1	–	–	~38	80	470	0.37	–
ZnLa-Zr ₃ SiBEA	0.14 ²⁾	0.96	4.3	0.17	–	~47 (5-80)	75 ^{d)}	425	0.43	–

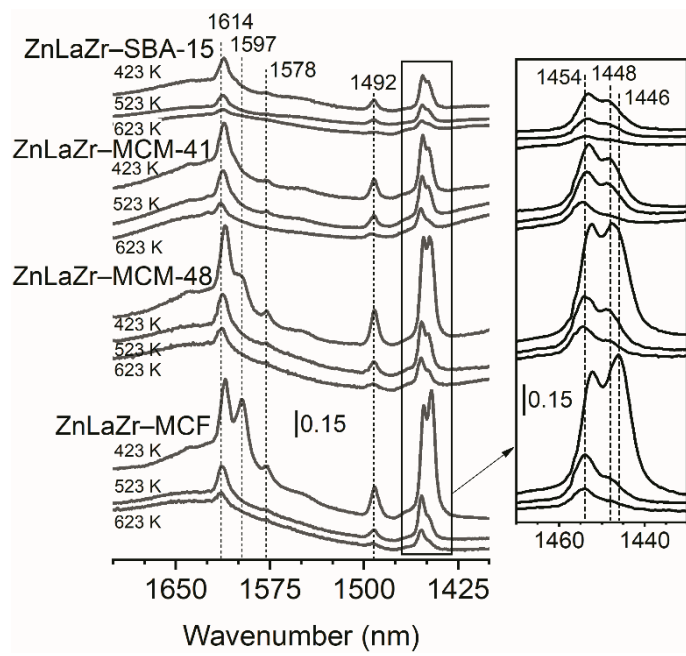
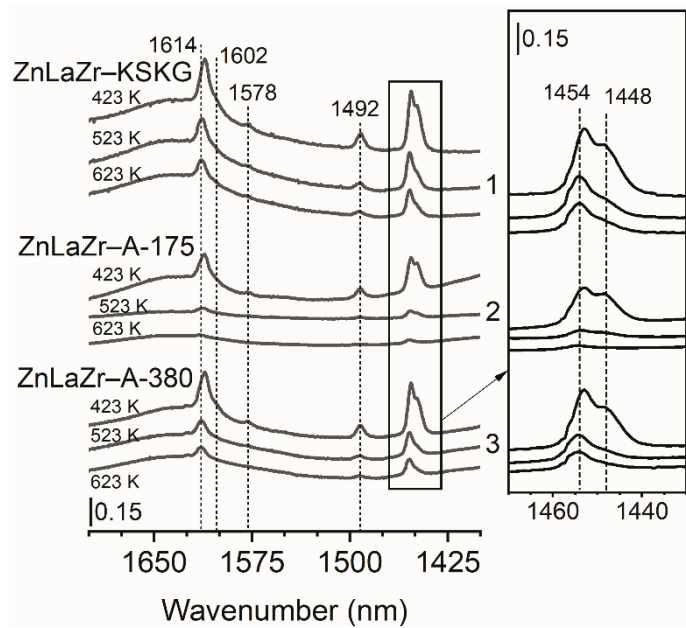
^a calculations using *t-plot* method, ^b micropore parameters according to Dubinin-Radushkevich equation TVFM, ^c total pore volume at p/p₀ = 1.0, ^d total mesopore and external surface area, * not determined.

The particle size and morphology of ZnLaZrSi oxide samples were determined by the SEM analysis and are shown in Figure S2. The samples based on KSKG, A-175 and A-380 contain randomly amorphous particles and small sphere-like primary particles with particle diameters of around 10–20 nm. SEM image of ZnLaZr-SBA-15 shows a typical hexagonal rod-like SBA-15 morphology with a primary particle size of around 1 μm. ZnLaZr-MCM-41 appears as spherical particles with a regular surface. ZnLaZr-MCM-48 possesses the particles of an irregular shape. In turn, ZnLaZr-MCF consists of mainly spherical particles with a partial preservation of the cellular morphology. The images of the samples based on dealuminated BEA zeolite (ZnLa-Zr₁SiBEA and ZnLa-Zr₃SiBEA) indicate a macrostructure similar for BEA zeolites;³⁹ size of the

spherical aggregates is about 100-200 nm. Thereby, BEA zeolite does not deteriorate after treatment with nitric acid and introduction of the heteroelement (Zr) into SiBEA.

Analysis of the nitrogen ad(de)sorption measurements (Table 1, Figure S3) shows that the prepared catalysts are characterized by the developed surface area (up to 615 m²/g) and high total adsorption up to ~ 1.15 cm³/g (at p/p₀ = 1.0). The porosity of KSKG and aerosil supported catalysts is formed as interparticle space between the aggregated silica particles. A-380 support consisting of smaller particles in comparison with A-175 provides the formation of smaller particles of the active phase. The resulting smaller particles lead to mesoporous structure with a visible bend on the isotherm in contrast to A-175 support providing larger particles of the active phase which is expressed by an adsorption increase in the region of higher relative pressures. Deposition of the active phase on silica MMS results in a decrease of the structural characteristics (specific surface area and total pore volume) of the prepared supported catalysts in comparison with the initial silica supports. The obtained materials are characterized by a comparatively homogeneous mesoporous structure indicating preferred uniform deposition of the active phase. For the samples based on SiBEA zeolite an increase in the characteristic adsorption energy in micropores E_0 is observed (3.9, 4.1, 4.3 kJ/mol, Table 1; Dubinin-Radushkevich equation, TVFM) with an increase in the content of active catalytic phase, which is also confirmed by a corresponding decrease in micropore volume (Table 1). It indicates the formation of a finely dispersed phase of the catalyst components inside the micropores, and not only on the surface and their corresponding blocking.

Acid characteristics of the catalysts



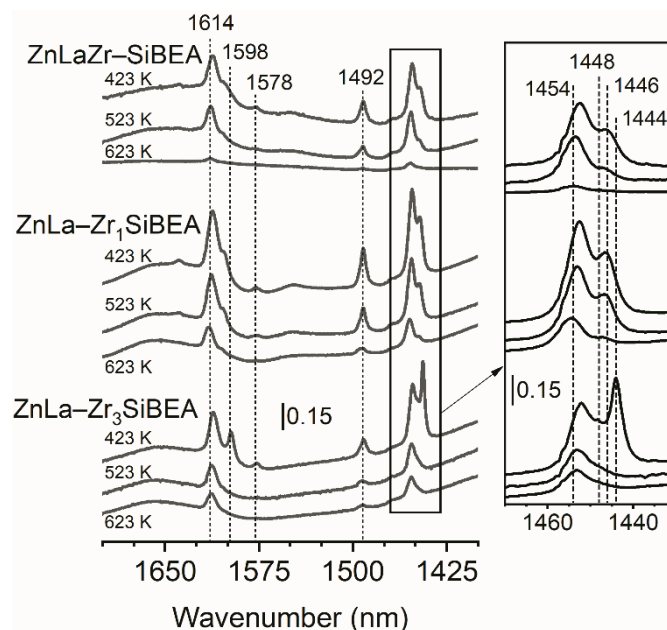


Figure 3. FTIR spectra of adsorbed pyridine the samples after their treatment at appropriate temperature.

As postulated in the Introduction section, selection of the catalyst for ETB process should be based on their acid–base properties. Figure 3 shows the results of investigation of acid sites on the catalyst surface by FTIR spectroscopy of adsorbed pyridine.

The IR adsorption bands are observed at 1614, 1602, 1578, 1492, 1454 and 1448 cm^{-1} for all samples. The additional bands at 1597 and 1446 cm^{-1} are found for ordered mesoporous silica and mesocellular siliceous foam. For dealuminated BEA zeolites one more band at 1444 cm^{-1} is detected. The bands at 1597, 1578 and 1444 cm^{-1} represent physically adsorbed and hydrogen-bonded pyridine, as proved by their almost complete disappearance after desorption of pyridine at 523 K. The band at 1578 cm^{-1} remaining at 523 K arises from coordinately bonded pyridine. The band at 1492 cm^{-1} corresponds to both Brønsted and Lewis acid sites.^{40,41} However, no evidence is found for the band at 1570-1540 cm^{-1} on the samples indicating that there are no Brønsted sites on the surface strong enough to react with pyridine. Intensive bands at 1614, 1454

and 1448 cm^{-1} and the shoulder at 1602 cm^{-1} correspond to pyridine interacting with Lewis acid sites.

Several types of Lewis acid sites (LAS) can be formed on the surface of ZnLaZrSi oxide systems: lanthanum-, zirconium-, and zinc-containing ones. We have previously shown that only weak acid sites with a pyridine desorption temperature of up to 423 K are formed on the surface of LaSi oxide catalysts (the bands at 1602 and 1446 cm^{-1}).³⁶ In support of this, the poorly marked shoulder at 1602 cm^{-1} disappears at a pyridine desorption temperature of 523 K. Therefore, for ZnLaZrSi oxide samples, the FTIR spectra of adsorbed pyridine after its desorption at 523 and 623 K contain IR adsorption bands of coordinatively bonded pyridine only with zinc- and zirconium-containing LAS. The bands at 1614 and 1454 cm^{-1} corresponds to Zn-containing acid sites, as proved by the absence of these bands in the FTIR spectra of adsorbed pyridine for the samples without zinc (Figure S4). The band at 1610 and 1448 cm^{-1} represents Zr-containing acid sites.^{42,43}

Catalytic properties of the ZnLaZrSi-oxide systems

Table 2. The indices of ETB process in the presence of ZnLaZrSi oxide systems prepared with amorphous silicon oxides^a

Catalysts	WHSV, h^{-1}	EtOH Conv., %	Selectivity, %				1,3-BD Yield, %	1,3-BD Product., $\text{g}_{1,3\text{-BD}} \cdot \text{g}_{\text{cat}}^{-1} \cdot \text{h}^{-1}$
			1,3-BD	AA	Ethylene +DEE	Others		
ZnLaZr-KSKG	0.4	58.4	68.0	7.3	18.1	6.7	39.7	0.093
	0.8	40.1	61.4	11.3	22.3	5.0	24.6	0.116
	1.2	33.8	65.0	10.8	19.8	4.4	22.0	0.155
ZnLaZr-A-175	0.4	63.6	65.5	9.9	18.6	6.0	41.7	0.098
	0.8	37.0	58.1	15.7	22.3	3.9	21.5	0.100

	1.2	26.1	57.4	16.5	22.9	3.2	15.0	0.106
ZnLaZr-A-380	0.4	46.1	60.7	13.7	21.0	4.6	28.0	0.066
	0.8	38.9	60.1	12.1	23.6	4.2	23.4	0.110
	1.2	28.5	60.0	12.5	24.2	3.3	17.1	0.120

^aT = 648 K, TOS = 1-5 h, 80 vol% EtOH in water as an EtOH source.

Table 2 presents catalytic performance of ZnLaZrSi oxide systems prepared with amorphous silicon oxides (KSKG, A-175, A-380) for EtOH conversion. At WHSV = 0.4 h⁻¹, the highest values of ethanol conversion (63.6 %) and 1,3-BD yield (41.7 %) are achieved in the presence of ZnLaZr-A-175 (the sample with the lowest S_{BET} and the highest S_{ext} among the studied systems, see Table 1). With an increase in WHSV, the target process indices (ethanol conversion and 1,3-BD selectivity) are decreased with a simultaneous increase in the selectivity of acetaldehyde (AA) and ethanol dehydration products (ethylene and diethyl ether, DEE). For ZnLaZr-A-175, a decrease in the target process indices is more significant than for the catalysts with a higher S_{BET} (ZnLaZr-KSKG, ZnLaZr-A-380). That may be a consequence of the fact that the concentration of Lewis acid sites on the ZnLaZr-A-175 surface is lower than for the samples based on KSKG and A-380 supports (Figure 3). At WHSV = 0.8–1.2 h⁻¹, in the presence of ZnLaZr-KSKG and ZnLaZr-A-380 ethanol conversion is higher than for ZnLaZr-A-175. At WHSV = 1.2 h⁻¹, the catalyst productivity of 1,3-BD formation changes in the following sequence: ZnLaZr-A-175 < ZnLaZr-A-380 < ZnLaZr-KSKG, which is consistent with a change in the values of S_{BET} .

At similar ethanol conversion values (34-39%), the most active catalyst is ZnLaZr-KSKG with 1,3-BD selectivity of 65% and the productivity of 0.155 g_{1,3-BD}·g_{cat}⁻¹·h⁻¹. The ZnLaZr-KSKG sample is also more selective for 1,3-BD formation at ethanol conversion values of 38.9–40.1% (at WHSV = 0.8 h⁻¹). In the ref.²³, higher 1,3-BD selectivity and yield are also achieved in the presence of the 1 %Ag/4 %ZrO₂/SiO₂ catalyst based on KSKG support (in comparison with the

samples based on A-380 or other types of silica). Wherein, ZnLaZr-KSKG catalyst has a maximum mesopore size of 10 nm. In the ref.²¹, the 150 Å ZnZrSi oxide system gave a high selectivity to 1,3-BD compared to the 40 and 60 Å ones at a similar ethanol conversion too.

Among the MMS-based catalysts (Table 3), at WHSV = 0.4 h⁻¹ the highest 1,3-BD selectivity of 63.1 % and the yield of 31.0 % are achieved over ZnLaZr-SBA-15, although ethanol conversion does not exceed 49.2 %. In the presence of ZnLaZr-MCM-48 (a sample with porous structure characteristics comparable to the ZnLaZr-SBA-15 catalyst, Table 1), a comparable 1,3-BD yield of 30.8 % is achieved. Although the ethanol conversion accounts for 69.0 %, the selectivity of 1,3-BD formation is 44.7 %, which is much lower than for ZnLaZr-SBA-15. In the presence of ZnLaZr-MCM-41 and ZnLaZr-MCF (the samples with the highest S_{BET} and pore diameters/volumes, respectively, Table 1) the selectivity of 1,3-BD formation and the yield are lower than for ZnLaZr-SBA-15, also with an increase in WHSV. It should be noted that higher selectivity values towards by-products of ethanol dehydration are achieved in the presence of the catalysts based on MCM-41, MCM-48 and MCF which surface is characterized by a higher amount of weak acid sites (Figure 3, the intensity of adsorption band at 1446 cm⁻¹ at 423 K). Thus, the target process indices are less affected by the structural-dimensional characteristics of ZnLaZrSi oxide systems based on MMS than acid properties of the catalysts.

At similar ethanol conversion (47.6-52.0%), the highest selectivity of 1,3-BD formation among the MMS-based catalysts is achieved in the presence of ZnLaZr-SBA-15. Wherein, 1,3-BD productivity of ZnLaZr-MCM-48 is the highest under such comparison conditions, but the main conversion products are ethylene and DEE.

Table 3 The indices of ETB process in the presence of ZnLaZr-MMS mesoporous amorphous systems^a

Catalysts	WHSV, h ⁻¹	EtOH Conv., %	Selectivity, %				1,3-BD Yield, %	1,3-BD Product., g _{1,3-BD} ·g _{cat} ⁻¹ ·h ⁻¹
			1,3-BD	AA	Ethylene +DEE	Others		
ZnLaZr-SBA-15	0.4	49.2	63.1	10.6	19.4	6.8	31.0	0.073
	0.8	39.8	58.9	13.4	23.1	4.6	23.4	0.11
	1.2	33.1	62.6	11.0	22.7	3.7	20.7	0.146
ZnLaZr-MCM-41	0.4	60.8	57.1	5.0	32.3	5.6	34.7	0.082
	0.8	51.3	55.4	6.3	34.1	4.2	28.4	0.133
ZnLaZr-MCM-48	0.4	69.0	44.7	4.8	45.0	5.5	30.8	0.072
	1.2	52.0	41.3	5.2	48.7	3.8	21.5	0.151
ZnLaZr-MCF	0.4	54.5	48.2	10.4	36.9	4.3	26.3	0.062
	0.8	47.6	45.4	9.5	41.1	4.0	21.6	0.101
	1.2	35.0	44.6	10.7	42.0	2.7	15.6	0.110

^a T = 648 K, TOS = 1-5 h, 80 vol% EtOH in water as an EtOH source.

The data of catalytic performance of ZnLaZrSi oxide systems prepared with dealuminated BEA zeolite for EtOH conversion are presented in Table 4. At WHSV = 0.4 h⁻¹ in the presence of ZnLaZr-SiBEA, the highest ethanol conversion of 67.4 % is achieved, while 1,3-BD selectivity is only 34.6 % because of the preferred occurrence of by-product dehydration. The selectivity of ethylene and diethyl ether formation accounts for 51.3 %. For the catalyst samples prepared by post-synthetic method of zirconium incorporation (ZnLa-Zr₁SiBEA, ZnLa-Zr₃SiBEA), the conversion of ethanol is lower (less than 58.0 %), but the selectivity of 1,3-BD formation is higher (more than 46.4 %). Such increase in 1,3-BD selectivity, despite a lower (several-fold) zirconium content in the ZnLa-Zr₁SiBEA catalyst sample, can be explained by the presence of

Lewis acid sites in the framework of zeolite as isolated tetrahedral Zr(IV) species. These sites are more active in aldol condensation of acetaldehyde than zirconium present as zirconium oxide.^{32,38}

Table 4 The indices of ETB process in the presence of ZnLa-ZrSiBEA microporous crystalline systems^a

Catalysts	WHSV, h ⁻¹	EtOH Conv., %	Selectivity, %				1,3-BD Yield, %	1,3-BD Product., g _{1,3-BD} ·g _{cat} ⁻¹ ·h ⁻¹
			1,3-BD	AA	Ethylene +DEE	Others		
ZnLaZr-SiBEA	0.4	67.4	34.6	7.5	51.3	6.1	23.3	0.055
	1.6	32.8	30.5	8.8	53.3	7.4	12.4	0.070
ZnLa-Zr ₁ SiBEA	0.4	58.0	46.4	12.3	31.6	9.7	24.1	0.057
	2.88	31.1	61.7	3.4	12.0	22.9	19.2	0.324
ZnLa-Zr ₃ SiBEA	0.4	36.7	50.8	19.9	11.8	17.5	18.6	0.044
	0.8	19.9	51.1	12.5	10.3	17.1	10.2	0.048
	1.2	14.5	52.0	21.7	9.9	16.1	7.5	0.053

^aT = 648 K, TOS = 1-5 h, 80 vol% EtOH in water as an EtOH source.

At similar ethanol conversion (32.8-36.7 %), the highest 1,3-BD selectivity among the microporous crystalline systems is achieved over ZnLa-Zr₁SiBEA. The selectivity of ethylene+DEE formation is comparable for both Zr-containing dealuminated BEA zeolites, but the residual quantity of AA is higher for ZnLa-Zr₃SiBEA. In contrast, in the presence of ZnLa-Zr₁SiBEA, higher selectivity towards others products is observed, probably, because of non-selective transformation of AA. In the presence of ZnLaZr-SiBEA, ethanol is preferably converted to ethylene and DEE.

It is worth emphasizing that ethanol conversion decreases with an increase in zirconium content in ZnLa-ZrSiBEA samples (Table 4). Such tendency may be due to an excess of

zirconium cations in comparison with vacant T-atom sites formed in the BEA zeolite framework upon dealumination step with nitric acid treatment and, as a result, subnanoclusters of zirconium oxide could be formed. This assumption is consistent with the results of FTIR spectroscopic measurements of adsorbed pyridine (Figure 3): at 523 K the intensity of the bands corresponding to Lewis acid sites is higher for ZnLa–Zr₁SiBEA than for ZnLa–Zr₃SiBEA. The formation of subnanoclusters with an increase of the heteroelement content in dealuminated BEA zeolite involves a decrease in ethanol-acetaldehyde mixture conversion (without a decrease in 1,3-BD selectivity) as was previously observed for Nb-containing catalysts.³⁰

With an increase in WHSV, a decrease in ethanol conversion is observed for both ZnLa–ZrSiBEA samples, but in the presence of ZnLa–Zr₁SiBEA 1,3-BD selectivity increases markedly (Figure 4). This may occur because, at a high linear flow rate, the reacting molecules primarily interact with open Zr⁴⁺ Lewis acid sites, represented by isolated Zr(IV) in tetrahedral positions of the zeolite crystalline structure and connected to three -O-Si linkages and one OH-group. According to Sushkevich et al.³², open Zr⁴⁺ Lewis acid sites are more active towards 1,3-BD formation.

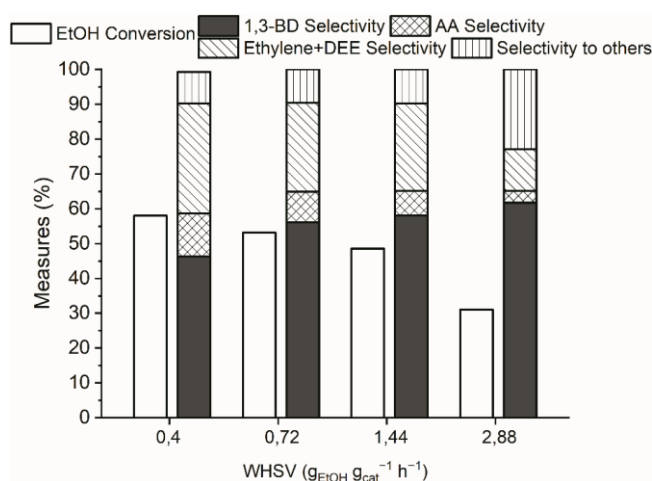


Figure 4. Effect of WHSV on ethanol conversion and selectivity to 1,3-butadiene, acetaldehyde, ethylene+diethyl ether and others products in ETB process over ZnLa–Zr₁SiBEA catalyst (T = 648 K; time-on-stream = 1–5 h; 80 vol% EtOH in water as an EtOH source).

The development of catalysts containing Lewis acid sites, represented by isolated Zr⁴⁺ species, is the pathway to obtain high performance zeolite catalysts. In the presence of ZnLa–Zr₁SiBEA, 1,3-BD formation rate reaches 1.67 $\mu\text{mol}\cdot\text{g}_{\text{cat}}^{-1}\cdot\text{sec}^{-1}$ during the conversion of 80 vol% EtOH aqueous mixture at WHSV = 2.88 h⁻¹, and accounts for 1.64, 1.23 and 0.79 $\mu\text{mol}\cdot\text{g}_{\text{cat}}^{-1}\cdot\text{sec}^{-1}$ during the conversion of 96, 80 and 50 vol% EtOH aqueous mixtures, respectively, at WHSV = 1.44 h⁻¹. These values are comparable with those obtained in the presence of 1 %Ag/Zr_{1,3}SiBEA catalyst prepared by post-synthetic incorporation of zirconium into dealuminated BEA zeolite matrix.³⁸ At the same time, the ZnLa–Zr₃SiBEA catalyst productivity of 1,3-BD formation does not reach the value for ZnLa–Zr₁SiBEA even with a triple increase in WHSV.

In general, among the studied ZnLaZrSi-oxide systems, during the conversion of 80 vol% EtOH aqueous mixture (WHSV=1.2 h⁻¹, T=648 K) the highest catalyst productivity of 1,3-BD formation (0.155 $\text{g}_{1,3\text{-BD}}\cdot\text{g}_{\text{cat}}^{-1}\cdot\text{h}^{-1}$) is achieved for ZnLaZr–KSKG sample and a slightly lower productivity (0.146 $\text{g}_{1,3\text{-BD}}\cdot\text{g}_{\text{cat}}^{-1}\cdot\text{h}^{-1}$) for ZnLaZr–SBA-15 sample. The productivity values for ZnLa–Zr₁SiBEA account for 0.239 $\text{g}_{1,3\text{-BD}}\cdot\text{g}_{\text{cat}}^{-1}\cdot\text{h}^{-1}$ at WHSV = 1.44 h⁻¹ and 0.324 $\text{g}_{1,3\text{-BD}}\cdot\text{g}_{\text{cat}}^{-1}\cdot\text{h}^{-1}$ at WHSV = 2.88 h⁻¹. Analyzing the data presented in Tables 2-4, it is concluded that no significant change in the target product selectivity, i.e. in a relative change in 1,3-BD selectivity within 10 % for one catalyst, with the ethanol conversion variations in the range of 20-60 %. This is consistent with the data obtained by Baba et al.^{44,45} for Zn-talk and ZnGe-talk catalysts in the mentioned ethanol conversion range.

For ZnLaZr–KSKG and ZnLaZr–SBA-15 catalysts, EtOH conversion has been studied as a function of time-on-stream (Figure 5). In the presence of the catalysts, the values of EtOH conversion and 1,3-BD yield do not reduce significantly with time-on-stream up to 700 min. The constancy of 1,3-BD selectivity is also observed. A slight decrease in the selectivity to ethanol dehydration products with an increase in AA selectivity is observed after 9 h of operation.

The data in Figure 6 allow comparing the formation rates of 1,3-BD, acetaldehyde and ethylene+DEE, calculated per unit mass (r_m), per unit surface area (r_s) and per unit external surface area ($r_{\text{extern.}}$) of the catalyst. In the case of 1,3-BD formation, r_m values are higher for ZnLaZr–KSKG and ZnLaZr–SBA-15 catalysts. The highest r_s value of BD formation (as well as AA formation) is observed for ZnLaZr–A-175, which may be attributed to the lowest S_{BET} for this sample. Comparing the values of $r_{\text{extern.}}$, the degree of external surface use is supposed to be lower for ZnLaZr–A-175 as opposed to the samples based on KSKG and SBA-15. Furthermore, the $r_{\text{extern.}}$ value for ZnLaZr–A-175 is comparable to that for ZnLaZr–MCF probably because of the similar values of S_{ext} . Overall, the 1,3-BD formation rates are changed in the following sequences (the samples are marked as the silica support used):

$$r_m: \text{KSKG} > \text{SBA-15} > \text{MCF} \geq \text{A-175} > \text{Zr}_3\text{SiBEA},$$

$$r_s: \text{A-175} > \text{KSKG} > \text{SBA-15} > \text{MCF} > \text{Zr}_3\text{SiBEA},$$

$$r_{\text{extern.}}: \text{KSKG} > \text{SBA-15} > \text{MCF} \geq \text{A-175} > \text{Zr}_3\text{SiBEA}.$$

□ EtOH Conversion ★ 1,3-BD Selectivity ◆ AA Selectivity ▲ Ethylene+DEE Selectivity ○ 1,3-BD Yield

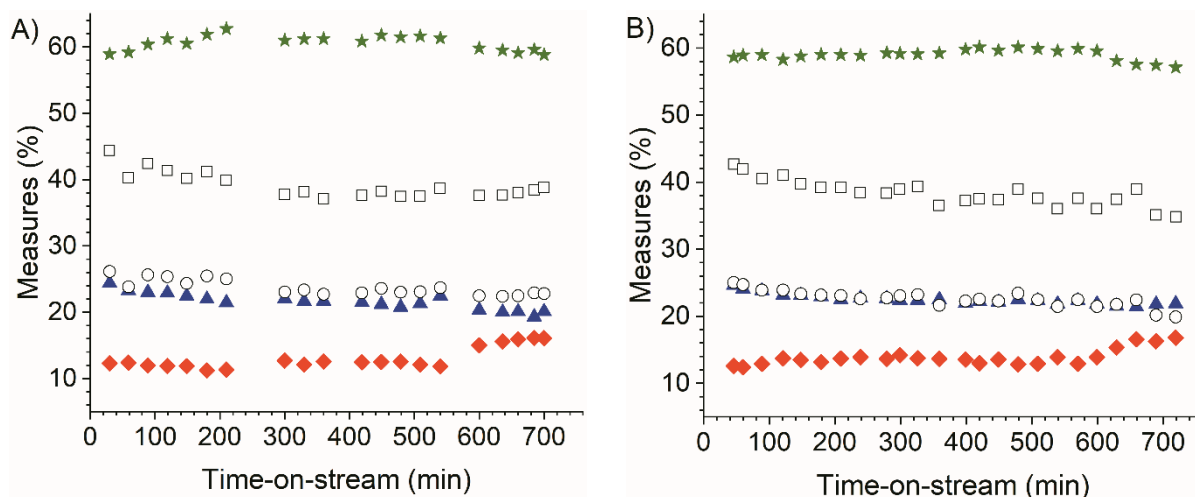


Figure 5. Ethanol conversion, 1,3-butadiene yield, and selectivity to products over ZnLaZr–KSKG (A) and ZnLaZr–SBA-15 (B) catalysts as a function of time-on-stream ($T = 648\text{ K}$; $\text{WHSV} = 0.8\text{ g}_{\text{EtOH}} \cdot \text{g}_{\text{cat}}^{-1} \cdot \text{h}^{-1}$; 80 vol% EtOH in water as an EtOH source).

The high r_m , r_s , and $r_{\text{extern.}}$ values of ethylene+DEE formation are achieved in the presence of ZnLaZr–MCF, while the r_s and $r_{\text{extern.}}$ values are generally low for ZnLa–Zr₃SiBEA. The lowest values of 1,3-BD formation over zeolite supported catalysts are obviously associated with the support microporosity which can be involved into the catalytic process only partially due to a limited micropore size and corresponding pore blocking. Therefore, acid characteristics of the resulting catalyst for ETB process are suggested to be more significant than high S_{BET} and $S_{\text{ext.}}$ values of the silica support used for preparation of ZnLaZrSi oxide systems.

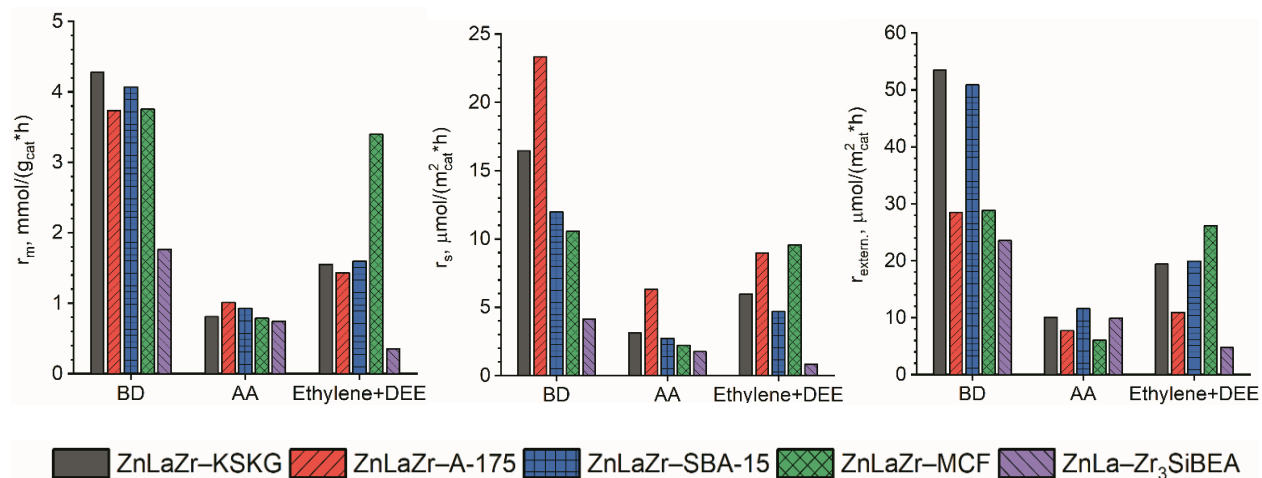
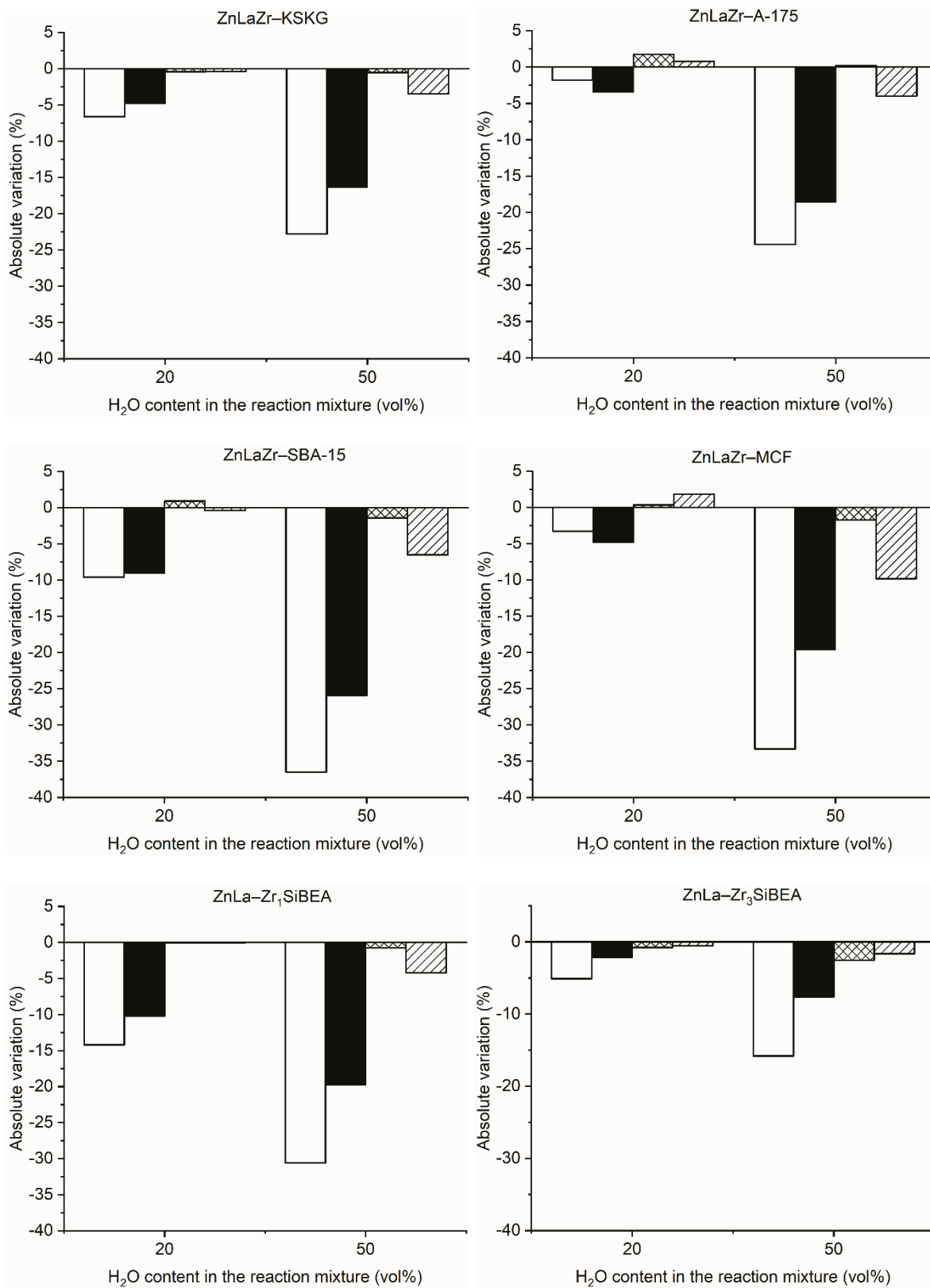


Figure 6. The formation rates of 1,3-butadiene, acetaldehyde and ethylene+diethyl ether, calculated per unit mass (r_m), per unit surface area (r_s) and per unit external surface area ($r_{extern.}$) of the catalyst samples under the following conditions: $T = 648$ K; time-on-stream = 1–5 h; $WHSV = 0.8 \text{ g}_{\text{EtOH}} \text{ g}_{\text{cat}}^{-1} \text{ h}^{-1}$; 80 vol% EtOH in water as an EtOH source.

Water vapor effect on catalytic performance of the ZnLaZrSi oxide systems for ETB-process

The water content in aqueous EtOH feed is a very important parameter affecting the target indices of ETB process.^{12,15,17,37} An absolute change in the values of ethanol conversion and the yields of 1,3-butadiene, acetaldehyde and ethylene+diethyl ether can be estimated using Figure 7. The complete data of catalytic performance of ZnLaZrSi-oxide systems in EtOH-water mixture conversion are summarised in Table S1.



EtOH Conversion
 1,3-BD Yield
 AA Yield
 Ethylene+DEE Yield

Figure 7. Effect of water vapor on absolute change of ethanol conversion and the yields of 1,3-butadiene, acetaldehyde, ethylene+diethyl ether in ETB process over ZnLaZrSi oxide systems ($T = 648$ K, time-on-stream = 1-5 h, WHSV = 0.8 h^{-1} ; for ZnLa–Zr₁SiBEA sample WHSV = 1.44 h^{-1}).

With the change of 96 vol% EtOH aqueous mixture feed to 80 vol% EtOH one, a noticeable decrease in both ethanol conversion and 1,3-BD selectivity is observed. 1,3-BD yield decreases by less than 5 % for the catalyst samples based on KSKG, A-175, MCF and Zr₃SiBEA supports, and by more than 10 % for ZnLaZr–SBA-15 and ZnLa–Zr₁SiBEA. At the same time, the selectivity to intermediate (AA) and side (ethylene+DEE) products rises to some extent. As a result, the yields of AA and ethylene+DEE are reduced by less than 2 %, and even rise for some catalysts. With the change of 96 vol% EtOH aqueous mixture feed to 50 vol% EtOH one, the values of ethanol conversion and 1,3-BD selectivity decreased more significantly. There is a substantial decrease in 1,3-BD yield. Wherein, the selectivity of ethylene+DEE formation increases by several percent, and AA selectivity almost doubles (see Table S1).

The obtained data are consistent with the results and conclusions of the ref.¹²: the reason for the deactivation of ETB process catalyst is water adsorption on the LAS responsible for aldol condensation of acetaldehyde. Let's consider the effect of water vapor on activity of the catalysts to dehydrogenation (selectivity to AA and 1,3-BD as the main product of AA aldol condensation) and dehydration (selectivity to ethylene and DEE). The selectivity of 1,3-BD formation declines however AA selectivity increases (see Table S1), and a fraction of non-condensed AA rises in the product stream. Therefore, the effect of water vapor on ethanol dehydrogenation is considered less significant than on aldol condensation. The selectivity of

ethylene+DEE formation rises noticeably, herewith, ethylene+DEE yield is reduced along with 1,3-BD one but to varying degrees (Figure 7).

In the refs.^{15,17} the reason for the deactivating effect of water vapor are proposed to be H₂O adsorption on LAS with their transformation to Brønsted acid sites (BAS) responsible for EtOH dehydration. However, this explanation seems to be incomplete. The water vapor released in the conversion of rectified EtOH into 1,3-BD and ethylene would be sufficient to hydrate the catalyst surface with the formation of low amount of BAS. Moreover, as mentioned in the ref.⁴⁶, only partial blocking of LAS occurs on the hydrated surface of ZrO₂. Lebedev et al.¹⁸ have noted the negative effect of the presence of water in EtOH feed on 1,3-BD yield in ETB process, whereas an increase in AA yield and a decrease in ethylene yield are observed. According to the ETB process mechanism proposed by Lebedev, C-C coupling occurs between ethylene and acetaldehyde via the Prins reaction. Therefore, the main reason for the decrease in 1,3-BD yield has been determined as a decrease in the rate of ethylene formation. Based on the thermodynamic calculations,³ in ETB process the formation of C₄ products via the Prins reaction has shown to be unlikely. An increase in AA yield and a decrease in ethylene yield can be explained by a deactivation of EtOH dehydration sites. A decrease in the selectivity towards dehydration products with an addition of water to the feed mixture of EtOH/AA is also observed in ref.^{19,47} Therefore, LASs responsible for aldol condensation of acetaldehyde just remain blocked by water without forming new EtOH dehydration sites.

The ZnLa–Zr₃SiBEA sample is an exception in this respect. In the presence of this catalyst 1,3-BD selectivity is even higher during the conversion of 50 vol% EtOH aqueous mixture as opposed to 96 vol% EtOH one, while AA selectivity remains almost unchanged (Table S1). Besides, in the presence of ZnLa–Zr₁SiBEA an absolute decrease of 1,3-BD selectivity is less

than for the studied non-zeolite catalysts. A similar tendency with a change in ethanol conversion, 1,3-BD and AA selectivity is also observed during the conversion of 50 vol% and 96 vol% EtOH aqueous mixtures in the presence of Zn₁Ta₁SiBEA catalyst (in contrast to Cu₁Ta₁SiBEA and Ag₁Ta₁SiBEA).³⁷ Perhaps, in the case of Zn-M-containing zeolite catalysts, the bifunctional sites like Zn-La, Zn-Zr, Zn-Ta species are formed in the SiBEA structure, as it has been previously noted for Zn-Y/Beta catalysts.⁴⁸ The intermediates react with each other on these sites to generate 1,3-BD rather than desorb as unwanted byproducts. Therefore, H₂O as a Lewis base is adsorbed on the catalyst surface blocking of a part of LASs responsible for aldol condensation of acetaldehyde without the formation of new EtOH dehydration sites.

It should be noted that higher 1,3-BD selectivity is achieved during the conversion of ethanol-aqueous mixtures in the presence of ZnLaZr-KSKG as opposed to ZnZr-KSKG (see Table S2, ¹⁶). This may be attributed to the fact that Brønsted basic sites being active in aldol condensation⁴⁹⁻⁵¹ (including ETB process)⁵² are formed during deposition of lanthanum oxide or hydroxide on SiO₂.⁵³ The similar lanthanum effect on EtOH conversion rates and 1,3-BD selectivity can be observed by comparing the data of the ref.²¹ (the conversion of 96 vol% EtOH aqueous mixture in the presence of the 1.5%Zr-0.5%Zn/SiO₂ catalyst) and the ref.¹³ (the conversion of 92 vol% EtOH aqueous mixture in the presence of the ZrZn-La/SiO₂ catalyst). The addition of 2 wt % lanthanum to the ZrZn/SiO₂ system promotes an increase in 1,3-BD selectivity, and an increase in the lanthanum content up to 3 wt % increases 1,3-BD yield. However, the productivity of 1,3-BD formation is higher for the unpromoted ZrZn/SiO₂ catalyst, as in our case.

In the lattice of dealuminated BEA zeolite, there are many unsaturated vacancies. Therefore, in the presence of ZnLa-ZrSiBEA catalysts unvaried high 1,3-BD selectivity with the change of

96 vol% EtOH aqueous mixture feed to 50-80 vol% EtOH one (see Table S1) can be explained by a larger number of Si-O-La-OH sites formed in the dealuminated BEA zeolite.

There is another problem of slow deactivation of ZnLaZrSi oxide systems in ETB process at long time-on-stream,²⁰ e.g. due to the formation of heavy by-products via condensation reactions. ¹H-¹³C CP/MAS NMR investigations of organic deposits on a spent catalyst allow finding out the nature of occluded organic compounds and evaluating mechanistic aspects of deactivation.^{19,54}

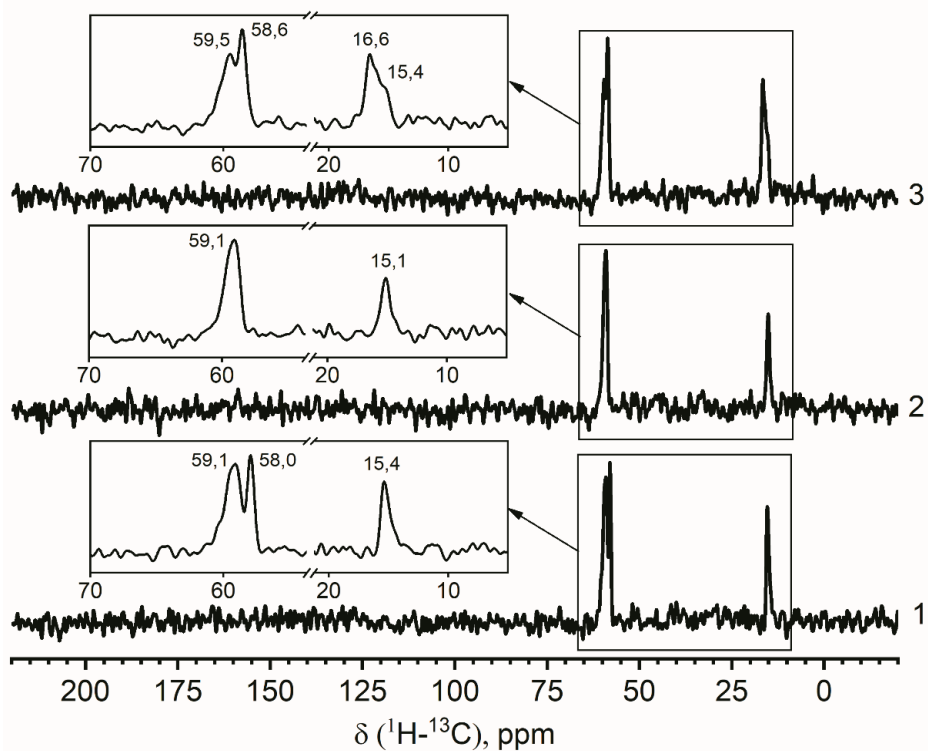


Figure 8. ¹H – ¹³C CP/MAS NMR spectra of the samples after catalysis: 1 – ZnLaZr–KSKG; 2 – ZnLaZr–SBA-15; 3 – ZnLa–Zr₃SiBEA.

The ¹H-¹³C CP/MAS NMR spectra of ZnLaZr–KSKG, ZnLaZr–SBA-15 and ZnLa–Zr₃SiBEA catalysts obtained after operation in the ethanol to 1,3-butadiene conversion (T=648 K, WHSV = 0.8 g_{EtOH} g_{cat}⁻¹ h⁻¹, 80 vol% EtOH in water as an EtOH source, time-on-stream = 5 h) are shown in

Figure 8. There are ^{13}C NMR signals at 15–17 ppm due to primary carbon atoms in $\text{CH}_3\text{-R}$ compounds. The signals at 58-60 ppm can reflect the presence of carbon atoms in $\text{-CH}_2\text{-O-}$ and $\text{CH}_3\text{-O-}$ functional groups, and/or tertiary carbon atoms like CH(-R)_3 .⁵⁵ The signals of alkenes, alkynes, aromatic compounds, esters, acids, aldehydes and ketones are absent, likely because these products are not formed during ETB process or are easily desorbed from the catalyst surface. Thus, the presence of water vapor in the feed prevents carburizing of the catalyst during its operation in the ETB process,¹⁹ in particular by reducing the formation of C_{6+} compounds.¹⁷ In the reaction conditions water vapor, being more efficient oxidizing agent than O_2 ,⁵⁶ can oxidize C_{6+} compounds at the time of their formation, when their reactivity is still high. Also, the blocking of strong acid sites being active in the formation of C_{6+} compounds,^{19,57} as a result of water vapor adsorption, helps to reduce the formation rate of these compounds. However, this topic will be the subject of a separate study.

CONCLUSIONS

The ZnLaZrSi oxide systems prepared with a silica component of the different type (nature) have been studied in the 1,3-BD production from aqueous ethanol. The commercially available silica (KSKG, A-175, A-380), ordered mesoporous silica (SBA-15, MCM-41, MCM-48), mesocellular siliceous foam (MCF) and dealuminated BEA zeolites have been used. The characteristics of the silica support porous structure, such as total porosity, pore size distribution, specific and external surface areas, are found not to be critical parameters for achieving a high 1,3-BD yield during the EtOH- H_2O mixture conversion in the presence of ZnLaZrSi oxide catalysts. The quantity and strength of Lewis acid sites, which in turn differ depending on the choice of the silica material, have a significant impact on 1,3-BD selectivity and yield. Ethylene

and diethyl ether selectivity is increased with the concentration of weak Lewis acid sites, thereby a decrease in 1,3-butadiene selectivity is observed.

The highest values of the selectivity of 1,3-BD formation and yield as well as stability toward deactivation in the presence of water vapor are achieved over ZnLa-(6%)Zr-KSKG, ZnLa-(6%)Zr-SBA-15 and ZnLa-Zr₁SiBEA prepared by incorporation of zirconium cations in the framework of BEA zeolite as isolated tetrahedral Zr(IV) species. In the presence of ZnLaZrSi oxide catalysts based on MCM-41, MCM-48 and MCF with a developed porous structure, a high selectivity to ethanol dehydration products is observed, which is caused by a high amount of weak Lewis acid sites on the surface.

The main reason for the decrease in 1,3-BD yield in the presence of water vapor in the reaction mixture is shown to be the deactivation of acetaldehyde condensation sites on the catalyst surface, while the rate of acetaldehyde formation decreases slightly. In the presence of lanthanum-containing ZnLaZrSi oxide system the selectivity of 1,3-BD formation is enhanced. According to ¹H-¹³C CP/MAS NMR spectroscopic results, the use of aqueous ethanol as the feed for ETB process is found to prevent carburization of the catalysts.

To prepare the ZnLa-(6%)Zr-KSKG catalyst, KSKG industrial silica gel (silica gel granular with large pores) has been used; it is the cheapest silica material among the used ones. Therefore, a highly efficient catalyst for ETB process can be prepared in a simple method (by wet kneading) using the mentioned inexpensive material, that will enhance the prospects for successful implementation of 1,3-butadiene production from ethanol in the future. However, the constancy of 1,3-BD selectivity is observed in the presence of the systems based on ZrSiBEA with isolated tetrahedral Zr(IV) species even using 50 vol% EtOH aqueous mixture as a feed. Moreover, the productivity of ZnLa-Zr₁SiBEA catalyst accounts for 0.324 g_{1,3-BD}·g_{cat}⁻¹·h⁻¹ (T = 648 K, WHSV

= 2.88 h⁻¹, 80 vol% EtOH in water as an EtOH source). Therefore, for industrial realization of ETB process the development of the effective catalysts based on dealuminated BEA zeolites is also promising.

ASSOCIATED CONTENT

Supporting Information. The following file is available free of charge.

²⁹Si and ¹H – ²⁹Si CP MAS NMR spectra of ZnLa–Zr₃SiBEA sample, SEM images of the ZnLaZrSi samples at three magnifications, the indices of EtOH-water mixture conversion in the presence of ZnLaZrSi oxide systems, the indices of ETB process in the presence of 2 %ZnO/6 %ZrO₂-KSKG (DOCX).

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The manuscript was written through contributions of all authors. All authors have given approval to the final version of the manuscript.

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Graphical Abstract

What is the better silica support
for the catalyst of
1,3-butadiene production
from ethanol-water mixture?

